



# Enhanced performance of dye-sensitized solar cell with thermally stable natural dye-assisted TiO<sub>2</sub>/MnO<sub>2</sub> bilayer-assembled photoanode

Shyamal Datta<sup>1</sup> · Argha Dey<sup>1</sup> · Nayan Ranjan Singha<sup>2</sup> · Subhasis Roy<sup>1</sup>

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## Abstract

This study reports the performance analysis of an organic dye-sensitized solar cell (DSSC), introducing MnO<sub>2</sub> as an electron transport layer in TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly. The DSSCs have been fabricated using TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> layer-by-layer architecture films onto fluorine-doped tin oxide (FTO) glass and sensitized with natural dye extracted from *Malvaviscus penduliflorus* flower in ethanol medium. The counter electrode was prepared to layer copper powder containing paste onto FTO's conductive side by the doctor's blade method. The optical, morphological, and structural properties of photoanodes were explored via ultraviolet–visible, field emission scanning electron microscopy, and X-ray diffraction analyses. Moreover, dye complexity and thermostability of dyes were characterized via Fourier-transform infrared spectroscopy and thermogravimetric analyses. The iodide/triiodide (i.e., I<sup>-</sup>/I<sub>3</sub><sup>-</sup>) redox couple of electrolyte solution was employed as a charge transport medium between the electrodes. Finally, photoanode and counter electrode sandwiches were assembled to envisage the photovoltaic performance potential under simulated AM 1.5G solar illumination using 100 mW cm<sup>-2</sup> light intensity. The as-fabricated DSSC comprising TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly exhibited 6.02 mA cm<sup>-2</sup> short circuit current density ( $J_{sc}$ ), 0.38 V open-circuit voltage ( $V_{oc}$ ), 40.38% fill factor, and 0.92% conversion efficiency, which is about 200% higher compared to the assembly devoid of MnO<sub>2</sub> layer.

**Keywords** Low cost dye-sensitized solar cells · Organic dye · Bilayer structure · Dye stability analysis

## Introduction

Nowadays, clean, near-zero-emission and sustainable renewable energy production are vital concerns for policymakers worldwide. The energy consumption demand is expected to be doubled in 2050 [1, 2]. Notably, a significant percentage of energy demand is backed by fossil fuel-based resources, implying an adverse impact on global warming, environment, and ecosystem. For circumventing the problems related to the limited amount of fossil fuel and adverse

environmental concerns, present researchers are focused on renewable energy sources, i.e., solar energy, which is abundant and cost-effective.

Of variable third-generation solar cells, DSSC is low cost and easy to fabricate. The DSSC module works as a photoelectrochemical device in which electron–hole pair is generated via light-induced exciton [3]. The dye-anchored metal oxide semiconductor, i.e., TiO<sub>2</sub>, works as an electron transport layer, and electrolyte functions as a charge transport media and electron recombination suppression media between photoanode and cathode [4, 5].

The sensitizer mostly influences the performance potential of DSSC. The absorption of a wider range of light spectrum and a high molar absorption coefficient elevate the light-harvesting ability of TiO<sub>2</sub> film. The natural color pigments of plant leaves, fruits, and flowers can be used as a sensitizer for DSSC application(s). Though DSSC has impressive performance under indoor lighting, the long-term stability and outdoor performance under high-temperature ambient conditions are yet to be achieved to date [6]. Moreover, the DSSC conversion efficiency is much lower than that

✉ Subhasis Roy  
srchemengg@caluniv.ac.in; subhasis1093@gmail.com

<sup>1</sup> Department of Chemical Engineering, University of Calcutta, 92 APC Road, Kolkata, West Bengal 700009, India

<sup>2</sup> Advanced Polymer Laboratory, Department of Polymer Science and Technology, Government College of Engineering and Leather Technology (Post Graduate), Maulana Abul Kalam Azad University of Technology, Salt Lake City, Kolkata, West Bengal 700106, India

of the commercially available silicon-based solar cells. The chlorophyll extract from various leaves has been used as a natural sensitizer in DSSC.

Moreover, in DSSC, red/purple pigment in various leaves/flowers has been used as a sensitizer [7]. Notably, an abundantly available organic dye is easily extracted from flowers and leaves, mainly responsible for light absorption in DSSC. It is expected that the dye should absorb ultraviolet–visible (UV–Vis) and near-infrared (NIR) regions of the light spectrum. The hydrophobic organic dye is selected for long-term stability. Hydrophobic dye reduces the possibility of direct contact among photoanode and liquid electrolyte, retarding water-induced dye degradation for the cell's long-term stability [8, 9]. The natural color pigments originated from organic dyes impart an anthocyanins group present in the different parts (e.g., flower, leaves) of the plant [10–12]. Red pigment-containing flower *Hibiscus rosa-sinensis* possessing a higher concentration of anthocyanins group is used for a natural dye-sensitized solar cell [13, 14]. In fact, *Malvaviscus penduliflorus* flower is closely related to the Hibiscus family. However, potential research based on *M. penduliflorus* flower -extracted dye in DSSC is still lacking. The broader absorption of *Hibiscus*-extracted dye within 400–500 nm can further be enhanced using either concentrated dye solution or operating the sensitization process at elevated temperature [15–17]. The absorption defect of the *M. penduliflorus* flower-extracted dye starts from 550 nm onward [18]. The performance of the DSSC can be compensated introducing a scattering layer or interfacial modification in the photoanode and concomitantly improves the broader spectrum wavelength range of light absorption to make it suitable for outdoor application(s) [19, 20]. The front side and back side illumination for photovoltaic devices originate from the cost-effective perspective, as the higher percentage of the incident solar spectrum can be harvested without altering the cell area. The light-harvesting property, along with the back side illumination, of the device can be enhanced by integrating porous scattering sensitized film, i.e., mesoporous TiO<sub>2</sub>, with a semi-transparent counter electrode. TiO<sub>2</sub> has a wide bandgap of around 3.2 eV and is used widely in DSSC photoanode fabrication [21]. The tuning of optical and electrical properties and device performance enhancement is possible via suitable semiconductor material doping in the layered architecture [22, 23]. In this regard, semiconductor and bilayer assembly with MnO<sub>2</sub>/TiO<sub>2</sub> can be employed for DSSC application(s).

