

Nucleation and Growth Mechanisms for Pd–Pt Bimetallic Nanodendrites and Their Electrocatalytic Properties

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Received: 14 November 2009 / Revised: 25 November 2009 / Accepted: 27 November 2009

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ABSTRACT

In a seed-mediated synthesis, nanocrystal growth is often described by assuming the absence of homogeneous nucleation in the solution. Here we provide new insights into the nucleation and growth mechanisms underlying the formation of bimetallic nanodendrites that are characterized by a dense array of Pt branches anchored to a Pd nanocrystal core. These nanostructures can be easily prepared by a one-step, seeded growth method that involves the reduction of K_2PtCl_4 by *L*-ascorbic acid in the presence of 9-nm truncated octahedral Pd seeds in an aqueous solution. Transmission electron microscopy (TEM) and high-resolution TEM analyses revealed that both homogeneous and heterogeneous nucleation of Pt occurred at the very early stages of the synthesis and the Pt branches grew through oriented attachment of small Pt particles that had been formed via homogeneous nucleation. These new findings contradict the generally accepted mechanism for seeded growth that only involves heterogeneous nucleation and simple growth via atomic addition. We have also investigated the electrocatalytic properties of the Pd–Pt nanodendrites for the oxygen reduction and formic acid oxidation reactions by conducting a comparative study with foam-like Pt nanostructures prepared in the absence of Pd seeds under otherwise identical conditions.

KEYWORDS

Palladium, platinum, seeded growth, oxygen reduction, formic acid oxidation

Introduction

Understanding of nucleation and growth mechanisms is a prerequisite for the rational design of metal nanostructures with desired properties. Recently, heterogeneous, seeded growth has emerged as a powerful tool for precisely controlling the morphology and composition of bimetallic nanostructures [1–13]. This technical approach has enabled the preparation of a wide variety of bimetallic nanostructures that cannot be achieved otherwise. Notable examples

include nanocubes [1–3] and nanoplates [4–6] with well-defined core–shell structures, highly anisotropic nanorods consisting of a secondary metal deposited at the tips of pre-formed nanorods [7–9], and nanodendrites characterized by a dense array of branched arms made of one metal supported on a core of another metal [10–13].

In heterogeneous, seeded growth, pre-formed nanocrystals are generally considered to serve as primary sites (i.e., seeds) for the nucleation of another metal with a different chemical identity and nanocrystal

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growth is often described by assuming the absence of homogeneous nucleation in the solution. Experimentally, this postulation has been supported by observations that include the absence of isolated particles made of a secondary metal in the final product and/or the epitaxial relation between the core and the shell. A recent theoretical study based on Monte Carlo simulations of heterogeneous nucleation has shown that foreign seeds larger than a critical size can indeed lower the free-energy barrier for a nucleation process and the rate of homogeneous nucleation is vanishingly small [14]. Nevertheless, the exact mechanism of seeding by a foreign nanocrystal still remains largely unresolved due to the lack of evidence for nucleation and growth occurring at early stages of a synthesis.

Recently, Pd–Pt and Au–Pt bimetallic nanodendrites have been synthesized using the seeded growth method by our and other research groups [10–13]. These bimetallic nanodendrites are of particular interest due to their high specific surface areas, and have shown great promise for use as catalysts or electrocatalysts with high activity. In previous studies, however, a mechanism that only considers heterogeneous nucleation and simple growth via atomic addition was suggested to account for the formation of such bimetallic nanostructures without direct evidence for nanocrystal growth in solution. Here we provide new insights into the nucleation and growth mechanisms underlying the formation of a bimetallic nanodendrite that consists of Pt branches anchored to a Pd nanocrystal core. In this new work, we discovered that both homogeneous and heterogeneous nucleation of Pt occurred at the very early stages of seeded growth, and led to the formation of discrete small Pt particles in the solution and Pt bumps at multiple sites on the Pd seed, respectively. Furthermore, we also found that the growth of Pt branches mainly proceeded via oriented attachment of the initially formed Pt particles. The new results presented herein contradict the previously reported mechanisms and provide clear evidence for the involvement of homogeneous nucleation in the heterogeneous, seeded growth of bimetallic nanodendrites. In addition to the mechanistic study, we systematically investigated the electrocatalytic properties of the Pd–Pt nanodendrites

by conducting a comparative study with foam-like Pt nanostructures prepared in the absence of Pd seeds under otherwise identical conditions. We have applied these nanostructures not only as a cathode catalyst for the oxygen reduction reaction (ORR) but also as an anode catalyst for the formic acid oxidation reaction in a proton-exchange membrane (PEM) fuel cell.

The Pd–Pt nanodendrites were synthesized in an aqueous solution by reducing K_2PtCl_4 with *L*-ascorbic acid in the presence of 9-nm truncated octahedral Pd seeds, with the addition of poly(vinyl pyrrolidone) (PVP) as a stabilizer [12]. Transmission electron microscopy (TEM) and high-resolution TEM images of the Pd seeds are shown in Fig. S-1 of the Electronic Supplementary Material (ESM). To investigate the morphological evolution of the Pd–Pt nanodendrites, aliquots of the reaction solution were taken out at various stages and examined using electron microscopy. Figures 1(a)–1(d) show typical TEM images of the samples that were taken out at 1, 2, 5, and 10 min, respectively. For comparison, we also studied the morphological evolution of Pt nanostructures in the absence of Pd seeds under otherwise identical conditions (Figs. 1(e)–1(h)). At $t=1$ min, the Pd–Pt sample contained a large number of small Pt particles with sizes less than 3 nm in addition to Pt-decorated Pd seeds (Fig. 1(a)), indicating that both homogeneous and heterogeneous nucleation of Pt occurred at the early stages of the reaction. Figure 2(a) gives a high-resolution TEM image of a single Pd–Pt particle, clearly showing the formation of Pt bumps via heterogeneous nucleation at multiple sites on the truncated octahedral Pd seed. The small Pt particles formed via homogeneous nucleation had a single-crystal structure, as revealed by the high-resolution TEM image in Fig. 2(b). As the reaction proceeded to $t=5$ min, the number of small Pt particles rapidly decreased while a number of Pt branches started to appear on each Pd seed in random directions (Figs. 1(b) and 1(c)). These observations indicate that the Pt branches were formed through attachment of the initially formed, small Pt particles. Some secondary branches also began to grow from the main branches during this period. In the following 5 min, essentially all the small Pt particles disappeared and the Pd–Pt

