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Hydrothermal synthesis and characterization of straw bundle-like lithium sodium disilicate (silinaite) micro-rods

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Abstract

Highly mono-shaped straw bundle-like silinaite micro-rods have been synthesized using an aqueous mixture of silicic acid, lithium sulfate, and sodium hydroxide under hydrothermal conditions at 180°C. The synthesized materials were characterized by powder X-ray diffraction technique and Fourier transform infrared spectroscopy. The morphologies of the synthesized materials were studied by field emission scanning electron microscopy technique. The synthesized micro-rods are composed of long, fine nanosheets. Electronic absorption and photoluminescence spectra of the synthesized materials were also studied. With increasing reaction time, a blueshift was observed in the UV-vis electronic absorption and emission spectra of the synthesized materials.

Keywords: Hydrothermal method, Lithium sodium silicate, Optical properties, Micro-rods

Background

Layered compounds are of scientific and technological interest. They act not only as exchanger and fillers in detergents and nanocomposites [1,2] but also as hosts for intercalation of functional molecules or ions to yield hybrid materials [3]. Silinaite (lithium sodium silicate, $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$) is a monophyllosilicate consisting of layers of $[\text{SiO}_4]$ tetrahedra parallel to (001). The lithium ions are held in tetrahedral coordination by four oxygen ions from four $[\text{SiO}_4]$ tetrahedra and pull the layers together. The octahedra around the sodium ions consist of four oxygen atoms from water molecules and two oxygen ions which also belong to $[\text{SiO}_4]$ and $[\text{LiO}_4]$ tetrahedra. The $[\text{NaO}_6]$ octahedra bulge the layers somewhat apart so that the layers are slightly undulated [4].

Natural silinaite was identified and separated from minerals in 1990 [5,6]. Also, the crystal structure of silinaite was determined in 1991 [7]. Bulk silinaite was synthesized by dispersion of silica in aqueous solutions

of NaOH and LiOH at 90°C to 150°C [4]. Also, the compound is synthesized by hydrothermal crystallization of the gel formed by an aqueous mixture of water glass, NaOH, and LiOH [8]. Among the low-temperature methods, the hydrothermal method is very versatile for the synthesis of nanophase materials and has been well established. This is an environment friendly method because the reactions take place in closed-system conditions and save energy [9-11]. In this research work, we report a simple hydrothermal approach to the synthesis of new multi-layer straw bundle-like silinaite micro-rods with a regular new morphology. The morphology and optical properties of the obtained materials are investigated. Also, cell parameter refinement performed by Celref software version 3 (Laboratoire des Matériaux et du Génie Physique de l'Ecole Supérieure de Physique de Grenoble, Grenoble, France) shows the preferential orientation of crystals in the products due to changing the reaction time.

Methods

All chemicals were of analytical grade, obtained from commercial sources, and used without further purification. Phase identifications were performed on a Siemens D5000 powder X-ray diffractometer (Siemens,

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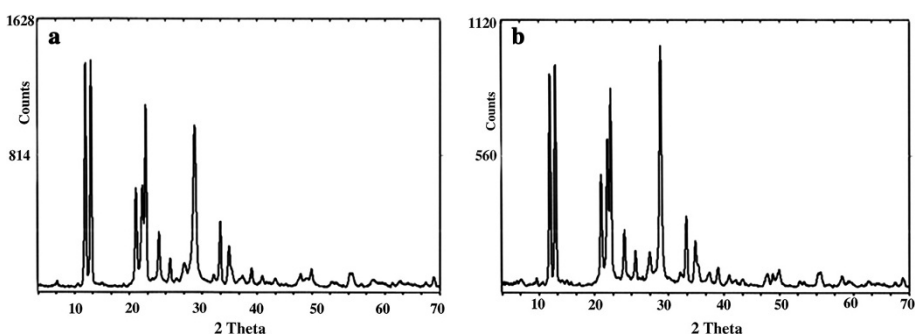


Figure 1 PXRD patterns of the $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ micro-rods hydrothermally synthesized at 180°C . After (a) 48 and (b) 72 h.