Moreover, the bandgap of MnO<sub>2</sub> around 2.57 eV makes it an ideal material for MnO<sub>2</sub>/TiO<sub>2</sub> bilayer photoanode fabrication [24]. In DSSC fabrication, most of the research works use expensive platinum as a counter electrode, limiting the employability toward large-scale production. Thus, less expensive copper as a counter electrode is expected to facilitate the large-scale industrial application(s). This research's

main objective is to elevate the output performance of cost-friendly DSSC by incorporating bilayer TiO<sub>2</sub>/MnO<sub>2</sub> in the working electrode. Moreover, the present work aims to draw attractive attention toward cost effective DSSC fabrication. Accordingly, an attempt was made to extract the dye, i.e., sensitizer, from natural origin, i.e., *Malvaviscus penduliflorus* flower layer-by-layer architecture of TiO<sub>2</sub>/MnO<sub>2</sub> toward the fabrication of photoelectrode for cost-friendly DSSC fabrication/ application.

## Experimental details

### Materials and chemicals

All the chemicals were purchased from commercial sources and used without purification. Natural dye was extracted from *Malvaviscus penduliflorus* flower using ethanol. Fluorine-doped tin oxide (FTO) glass substrate of 12–15 Ω sq<sup>-1</sup> surface resistivity and 2.3 mm thickness purchased from Pilkington (India) was used as a conductive surface in the working and counter electrodes. Scotch tape (25 μm mm thick, Solaronix) was used in the doctor's blade method to control the film's thickness. TiO<sub>2</sub> nanoparticle paste (Solaronix, 18 wt%) was used for thin film deposition on the working electrode. The polyvinyl alcohol adhesive was used as a binder for preparing MnO<sub>2</sub> powder paste and copper powder paste. MnO<sub>2</sub> powder (Fisher Scientific, India) was used for bilayer assembly preparation in photoanode. Copper powder (Loba Chemie, India) was used to prepare a counter electrode with copper paste. Potassium iodide and iodine were used to prepare the iodide/ triiodide electrolyte solution and acted as a charge transport media between the two electrodes. Paraffin spacer (50 μm thick, Merck, Germany) was used as a sealant and separation layer between the two electrodes during device assembling.

### Cleaning of glass substrates

FTO glass substrates were appropriately cleaned with distilled water, followed by ethanol in an ultrasonic bath for 10 min before the material deposition.

### Fabrication of working electrode

A scotch tape of 25 μm thickness was used to cover the cleaned FTO glass substrate's edges for making bottom contact. The thickness of TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer photoanodes was controlled to approximately 25 μm and 50 μm, respectively. A transparent TiO<sub>2</sub> nanoparticle paste was layered on the FTO glass substrate's conductive side by the doctor blade method. The as-prepared TiO<sub>2</sub> film was dried in a hot air oven dryer at 70 °C for

approximately 15 min before  $\text{MnO}_2$  layer deposition. After that, 0.1 g  $\text{MnO}_2$  powder was mixed with 1 g adhesive to prepare a uniformly dispersed  $\text{MnO}_2$  powder paste.  $\text{MnO}_2$  paste was similarly deposited onto the  $\text{TiO}_2$  layer, followed by drying at around  $70^\circ\text{C}$  for 15 min. The layer-by-layer architecture deposition method was employed for  $\text{TiO}_2/\text{MnO}_2$  bilayer assembly preparation. The thickness of the  $\text{TiO}_2$  and  $\text{TiO}_2/\text{MnO}_2$  bilayer photoanodes was maintained at  $25\ \mu\text{m}$  and  $50\ \mu\text{m}$ , respectively. Finally,  $\text{TiO}_2$  and  $\text{TiO}_2/\text{MnO}_2$  bilayer structure samples were annealed in a furnace at  $450^\circ\text{C}$  for 45 min to improve the deposited surface morphology and crystallinity material. The samples were cooled down slowly to avoid any surface defects or cracks before dipping the samples in a natural dye solution. These samples were immersed in a dye solution for 36 h and kept in a dark place at room temperature. The electrode samples were gently washed with distilled water to remove any unadsorbed excess dye and dried in a hot oven air dryer at  $40^\circ\text{C}$  for 5 min.

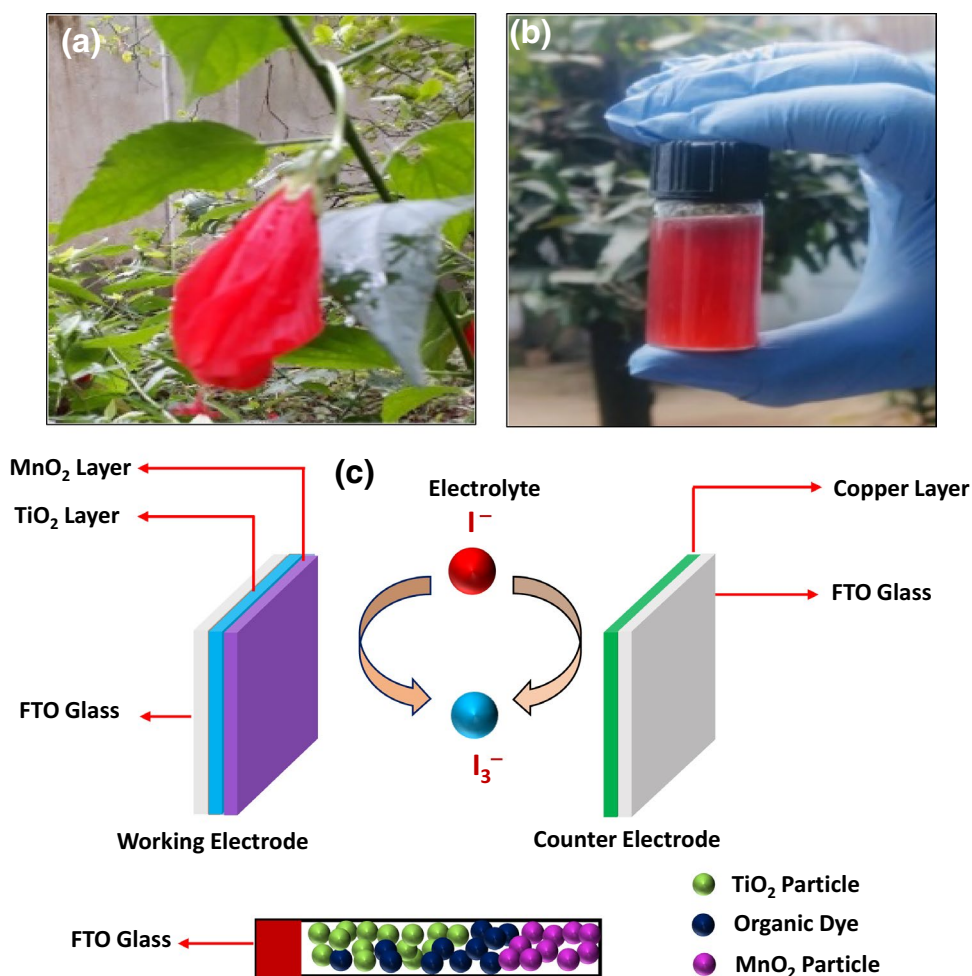
### Fabrication of the counter electrode

