

Strong Enhancement of deSoot Activity of Transition Metal Oxides by Alkali Doping: Additive Effects of Potassium and Nitric Oxide

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Abstract A series of potassium-promoted spinels (Mn, Fe, Co) were prepared with various K⁺ promoter locations: on the surface (surface promotion) or in the bulk (formation of new layered and tunneled nanostructures via solid state reaction). All prepared samples were characterized by means of X-ray diffraction, Raman spectroscopy, X-ray fluorescence and N₂-BET specific surface area analysis. Catalytic activity in soot combustion in different reaction conditions was investigated (*tight contact*, *loose contact*, *loose contact* with NO addition). It was shown that in all cases the nanostructuring is more effective than the surface promotion, with the layered structures of KCo₄O₈, KMn₄O₈ being the most catalytically active phases, lowering the soot combustion down to 250 °C. The difference in activity between *tight* and *loose contacts* can be bridged in the presence of NO due to its transformation into NO₂, which acts as the oxygen carrier from the catalyst surface into soot particles, eliminating the soot-catalyst contact difference.

Keywords Soot combustion · NO · Potassium promotion · Spinel · Nanostructuring

1 Introduction

During pyrolysis and the incomplete combustion of fuels (such as wood, coal, lignite or different fractions of petroleum) agglomerated soot particles of different carbon

forms are produced. They are one of the main components of particulate matter (PM). Long exposition of the human organism to PM leads to many illnesses like cardiovascular and lung diseases, heart attacks or arrhythmia. The nanometric size of the soot particles allows for their high penetration into the human body causing DNA code modifications that result in mutations and possible cancerogenic changes. Moreover, PM accumulation leads to atherosclerosis, antenatal complications, and respiratory system diseases. Deposition of PM, especially soot, on snow and ice, lowers its ability to reflect sunlight. The ice or snow, absorbing solar energy more readily, melts and in further effect has an impact on global warming [1].

The most important anthropogenic sources are coal combustion in electric plants and home heating systems, waste incineration, and car emissions—soot from fuel combustion and dust from tires wear. Since the main part of anthropogenic particulate matter is soot, it is an important task to decrease its emissions from diesel engines [1]. Many governments have introduced strict norms for pollutant emission [2]. These have led to the development of new technologies for soot emission reduction. It seems that catalytic diesel particulate filters (CDPF) are the most promising technology. DPFs can lower soot emission by up to 99 % but soot accumulates inside them and regeneration is needed. Catalysts can regenerate DPFs in the temperature of exhaust gasses while a temperature of up to 600 °C is needed for regeneration without any catalyst [3, 4].

A lot of different groups of catalysts were investigated for soot combustion—noble metal-based, perovskites, spinels, ceria catalysts, chlorides, low melting point catalysts. The catalytically active phases were recently reviewed in [3, 5–8].

The alkali-containing systems form an interesting group of soot combustion catalysts. Alkali metals (Li, Na, K, Rb,

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Cs) can be used as surface or bulk dopants in the metal oxide matrix. Surface promotion of CeO_2 by a series of alkali metals showed that the catalytic activity depends both on the dopant nature and its concentration [9, 10]. In [11] and [12] structural and surface alkali promotion of SrTiO_3 was compared. The results showed that all dopants increased the activity with potassium promotion being the most effective. It was also shown that structural promotion was more effective than the surface one. Details concerning alkali promotion can be found in [13]. Systematic investigations of perovskite $\text{K}_x\text{La}_{1-x}\text{FeO}_3$, spinels $\text{A}_{2x}\text{Cu}_{1-x}\text{Fe}_2\text{O}_4$ and $\text{A}_{2x}\text{Co}_{1-x}\text{Fe}_2\text{O}_4$ (where A is an alkali metal) led to the conclusion that the most effective promoter is potassium and the activity depends on the promoter concentration. More comprehensive studies on the surface alkali promotion of d- and f-electron metal oxides such as cobalt, lanthanum, iron, chromium, manganese, niobium, vanadium, and cerium are presented in [14]. Examples of the use of mixed oxides (potassium bulk promotion) as catalysts include $\text{K}_6\text{Ti}_4\text{O}_{11}$, $\text{K}_2\text{Ti}_2\text{O}_5$, $\text{K}_2\text{Ti}_4\text{O}_9$ [15], KNbO_3 [16], KFeO_2 and $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ [17] or KVO_3 and $\text{K}_4\text{V}_2\text{O}_7$ [18]. Potassium containing glasses [19, 20] have also been found recently as a component of the most active soot combustion catalysts [21].

Since nitrogen oxides are present in diesel exhaust, their influence on catalytic activity in the soot oxidation process has been investigated. It was found that NO has a positive effect on catalyst activity in soot combustion as it forms NO_2 , which is a stronger oxidant for carbon particles than the oxygen [22]. A systematic shift of the temperature window of soot combustion on potassium-promoted spinels in the presence of NO was also reported [23]. It was found that potassium-containing catalysts can be used for the simultaneous removal of NO_x and soot [24–26], so similar catalytic systems were also investigated in bench-scale experiments [26–28].