particles evolved into a highly branched structure with an overall dendritic morphology (Fig. 1(d)). The Pt branches were distributed over the entire surface of the Pd core and did not overlap appreciably. Growth via particle attachment was also observed in the absence of Pd seeds (Figs. 1(e)–1(h)). In this case, however, the small Pt particles that had been formed via homogenous nucleation self-aggregated into a denser nanostructure with a foam-like morphology. Both the nanostructures, i.e., Pd–Pt nanodendrites and foam-like Pt aggregates, did not exhibit any morphological change during the next 3 h (Fig. S-2 in the ESM).

We suggest that in the synthesis of the Pd–Pt nanodendrites, homogeneous nucleation of the Pt occurred as follows. In the initial stage of synthesis, the degree of supersaturation for Pt atoms increased rapidly due to the fast reduction of K_2PtCl_4 precursor by *L*-ascorbic acid. However, the number of possible nucleation sites provided by the Pd seeds might not be sufficient to accommodate all the Pt atoms generated from the reduction and, as a result, the supersaturation would be kept at a high level, facilitating homogeneous nucleation in the solution [15, 16]. Accordingly, the Pt atoms that could not nucleate on the surfaces of Pd seeds would self-nucleate to form small Pt particles.

The observed attachment of Pt particles might be related to the ineffective stabilization by PVP. It has been shown that PVP is less effective as a stabilizer in a polar medium such as water [17–19]. The small particles have a higher chemical potential due to a large surface area-to-volume ratio, and particle attachment can reduce surface energy by reducing the surface area. We attribute the formation of Pt branches to the oriented attachment mechanism

[20–27], where adjacent particles self-organize so that they share a common crystallographic orientation. Figure 2(c) shows a high-resolution TEM image of a single Pd–Pt nanodendrite. The lattice fringes were