Here, a cleaned FTO glass substrate was used for counter electrode fabrication. Scotch tape was used to cover the edges of the FTO glass, similar to working electrode preparation. 0.12 g copper powder was mixed with 1 g adhesive to prepare a uniformly dispersed copper powder paste. The as-prepared copper powder paste was deposited on the pre-cleaned FTO glass's conductive side by the doctor blade method. Finally, the sample was placed in a furnace at around  $200^\circ\text{C}$  for 1 h. Before removing the furnace sample, the copper-coated FTO counter electrode was cooled slowly to avoid any surface cracks on the thin film.

### Preparation of natural dye solution

Herein, 5 g of fresh *Malvaviscus penduliflorus* flower (Fig. 1a) was washed with distilled water and cleaned flower petals were cut into small pieces. After that, the flower petals in 100 ml ethanol were poured in a mixer and ground for 3 min. The process was repeated thrice for preparing a thick solution. Furthermore, the solution was

**Fig. 1** a Picture of *Malvaviscus penduliflorus* flower, b dye solution extracted from *Malvaviscus penduliflorus* flower, and c schematic of  $\text{TiO}_2/\text{MnO}_2$  bilayer-assembled solar cell



stirred continuously in a magnetic stirrer for 2 h to get a uniformly dispersed solution (Fig. 1b). The solution was kept in a cold and dark place for 36 h before filtration. Finally, the filtrate dye solution was collected and the residue was filtered out. The as-prepared dye solution, i.e., sensitizer, was safely stored in a cold and dark place.

### Preparation of electrolyte

Herein, 80 mg potassium iodide and 13 mg iodine were mixed with ethanol in a culture tube and placed under continuous stirring for 5 min. The electrolyte solution containing a culture tube was stored in a dark place.

### DSSC device assembling

The photoanodes and the copper-coated counter electrode were separated by a paraffin spacer and clamped together, keeping both the conducting surface facing each other (Fig. 1c). An electrolyte solution containing  $I^-/I_3^-$  was injected in the hollow spacer area between the working electrode and counter electrode before assembling. As the electrolyte solution spread over and covered the whole active area via capillary action, two binder clips were used to hold the sandwich-like assembly for further electrical characterization.