The aim of the study is to show how the doping of transition metal (Mn, Fe, Co) oxides by alkali facilitates the soot combustion process, by shifting the combustion temperature window below the target 400 °C. This issue was addressed by systematic studies of catalytic activity in the soot combustion of a series of surface and structural potassium-promoted transition metal spinels in *tight*, *loose* contacts and various gas environments (with and without NO).

2 Experimental

2.1 Catalyst Preparation

The surface doped catalysts were prepared by the wet impregnation of 1 g of the iron (Aldrich) and cobalt spinels (Fluka) with a potassium carbonate solution, to achieve the

optimal alkali loading corresponding to the theoretical 0.5 monolayer, as described in [23]. The impregnated spinel was dried at 100 °C for 1 h and subsequently calcined at 400 °C for 4 h. These samples were further designated as $\text{K/Fe}_3\text{O}_4$ and $\text{K/Co}_3\text{O}_4$, respectively.

The samples of $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ ferrite were prepared by the solid-state reaction of a stoichiometric mixture of K_2CO_3 (POCh) with $\alpha\text{-Fe}_2\text{O}_3$ (Merck) at 1200 °C as described previously [29]. The KFeO_2 sample was synthesized from K_2CO_3 with $\alpha\text{-Fe}_2\text{O}_3$ according to the procedure described in [30]. In brief, the mixtures of finely ground powders were placed in a porcelain crucible and heated at the rate of 5 °C/min in the static air up to the final temperatures of 800 °C for KFeO_2 and 1200 °C for $\text{K}_2\text{Fe}_{22}\text{O}_{34}$.

The cryptomelane phase was prepared as described in [31]. The mass of 11 g of manganese(II) acetate was dissolved in 40 ml of distilled water and 5 ml of glacial acetic acid was added to achieve a pH = 5. The solution was heated and stirred under reflux for 30 min. 6.5 g of potassium permanganate was dissolved in 150 ml of distilled water, then this solution was added to the previously prepared solution of Mn^{2+} . The mixture was heated and stirred for 24 h under reflux. The solid residue was washed with water (for neutral pH), dried overnight at 110 °C, and calcined at 450 °C for 2 h.

Birnessite was synthesized as described in [32]. A glucose solution (1.4 mol/dm^3) and KMnO_4 solution (0.27 mol/dm^3) were mixed, with a glucose to KMnO_4 molar ratio equal 1.5. The solution was stirred at room temperature until a xerogel was formed. The xerogel was washed with water and dried at 110 °C overnight. The powder was calcined at 450 °C for 2 h.

The cobalt oxide nanostructured by potassium (KCo_4O_8) was prepared by grinding 2.13 g of the cobalt spinel with 1.1 g of potassium hydroxide (POCh). The produced powder was placed in a preheated oven and calcined at 700 °C for 12 h. The pre-calcined bulk materials were placed in preheated ovens to avoid alkali loss due to thermal desorption, further compensated by the use of excessive amounts of alkali substrate (three times excess of potassium hydroxide). The synthesized bulk-doped sample was thoroughly washed with distilled water to remove excess potassium and ground in an agate mortar into a fine powder for further studies [33].

2.2 Catalyst Characterization

XRD patterns were recorded on a Rigaku MiniFlex powder diffractometer with Cu $\text{K}\alpha$ radiation, 2 θ step scans of 0.02° and a counting time of 1 s per step. The micro-Raman spectra were recorded under ambient conditions using a Renishaw InVia spectrometer equipped with a Leica DMLM confocal microscope and a CCD detector, with an

excitation wavelength of 785 nm. The laser power at the sample position was 1.5 mW (0.5 % of total power) with a magnification of 20 \times . The Raman scattered light was collected in the spectral range of 150–800 cm^{-1} . At least 6 scans, 10 s each, were accumulated to ensure a sufficient signal to noise ratio. For XRF analysis, ARL Quant'X spectrometer with Uniquant software, calibrated with metal standards series, was used. Specific surface area (SSA) was determined by three-point nitrogen adsorption/desorption with BET isotherm method with Quantasorb Junior (Quantachrome). The X-ray photoelectron spectra (XPS) of the catalyst samples were measured with a Prevac photoelectron spectrometer equipped with a hemispherical VG SCIENTA R3000 analyzer. The photoelectron spectra were measured using a monochromatized aluminum $\text{AlK}\alpha$ source ($E = 1486.6$ eV). The base pressure in the analysis chamber during the measurements was 5×10^{-9} mbar. Spectra were recorded with a constant pass energy of 100 eV for the survey and narrow scan spectra. The binding energy scale was aligned with the main C1 s band set to 295 eV.