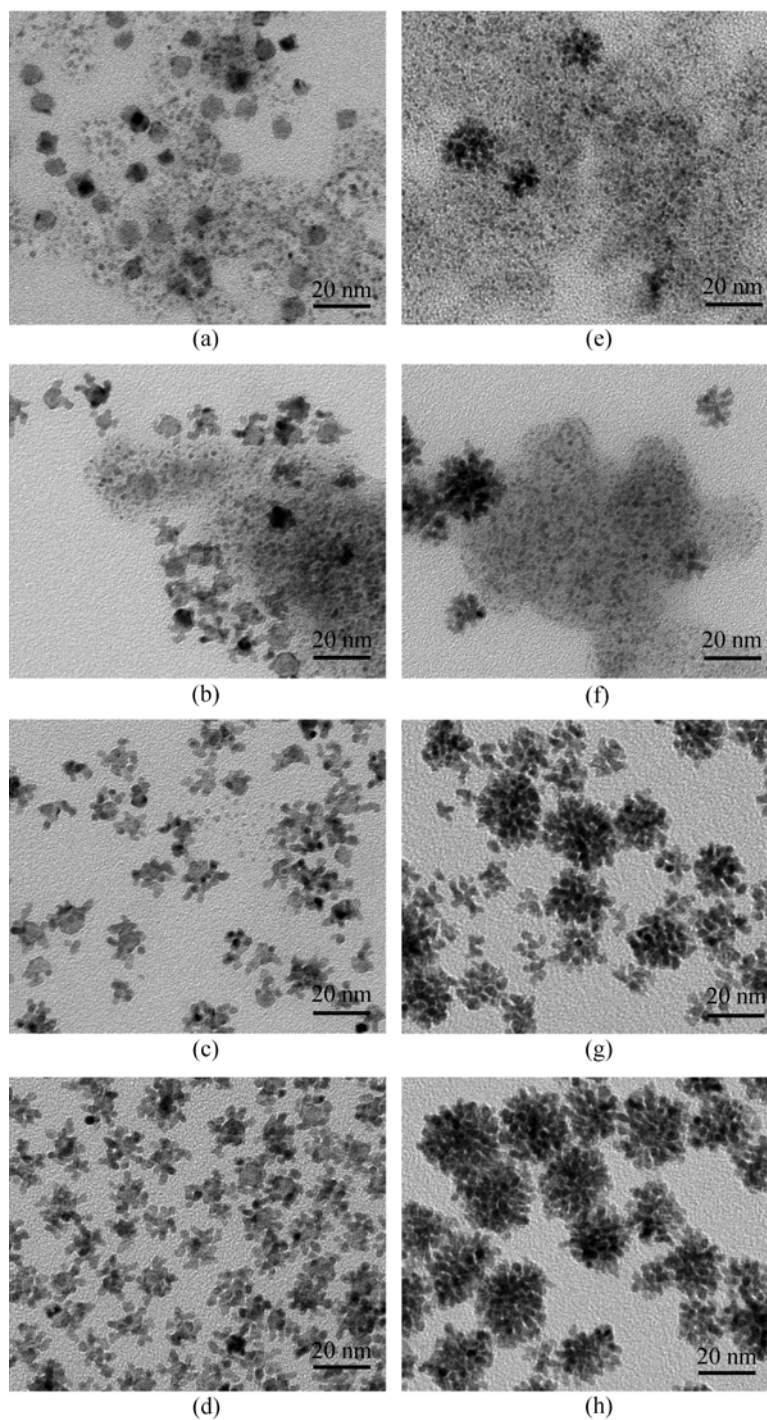


Figure 1 TEM images showing the morphological evolution of (a–d) Pd–Pt nanodendrites and (e–h) foam-like Pt aggregates synthesized in the presence and absence of Pd seeds, respectively. The reaction time was (a, e) 1 min, (b, f) 2 min, (c, g) 5 min, and (d, h) 10 min

coherently extended from the Pd core to the Pt branches, suggesting that the Pt particles were epitaxially fused into the Pt bumps nucleated on the Pd seed. Most of the Pt branches exhibited a single-crystal structure with a highly ordered, continuous fringe pattern, which supports the growth mechanism based on oriented attachment. However, we also found examples of the formation of a twin boundary at the attachment interface (Fig. 2(d)). In an oriented attachment process, twins can be formed when the particles on opposite sides of the interface are related by mirror symmetry [26]. High-resolution TEM analysis also revealed that the Pt branches were predominantly bounded by {111} facets in addition to some {110} facets (Fig. S-3 in the ESM).

Figures 3(a)–3(c) show high-resolution TEM images of a series of small Pt aggregates obtained at 5 min

into the reaction carried out in the absence of Pd seeds. The aggregates had either a linear or a branched structure. The lattice fringes of these aggregates were perfectly aligned along the long axis, indicating that they were also formed via the oriented attachment mechanism. The similarity in morphology between the Pd-supported Pt branches and the small Pt aggregates strongly suggests that both nanostructures originated from the same growth mechanism based on oriented attachment. Although the initial Pt aggregates were defect free, the lattice fringes of the final, foam-like aggregates were not perfectly aligned across the entire particle (Fig. 3(d); also see Fig. S-4 in the ESM). This phenomenon is likely due to the involvement of twinning and/or imperfect oriented attachment characterized by a minor mis-orientation at the interface [28]. In the high-resolution TEM

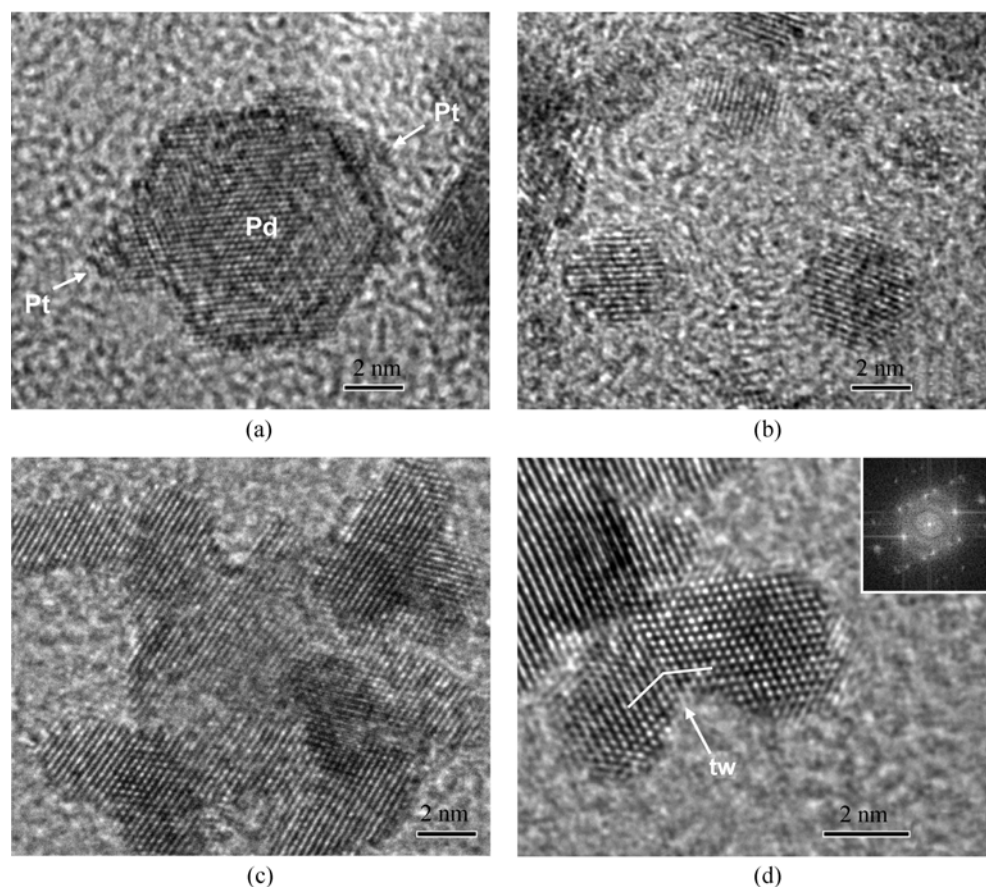


Figure 2 High-resolution TEM images of (a) an individual Pd–Pt particle and (b) the small Pt particles shown in Fig. 1(a). In (a), the Pt bumps on the Pd seed are indicated by arrows. (c) High-resolution TEM image of a single Pd–Pt nanodendrite. It can be seen that the lattice fringes are coherently extended from the Pd core to the Pt branches. (d) High-resolution TEM image taken from a Pt branch containing a single twin plane in its structure and the corresponding Fourier transform (FT) pattern (inset). A twin plane is indicated by tw

