



Carbon skeleton doped with Co, N, S and P as efficient electrocatalyst for oxygen evolution reaction

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ABSTRACT A new strategy for the preparation of highly efficient catalyst used in oxygen evolution reaction (OER) in alkaline media was developed. A Co-containing carbonitride polymer network (CoCN) was selected as a structural-directing template and a hypercross-linked polymer containing S and P, which formed on CoCN skeleton *in situ*, was used as a cover. After calcination at 450°C for 2 h, an interconnected nanostructure was obtained and showed excellent activity and high stability for electrochemical water splitting. Trace amount of Co and other heteroatoms including N, S, P and the formed Co–N and Co–O species are essential for the impressive catalysis performance. The calcination temperature of 450°C is optimal to the catalysis performance. These results suggest that Co in addition to heteroatom-doped (S, P) carbonitride could be used as a supplement and/or an alternative to noble metal oxides for water splitting.

Keywords: electrocatalyst, oxygen evolution, carbonitride, calcination, alkaline

INTRODUCTION

The aggravating energy and environment issues have stimulated intense focus on searching renewable and clean energy. Among various candidates proposed to date, hydrogen appears as a promising option. The state-of-the-art method for hydrogen generation is water splitting triggered either by electricity or light, which contains two complementary processes, i.e., the hydrogen evolution process [1] and the oxygen evolution reaction (OER) [2–4]. To overcome sluggish kinetics involved in OER relies on the development of efficient anodic electrocatalysts.

Up to date, the most active OER electrocatalysts are still precious metals and their oxides, such as Ru, Ir and RuO₂.

Their high costs and low abundance hindered their large-scale application. Alternatively, Co-based catalysts have received much attention due to facile preparation and good performance, such as Co₃O₄ [5–10], Co₉S₈ [11,12], CoSe [13] or CoSe₂ [14], CoP [15,16], Co(OH)₂ [17], Co_xN (x=2, 3, 4) [18] and various hybridized derivatives [19–29]. The development of metal-free OER catalyst [2,3] is also purposed. However, up to now this method is concentrated on graphitic carbonitride (g-C₃N₄) [30–33] and modified carbon nanostructures such as carbon nanotubes (CNTs) [33–35]. The performance of other metal-free materials which are entirely composed of light-weight elements is not good in most cases. To integrate the advantages of both transition metals and light-weight elements into one catalyst, a promising strategy containing carbon materials with metallic and/or heteroatomic dopants is proposed. In addition, the stability of the catalysts could be greatly improved due to the protection from the carbon materials in harsh conditions. Moreover, the performance of the hybridized materials could be further improved by the introduction of heteroatoms to modify the electronic structure of the adjacent carbon atoms.

We present the preparation of a new carbon-based catalyst doped with Co, N, S and P (Scheme 1a). Only trace level of Co (<5% by mass) is detected in the final catalyst. The catalyst shows impressive performance in OER in alkaline media. To optimize the performance of the catalysts, we also investigate the influences of other transition metals and calcination conditions.

EXPERIMENTAL SECTION

Materials

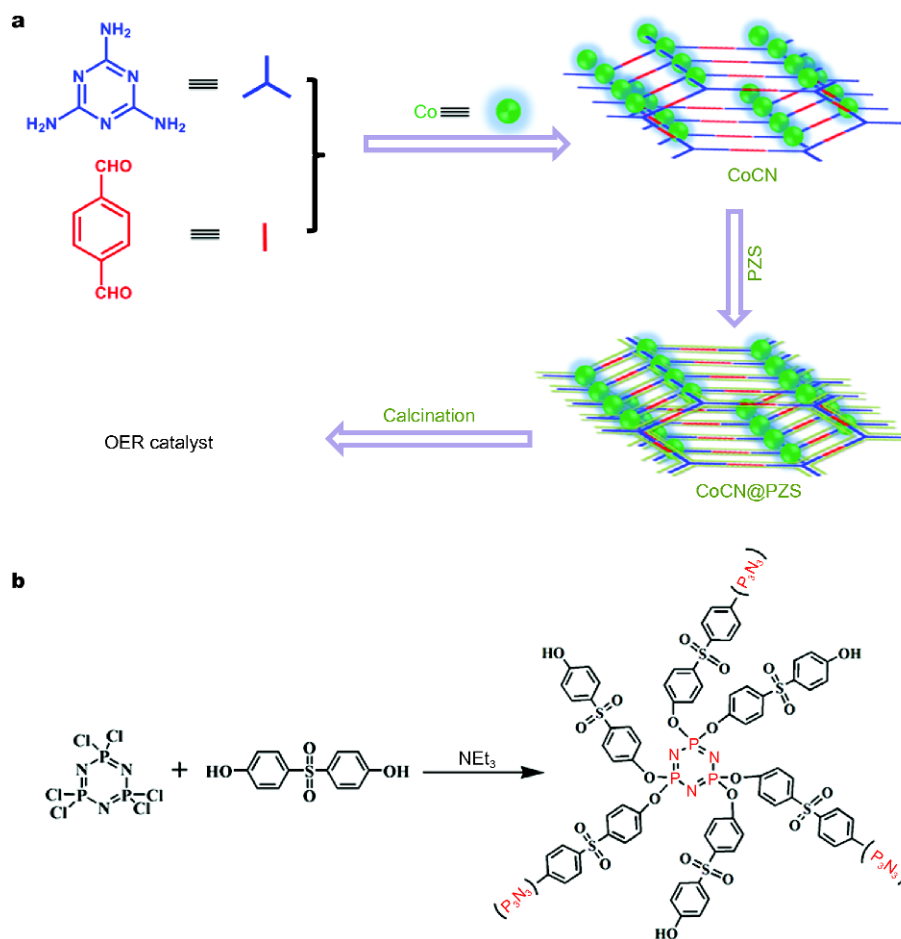
All reagents were of analytical grade and used as received.

