



# Isotherm kinetics of PIP2 bound gelsolin inactivation

Dávid Szatmári<sup>1</sup> · Dénes Lőrinczy<sup>1</sup>

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## Abstract

Actin monomers (G-actin) and filaments (F-actin) have dynamical rearrangement thus manage cellular motility, division and transport processes. The gelsolin (GSN) regulates the remodeling of cytoskeleton. After the activation of GSN by calcium ions, it can sever actin filaments then capped at its barbed end. In the cytoplasm, GSN manages the cellular motions and morphology. Phosphatidylinositol 4,5-bisphosphate (PIP2) is involved in signal transduction and the regulation of the actin cytoskeleton by regulation of actin-binding proteins. GSN can bind to PIP2 and thus can be localized in the near of the plasma membrane and released from the end of F-actin. We test here with isoperibol calorimetry the enthalpy change, within the interplay between GSN and F-actin under nano-, micro- and millimolar calcium concentrations and express the importance of PIP2 binding for the inactivation of GSN. As we have demonstrated here that PIP2 binding stabilizes the structure of gelsolin and reduces its actin monomer binding activity under nanomolar calcium as the typical cytoplasmic calcium concentration of resting cells. The gelsolin shows partial activity under micromolar and total activity with strong responses under millimolar calcium. If gelsolin-capped filaments point at the plasma membrane helps the binding between gelsolin and PIP2, and hence, filament uncapping in case of resting cells. We presume that the low free calcium concentration keeps on the structure of gelsolin which is able to bind actin within the cooperativity of actin bound calcium. Gelsolin can help to manage monomer pool far from the membrane and it can be linked to a basic sensory mechanism which drives the direction of filament growth in the near of the membrane.

**Keywords** Actin · Gelsolin · PIP2 · Isothermal kinetics

## Introduction

The actin is an essential unit of the eukaryotic cytoskeleton and muscle sarcomeres with highly conservative sequences [1]. The specialized intracellular functions driven to low variety of proteins in their phylogenetic development [2–4]. Actin monomers (G-actin) and filaments (F-actin) have dynamical rearrangement thus manage cellular motility, division and transport processes [5–11]. The remodeling of actin filaments highly regulated by divalent cations ( $\text{Ca}^{2+}$  or  $\text{Mg}^{2+}$ ) [12] and nucleotides (ATP, ADP) [13–20]. The stability of G-actin depends on the bound cations and nucleotides but the majority in cytoplasm is in the ATP-G-actin form [21, 22] then subsequently hydrolyses to ADP and Pi with their polymerization [13–20]. The gelsolin regulates

the rearrangement of cytoskeleton [21]. After the activation of gelsolin (GSN) by calcium ions, it can sever actin filaments then capped at its barbed end. In the cytoplasm nearby of the plasma membrane, GSN manages the cellular motions and morphology by capping, severing and uncapping of F-actin [22–24]. The process of uncapping is a crucial step to help the actin polymerization with an indirect influence on the membrane remodeling. Phosphatidylinositol 4,5-bisphosphate (PIP2) is involved in signal transduction and the regulation of the actin cytoskeleton by regulation of actin-binding proteins [25–42]. GSN can bind to PIP2 [43] and thus can be localized in the near of the plasma membrane when it is released from the end of F-actin [29, 44–49]. The PIP2-binding sites on gelsolin overlaps with the ATP binding site [47, 50–52]. The structural change of gelsolin with the mechanism of uncapping is not well known but in the absence of calcium the PIP2 competes with actin for binding to GSN [46, 49, 52–54]. However, we interpreted earlier that calcium activated GSN cannot bind PIP2 [49].

✉ Dénes Lőrinczy  
denes.lorinczy@aok.pte.hu

<sup>1</sup> Department of Biophysics, Medical School, University of Pécs, Szigeti Str. 12, Pécs 7624, Hungary

As previously shown the free calcium concentration in resting cells changing on the nanomolar level [55, 56], and in non-resting state can reach the micromolar levels [57, 58]. Cytoplasmic free magnesium levels are regulated in the nanomolar range but it does not have effect on GSN activity [49, 59, 60]. Interestingly, the reduction of calcium levels cannot cause directly the GSN release from actin, but a single calcium ion is still trapped in the complex [61]. The main concern with GSN inactivation is that it still depolymerizing actin filaments in the presence of millimolar EGTA provided nanomolar free calcium level [62, 63]. Possibly, only the PIP2 binding can fully inhibit any activity of GSN [49]. Structural rearrangement of F-actin and its binding molecules can be studied by DSC [64–66]. We probe here the isothermal enthalpy change within the interplay between GSN and F-actin under nano-, micro- and millimolar calcium concentrations and express the importance of PIP2 binding for the total inactivation of GSN. We can postulate that GSN shows remarkable thermal response to actin, addition under nanomolar calcium concentration which can be reduced by PIP2 binding.