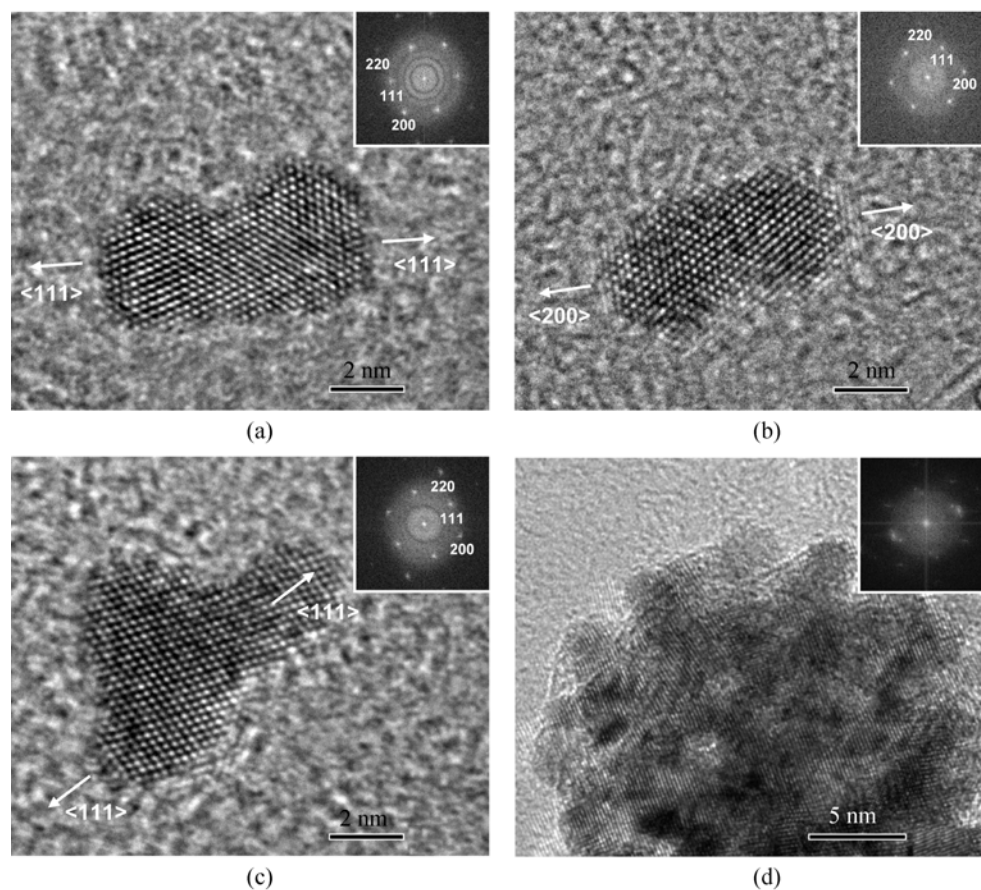


Figure 3 (a–c) High-resolution TEM images of small Pt aggregates obtained at 5 min into the reaction carried out in the absence of Pd seeds. (d) High-resolution TEM image of a foam-like Pt aggregate. The insets show the corresponding FT patterns

image, the foam-like Pt aggregates did not appear to have preferred facets on their surfaces.

In the absence of strong interparticle interactions (e.g., dipole–dipole interactions) [29–31] or a templating effect provided by self-assembled molecules [32, 33] the particles tend to aggregate in a diffusion-controlled manner and accordingly dendritic growth can occur [34, 35]. In previous studies, a mechanism based on slow nucleation and fast growth mediated by autocatalytic reduction has been proposed to explain the formation of foam-like Pt nanostructures prepared with *L*-ascorbic acid as a reducing agent in an aqueous solution [36, 37]. As demonstrated in Figs. 1(e)–1(h), however, this type of nanostructures is a result of extensive self-aggregation, which has also been observed in the synthesis of nanoparticles of other metals such as Ni, Ru, and Rh [38–40]. The truncated octahedral Pd seeds play a pivotal role in achieving an open, dendritic structure by providing

multiple nucleation sites for Pt. In this case, attachment is initiated at the Pt bumps on the Pd seed that are spatially separated from each other and, therefore, the Pt branches can grow without significant overlap and fusion between them. In this way, a large number of attachment sites provided by the growing Pd–Pt particles could also enable to avoid the formation of pure, foam-like Pt aggregates.

If two metals have a very close lattice match, the mode of heterogeneous nucleation is mainly determined by the interrelationship of the specific surface energies of the substrate (γ_s), the deposit (γ_d), and the substrate–deposit interface (γ_i). This condition applies to the Pd–Pt system because Pd and Pt have a lattice mismatch of only 0.77%. According to the thermodynamic analysis established by Bauer [41], the deposition proceeds in an island growth mode (Volmer–Weber mechanism) when $\gamma_s < \gamma_d + \gamma_i$ while the layered growth (Frank–van der Merwe mechanism) takes place when

$\gamma_s = \gamma_d + \gamma_i$. Bauer's conditions predict an island growth mode for Pt deposition on a Pd surface because Pt has a relatively higher surface energy than Pd (Table S-1 in the ESM) [42], which is also consistent with our observation; that is, the formation of Pt bumps at multiple sites on the Pd seed. In a previous study, we observed the conformal growth of thin Pt layers on the surface of Pd nanoplates in the presence of citric acid as both the reducing and capping agent [5]. Citric acid seems to play a significant role in stabilizing the {111} facets of the deposited Pt layers and lowering their surface energy, thus promoting the layered growth. In the current system, the stabilizing effect of PVP is relatively weak and, therefore, the layered growth is not favored.