2.3 Soot Combustion Tests

The activity in soot combustion of the catalysts was measured by means of TPO (temperature-programmed oxidation). The quartz fixed-bed reactor was heated (10 $^{\circ}\text{C}/\text{min}$) from RT to 750 $^{\circ}\text{C}$ while the concentrations of the released gases were measured by a quadrupole mass spectrometer (SRS RGA 200). Two gas mixtures were used in the tests: 5 % O_2 in He, as well as 3.3 % O_2 , 0.3 % NO in He, both at a 60 ml/min flow. The samples for the combustion test were prepared by weighing 0.05 g of soot (Degussa–Printex80, with the SSA of 230 $\text{m}^2 \text{g}^{-1}$ and average composition of 98 % carbon and 1 % oxygen, 1 % of hydrogen, and other mineral components (of which 80 % sulfur, 10 % phosphorus and 10 % potassium), examined by XRF and XPS) and the catalyst with an 8:1 ratio and then grinding them in a mortar for 10 min following the tight contact method or gently mixing in an Eppendorf tube following the *loose contact* protocol. The shifts of the temperature of 50 % soot conversion (ΔT_{50}) was defined as T_{50} for unpromoted spinel oxides minus T_{50} for potassium-promoted samples. We also assessed the value of the reaction onset. For this purpose we used of the Arrhenius plot, assuming that CO_2 concentration is proportional to the reaction rate constant (at the beginning of the reaction), and plot $\log(p_{\text{CO}_2}/\text{Pa})$ as a function of T^{-1} . On such Arrhenius-like plot the intercept between the line associated with the p_{CO_2} background and the line corresponding to CO_2 evolution due to soot combustion reaction is taken as an indicator of the ignition temperature. We found this value

to be independent on the amount of soot, thus more reliable than T_{10} , and take it as a measure of the reaction onset.

3 Result and Discussion

3.1 Characterization

The stoichiometry of all the investigated catalysts was assessed by XRF analysis (Table 1). An excess of potassium (observed also as potassium carbonate segregation on XRD results, see below) was confirmed in the birnessite and β -ferrite samples. The specific surface area of the investigated samples varied from 4 to 70 m^2/g . Separate catalytic experiments with the same phase composition (birnessite) with different specific surface areas showed no significant difference in their reactivity.

Diffraction patterns obtained for the investigated samples are shown in Fig. 1—manganese-based in Fig. 1a, iron-based in Fig. 1b, and cobalt-based in Fig. 1c. For all samples, the diffraction peaks were labeled with Miller indexes. The manganese catalysts, i.e., spinel, cryptomelane (tunneled structure), and birnessite (layered structure), were indexed in the $Pmab$ (ICSD-30005 [34]), $I4/m$ (ICSD-59159 [35]) and $R\bar{3}m$ (ICSD-152290 [36]) space groups, respectively. Some reflections originating from an excess of potassium carbonate are only detectable in the case of the birnessite phase (marked with violet diamonds). Likewise, the iron catalysts were indexed, with the spinel ($Pmc2_1$ (ICSD-35002 [37]), β -ferrite $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ (tunneled structure) ($P6_3/mmc$ (ICSD-83285 [29]) and monoferrite KFeO_2 (layered structure), ($Pbca$ (ICSD-94467 [38]) space groups. Some minor reflections from excess hematite $\alpha\text{-Fe}_2\text{O}_3$ are visible for the surface promoted and β -ferrite samples (marked with orange squares). The cobalt spinel Co_3O_4 was indexed in the $Fd\bar{3}m$ (ICSD-150805 [39]) space group. Potassium cobaltate KCo_4O_8 (tunneled structure) was indexed in the $R3m$ space group (ICSD-89454 [40]).

Table 1 Actual formulas from XRF analysis and specific surface areas (SSA) of the investigated samples

	Formula from XRF	SSA ($\text{m}^2 \text{g}^{-1}$)
Mn_3O_4	–	6
$\text{KMn}_8\text{O}_{16}$	$\text{K}_{0.93}\text{Mn}_8\text{O}_{16}$	70
KMn_4O_8	$\text{K}_{1.28}\text{Mn}_4\text{O}_8$	5.6
Fe_3O_4	–	35
$\text{K}_2\text{Fe}_{22}\text{O}_{34}$	$\text{K}_{2.08}\text{Fe}_{22}\text{O}_{34}$	4
KFeO_2	$\text{K}_{0.94}\text{FeO}_2$	17
Co_3O_4	–	16
KCo_4O_8	$\text{K}_{0.76}\text{Co}_4\text{O}_8$	5