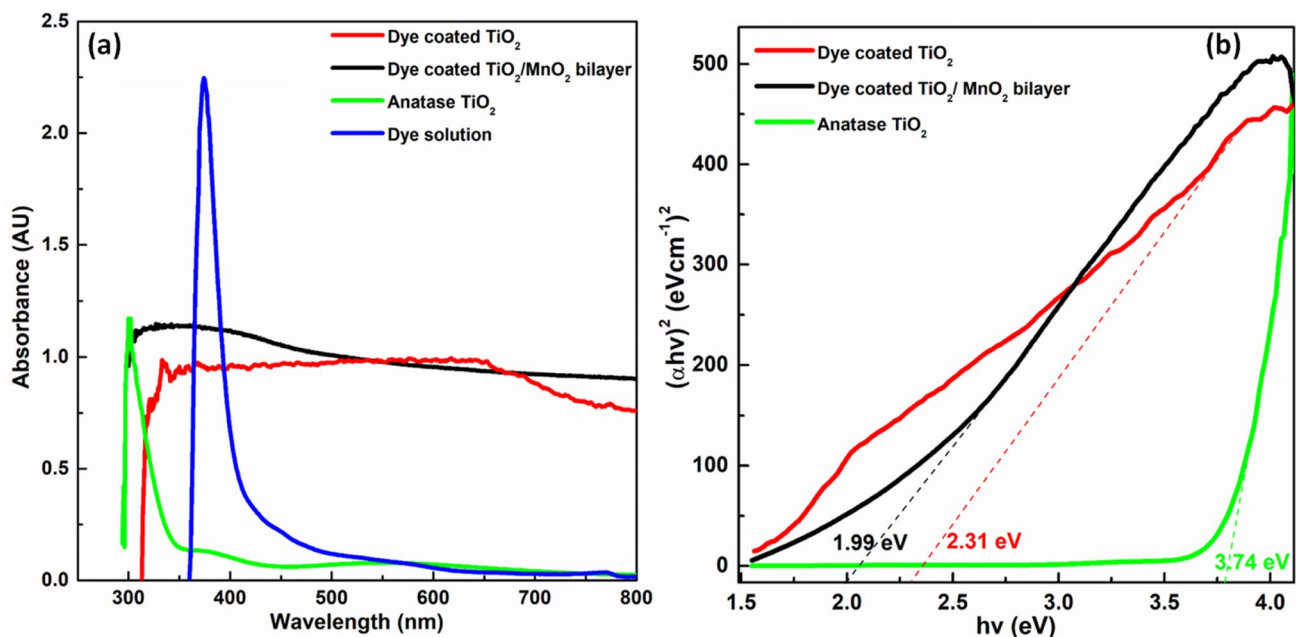
### Measurement and characterization

The light absorption property of the dye solution, dye-adsorbed  $TiO_2$ , and  $TiO_2/MnO_2$  bilayer assembly of photoanode was analyzed and recorded using a UV–Vis spectrophotometer (Hitachi Spectrophotometer, Japan). In dye solutions, functionalities and complex molecular bindings were studied within  $500\text{--}4000\text{ cm}^{-1}$  using Fourier transform infrared spectroscopy (PerkinElmer, USA). The nanostructure's surface topography on photoanode was investigated by field emission scanning electron microscopy (JEOL, Japan). The diffractogram was studied via XRD diffractometer (XPERT PRO, USA) using a  $CuK_{\alpha}$  source ( $1.54\text{ \AA}$ ). 40 kV and 30 mA sources were used to measure  $2\theta$  within  $10^{\circ}\text{--}80^{\circ}$  using nickel-filtered  $CuK_{\alpha}$  radiation. The temperature-induced degradation of the natural dye material was characterized by TGA (PerkinElmer, USA). The fabricated DSSC performance was analyzed using an AM 1.5 G sun simulator exhibiting irradiation of  $100\text{ mW cm}^{-2}$  intensity of light (Compact 150WSolar Simulator, Zolix Instruments, China) using source meter (model: 2450, Keithley, USA).

### Results and discussion

#### UV–Vis analysis

Figure 2 envisages the absorption spectra of thin both films of natural dy-coated  $TiO_2/MnO_2$  bilayer assembly,



**Fig. 2** **a** Optical absorption spectra of dye-coated  $TiO_2/MnO_2$  bilayer, dye-coated  $TiO_2$ , anatase  $TiO_2$ , and dye solution and **b** extrapolation of the Tauc plot for bandgap calculation

dye-coated TiO<sub>2</sub>, anatase TiO<sub>2</sub>, and dye solution. Though absorption band edges of TiO<sub>2</sub> thin film and dye solution were observed at 300 nm and 375 nm, respectively, no significant peak was found for TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly and dye-coated TiO<sub>2</sub>. Interestingly, TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly and dye-coated TiO<sub>2</sub> envisaged higher absorbance and long UV–Vis spectrum than the rest of the samples. From the Tauc plot, anatase TiO<sub>2</sub> envisaged a high bandgap of 3.74 eV, whereas dye-coated TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly presented 2.31 and 1.99 eV bandgaps, respectively. Thus, the photoanode fabricated with dye-coated TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly elevated the absorption property throughout the longer UV–Vis spectrum via layer-by-layer photoanode architecture fabrication, suitable for solar cell application [25].

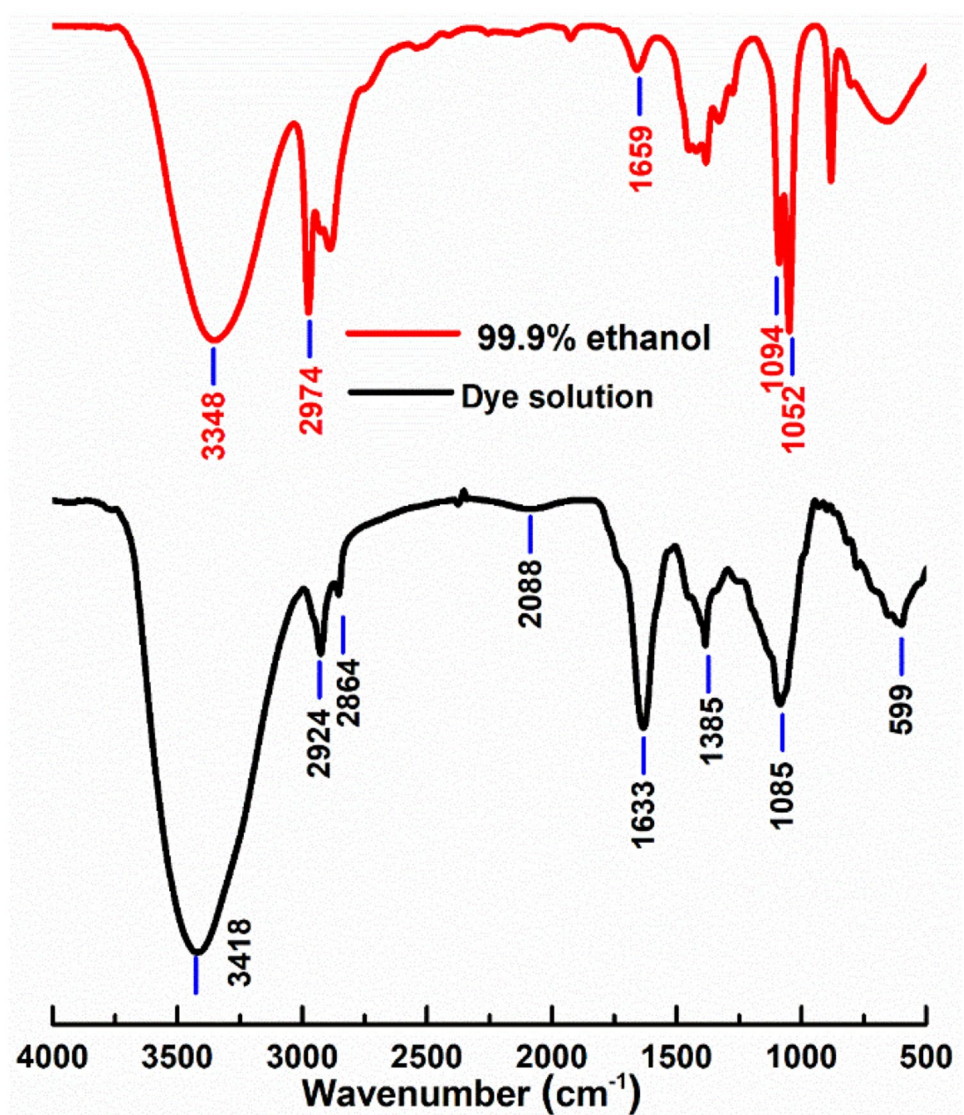
### FTIR analysis

The Fourier-transform infrared spectroscopy (FTIR) peak at 3418 cm<sup>-1</sup> was ascribed to the O–H *str.* of dye [26]. Moreover, peaks at 2925 and 2864 cm<sup>-1</sup> were related to the C–H *str.* of –CH<sub>2</sub>– and –CH<sub>3</sub> functionalities, respectively [27]. The –C≡N, C–N and N–H *str.* of amide, –CH<sub>3</sub>, and –O–C≡N functionalities in dye were confirmed from peaks at 2088, 1633, 1385, and 1085 cm<sup>-1</sup>, respectively [28, 29]. In ETOH, the change in C–O *str.* intensity was substantiated via slight displacement of the peak from 1094 to 1085 cm<sup>-1</sup> in ethanol–dye solution (see Fig. 3).