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Scheme 1 Illustration of the synthetic route of the OER catalyst (a) and the hyper cross-linked polymer PZS (b).

Melamine (>99%), terephthalaldehyde (>99%), 4,4'-sulfonyldiphenol (>99%), phosphonitrilic chloride trimer (>98%) were purchased from J&K Chemical Reagent Co., Ltd. Cobalt nitrate hexahydrate (>99%), nickel nitrate hexahydrate (>98%) and copper (II) nitrate trihydrate (>98%) were obtained from Kermel Chemical Reagent Co., Ltd. Distilled water was utilized in all the experiments.

Synthesis of MCN (M=Co, Ni, Cu)

For the synthesis of CoCN, 4 mmol of melamine and 7 mmol of terephthalaldehyde were dissolved in 50 mL of dichloromethane (DCM). The mixture was stirred at 70°C for 30 min, and then 6 mmol of cobalt nitrate hexahydrate ($\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$) was added and stirred for 24 h. Finally, the system was cooled down and the product was isolated by filtration, washed successively with distilled water and ethanol. The as-prepared product was

further washed with distilled water in a Soxhlet apparatus for 24 h, and dried at 55°C under vacuum until constant weight with a quantitative yield (62% based on the conversion of melamine). NiCN and CuCN were prepared using similar procedure.

Synthesis of MCN@PZS (M=Co, Ni, Cu)

CoCN (400 mg) and 80 mL of methanol were mixed by sonication for 5 min. Then another portion of methanol solution (20 mL) containing 300 mg of phosphonitrilic chloride trimer and 675 mg of 4,4'-sulfonyldiphenol was added dropwise under stirring. After 5 min, 1 mL of triethylamine was added dropwise and the solution was further stirred for 18 h. After that, the celadon precipitates were collected and washed with methanol three times and dried under vacuum at room temperature for 12 h to obtain CoCN@PZS. The preparation of NiCN@PZS and CuCN@PZS were carried out in a similar pro-

cedure. For CN@PZS the process was similar except the addition of metal nitrate salts.

Synthesis of MCN@PZS-450°C (M=Co, Ni, Cu)

The dried powder of MCN@PZS (M= Co, Ni, Cu) was calcined at 450°C under N₂ in a quartz tube furnace for 2 h with a ramping rate of 5°C min⁻¹. After cooling to room temperature, the product was collected. The obtained products were denoted as MCN@PZS-450°C (M= Co, Ni, Cu). To investigate the temperature effect, the temperature during calcination was also varied.

Structural characterizations

X-ray diffraction (XRD) patterns were obtained from a Rigaku D/max-2400 diffractometer (Japan) with Cu K α radiation ($\lambda=0.15418$ nm). The microstructure and morphology of materials were characterized using a transmission electron microscope (TEM; TF20) and a field-emission scanning electron microscope (FESEM; JSM-6701F, JEOL). Energy-dispersive X-ray spectroscopy (EDS) attached to the FESEM was utilized to analyze the composition. X-ray photoelectron spectroscopy (XPS) was performed on an ESCALAB 250Xi and Raman spectra were recorded on HR-800, Jobin Yvon with a 514 nm Ar-ion laser as excitation source. Fourier transform-infrared spectroscopy (FTIR) was carried out on a VERTEX-70/70v spectrometer (Bruker Optics, Germany).

Electrochemical measurements

Electrochemical measurements were carried out using a three-electrode system at an electrochemical workstation (CHI660E, Shanghai Chenhua, China) in an aqueous solution containing 1.0 mol L⁻¹ KOH. Ag/AgCl was selected as the reference electrode and a Pt wire was used as the counter electrode. The working electrode was a glassy carbon with a diameter of 3 mm. To prepare the working electrode for OER test, 4 mg of the prepared catalyst was dispersed in 1 mL of ethanol and 10 μ L of nafion (5%) by sonication for 5 min. The suspension was then carefully dropped onto the electrode. The OER activity of the catalysts was measured by linear sweep voltammetry (LSV) from 1.20 to 1.80 V vs. reversible hydrogen electrode (RHE) at a scan rate of 5 mV s⁻¹. The essential stability tests were measured with cyclic voltammetry (CV) between 1.02 and 1.32 V at 100 mV s⁻¹ for 1,000 cycles. The surface active area of CoCN@PZS-450°C electrode was measured by CV from 1.02 to 1.32 V to further analyze their catalytic activity. All the measured potentials were calibrated with the RHE scale according

to the Nernst equation ($E_{\text{RHE}}=E_{\text{Ag/AgCl}}+0.059\text{pH}+0.197$) without any *iR*-correction.