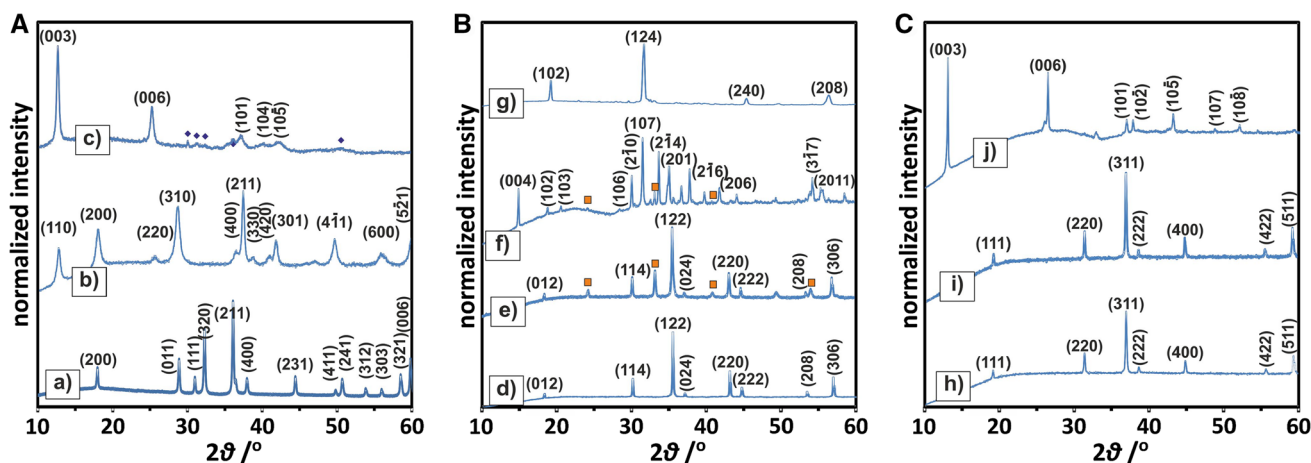


Fig. 1 Diffraction patterns of manganese-based (a), iron-based (b) and cobalt-based (c) catalysts. Manganese spinel Mn_3O_4 (a), cryptomelane $\text{KMn}_8\text{O}_{16}$ (b), birnessite KMn_4O_8 (c), iron spinel Fe_3O_4 (d), potassium-doped Fe_3O_4 (e), β -ferrite $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ (f), monoferrite

KFeO_2 (g), cobalt spinel Co_3O_4 (h), potassium-doped Co_3O_4 (i), potassium cobaltate KCo_4O_8 (j). Hematite Fe_2O_3 phase is marked by orange squares

Raman spectra obtained for investigated samples were shown in Fig. 2—manganese-based in Fig. 2a, iron-based in Fig. 2b, and cobalt-based in Fig. 2c. Where possible the symmetry was assigned based on the supporting literature. The Mn_3O_4 spectra showed six visible bands: 175, 289, 318, 374, 477, and 658 cm^{-1} [41]. Bands at 186, 285, 393, 479, 514, 582, 639 and 750 cm^{-1} were visible on the cryptomelane spectra [42]. On the birnessite spectra, bands at 280, 409, 477, 506, 577, 634 and 728 cm^{-1} are present. The same bands were observed in [43].

The iron spinel Fe_3O_4 spectrum shows three visible bands: 304, 541 and 665 cm^{-1} [44]. In the potassium doped Fe_3O_4 spectrum, some additional peaks (marked with orange squares) were assigned to hematite, formed

during calcination. On the other hand potassium β -ferrite displays rich Raman spectrum. There is no literature concerning this material, but using detailed studies on $\text{BaFe}_{12}\text{O}_{19}$ spectrum, based on the vibrations of β -alumina (isostructural with β -ferrite) and other ferrites described in [45], allowed for some interpretation. The most intensive band, observed at 414 cm^{-1} , was assigned to the vibration of the octahedral group FeO_6 in the A_{1g} symmetry, as were the intensive bands at 517 and 618 cm^{-1} and the less intensive bands at 445 and 489 cm^{-1} . Relatively intensive peaks at 680 and 719 cm^{-1} were assigned to the vibrations of the FeO_5 bipyramid and FeO_4 tetrahedral group, respectively, both in the A_{1g} symmetry. The less intensive band at 221 cm^{-1} was assigned to vibrations of the O–Fe–