### FESEM analysis

Field emission scanning electron microscopy (FESEM) images envisaged uniform particle size distribution of

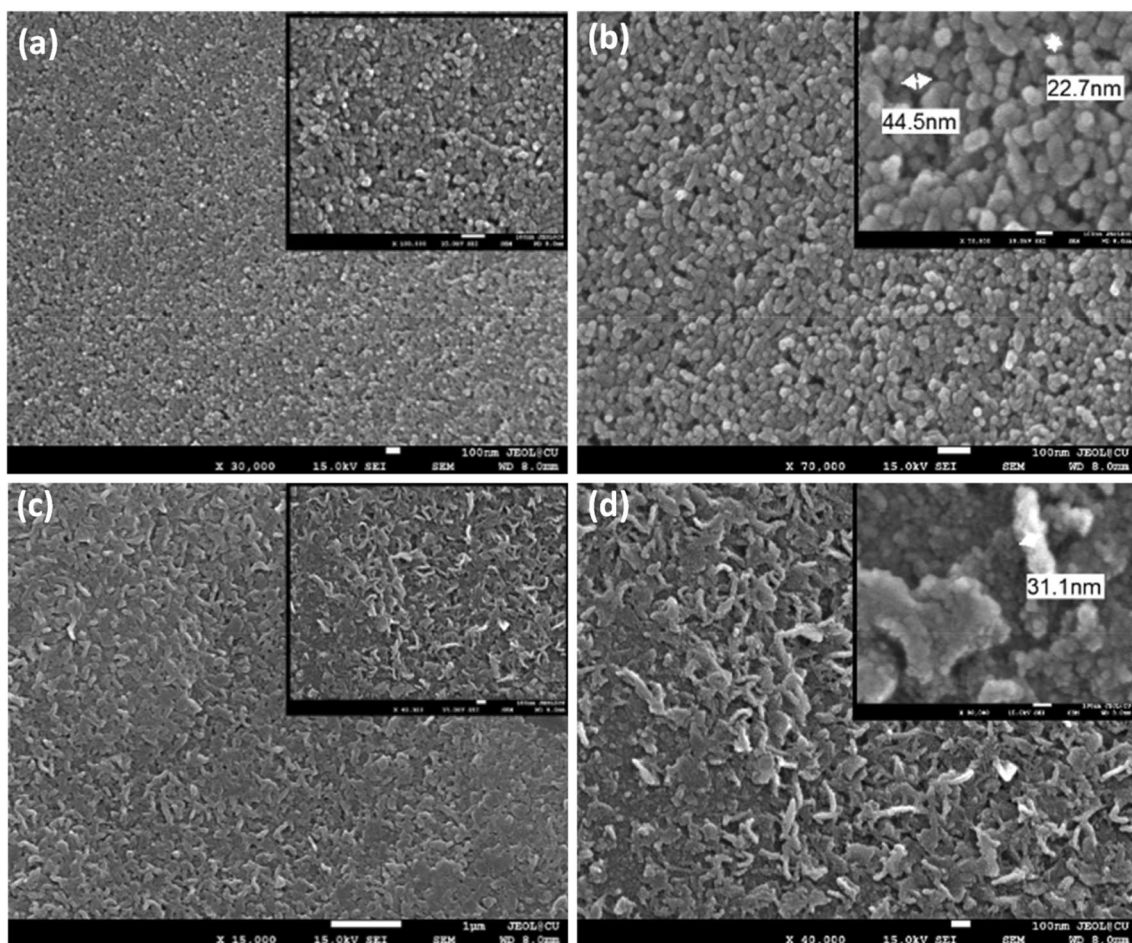
**Fig. 3** FTIR spectra of 99.9% ethanol and *Malvaviscus penduliflorus* flower-extracted dye in ethanol solution



TiO<sub>2</sub> within 20–40 nm and resembled a distributed porous nanostructure network. TiO<sub>2</sub>/MnO<sub>2</sub> bilayer nanostructure was more compact than TiO<sub>2</sub>. The surface morphology altered from the granular porous network of TiO<sub>2</sub> to a more compact and relatively larger grain size of MnO<sub>2</sub> in the TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly. The incorporation of MnO<sub>2</sub> in a bilayer assembly significantly reduced the void fractions in the photoanode sample. The larger surface area of layer-by-layer architecture of TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly compared to other photoanode counterparts resulted in higher dye loading capability in the nanostructure [30]. The surface roughness and void fractions on the photoanode samples should substantially impact the electrical performances of the photoanode [31]. In this regard, the nanomaterials' electrical performance is significantly influenced via surface morphology linked directly with the specific surface area interaction with electrolyte [32] (see Fig. 4).