Platinum is one of the most widely used elements in electrocatalysis because of its excellent properties in the adsorption and dissociation of hydrogen, oxygen, and various other molecules [43–48]. The performance of a Pt or Pt-based electrocatalyst is highly dependent upon the electrochemically active surface area (ECSA), which determines the number of catalytically active sites available for an electrochemical reaction. We first compared the ECSAs of the Pd–Pt nanodendrites and the foam-like Pt aggregates using cyclic voltammetry (CV). Figure 4(a) shows CV curves recorded at room temperature in N₂-purged 0.1 mol/L HClO₄ solutions with a sweep rate of 50 mV/s. The ECSA was calculated by measuring the charge collected in the H_(upd) adsorption/desorption region after double-layer correction [49], where H_(upd) refers to the underpotentially deposited hydrogen. The ECSA was then normalized to the loading amount of metal in order to obtain the specific ECSA (i.e., ECSA per unit mass of metal). The specific ECSA of the Pd–Pt nanodendrites was 52.7 m²/g_{Pd+Pt}, which was almost twice that of the foam-like Pt aggregates (28.8 m²/g_{Pt}). If the Pt mass was solely taken into account (85 wt.% based on inductively coupled plasma mass spectrometry (ICP-MS) measurements), the specific ECSA of the Pd–Pt nanodendrites (62.0 m²/g_{Pt}) was more than twice that of the foam-like Pt aggregates, demonstrating that the open dendritic structure provides a higher surface area.

We employed these two nanostructures as electrocatalysts for the ORR, the rate-determining step in a

PEM fuel cell [50]. The ORR measurements were performed at room temperature in O₂-saturated 0.1 mol/L HClO₄ solutions using a glassy carbon rotating disk electrode (RDE) with a sweep rate of 10 mV/s and a rotation rate of 1600 r/min. Figure 4(b) presents ORR polarization curves for these two catalysts. The half-wave potentials for the Pd–Pt nanodendrites and the foam-like Pt aggregates were 867 and 834 mV versus a reversible hydrogen electrode (RHE), respectively, indicating a higher ORR activity for the Pd–Pt nanodendrites. The kinetic current was calculated from the ORR polarization curve by using mass-transport correction and then normalized to the loading amount of metal in order to obtain the mass activity. Figure 4(c) shows a quantitative comparison of the mass activities of the Pd–Pt nanodendrites and the foam-like Pt aggregates. In a mixed kinetic-diffusion control region between 0.80 and 0.95 V versus RHE, the Pd–Pt nanodendrites outperformed the foam-like Pt aggregates by a factor of three on the basis of equivalent Pt mass.

In order to achieve a better understanding of the observed difference in ORR activity, we also compared the specific activities (i.e., kinetic current per unit surface area of catalyst) of these two catalysts. As illustrated in Fig. 4(d), the specific activity of the Pd–Pt nanodendrites was 1.4–1.6 times greater than that of the foam-like Pt aggregates in the potential region between 0.80 and 0.95 V versus RHE, indicating the accelerated ORR kinetics on the surface of the Pd–Pt nanodendrites. It is known that the difference in ORR activity on low-index Pt surfaces most likely arises from the structure-sensitive inhibiting effect of OH_(ad) species on Pt(*hkl*) and, in a non-adsorbing electrolyte such as perchloric acid, the activity increases on the order of Pt(100) << Pt(111) < Pt(110), with the difference in activity between Pt(111) and Pt(110) being quite small [51]. For the Pd–Pt nanodendrites, the preferential exposure of {111} facets in addition to some {110} facets on the Pd-supported Pt branches may account for their higher specific activity as compared to the foam-like Pt aggregates with poorly-defined surface facets.

The Pd–Pt nanodendrites have also been evaluated as electrocatalysts for the oxidation of formic acid. We performed potentiodynamic experiments in order