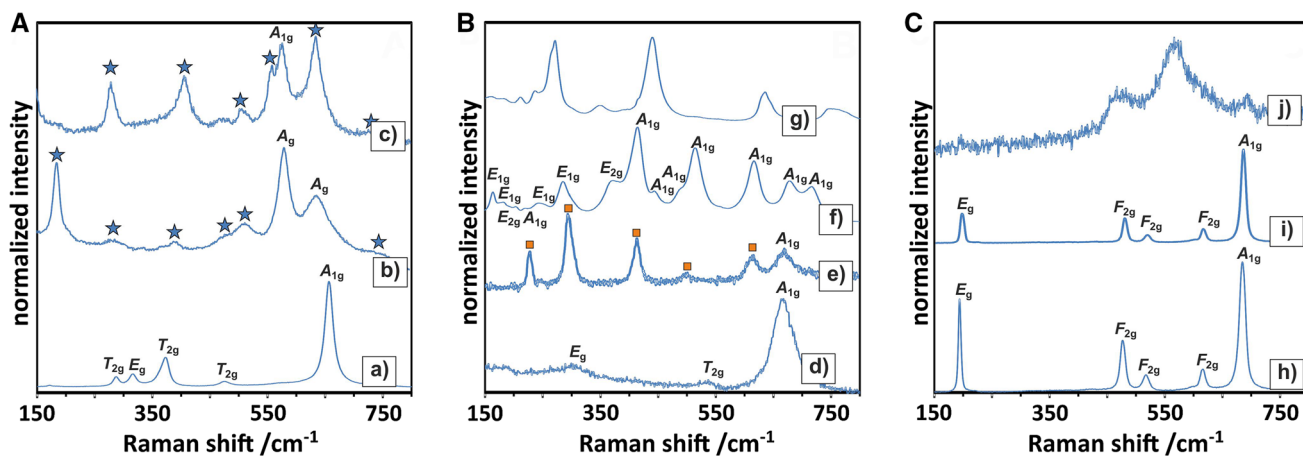


Fig. 2 Raman spectra of manganese-based (a), iron-based (b) and cobalt-based (c) catalysts. Manganese spinel Mn_3O_4 (a), cryptomelane $\text{KMn}_8\text{O}_{16}$ (b), birnessite KMn_4O_8 (c), iron spinel Fe_3O_4 (d), potassium-doped Fe_3O_4 (e), β -ferrite $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ (f), monoferrite KFeO_2 (g), cobalt spinel Co_3O_4 (h), potassium-doped Co_3O_4 (i),

potassium cobaltate KCo_4O_8 (j). Fingerprints without assigned symmetry are marked by blue stars (for KFeO_2 , $\text{K}_2\text{Fe}_{22}\text{O}_{34}$ Raman spectra are reported for the first time in the literature). Hematite α - Fe_2O_3 phase is marked by orange squares

O bridge in the A_{1g} symmetry. Relatively intensive and less intensive bands at 166 and 185 cm^{-1} were assigned to vibrations of spinel blocks in E_{1g} . The rest of the bands were not assigned to specific vibrations but symmetry can be assigned to -206 cm^{-1} (E_{2g}), 245 cm^{-1} (E_{1g}), 288 cm^{-1} (E_{1g}) and 373 cm^{-1} (E_{2g}) Potassium monoferrite spectra had fewer peaks than the β -ferrite spectra. It was related to a higher symmetry of the KFeO_2 primitive cell than that of $\text{K}_2\text{Fe}_{22}\text{O}_{34}$. This spectrum was different from other simple and mixed iron oxides so it can be used for identification but the lack of literature on this material prevented vibration assignment.

The Co_3O_4 spectrum displayed five bands: 193, 476, 518, 617 and 685 cm^{-1} [46]. No additional peaks were visible on potassium-doped Co_3O_4 Raman spectrum. On the KCo_4O_8 spectrum, three broad bands were visible: 464, 575, and 695 cm^{-1} .

Binding energies for K $2p_{3/2}$, O 1s and $2p_{3/2}$ maxima for transition metals are summarized in Table 2. In the case of birnessite and cryptomelane peaks assigned to both Mn(III) and Mn(IV) were clearly visible and shifted to higher energies due to potassium promotion. In the case of Fe-based samples, potassium promotion led to an increase in the iron binding energy. As described in [48] it is connected with the increase of Fe(III) content. The shift in Co $2p_{3/2}$ peak position to a higher energy can be interpreted as an increase in the Co(IV) content, as discussed in [33]. The highest differences between the binding energies of K $2p_{3/2}$ for surface and bulk promotion were seen for Fe-based samples 1.1–1.3 eV, whereas for Co-based catalysts it was only 0.3 eV. It showed a remarkable difference of the potassium state on the surface and in the bulk of the iron oxide matrix, whereas the difference in the case of the cobalt-oxide matrix is much lower. The difference of

potassium binding energy between the tunneled cryptomelane and layered birnessite phases is 0.7 eV, whereas for the tunneled monoferrite and layered β -ferrite it is only 0.1–0.3 eV. Oxygen 1 s peak energy was slightly increased in the case of Mn- and Co-based samples, whereas for Fe-based samples the observed shift was much higher.

In summary, the performed characterization of the obtained materials revealed that the catalysts stoichiometry was close to the assumed. The samples are essentially monophasic—surface promotion did not alter the diffraction patterns—whereas for bulk promoted samples the presence of minor spurious phases undetectable by XRD was also excluded by the locally more sensitive Raman probe. For all surface promoted materials the XPS analysis confirmed the presence of potassium. The surface or bulk (nano-structuration) location of the promoter was confirmed in the case of iron and cobalt. For the manganese spinel, due to a very low temperature of intercalation, nanostructured phases alone were observed [49].