RESULTS AND DISCUSSION

A well-defined Co-doped carbonitride framework (i.e., CoCN) was prepared first following a modified method [36]. In brief, melamine was reacted with terphthalaldehyde at the presence of trace cobalt nitrate *via* a one-step polycondensation reaction. SEM (Fig. 1a and Fig. S1) and TEM (Fig. S2) images show that the as-obtained CoCN exhibits interconnected nanostructures. This structural feature makes CoCN an ideal catalyst, because the nanostructure can facilitate the exposure of the active sites. Without cobalt nitrate only micrometer-sized blocks are obtained (Fig. S3). Besides C, N, and O, evenly-distributed Co of about 1.15% (by atomic ratio) was detected (Fig. S4). The FTIR (Fig. 2) bands at 3,410 and 1,150 cm⁻¹ belong to the stretching vibrations of the secondary amine (-NH-), indicating the formation of amination linkages. The bands at 2,965 and 2,911 cm⁻¹ can be assigned to the methyne (-CH) groups, and the characteristic absorptions of triazinering are observed at 1,562 and 1,468 cm⁻¹ [37]. To enrich heteroatom-doping, hyper-crosslinked polymer (PZS [38]) was then coated on the surface of CoCN by *in situ* polymerization between phosphonitrilic chloridotrimer and 4, 4'-sulfonyldiphenol at the presence of triethylamine. With the PZS coating, the material surface becomes smooth (Fig. 1b and Fig. S5a). Typical SEM and TEM images for the catalyst obtained at 450°C for 2 h (CoCN@PZS-450°C) show the catalyst becomes porous after calcination in N₂ atmosphere (Fig. 1c and Fig. S5b). EDS mapping reveals that all the elements are uniformly distributed in the sample (Fig. 1d, e, Fig. S6 and Table S1).

XPS spectra (Fig. S7) and high resolution XPS spectra (Fig. 3a-d and Fig. S8) show the Co 2p peaks at about 780 and 786 eV assigned to Co-N specie, while the peaks at 796.5 and 804.6 eV originate from Co-O species. The N 1s spectra can be fit well with four kinds of N species. Peaks at about 398.8, 400.0, 401.0 and 404.9 eV correspond to pyridinic nitrogen, pyrrolic nitrogen, graphitic nitrogen, and pyridinic N⁺-O⁻, respectively [39,40]. Besides these four peaks, an additional peak locates at 398.6 eV, which is ascribed to the pyridinic nitrogen or P=N/P-N bond [41,42]. The S 2p region displays three peaks at 164.0, 165.2 and 168.3 eV, respectively, which are ascribed to 2p_{3/2} and 2p_{1/2} splitting of the S 2p spin orbital (-C-S-C-) and oxidized S, respectively. The P 2p spectrum shows two peaks at 132.7 and 133.9 eV, corresponding to P-C and P-O, respectively. The C 1s