Munich, Germany) using CuK_α radiation. The morphology of the obtained materials was examined with a Philips XL30 scanning electron microscope (Philips, Amsterdam, Netherlands). Absorption and photoluminescence spectra were recorded on a Jena Analytik Specord 40 (Jena, Germany) and a Perkin Elmer LF-5 (Waltham, MA, USA) spectrometer, respectively. Cell parameter refinement was performed by Celref version 3 software.

Synthesis of lithium sodium disilicate (silinaite)

In a typical synthetic experiment, 0.3 g (2.3 mmol) of $\text{Li}_2\text{SO}_4 \cdot \text{H}_2\text{O}$ ($M_w = 127.95 \text{ gmol}^{-1}$) and 0.44 g (4.6 mmol)

of $\text{SiO}_2 \cdot \text{H}_2\text{O}$ ($M_w = 96.11 \text{ gmol}^{-1}$) were added to 30 mL of hot aqueous solution of NaOH (0.23 M) under magnetic stirring at 80°C . The resultant solution was stirred further for 15 min and then diluted to 60 mL. The resultant solution was transferred into a 100-mL Teflon-lined stainless steel autoclave. The autoclave was sealed and heated at 180°C for 48 h (sample a) or 72 h (sample b) and then cooled to room temperature by water immediately. When the reaction was completed, a white powder was collected from the solution and washed with distilled water. The prepared powder was dried at 110°C for 20 min under normal atmospheric conditions.

Table 1 Crystallographic data of the hydrothermally synthesized $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ nanomaterials obtained after 48 h at 180°C

Number	2θ	Intensity	Inter-planar space
1	11.86	1,282	7.456
2	12.77	1,300	6.928
3	20.13	506	4.408
4	21.19	303	4.189
5	21.73	803	4.087
6	23.90	284	3.719
7	25.75	142	3.457
8	28.13	75	3.169
9	29.77	901	2.999
10	32.93	39	2.717
11	33.98	360	2.636
12	35.42	166	2.532
13	39.13	94	2.300
14	40.90	45	2.204
15	43.02	36	2.101
16	47.21	41	1.924
17	48.97	71	1.859
18	69.12	50	1.358

Table 2 Crystallographic data of the hydrothermally synthesized $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ nanomaterials obtained after 72 h at 180°C

Number	2θ	Intensity	Inter-planar space
1	11.83	837	7.471
2	12.72	877	6.956
3	20.10	401	4.414
4	21.16	359	4.195
5	21.68	572	4.096
6	23.89	194	3.722
7	25.68	114	3.466
8	27.99	96	3.185
9	29.76	942	3.000
10	33.92	259	2.641
11	35.41	124	2.533
12	37.59	37	2.391
13	39.05	67	2.304
14	47.05	37	1.930
15	47.99	29	1.894
16	48.91	44	1.861
17	59.22	34	1.559
18	69.06	52	1.358

Table 3 Refined cell parameter data of the LiNaSi₂O₅·2H₂O nanomaterials hydrothermally synthesized at 180°C

Refined cell parameters	Reaction times	
	48 h	72 h
<i>a</i>	5.021	5.077
<i>b</i>	8.2895	8.3118
<i>c</i>	14.5418	14.4189

Results and discussion

PXRD analysis

The crystal phase and chemical composition of the synthesized samples were examined by powder X-ray diffraction (PXRD) technique. The PXRD patterns of the synthesized LiNaSi₂O₅·2H₂O samples, obtained after 48 and 72 h, are shown in Figure 1. All diffraction peaks in the PXRD patterns of the as-prepared LiNaSi₂O₅·2H₂O materials can be clearly indexed to the monoclinic phase of the natural silinaite with space group A2/n. There are four molecules per unit cell, *Z* = 4 [7], which is identified by a group of the reflections around $2\theta = 21.3^\circ$ in the X-ray diffraction diagram [4,6-8,12]. The corresponding cell parameters are *a* = 5.06 Å, *b* = 8.33 Å, *c* = 14.38 Å, and $\beta = 96.67^\circ$ [13]. The measured PXRD data show no impurity peaks, indicating the high purity of the as-obtained samples. Because of the different coordination modes of Na⁺ (octahedral) and Li⁺ (tetrahedral) ions in the structure,