3.2 Catalytic Activity

Detailed investigations of catalytic activity in different reaction conditions of the investigated samples were performed and the results are shown in Fig. 3 for *tight contact*, *loose contact* and *loose contact* with NO addition. As a reference, a conversion curve of the non-catalytic soot combustion is also presented (dotted line). The arrows represent the promotional effect as a shift of the conversion curves towards lower temperature around the T_{50} . The results clearly reveal that the soot combustion progress depends strongly on both the contact mode and also the type of transition metal and potassium location (bulk or surface), as well as the presence of NO.

Table 2 XPS binding energies for K 2p, O 1s and transition metal (Mn $2p_{3/2}$, Fe $2p_{3/2}$, Co $2p_{3/2}$) in the synthesized samples (this work) and the parent spinels (literature references)

Sample	Binding energy/eV		
	K $2p_{3/2}$	Transition metal	O 1s
		Mn $2p_{3/2}$	
Mn_3O_4	–	641.2	530.6 [47]
KMn_4O_8	293	641.6 (Mn^{3+}), 642.9 (Mn^{4+})	529.9, 531.8, 533.0
$\text{KMn}_8\text{O}_{16}$	292.3	644.1, 642.6	530.1, 530.6, 531.1
		Fe $2p_{3/2}$	
Fe_3O_4	–	709.0	528.3, 529.6 [48]
K/ Fe_3O_4	293.6	712.1	532.3, 530.1
$\text{K}_2\text{Fe}_{22}\text{O}_{34}$	292.5	710.3	529.6
KFeO_2	292.2–292.4	709.7–710.0	530.4–530.8
		Co $2p_{3/2}$	
Co_3O_4	–	779.4	529.4, 531.5 [33]
K/ Co_3O_4	292.7	779.2	529.2, 531.0
KCo_4O_8	293.0	779.7	531.1