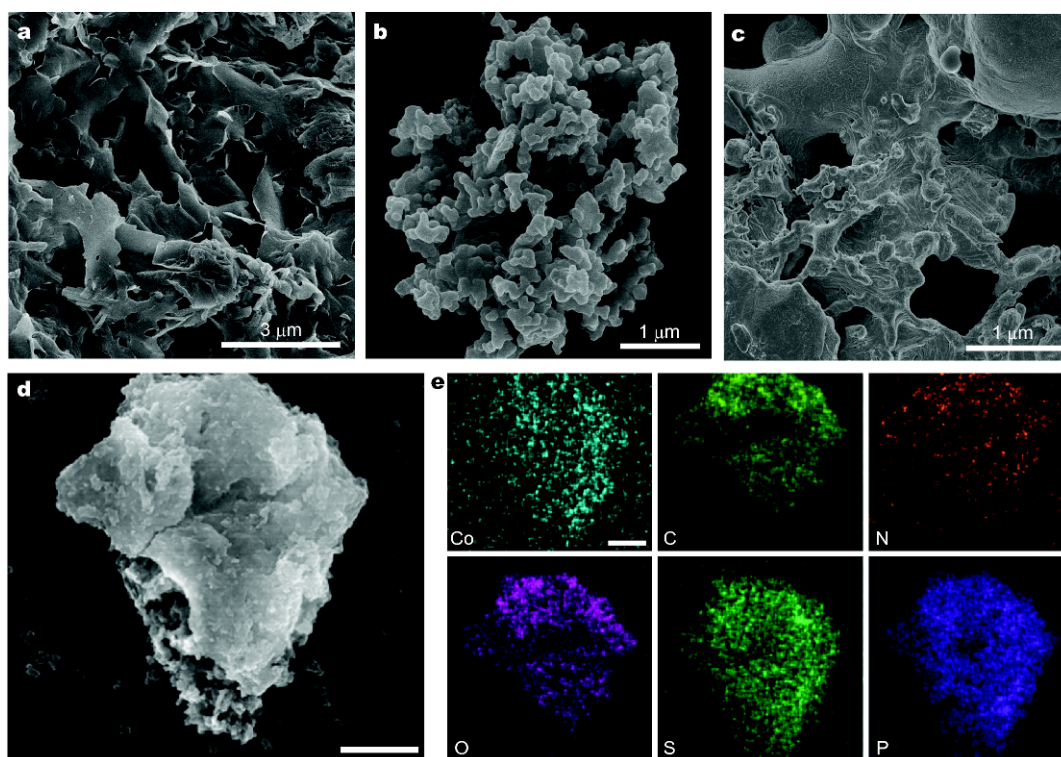


Figure 1 SEM images of (a) the as-synthesized CoCN, (b) CoCN@PZS, (c) CoCN@PZS-450°C. (d) The SEM image of CoCN@PZS-450°C investigated by EDS mapping; (e) EDS mapping of Co, C, N, O, S and P. The scale bar in (d, e) is 500 nm.

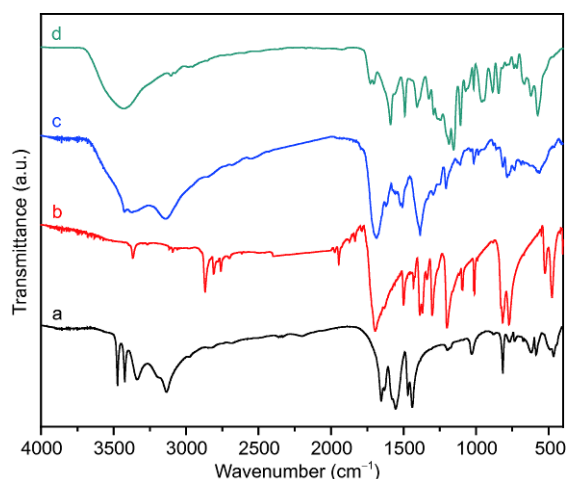


Figure 2 FTIR spectra of (a) melamine, (b) terephthalaldehyde, (c) CoCN and (d) CoCN@PZS (before calcination).

spectrum exhibits a C–C peak at 284.7 eV, a C–N peak at 285.8 eV and a C=O peak at 287.8 eV. For O 1s, the peaks centered at 531.5 and 533 eV can be assigned to the O 1s core levels in highly cross-linked PZS and adsorbed water, respectively. The atomic percentage of C, O, Co, N, S and

P is 73.99%, 14.13%, 0.80%, 4.60%, 1.18% and 5.30%, respectively (Table S1).

The wide-angle XRD pattern (Fig. 3e) of the catalyst (CoCN@PZS-450°C) is similar to the standard patterns (JCPDF No. 76-1125). In Raman spectrum (Fig. 3f), two peaks at 1,341 and 1,598 cm^{-1} are assigned to D and G bands, respectively. The broad D-band indicates the presence of substantial defects or disorder-induced mode in the catalyst and the G-band is attributed to the graphitic structure. The structural defects are mainly caused by the doping of N, S, P and concomitant presence of Co atom. The high intensity ratio of D-band and G-band (I_D/I_G) suggests a large number of defects that might act as active catalytic sites [43].

Cu and Ni doped catalysts were prepared with the same procedure to investigate the influence of the type of doped metal on the structure and properties of the catalyst. The structural characterizations including SEM observations, XRD and Raman spectra are summarized in Fig. 4, and the contents of the elements both in the bulk and on the surface are listed in Table S1. EDS mapping of CuCN@PZS-450°C shows a significant increase of the metal content, which is up to 15.25% (Table S1). The