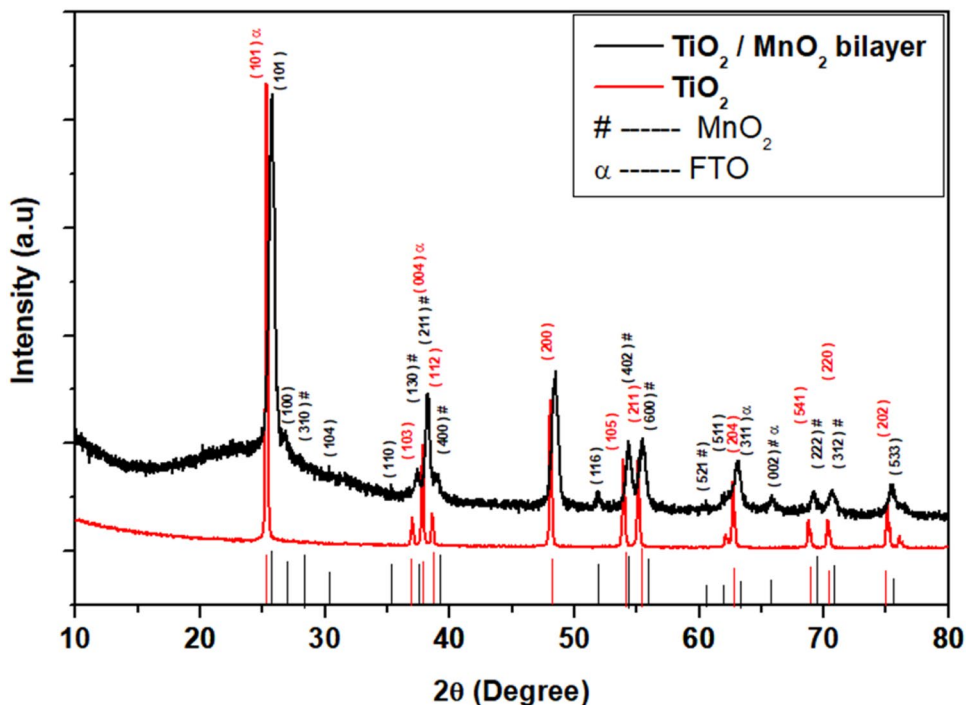
## Structural analysis

The extent of orderliness, i.e., crystalline nature, and phase formation of the dye-adsorbed TiO<sub>2</sub>/MnO<sub>2</sub> bilayer photoanode and TiO<sub>2</sub> film were characterized by thin film XRD analysis (Fig. 5) in which the characteristic diffraction peaks recorded at  $2\theta = 25.78^\circ, 37.6^\circ, 38.18^\circ, 48.38^\circ, 54.37^\circ, 63.22^\circ, \text{ and } 69.21^\circ$  corresponding to crystallographic planes (101), (103), (004), (200), (105), (204), and (541) of the anatase phase of TiO<sub>2</sub>, respectively [33]. The complexation among TiO<sub>2</sub> and MnO<sub>2</sub> was substantiated by the presence of MnTiO<sub>3</sub> in bilayer structure photoanode [34], confirmed by the diffraction peaks of MnTiO<sub>3</sub> at  $2\theta = 31.64^\circ, 35.34^\circ, 51.8^\circ, \text{ and } 60.47^\circ$  for (104), (110), (116), and (124) crystallographic planes, respectively. Moreover, the peaks at  $2\theta = 28.35^\circ, 37.55^\circ, 60.59^\circ, \text{ and } 65.86^\circ$  for (310), (211), (521), and (002) planes confirmed the presence of MnO<sub>2</sub> [35]. The additional characteristic peaks were correlated to the prevalent FTO and Ti [36]. In the TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly, a slight change in the



**Fig. 4** FESEM images of **a** TiO<sub>2</sub> photoanode, **b** particle diameter and pore size in TiO<sub>2</sub> sample, **c** TiO<sub>2</sub>/MnO<sub>2</sub> bilayer photoanode, and **d** MnO<sub>2</sub> particle size

**Fig. 5** XRD pattern of TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer films

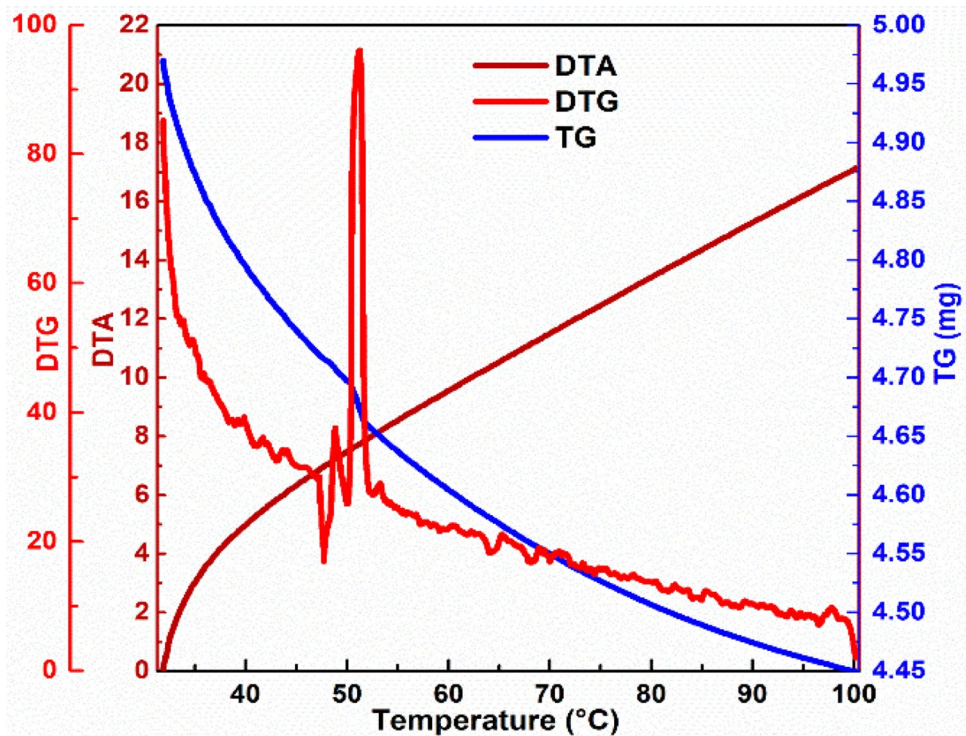


main peak position of TiO<sub>2</sub> from 25.32° to 25.78° was ascribed to the change in TiO<sub>2</sub> crystal lattice due to the diffusion of Mn<sup>2+</sup> ion in the TiO<sub>2</sub> crystal.

**TG analysis**

Figure 6 shows the differential thermal analysis (DTA), derivative thermogravimetry (DTG) and TG curves of the natural dye extracts of *Malvaviscus penduliflorus* flower. From the dye’s TG curve (Fig. 6), the weight losses within

**Fig. 6** TG analysis of *Malvaviscus penduliflorus* flower extract



50–100 °C corresponded to the loss of unbound water in the dye. Notably, the DTA curve was devoid of exothermic decomposition peak(s) or the endothermic dehydration peak(s) within 50–100 °C. Thus, this dye was suitable for DSSC solar cell application(s), confirming the dye’s suitability in extreme climate conditions [6].