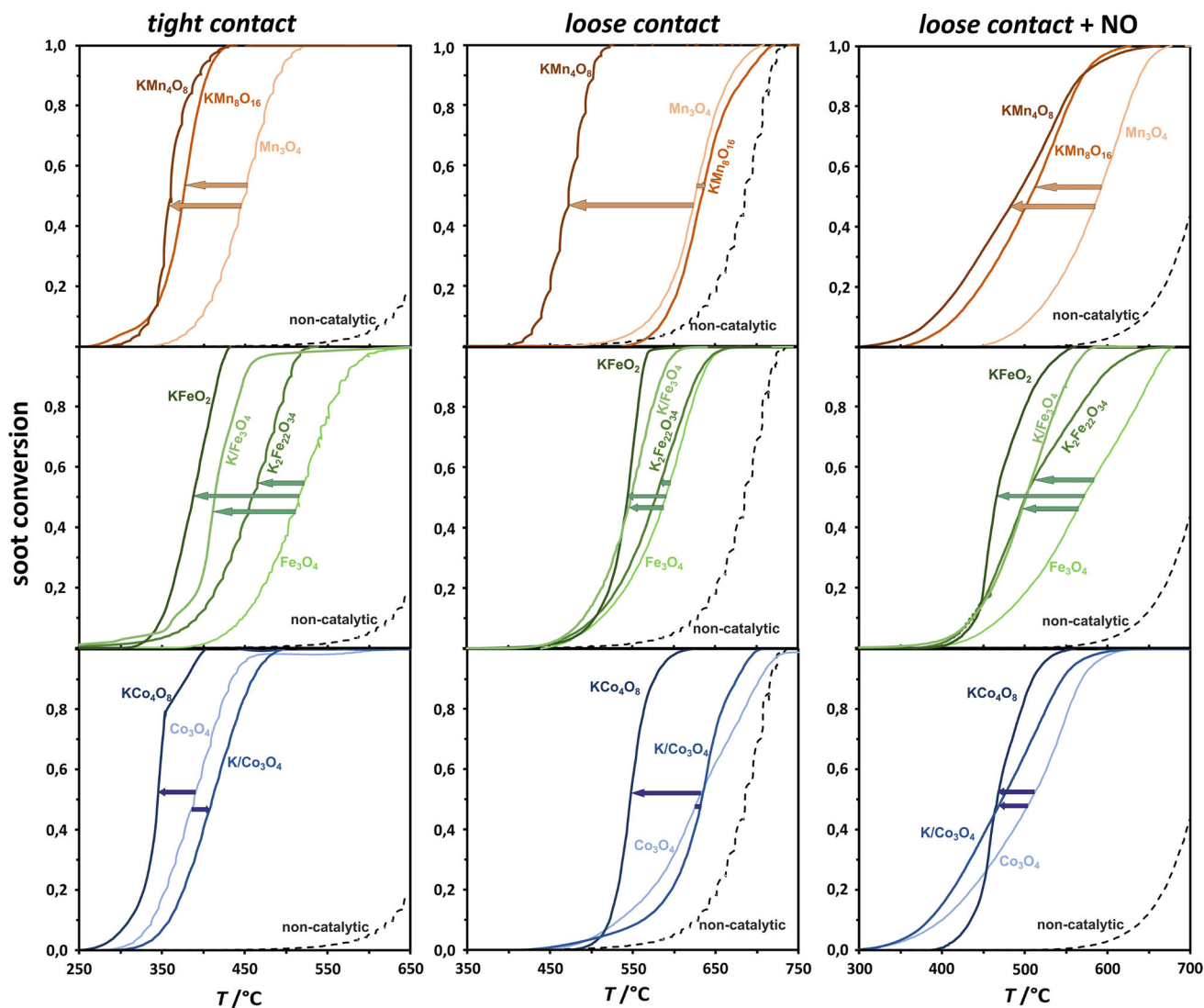


Fig. 3 Soot combustion tests on manganese, iron and cobalt spinels unpromoted and promoted surficial and structural by potassium in different reaction conditions—*tight contact*, *loose contact*, *loose contact* with NO addition

In the *tight contact* mode, the most active phases were potassium cobaltate (KCo_4O_8), birnessite (KMn_4O_8) and monoferrite (KFeO_2)—the temperature of 50 % conversion (T_{50}) of soot were shifted by about 45, 90, and 130 °C, respectively, in comparison to the unpromoted spinels.

Surface potassium promotion is beneficial for soot combustion, however, only in the case of iron ($\Delta T_{50} \sim 105$ °C), while for cobalt it has a slightly negative effect ($\Delta T_{50} \sim -20$ °C). A proposed explanation for these observations is the different interaction with oxygen. Whereas magnetite is readily oxidized into the corresponding sesquioxide ($\gamma\text{-Fe}_2\text{O}_3/\alpha\text{-Fe}_2\text{O}_3$) in the reaction conditions, cobalt spinel is oxidation resistant [50]. The interaction of the unpromoted Co_3O_4 with oxygen leads to the formation of surface oxygen species, which participate in the soot oxidation. However, the presence of potassium

leads to facile recombination of the surface oxygen species reducing their availability in the course of soot combustion [33]. The highest activity is exhibited by KCo_4O_8 , and KMn_4O_8 catalysts which lower down the reaction onset (determined from Arrhenius-like plot) to 255 and 250 °C, respectively, in comparison to non-catalytic soot combustion, beginning above 550 °C.

Summarizing, for all the investigated catalyst potassium promotion has a dramatic effect on catalytic activity, and the best results were observed for bulk promotion in each case. The highest activity is exhibited by KCo_4O_8 , and KMn_4O_8 catalysts, which lower the onset of the soot oxidation down to the spectacular value of ~ 250 °C.

In the *loose contact* mode the reaction, as expected, shifts to much higher temperatures (by ~ 150 °C), but the observed changes are not that regular as those in the *tight*

contact mode. The most active phases were the same as in *tight contact* mode: KMn_4O_8 , KFeO_2 , and KCo_4O_8 . In the case of the manganese-based samples, the largest activity increase for the KMn_4O_8 catalyst was revealed ($\Delta T_{50} \sim 150^\circ\text{C}$). Surprisingly, the second structurally promoted sample ($\text{KMn}_8\text{O}_{16}$) was less active than the unpromoted manganese spinel ($\Delta T_{50} \sim -10^\circ\text{C}$), in contrast to the *tight contact* mode. Only for iron catalyst the activity sequence $\text{KFeO}_2 > \text{K/Fe}_3\text{O}_4 > \text{K}_2\text{Fe}_{22}\text{O}_{34} > \text{Fe}_3\text{O}_4/\text{Fe}_2\text{O}_3$ is univocally preserved, though the differences are much smaller.

In the *loose contact* mode, the presence of NO in the gas mixture leads to the substantial decrease in the reaction temperature, making the situation quite similar to that observed for the *tight contact* mode measurements. In these conditions the characteristics of the conversion curves are different from those of *tight* and *loose contact* conditions—they exhibit a smaller slope, implying a lower activation energy for the reaction. The most striking observation was that for all the samples potassium promotion of bare spinel oxides decreased the T_{50} for the same values as in the *tight contact* conditions. The only exception was the surface-promoted cobalt spinel, which had a lower activity than the unpromoted spinel in *tight contact* conditions, and in *loose contact* with NO addition became the most active.

The shifts of the temperature of 50 % of soot conversion for all investigated samples are collected in Fig. 4. Inspection of the presented results clearly shows that the potassium promotional effects (shift in ΔT_{50} values with respect to the parent spinels) are strikingly similar to the *tight contact* conditions and the *loose contact* with NO addition. At the same time, they are dramatically different from the results obtained in the *loose contact* mode. This indicates that the loss of the mutual contact between the

catalyst surface and the soot particles upon passing from *tight* to *loose contact* can be to a large extent compensated in the presence of NO. It may be accounted for by the crucial role of surface mobility of oxygen species in soot oxidation [3, 7].