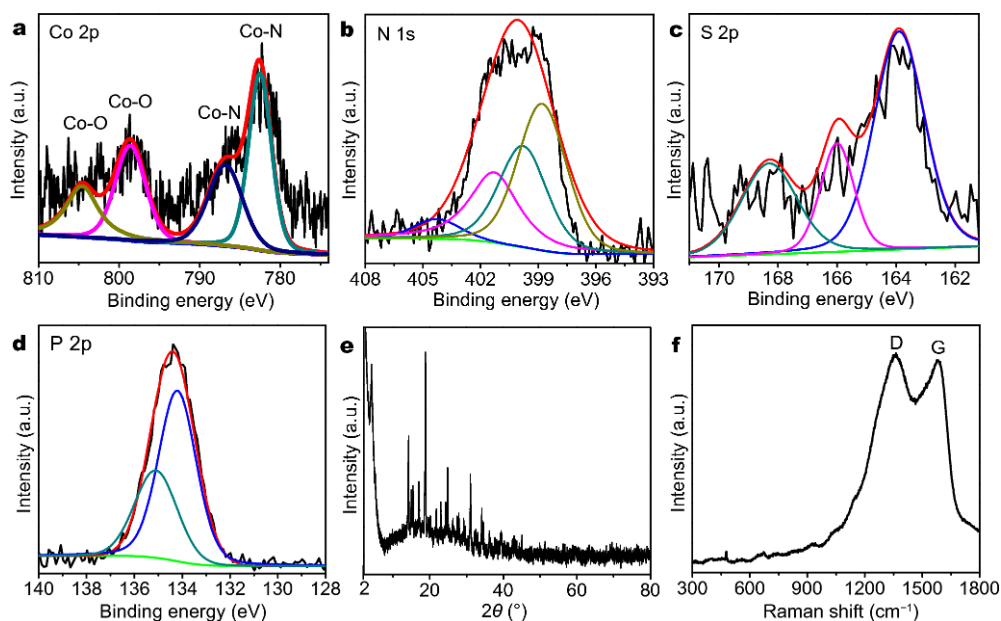


Figure 3 Structural characterizations of CoCN@PZS-450°C. High-resolution XPS spectra of (a) Co 2p, (b) N 1s, (c) S 2p and (d) P 2p. (e) XRD pattern; (f) Raman spectrum.

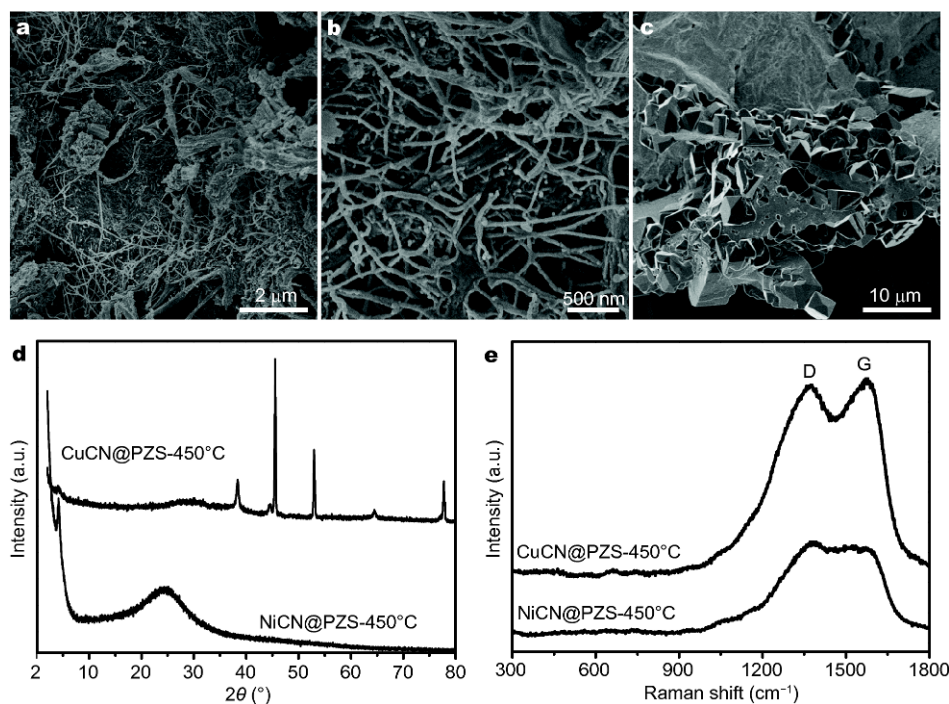


Figure 4 (a, b) SEM images at different magnifications of CuCN@PZS-450°C; (c) a typical SEM image of NiCN@PZS-450°C; (d) XRD pattern and (e) Raman spectra.

large content of Cu can be also verified by the sharp Cu 2p signal in XPS survey (Fig. S9), which gives an atomic percentage of 10.42% (Table S1). Besides C, the

amount of S on the catalyst surface is high (Table S1). SEM images reveal the threadlike aggregates together with some amorphous structures (Fig. 4a, b). XRD pat-