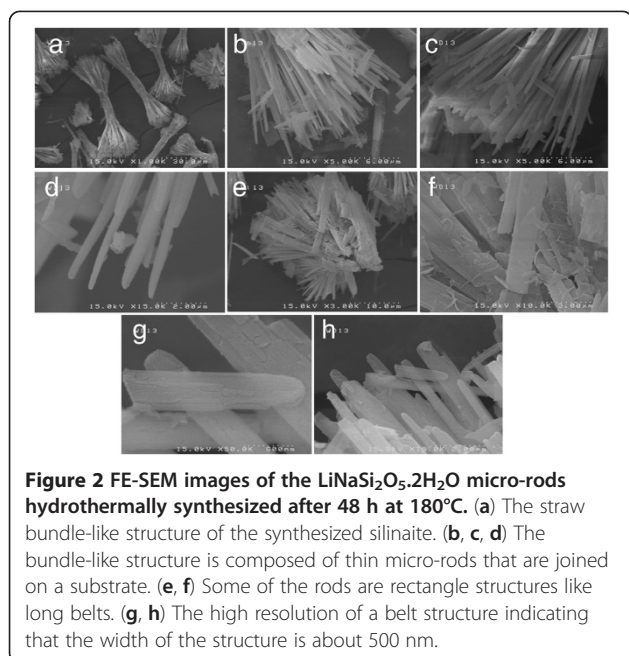


Figure 2 FE-SEM images of the LiNaSi₂O₅·2H₂O micro-rods hydrothermally synthesized after 48 h at 180°C. (a) The straw bundle-like structure of the synthesized silinaite. (b, c, d) The bundle-like structure is composed of thin micro-rods that are joined on a substrate. (e, f) Some of the rods are rectangle structures like long belts. (g, h) The high resolution of a belt structure indicating that the width of the structure is about 500 nm.

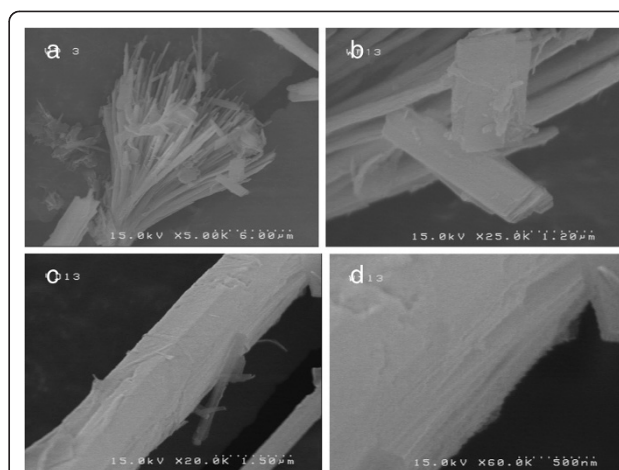


Figure 3 FE-SEM images of the LiNaSi₂O₅·2H₂O micro-rods hydrothermally synthesized after 72 h at 180°C. (a, b) The straw bundle-like structure of the synthesized silinaite composed of thin micro-rods, some of which are belt-like structures. (c, d) The trunks themselves are composed of sheets that are compacted tightly.

substitution of sodium ions by lithium ions cannot occur without structural rearrangements [4]. Also, inter-planar spacing in the crystalline material is calculated via Bragg's law ($n\lambda = 2d_{hkl} \sin \theta$) with the strongest peak in 48 h (12.77 with peak intensity of 1,300) selected as reference to calculate the parameter *d*. Thus, ($\Delta 2\theta = 12.77$ (48 h) – 12.72 (72 h) = 0.05°), ($\Delta d = 6.956$ (72 h) – 6.928 (48 h) = 0.028 Å). The data show that there is a shift to lower 2θ value, and so with increasing the reaction time, there is an expansion in unit cell. To calculate the crystal sizes, we used XRD pattern peaks at $2\theta = 12.77^\circ$ for the 48-h reaction time and $2\theta = 12.72^\circ$ for the 72-h reaction time. Crystal sizes that were measured via the Debye-Scherrer equation are as follows: 35.88 nm (48 h) and 35.91 nm (72 h). In addition, the intense and sharp diffraction peaks suggest that the as-synthesized products are well crystallized. In