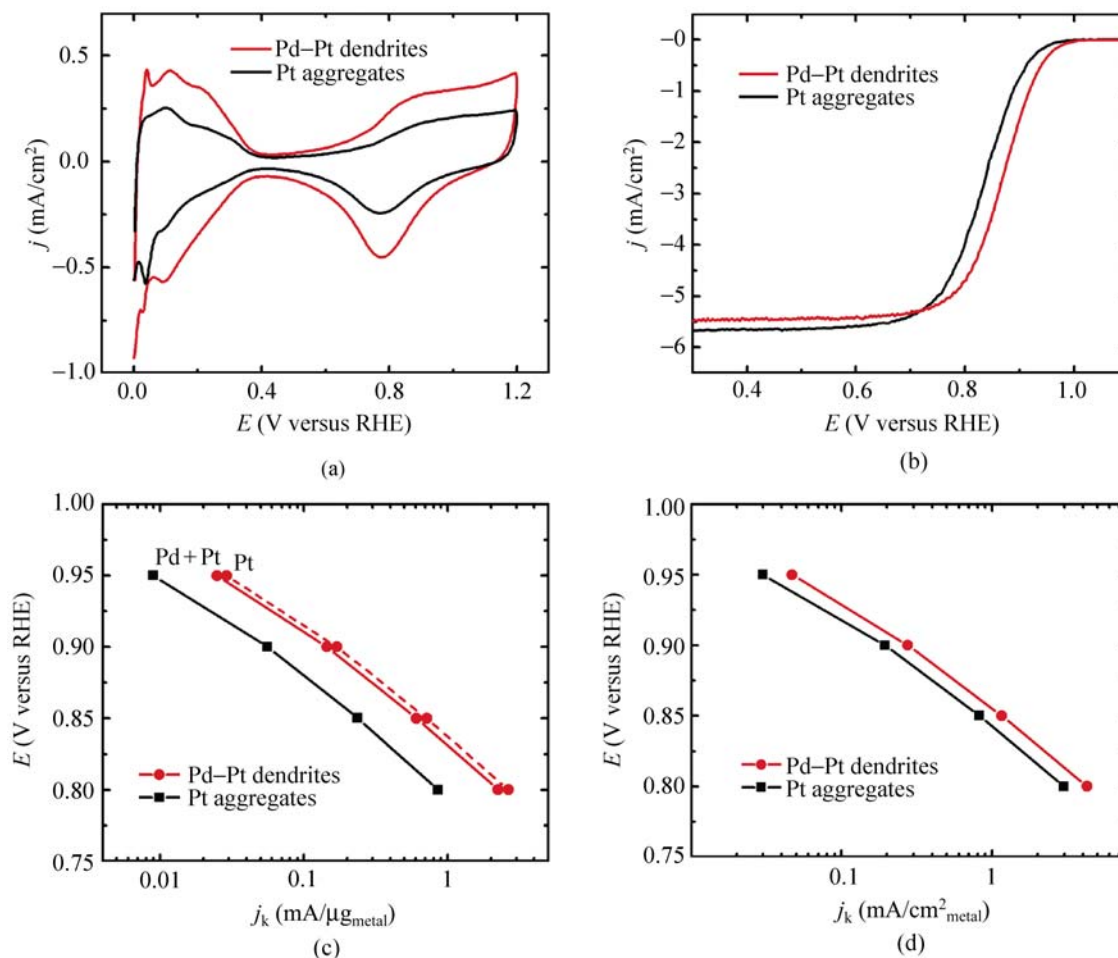


Figure 4 Comparison of electrocatalytic properties of the Pd–Pt nanodendrites and foam-like Pt aggregates for the ORR: (a) CV curves recorded at room temperature in N_2 -purged 0.1 mol/L $HClO_4$ solutions at a sweep rate of 50 mV/s; (b) ORR polarization curves recorded at room temperature in O_2 -saturated 0.1 mol/L $HClO_4$ solutions at a sweep rate of 10 mV/s and a rotation rate of 1600 r/min; (c, d) Mass and specific activities given as kinetic current densities (j_k) normalized against the loading amount of metal and the ECSA of catalyst, respectively (red solid line is the activity based on the total mass of Pd and Pt (i.e., Pd + Pt); red break line is the activity based on the mass of Pt). For all electrodes, the metal loading on a glassy carbon electrode was $15.3 \mu\text{g}/\text{cm}^2$. In (a) and (b), the current densities were normalized against the geometric area of the electrode (0.196 cm^2)

to investigate the individual reaction steps and adsorbed intermediates involved in the formic acid oxidation. Figure 5(a) shows CV curves for the formic acid oxidation catalyzed by the Pd–Pt nanodendrites and the foam-like Pt aggregates. For the foam-like Pt aggregates, only a minimal current associated with formic acid oxidation was observed between 0.30 and 0.70 V in the positive scan, implying that the surface was largely deactivated by the formation of a self-poisoning species (e.g., $CO_{(ad)}$). The oxidative elimination of the adsorbed CO at higher overpotentials ($>0.80 \text{ V}$) regenerates a bare Pt surface, which accounts for the higher currents observed in

the negative scan. On the other hand, the Pd–Pt nanodendrites showed less poisoning in the positive scan, with relatively higher currents between 0.30 and 0.70 V, suggesting that they provide a direct path for formic acid oxidation that does not form a self-poisoning species.

For a quantitative assessment of electrocatalytic activity, potentiostatic experiments were conducted, in which the potential was stepped to a target potential that was varied from 0.55 to 0.70 V and held for 60 s to obtain a steady-state current. Figure 5(b) shows transient current density curves for the formic acid oxidation on the Pd–Pt nanodendrites and the