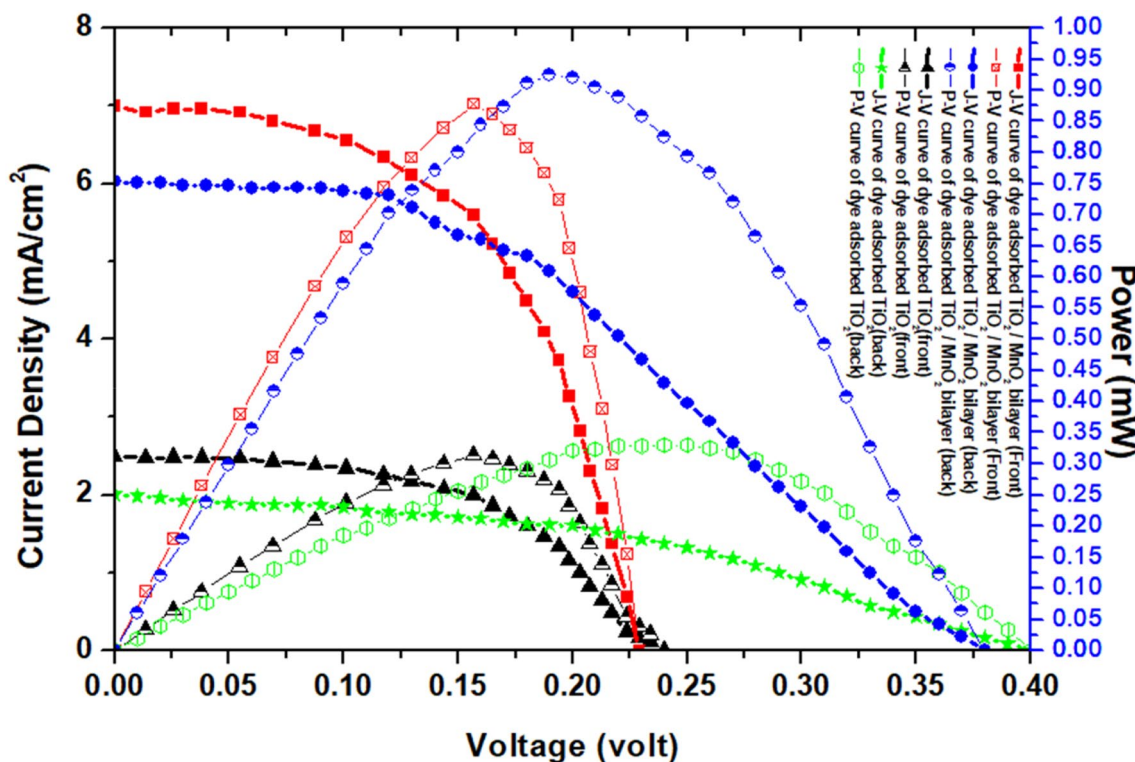
**Photovoltaic study**

The photovoltaic characteristics of the natural dye-sensitized DSSCs were tested via current density–voltage (*J–V*) and power–voltage (*P–V*) measurements. TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer-assembled DSSCs were tested under simulated AM 1.5G solar illumination having 100 mW cm<sup>-2</sup> light intensity. The performance was measured using a Keithley source meter at 27 °C. The output result of photovoltaic performance and different parameters of DSSCs, i.e., open-circuit

voltage (*V<sub>oc</sub>*), short circuit current density (*J<sub>sc</sub>*), fill factor (FF), efficiency (*η*), and maximum power output (*P<sub>max</sub>*), are given in Table 1. Because of the higher dye loading in TiO<sub>2</sub>/MnO<sub>2</sub> bilayer nanostructure, the dye-adsorbed TiO<sub>2</sub>/MnO<sub>2</sub> bilayer assembly showed more conversion efficiency compared to that of TiO<sub>2</sub> in the photoanode (Fig. 7) [37]. Indeed, the bilayer assembly imparted an additional surface area for a higher amount of dye loading [38]. The decrease in open-circuit voltage (*V<sub>oc</sub>*) in the bilayer assembly could be attributed to the nanoparticles’ elevated interfacial resistance. The *J<sub>sc</sub>* improved significantly in the bilayer assembly due to increased spectral response with a longer wavelength. Interestingly, for both the photoanode samples, an increase in *V<sub>oc</sub>* and a decrease in *J<sub>sc</sub>* were noted in the back side illumination. The rise in *V<sub>oc</sub>* was ascribed to the slow electron recombination, whereas the decrease in *J<sub>sc</sub>* was related to reducing light penetration intensity in the back

**Table 1** Photovoltaic parameters of TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer DSSC cells (active area of cell was 1 cm<sup>2</sup>)

	Materials for photoanode	<i>V<sub>oc</sub></i> (V)	<i>J<sub>sc</sub></i> (mA cm <sup>-2</sup> )	FF (%)	PCE (%)	<i>P<sub>max</sub></i> (mW)
Front side illumination	Dye-adsorbed TiO <sub>2</sub>	0.24	2.50	52.20	0.31	0.31
	Dye-adsorbed TiO <sub>2</sub> / MnO <sub>2</sub> bilayer	0.23	7.00	54.70	0.88	0.88
Back side illumination	Dye-adsorbed TiO <sub>2</sub>	0.40	2.00	41.10	0.32	0.32
	Dye-adsorbed TiO <sub>2</sub> / MnO <sub>2</sub> bilayer	0.38	6.02	40.30	0.92	0.92



**Fig. 7** Current density–voltage (*J–V*) and power voltage (*P–V*) plots of TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer solar cells



side illumination direction [39]. Notably, both light scattering effects and light reabsorbing ability of dye molecule reduced in the back side illumination.