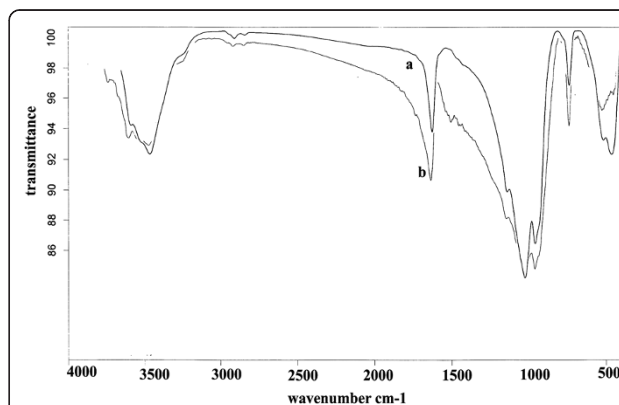
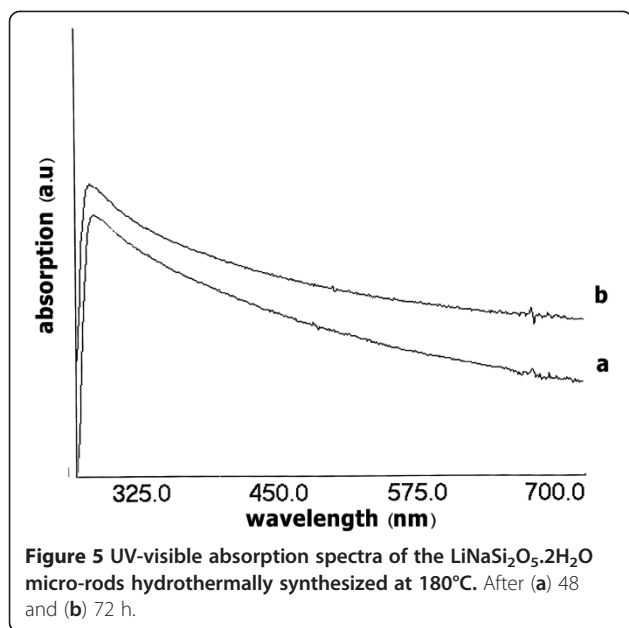


Figure 4 FT-IR spectra of the LiNaSi₂O₅·2H₂O micro-rods hydrothermally synthesized at 180°C. After (a) 48 and (b) 72 h.