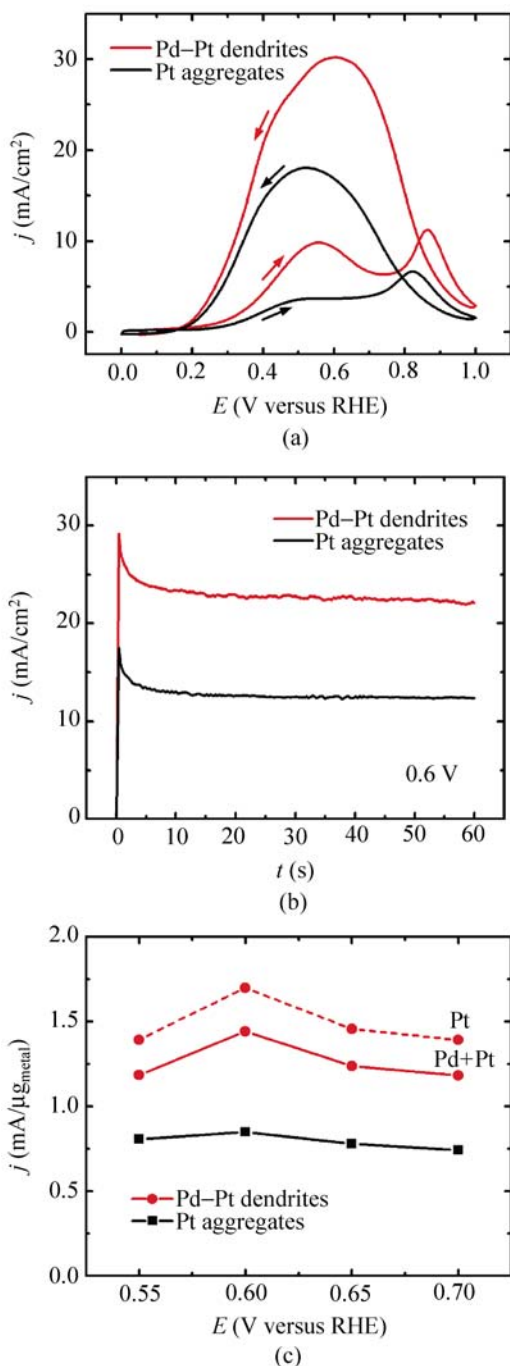
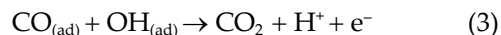
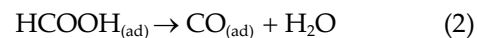
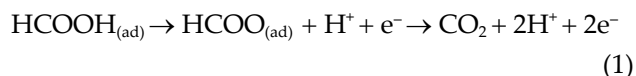


Figure 5 Comparison of electrocatalytic properties of the Pd–Pt nanodendrites and foam-like Pt aggregates for the formic acid oxidation reaction: (a) CV curves recorded at room temperature in 0.25 mol/L HCOOH + 0.5 mol/L H₂SO₄ solutions with a sweep rate of 50 mV/s; (b) transient current density curves at 0.6 V versus RHE; (c) mass activities given as the steady-state current densities (recorded at 60 s) normalized against the loading amount of metal (red solid line is the activity based on the total mass of Pd and Pt (i.e., Pd + Pt); red break line is the activity based on the mass of Pt). For all electrodes, the metal loading on a glassy carbon electrode was 15.3 μg/cm². In (a) and (b), the current densities were normalized against the geometric area of the electrode (0.196 cm²)

foam-like Pt aggregates recorded at 0.60 V. The steady-state current recorded at 60 s has been normalized to the loading amount of metal in order to compare the mass activity. As shown in Fig. 5(c), the Pd–Pt nanodendrites exhibited higher mass activities than the foam-like Pt aggregates in the potential region of 0.55–0.70 V versus RHE. At a peak potential of 0.60 V versus RHE, the Pd–Pt nanodendrites exhibited a two-fold improvement in Pt mass activity over the foam-like Pt aggregates.

In the electro-oxidation of formic acid, CO_(ad) formed by dehydration of formic acid has been identified as the main poisoning species on Pt surfaces. Formic acid oxidation on Pt surfaces proceeds via the so-called dual-path mechanism as shown in the following reaction schemes [52]:



The direct path involves the dehydrogenation of formic acid to CO₂ (reaction 1), while a dehydration reaction produces CO_(ad) and H₂O (reaction 2). The adsorbed CO may have two modes of action on the formic acid oxidation depending on the overpotential. At relatively low overpotentials, CO_(ad) acts as a poison and thus impedes the formic acid oxidation, whereas at higher overpotentials (typically > 0.80 V versus RHE) CO_(ad) can serve as a reactive intermediate by reacting with OH_(ad) formed from the dissociation of water and thus provide a reaction channel to CO₂ (reaction 3). It has been shown that for the formic acid oxidation on low-index Pt surfaces, the dehydrogenation and dehydration pathways are predominant on the Pt(111) and Pt(100) surfaces, respectively, while the Pt(110) surface exhibits an intermediate behavior [53]. The predominant exposure of {111} facets on the Pt branches of the Pd–Pt nanodendrites could allow the direct dehydrogenation path for formic acid oxidation and thus lead to reduced poisoning by CO_(ad) species, which may contribute to their higher activity.

In summary, we have resolved the nucleation and growth mechanisms involved in the seed-mediated synthesis of Pd–Pt bimetallic nanodendrites. Our results indicate that both homogeneous and heterogeneous nucleation of Pt occurred at very early stages of the synthesis and the growth proceeded via particle attachment. The Pd seeds played a crucial role in forming an open, dendritic structure by providing multiple nucleation sites for Pt that were spatially separated from each other, helping to avoid overlap and fusion between the Pt branches during the growth process. In the absence of Pd seeds, extensive self-aggregation of small Pt particles led to the formation of Pt nanostructures with a foam-like morphology. Due to their open, dendritic structure, the Pd–Pt nanodendrites exhibited a higher surface area and much enhanced activity for both the oxygen reduction and formic acid oxidation reactions as compared to the foam-like Pt aggregates. This work not only greatly advances our understanding of the dendritic growth of Pt nanostructures but also provides new insights into the mechanisms involved in heterogeneous, seeded growth. It is worth pointing out that foam-like Pt nanostructures have been reported by a number of groups (including us) and suggested as a promising candidate for a fuel cell electrocatalyst. The comparative studies described herein clearly contradict that suggestion. Our results show that a synthetic methodology based on heterogeneous, seeded growth provides a simple and more efficient way for generating Pt-based electrocatalysts with improved activities.