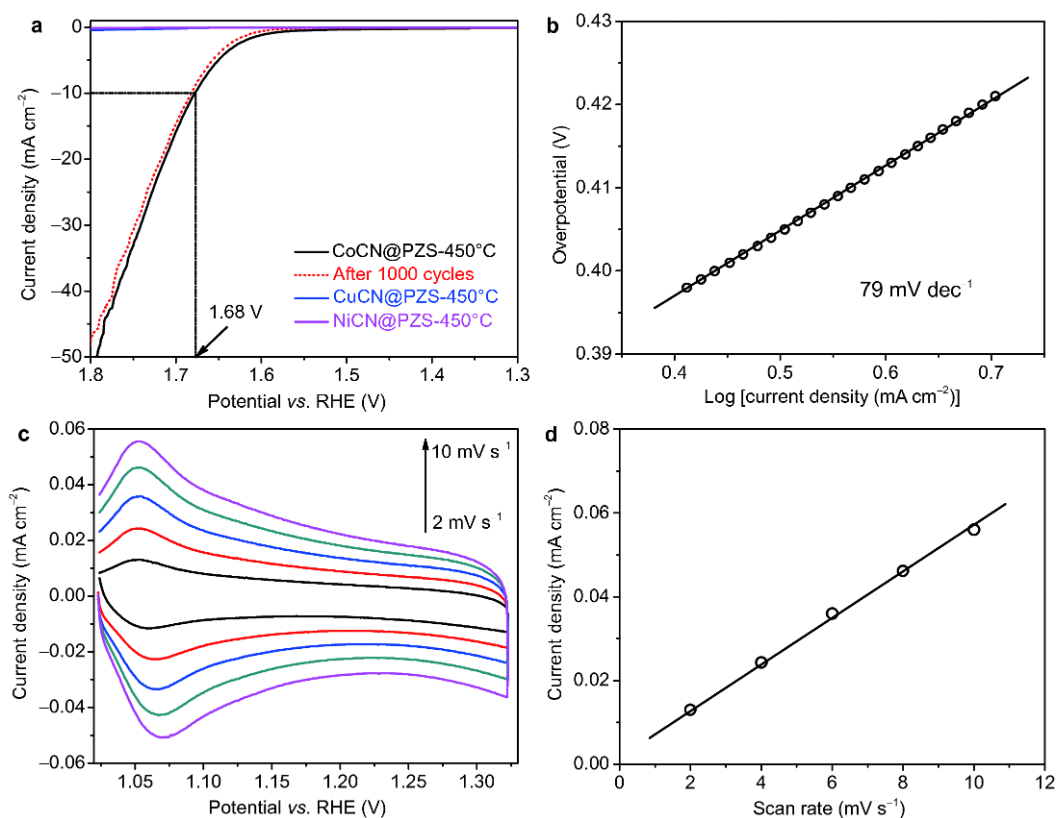


Figure 5 (a) OER polarization curves obtained in 1.0 mol L^{-1} KOH at a scan rate of 5 mV s^{-1} for CoCN@PZS-450°C, CuCN@PZS-450°C and NiCN@PZS-450°C. The curve after 1,000 cycles is also included, which is presented as the dotted line. (b) Tafel plot of CoCN@PZS-450°C. The slope obtained from the linear fit (the solid line) is 79 mV dec^{-1} ; (c) CV of CoCN@PZS-450°C measured at different scan rates in 1.0 mol L^{-1} KOH; (d) plot of current density at 1.054 V as a function of the scan rate.

tern (Fig. 4d) shows an overlying of face-centred cubic lattice (JCPDF No.85-1326) and primitive cubic lattice (JCPDF No.74-1230), indicating good polycrystallinity. The Raman spectrum (Fig. 4e) looks similar with that of CoCN@PZS-450°C with a slight decrease of I_D/I_G . In case of NiCN@PZS-450°C, the doping is unsuccessful and only trace of Ni (0.03%) is detected by EDS mapping (Table S1). XPS survey gives a slightly higher content of Ni, which is different from CoCN@PZS-450°C and CuCN@PZS-450°C where the metals tend to stay inside. However, the peak belonging to Ni is hardly seen due to the low content (0.27%) (Fig. S10).

The catalysts' activities in OER were evaluated by LSV using a three-electrode system with a Pt foil as the counter electrode, Ag/AgCl electrode as the reference electrode and a glassy carbon electrode loaded with catalyst as the working electrode in 1.0 mol L^{-1} KOH aqueous solution. As shown in Fig. 5a, CoCN@PZS-450°C was highly active for OER with a small overpotential of 450 mV to achieve the current density of 10 mA cm^{-2}

(η_{10}). The small Tafel slope of 79 mV dec^{-1} (Fig. 5b) demonstrates the superior catalytic activity and efficient electron/mass transfer in CoCN@PZS-450°C. The performance of CoCN@PZS-450°C is better than those reported recently under the same conditions [44,45] and comparable to those performed in 0.1 mol L^{-1} KOH aqueous solutions [46–49] (Table 1). Moreover, a very small decay of the current density after continuous CV scanning for 1,000 cycles demonstrates very good electrochemical stability, partially ascribed to the presence of an outer C-rich protecting layer.

