ERRATA

Erratum to: Solubility of 2-Chloro-3-(trifluoromethyl)pyridine and Correlation with the Polarity in Organic Solvents at 273–303 K¹

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Tables 1 and 2 should be as follows:

Table 1. Experimental and calculated solubility values for 2-chloro-3-(trifluoromethyl)pyridine in different pure solvents from $T = 273.15$ to 303.15 K at atmospheric pressure^a

T, K	x	Wilson model		$NRTL$ model	
		x^{cal}	$RD, \%$	x^{cal}	$RD, \%$
Ethyl acetate					
273.15	0.3951	0.3978	−0.68	0.3979	−0.70
278.15	0.4516	0.4531	−0.33	0.4531	−0.33
283.15	0.5104	0.5133	−0.57	0.5132	−0.55
288.15	0.5836	0.5789	0.80	0.5786	0.85
293.15	0.6567	0.6502	0.99	0.6498	1.05
298.15	0.7378	0.7282	1.31	0.7276	1.39
303.15	0.8227	0.8140	1.05	0.8135	1.11
n -Butanol					
273.15	0.1449	0.1530	−9.97	0.1529	−5.55
278.15	0.1875	0.1849	−1.16	0.1849	1.40
283.15	0.2308	0.2242	2.00	0.2241	2.91
288.15	0.2858	0.2733	5.41	0.2731	4.43
293.15	0.3480	0.3367	6.67	0.3363	3.37
298.15	0.4235	0.4226	6.88	0.4229	0.15
303.15	0.5002	0.5508	2.29	0.5581	−11.58
2-Propanol					
273.15	0.1678	0.1742	−3.81	0.1739	−3.63
278.15	0.2093	0.2101	−0.36	0.2101	−0.36
283.15	0.2597	0.2544	2.06	0.2545	2.02
288.15	0.3253	0.3099	4.73	0.3101	4.66
293.15	0.4005	0.3819	4.64	0.3821	4.59

¹ The article is published in the original.

Table 1. (Contd.)

T, K	x	Wilson model		$NRTL$ model	
		x^{cal}	$RD, \%$	x^{cal}	$RD, \%$
2-Propanol					
298.15	0.4783	0.4803	−0.41	0.4814	−0.64
303.15	0.5628	0.6267	−11.35	0.6344	−12.72
Methanol					
273.15	0.1964	0.2017	−2.71	0.2007	−2.20
278.15	0.2397	0.2413	−0.66	0.2415	−0.75
283.15	0.2957	0.2900	1.93	0.2912	1.52
288.15	0.3651	0.3513	3.77	0.3529	3.33
293.15	0.4471	0.4309	3.66	0.4319	3.41
298.15	0.5355	0.5388	−0.62	0.5383	−0.53
303.15	0.6302	0.6910	−9.65	0.6898	−9.46
Chloroform					
273.15	0.3588	0.3610	3.65	0.3448	3.90
278.15	0.4111	0.4091	2.04	0.4018	2.26
283.15	0.4622	0.4623	−0.94	0.4656	−0.74
288.15	0.5254	0.5216	−2.32	0.5368	−2.16
293.15	0.5861	0.5884	−5.18	0.6159	−5.08
298.15	0.6635	0.6654	−6.09	0.7035	−6.03
303.15	0.7576	0.7576	−5.65	0.8001	−5.61
Acetonitrile					
273.15	0.2935	0.2957	−0.74	0.2959	3.32
278.15	0.3410	0.3393	0.49	0.3394	2.19
283.15	0.3883	0.3891	−0.18	0.3889	−2.11
288.15	0.4486	0.4465	0.47	0.4459	−5.81
293.15	0.5176	0.5141	0.68	0.5133	−9.73
298.15	0.5963	0.5960	0.06	0.5960	−12.89
303.15	0.6933	0.7008	−1.08	0.7056	−13.40
Acetone					
273.15	0.4224	0.4366	−3.07	0.4347	−3.31
278.15	0.4808	0.4812	−1.69	0.4885	−0.13
283.15	0.5489	0.5315	0.47	0.5460	3.09
288.15	0.6184	0.5883	1.71	0.6075	4.80
293.15	0.7001	0.6527	3.72	0.6736	6.70
298.15	0.7777	0.7264	4.10	0.7452	6.57
303.15	0.8524	0.8110	3.27	0.8239	4.84
Cyclohexane					
283.15	0.2077	0.2010	2.45	0.2068	0.42
288.15	0.2454	0.2437	0.36	0.2472	−0.72
293.15	0.2969	0.2976	−0.07	0.2963	0.19
298.15	0.3579	0.3684	−2.01	0.3570	0.25
303.15	0.4338	0.4681	−5.78	0.4350	−0.29

^a The standard uncertainty for the temperatures $u(T)$ is 0.01 K. The relative standard uncertainty of solubility $u_r(x^{\text{cal}})$ is 0.06. The relative standard uncertainty of pressure is $u_r(P) = 0.03$.

Table 2. The corresponding values of $\Delta\lambda_{ij}$, Δg_{ij} in the pure solvents^a

Solvent	Wilson model			<i>NRTL</i> model		
	$\Delta\lambda_{12}$	$\Delta\lambda_{21}$	$10^2 RMSD$	Δg_{12}	Δg_{21}	$10^2 RMSD$
Ethyl acetate	−2141.67	1883.39	0.60	1293.80	−1673.89	0.63
<i>n</i> -Butanol	−454.045	6413.79	2.06	4661.73	394.668	2.32
2-Propanol	−1038.40	5556.63	2.60	3985.26	111.082	2.87
Methanol	−2706.90	6134.84	2.46	3522.94	−213.447	2.39
Trichloromethane	−3360.76	6952.88	0.22	157.620	−187.737	2.60
Acetonitrile	−3672.27	8063.21	0.34	5232.83	−1718.79	0.51
Acetone	22200.80	−2877.02	3.38	1725.83	−2298.60	2.04
Cyclohexane	−0.03530	10255.3	1.19	9234.35	955.597	0.12

^a Relative standard uncertainty $u_r(\Delta\lambda_{12}) = u_r(\Delta\lambda_{21}) = u_r(\Delta g_{12}) = u_r(\Delta g_{21}) = 0.02$.

The original article can be found online at <https://doi.org/10.1134/S0036024418110237>.