In the case of mobile surface oxygen species, the number of direct soot-catalyst contact is of less importance than for rigid oxygen. Indeed, spillover of mobile reactive oxygen species allows for oxidation of remote soot particles, whereas for immobile species their direct contact with the soot particles is required. This effect implies a large difference between *tight* and *loose contact* as observed in the case of e.g., Mn_3O_4 (immobile surface oxygen) and a smaller effect for KMn_4O_8 (mobile surface oxygen). As reported in detailed investigations, in the catalytic cycle of soot combustion NO can be oxidized to NO_2 , which in turn oxidizes soot more effectively [51, 52]. In such cases, it can be supposed that NO is preferentially oxidized on the investigated catalysts, transferring effectively the catalyst generated oxygen species into the soot particles. Thus, the difference in contact between soot and the catalyst resulting from the different mixing of the solids is substantially attenuated since the NO/ NO_2 couple acts as an oxygen carrier. Such a role of NO_2 as an oxygen carrier has also been reported in the case of deNO_x reactions over metallozeolites [53].

Summarizing, for Mn, Fe spinels the potassium promotion leads to pronounced enhancement in soot combustion activity in *tight* and *loose contacts*, however, for Co spinel the bulk promotion is positive, whereas surface promotion is slightly inhibitive. In all the investigated cases the nanostructuring of the parent spinel into the layered or tunneled structures results in the best catalytic performance. The difference in activity between *tight* and

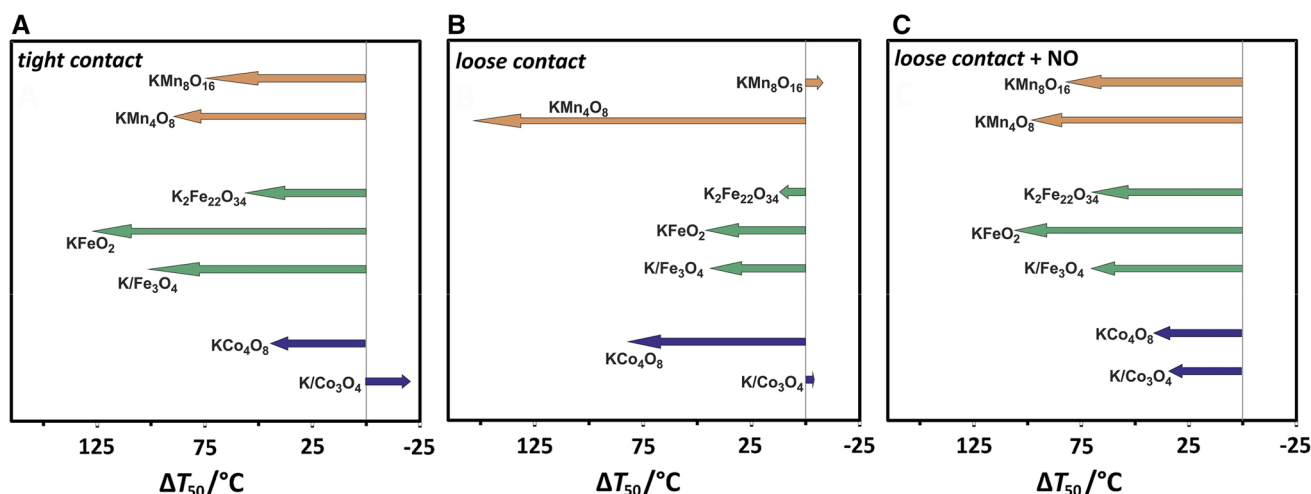


Fig. 4 Temperature shifts in soot combustion process caused by structural or surface potassium promotion of manganese, iron and cobalt spinels in *tight contact* (a), *loose contact* (b), *loose contact* and NO conditions (c)

loose contacts can be bridged in the presence of NO in the feed due to its transformation into NO₂, which acts as the oxygen carrier from the catalyst surface ($\text{NO} + \text{O}_{\text{surf}} \rightarrow \text{NO}_2$) into soot particles ($\text{NO}_2 + \text{C} \rightarrow \text{CO}_2 + \text{N}_2$).

4 Conclusions

The paper presents systematic studies of the promotional effect of potassium on the activity of transition metal (Mn, Fe, Co) spinels in catalytic soot combustion. The research presented in the paper reveals, that potassium addition can substantially improve the spinel activity in soot combustion, however, the effect strongly depends in each case on the promoter concentration but also on its location (bulk or surface). It was shown that the bulk promotion is more effective than the surface, with the layered structures of KCo₄O₈, KMn₄O₈ being the most active phases, which lower the onset of the soot oxidation down to the spectacular value of ~250 °C. The difference in activity between *tight* and *loose contacts* can be bridged in the presence of NO due to its transformation into NO₂, which acts as the oxygen carrier from the catalyst surface into soot particles, eliminating the soot-catalyst contact difference.

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