In Fig. 5a, the performance of CoCN@PZS-450°C is much better than its counterparts, i.e., CuCN@PZS-450°C and NiCN@PZS-450°C. Due to the ultralow current density in the range of high potentials ($> 1.45 \text{ V}$), the curves of CuCN@PZS-450°C and NiCN@PZS-450°C cannot be fitted to Tafel equation. Besides the intrinsic differences between the catalytic activity of Co, Cu and Ni, the different OER activity among these three catalysts could also be attributed to their different morphologies

Table 1 Comparison of several recently reported highly active noble metal-free OER catalysts

Catalyst	η_{10} (mV vs. RHE)	Tafel slope (mV dec ⁻¹)	Electrolyte	Substrate	Ref.
Co ₃ O ₄ -35	525	Not available	1 mol L ⁻¹ KOH	Gold disk	[44]
Co ₃ O ₄ /SWCNT	594	104	1 mol L ⁻¹ KOH	ITO	[45]
CoCN@PZS-450°C	450	79	1 mol L ⁻¹ KOH	Glassy carbon	This work
Co ₃ O ₄ /N-PC	390	72	0.1 mol L ⁻¹ KOH	Glassy carbon	[46]
CYM-350	~470	68	0.1 mol L ⁻¹ KOH	Glassy carbon	[47]
N/Co-doped PCP-RGO	430	292	0.1 mol L ⁻¹ KOH	Glassy carbon	[48]
N, S-codoped carbon nanosheets	~420	59	0.1 mol L ⁻¹ KOH	Glassy carbon	[49]

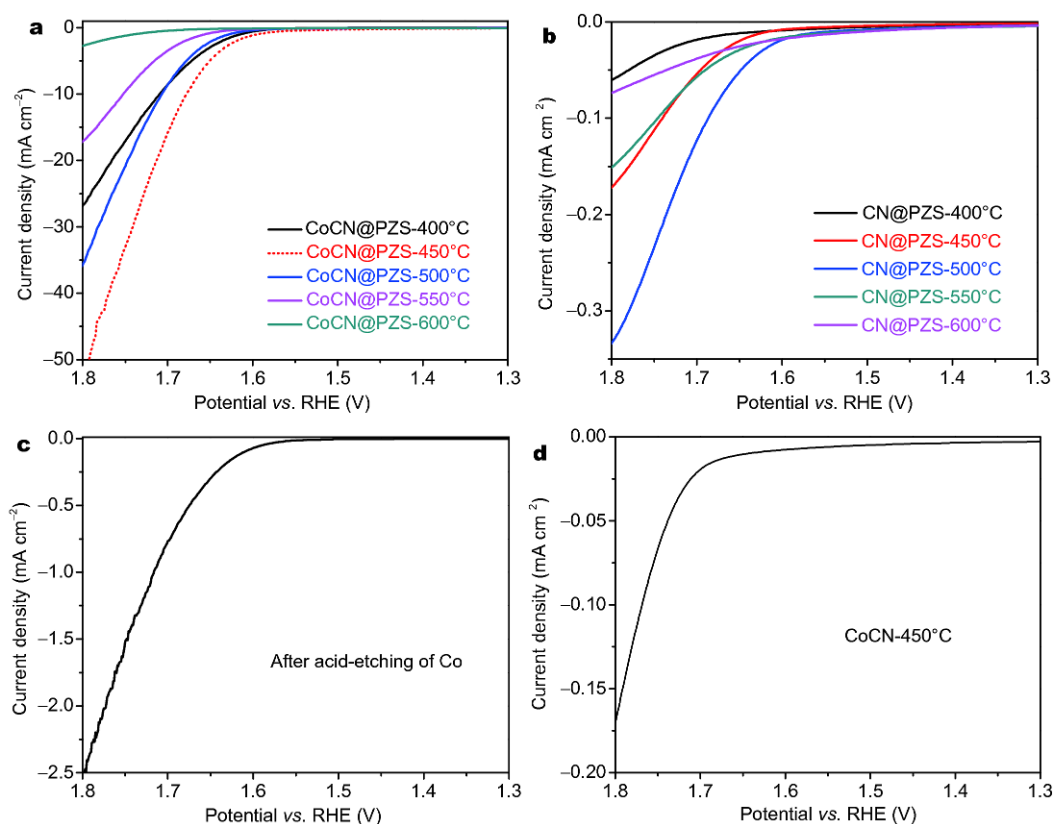


Figure 6 (a) OER polarization curves of CoCN@PZS calcinated at different temperatures. The curve of CoCN@PZS-450°C which shows the best performance is highlighted by the dotted line. (b) OER polarization curves of the catalysts without Co doping (CN@PZS) calcinated at different temperatures. (c) OER polarization curve of CoCN@PZS-450°C after removing Co by etching in 0.5 mol L⁻¹ H₂SO₄. (d) OER polarization curve of the catalyst without PZS coating calcinated at 450°C. Measurements were carried out in 1.0 mol L⁻¹ KOH at a scan rate of 5 mV s⁻¹.

and internal structures. Compared with the other two catalysts, CoCN@PZS-450°C with sheet-like nanostructure shows high crystallinity and provides more active sites. The electrochemically active surface area (ECSA) can be evaluated on the basis of the electrochemical double-layer capacitance (C_{dl}) [50–52], since C_{dl} is proportional to the number of active sites of the catalysts.