Experimental

Synthesis of Pd–Pt nanodendrites. For the preparation of 9-nm truncated octahedral Pd seeds, 11 mL of an aqueous solution containing PVP ($M_w \approx 55\,000$, 105 mg, Aldrich), *L*-ascorbic acid (60 mg, Aldrich), citric acid (60 mg, Fisher), and Na_2PdCl_4 (57 mg, Aldrich) was heated at 100 °C in air under magnetic stirring for 3 h and cooled down to room temperature [54]. The Pd–Pt nanodendrites were prepared using a previously reported method [12]. In a typical synthesis, 1 mL of the as-prepared suspension of the Pd seeds and 6 mL of an aqueous solution containing PVP

(35 mg) and *L*-ascorbic acid (60 mg) were added into a 25-mL, three-necked flask. The mixture was heated to 90 °C in air under magnetic stirring. Meanwhile, K_2PtCl_4 (27 mg, Aldrich) was dissolved in 3 mL of deionized water at room temperature. The aqueous K_2PtCl_4 solution was then rapidly injected into the flask using a pipette. The reaction mixture was heated at 90 °C in air for 3 h, and then cooled down to room temperature. The product was collected by centrifugation and washed several times with ethanol and deionized water for further use in electrochemical measurements.

Samples for TEM and high-resolution TEM. In the preparation of samples for TEM and high-resolution TEM studies, aliquots of the reaction solution were rapidly cooled by adding pre-cooled acetone. The resultant solutions were dropped directly onto carbon-coated copper grids without centrifugation and dried inside the vacuum chamber of a freeze-drier. After drying, the grids were washed with ethanol to remove excess PVP. TEM images were captured using a Phillips 420 microscope operated at 120 kV. High-resolution TEM was carried out on a JEOL 2100F microscope operated at 200 kV.

Electrochemical measurements. Electrochemical measurements were performed using a glassy carbon RDE (Pine Research Instrumentation) connected to a PARSTAT 283 potentiostat (Princeton Applied Research). A leak-free $\text{AgCl}/\text{Ag}/\text{KCl}$ (3 mol/L) electrode (Warner Instrument) was used as the reference. The counter electrode was a platinum mesh (1 cm × 1 cm) attached to a platinum wire. All potentials were converted to values with reference to RHE. To prepare the working electrode, the sample was diluted to 0.15 $\mu\text{g}/\mu\text{L}$ (based on ICP-MS measurements) and 20 μL of the dispersion was transferred onto the glassy carbon RDE with a geometric area of 0.196 cm^2 . Therefore, the loading amount of metal for Pd–Pt nanodendrites and foam-like Pt aggregates was 3 μg (i.e., 15.3 $\mu\text{g}/\text{cm}^2$ based on the geometric electrode area). After water evaporation, the electrode was covered with 15 μL of 0.05 wt.% Nafion solution. In order to produce a clean electrode surface, several potential sweeps between 0.0 and 1.2 V versus RHE were applied to the electrode prior to the measurement.



CV measurements were carried out in 0.1 mol/L HClO₄ solutions under a flow of N₂ (Airgas, ultrahigh purity) at a sweep rate of 50 mV/s. The ECSA was estimated by measuring the charge associated with H_(upd) adsorption (Q_H) between 0.05 and 0.40 V and assuming a value of 210 μC/cm² for the adsorption of a monolayer of hydrogen on a Pt surface (q_H). The H_(upd) adsorption charge (Q_H) can be determined using $Q_H = 0.5 \times Q$, where Q is the charge in the H_(upd) adsorption/desorption area obtained after double-layer correction. Then, the specific ECSA was calculated based on the following relation: specific ECSA = $Q_H/(m \times q_H)$, where Q_H is the charge for H_(upd) adsorption, m is the loading amount of metal, and q_H is the charge required for the adsorption of a monolayer of hydrogen on a Pt surface.

The ORR measurements were performed at room temperature in 0.1 mol/L HClO₄ solutions under a flow of O₂ (Airgas, Research grade) using the glassy carbon RDE at a rotation rate of 1600 r/min and a sweep rate of 10 mV/s. The kinetic current was calculated based on the Koutecky–Levich equation as follows:

$$\frac{1}{i} = \frac{1}{i_k} + \frac{1}{i_d}$$

where i is the experimentally measured current, i_d is the diffusion-limiting current, and i_k is the kinetic current. For each catalyst, the kinetic current was normalized to the loading amount of the metal and the ECSA in order to obtain mass and specific activities, respectively.

The electrochemical measurements for the formic acid oxidation reaction were performed at room temperature in a solution containing 0.25 mol/L HCOOH and 0.5 mol/L H₂SO₄. CV curves were collected at a sweep rate of 50 mV/s. In the potentiostatic experiments, the potential was stepped from 0.0 V to the high potential that varies from 0.55 to 0.70 V, and the transient current density curve of the formic acid oxidation was recorded.

Acknowledgements

This work was supported in part by the National Science Foundation (NSF) (No. DMR-0804088) and

startup funds from Washington University in St. Louis. T.Y. was also partially supported by the National Research Foundation of Korea Grant funded by the Korean Government (No. NRF-2009-352-D00160). Pedro Henrique Cury Camargo (P.H.C.C.) was also partially supported by the Fulbright Program and the Brazilian Ministry of Education (CAPES). Part of the work was performed at the Nano Research Facility (NRF), a member of the National Nanotechnology Infrastructure Network (NNIN), which is supported by the National Science Foundation under award No. ECS-0335765. NRF is part of the School of Engineering and Applied Science at Washington University in St. Louis.

Electronic Supplementary Material: Additional TEM and high-resolution TEM images (Figs. S-1 to S-4) and the surface energies of Pd and Pt (Table S-1) are available in the online version of this article at <http://dx.doi.org/10.1007/s12274-010-1010-8> and are accessible free of charge.

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