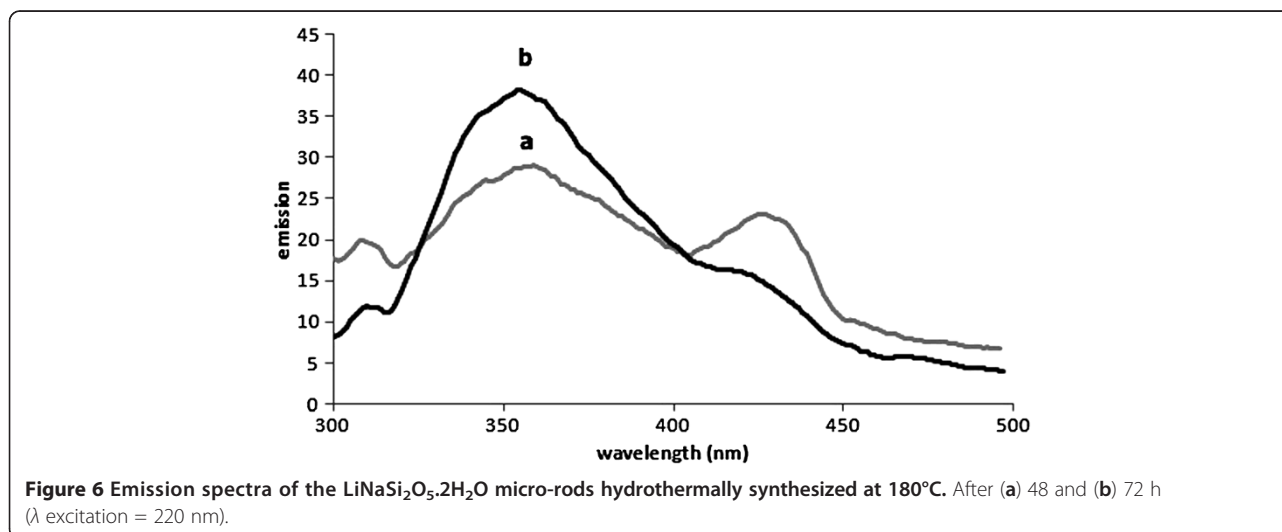


the PXRD pattern of sample b, compared to that of sample a, all peak intensities has decreased. However, in pattern b compared to that of a, the peak at $2\theta \approx 12.5^\circ$ shows a considerable decreased intensity, and the intensity of the peak at $2\theta \approx 30^\circ$ is significantly enhanced, indicating the preferred crystallographic orientation of the crystals in the products. Tables 1 and 2 show the 2θ values with their intensities and inter-planar spaces. Tables 1 and 2 show that with increasing reaction time, some peak intensities were decreased until hidden in the background. Table 3 shows the refined cell parameters calculated by Celref software version 3. As shown in Table 3, it is clear that with increasing reaction time, parameters *a* and *b* corresponding to reference [7] as the interlamellar and

intralaminar parameters in the direction of [100] and [010] increased from 5.021 to 5.077 Å and 8.289 to 8.3118 Å, respectively. Thus, we can conclude that the unit cell has expanded in the direction of [100] and [010]. However, parameter *c* that corresponds to the direction of [001] decreased from 14.542 to 14.419 Å, so in this direction, there is a contraction. Also, the volume of the unit cell was calculated by Celref software version 3 from 0.60273 nm³ that increased to 0.61014 nm³. Thus, we can conclude that there is an expansion in the unit cell.

Microstructure analysis

Figures 2 and 3 show the field emission scanning electron microscopy (FE-SEM) images of the $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ micro-rods hydrothermally synthesized, respectively, after 48 and 72 h at 180°C . The synthesized samples are composed of highly monodispersed straw bundle-like silinaite micro-rods with a length and width of about 10 to 15 μm and 400 nm, respectively (Figure 2). Figure 2a shows the straw bundle-like structure of the synthesized silinaite. Figure 2b,c,d shows that the bundle-like structure is composed of thin micro-rods that are joined on a substrate. Figure 2e,f shows that some of the rods are rectangle structures like long belts. Figure 2g,h shows the high resolution of a belt structure, indicating that the width of the structure is about 500 nm. Also, the vertical-section view of an individual straw bundle-like $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ (Figure 2d) reveals that the thickness of the rods is about 100 nm. High-magnification images of the as-synthesized material (see Figure 3) show that trunks split into long, fine nanosheets. Figure 3a,b shows that the straw bundle-like structure of the synthesized silinaite is composed of thin micro-rods some of which are belt-like structures. Figure 3c,d shows seemingly that the trunks themselves are composed of sheets that are compacted tightly.



Optical properties

Figure 4 shows the Fourier transform infrared (FT-IR) spectra of the synthesized $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ micro-rods. The broad absorption bands at $3,480$ and $3,270 \text{ cm}^{-1}$ are assigned as stretching vibrations of interlamellar water molecules [4,8]. The band at 456 cm^{-1} is characteristic of Li-O stretching vibration. The bands observed at $1,038$ and 965 cm^{-1} are assigned to the Si-O-Si asymmetric stretching vibration. Also, the symmetric stretching vibration mode of Si-O-Si is observed at 749 cm^{-1} . Moreover, the bending vibration mode of O-Si-O is observed at 534 and 467 cm^{-1} [4,8].