Fig. 5c displays the CV curves of CoCN@PZS-450°C in the range of 1.02–1.32 V vs. RHE at a scan rate ranging from 2 to 10 mV s⁻¹. A C_{dl} of 54 mF cm⁻² can be calculated from the linear relationship of the current density against the scan rate shown in Fig. 5d. This high value implies that CoCN@PZS-450°C has a large number of active sites, consistent with the excellent OER activity.

The calcination temperature has an obvious influence on the activity of the catalyst. The sample prepared at 450°C shows the best performance, and lowering or raising the calcination temperature leads to fast decrease of the catalytic activity (Fig. 6a), indicating sample prepared at 450°C has the optimum doping content. Another contribution could be the structural changes induced by temperature. While SEM image of the sample obtained at 550°C shows that the catalyst still contains discrete and interconnected particles (Fig. S11), and the sample obtained at 600°C shows totally loose porous structure (Fig. S12).

In order to further reveal the mechanism of good electrocatalytic activity of our catalyst, control experiments were carried out. First, samples without Co-doping (CN@PZS) calcinated at different temperatures were prepared and the OER activities exhibit poor performance (Fig. 6b). On the other hand, when the doped Co in CoCN@PZS-450°C was etched away by refluxing the catalyst in 0.5 mol L⁻¹ H₂SO₄ for 2 h, the OER activity of the Co-free catalyst became eclipsed (Fig. 6c). These observations clearly demonstrated that the trace Co in the form of Co–O and Co–N species in the catalyst is indispensable to get good catalytic performance. We also investigated the role of S, P and extra N introduced by coating PZS in the electrocatalytic process, and found that without PZS coating the catalyst exhibited unsatisfactory performance (Fig. 6d). Thus, it is clear that these heteroatoms in CoCN@PZS-450°C could take a synergistic catalytic effect on the OER activity by changing the electronic structure and spin density of the adjacent carbon atoms [53–55].

CONCLUSIONS

In summary, we demonstrated that OER catalyst using a Co-containing carbonitride polymer network as a template and a hyper cross-linked polymer containing S and P as a cover, can be facilely obtained by a simple *in situ* polymerization and subsequent calcination process. When used for electrochemical water splitting in 1.0 mol L⁻¹ KOH aqueous solution, the catalyst shows excellent activity and high stability. EDS mapping and XPS measurement show that the catalyst contains a catalytic amount of Co in Co–N and Co–O forms, which is essential for the impressive performance of the catalyst. Replacing Co by other transition metals such as Cu and Ni significantly lowers the activity of the catalyst. The performance of the catalyst is also affected by calcination procedure, with 450°C as the optimized temperature. The heteroatoms (N, S and P) could take a synergistic catalytic

effect on the OER activity. These results suggest that Co-containing, heteroatom-doped (S, P) carbonitride could be used as a supplement and/or an alternative to noble metal oxides for water splitting.

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Conflict of interest The authors declare that they have no conflict of interest.

Supplementary information Additional SEM and TEM images of CoCN and the carbonitride prepared at the absence of cobalt nitrate, EDS mapping of CoCN, TEM images of CoCN@PZS and CoCN@PZS-450°C, EDS analyses and XPS spectra for CoCN@PZS-450°C, CuCN@PZS-450°C and NiCN@PZS-450°C, SEM images of CoCN@PZS calcinated at 550°C and 600°C are available in the online version of the paper.



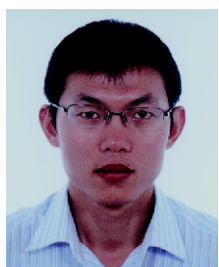
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掺杂钴、氮、硫、磷的碳骨架作为电化学析氧反应的高效催化剂

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摘要 本文提出了一种制备碱性介质中电化学析氧反应高效催化剂的新方法. 该方法选用一种含钴的碳-氮聚合物网络作为结构模板, 外面包裹一层原位制备的、含硫和磷的超支化交联聚合物. 450°C煅烧2 h后, 获得可用于电化学析氧反应的、内部交联、微观呈现层状结构的催化剂. 该催化剂在1.0 mol L⁻¹的氢氧化钾水溶液中表现出很好的电化学催化活性和高稳定性. 电子衍射图谱(EDS)和X-射线光电子能谱(XPS)研究表明该催化剂含有痕量钴及其他杂原子, 包括氮、硫、磷, 且证实了能够大幅提高催化活性的Co-N和Co-O活性物质的存在. 将钴替换为铜和镍之后, 催化剂的催化活性大大降低, 表明当前方法对钴基催化剂的制备最为有效. 煅烧过程中所选用的温度对催化剂的催化活性亦有显著影响, 450°C为最优温度. 这些结果表明, 含钴的、杂原子(硫、磷)掺杂的碳氮化物有望成为一类新的电解水催化剂, 以取代贵金属氧化物, 或作为其有益的补充.