The recombination kinetics of dye-coated TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer-assembled DSSC devices were analyzed using open-circuit voltage decay (OCVD) studies. From the as-obtained OCVD curve (Fig. 8), the lifetime of V<sub>oc</sub> prevailed in the region between steady state and dark equilibrium. Moreover, the devices' OCVD profiles presented a slow decay rate for dye-coated TiO<sub>2</sub>/MnO<sub>2</sub> bilayer devices compared to the device devoid of bilayer assembly, confirming a slow recombination rate in the bilayer structure of the device. Herein, the lifetime of the charge carrier is influenced by photoanode hierarchical architecture [40]. Notably, the lifetime of the charge carrier and V<sub>oc</sub> decay were correlated by the following reported equation [41]:

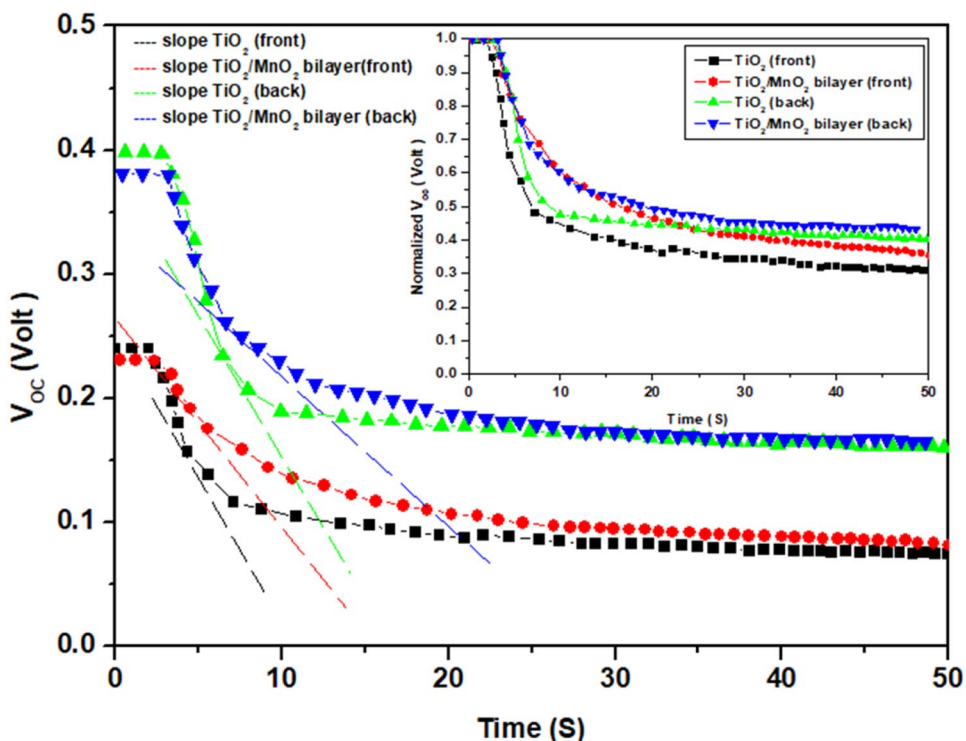
$$\tau = \frac{kT}{e} \left( \frac{dV_{oc}}{dt} \right)^{-1}, \tag{1}$$

Here, *k*, *T*, and *e* represent Boltzmann constant, absolute temperature, and charge of electron, respectively. The bilayer structure photoanode envisaged a slower slope than the device without bilayer, supporting a slower recombination rate and longer lifetime of charge carrier (see Table 2).

### Conclusions

Natural dye extract from *Malvaviscus penduliflorus* flower has been used for the fabrication of low-cost and eco-friendly DSSCs. The dye molecule's adsorption in the TiO<sub>2</sub> matrix facilitates the charge transfer from the dye molecule to the TiO<sub>2</sub> conduction band, elevating the efficiency. MnO<sub>2</sub> in the TiO<sub>2</sub>/MnO<sub>2</sub> bilayer structure reduces the bandgap, improving DSSC performance. The FTIR result indicates the interaction of dye molecule with alcohol. The change in intensity and shift of the peaks have revealed a charge-transfer complex between the dye molecule and alcohol. The hydroxyl anchored group's presence in the dye solution improves the light absorption coefficient in the visible

**Fig. 8** Open circuit voltage decay of TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer DSSC in the dark after an illumination period of 3 s at AM1.5 G



**Table 2** Lifetime of the charge carrier in TiO<sub>2</sub> and TiO<sub>2</sub>/MnO<sub>2</sub> bilayer DSSC cells

	Photoanode materials	Lifetime/(τ(s))
Front side illumination	Dye-adsorbed TiO <sub>2</sub>	0.289
	Dye-adsorbed TiO <sub>2</sub> / MnO <sub>2</sub> bilayer	0.344
Back side illumination	Dye-adsorbed TiO <sub>2</sub>	0.320
	Dye-adsorbed TiO <sub>2</sub> / MnO <sub>2</sub> bilayer	0.560

spectrum. FESEM images of bilayer structure photoanode comprise the maximum surface coverage with the minimum agglomeration and void portion. Incorporating the  $\text{MnO}_2$  layer works as an interface between the  $\text{TiO}_2$  and electrolyte layer, which improves the surface passivation and electron transport. The deposition of the  $\text{MnO}_2$  layer on the  $\text{TiO}_2$  photoanode has no significant impact on the phase change of the  $\text{TiO}_2$  material. Thus, the interaction between  $\text{TiO}_2$  and  $\text{MnO}_2$  layer has effects only on the surroundings optoelectronic change. It was evident that the minor shift of the main  $\text{TiO}_2$  peak position at a higher diffraction angle can be attributed to incorporating  $\text{Mn}^{2+}$  ion into the  $\text{TiO}_2$  lattice. TG analysis confirmed the thermostability of dye up to 100 °C. Combining an electron transport layer like  $\text{MnO}_2$  on  $\text{TiO}_2$  material significantly improves the optoelectronic property of the photoanode. It is also concluded that incorporating an electron transport layer like  $\text{MnO}_2$  on  $\text{TiO}_2$  material significantly improves the optoelectronic property of the photoanode due to a combination of several factors including the increase of dye absorption. The significant improvement of quality charge extraction increases the diffusional length of charge carriers and reduces surface recombination enhanced the performance of the DSSC.

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## Compliance with ethical standards

**Conflicts of interest** There are no conflicts to declare.

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