Figure 5 shows the UV-vis absorption spectra of the synthesized straw bundle-like $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ micro-rods, measured at room temperature. A sharp absorption band at 276 and 272 nm is observed respectively for samples a and b. According to the absorption electronic spectra, the bandgap of the synthesized $\text{LiNaSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$ micro-rods is calculated to be 4.49 eV (for sample a) and 4.56 eV (for sample b).

Figure 6 represents the room-temperature emission spectra of the as-synthesized materials under excitation at 220 nm . A strong broad emission band at 358 and 355 nm is observed for samples a and b, respectively.

Conclusions

In summary, uniform straw bundle-like silinaite micro-rods were successfully synthesized by a simple hydrothermal method. The synthesized micro-rods are composed of long, fine nanosheets. We found that the reaction time has no effect on the morphology of the products. However, the preferred crystallographic orientation of the crystals in the products changes with reaction time. UV-vis electronic absorption and emission spectra of the synthesized materials were investigated. Also, cell parameter refinement was calculated by Celref software version 3.

Competing interests

The authors declare that they have no competing interests.

Authors' contributions

AA, SK, SWJ, MD, AB, HM, and SS participated in the experiments. All authors read and approved the final manuscript.

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References

1. Murray, HH: Traditional and new applications for kaolin, smectite, and palygorskite: a general overview. *Appl. Clay. Sci.* **17**, 207 (2000)
2. Zhou, CH, Ge, ZH, Li, XN, Guo, HQ: Mesoporous silicoaluminum pillared clays with high thermal stability and Brønsted/Lewis acidity by one-pot synthesis. *J. Chin. Chem. Lett.* **14**, 1285 (2003)
3. Peter, C, Baron, L, Wang, Z, Pinnavaia, T: Polymer-layered silicate nanocomposites: an overview. *J. Appl. Clay. Sci.* **15**, 11 (1999)
4. Beneke, K, Thiesen, P, Lagaly, G: Synthesis and properties of the sodium lithium silicate silinaite. *Inorg. Chem.* **34**, 900 (1995)
5. Chao, GY, Conlon, RP: $(\text{Na,K})_2(\text{Mn,Fe,Ca,Ti,Al})_3(\text{Si,Al})_8\text{O}_{20} \cdot 8\text{H}_2\text{O}$; Mont St. Hilaire "UK38"; later equatzakharovite by original authors: Velthuisen J. V. *Mineral Record* **21**, 363 (1990)
6. Chao, GY, Grice, JD, Gault, RA: Silinaite, a new sodium lithium silicate hydrate mineral from Mont Saint-Hilaire, Quebec. *Can. Mineral.* **29**, 359 (1991)
7. Grice, JD: The crystal structure of silinaite, $\text{NaLiSi}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$: a monophyllosilicate. *Can. Mineral.* **29**, 363 (1991)
8. Zhou, CH, Guo, HQ, Cao, HJ, Li, XN, Ge, ZH: Li: Synthesis and characterization of novel layered lithium sodium silicate (Silinaite) and the mesoporous materials from silinaite. *J. Chinese Chem. Lett.* **17**, 525 (2006)
9. Stambaugh, EP, Miller, JP: Hydrothermal precipitation of high quality inorganic oxides. In: *Proceedings of the First International Symposium on Hydrothermal Reactions*, p. 859. Gakujutsu Bunken Fukyu-Kai, Tokyo (1982)
10. Barrer, RM: *Hydrothermal Chemistry of Zeolites*, p. 360. Academic, Waltham (1982)
11. Sridhar, K: Nanotechnology. *Current. Sci.* **85**, 25 (2003)
12. Chao, GY, Conlon, RP, Velthuisen, JV: Identical to UM1944-/-TeO:FeH; later described under the name poughite. *Mineral Record.* **21**, 363 (1990)
13. Kraehenbuehl, F, Stoeckli, HF, Brunner, F, Kahr, G, Mueller-Vonmoos, M: Study of the water-bentonite system by vapour adsorption, immersion calorimetry and X-ray techniques. I. Micropore volumes and internal surface areas, following Dubinin's theory. *Clay. Miner.* **22**, 1 (1987)

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