## Results

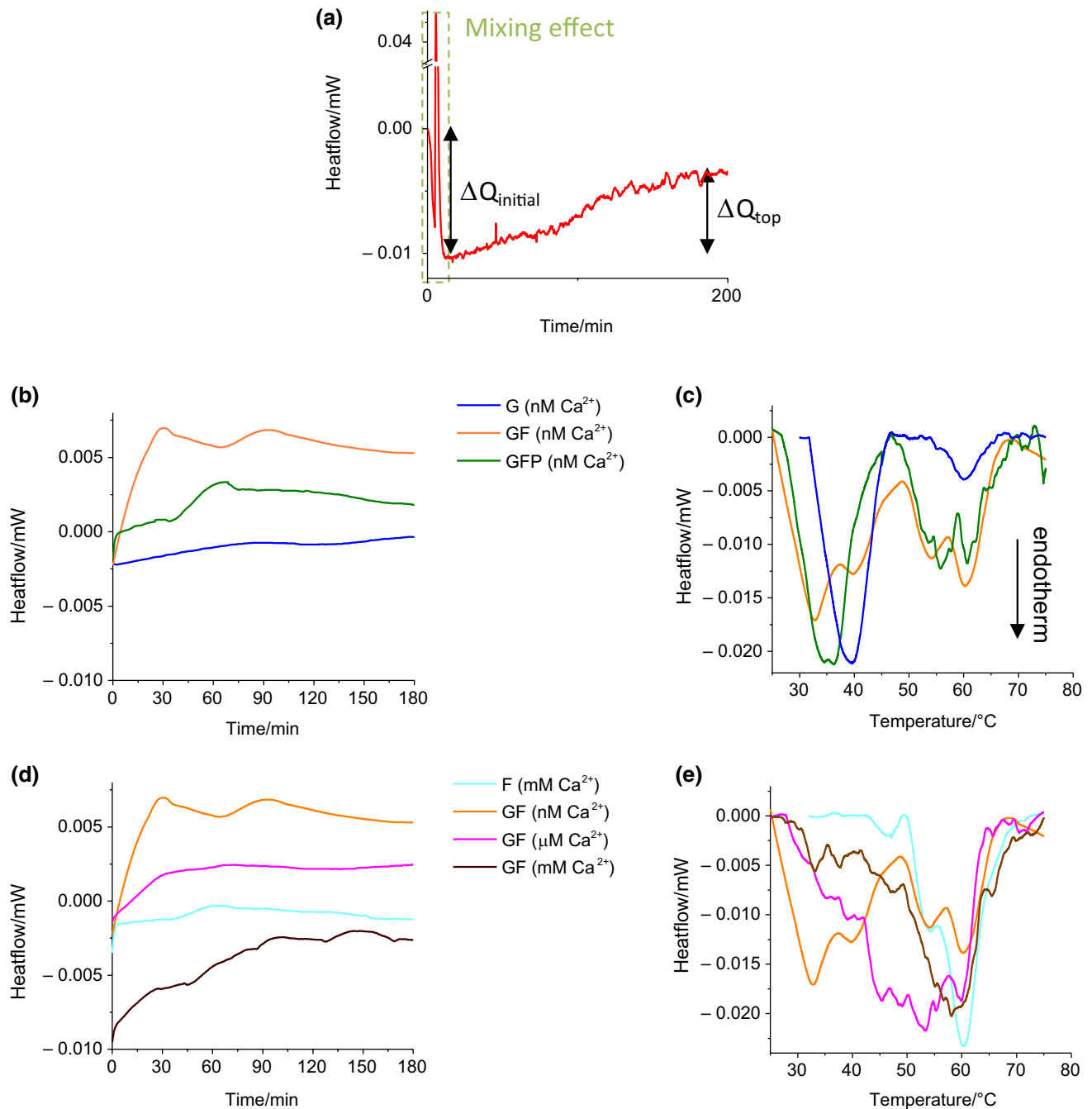
Figure 1 shows isothermal heatflow kinetics as well as the denaturation of mixing GSN with F-actin under different  $\text{Ca}^{2+}$  conditions, or in the presence of PIP2. All isotherm assays were taken 16 h but we focused on the first relevant 180 min after the relapse of the heatflow subsequently after the mixing effect, between the minimum ( $\Delta Q_{\text{initial}}$ ) and the maximum points ( $\Delta Q_{\text{top}}$ ) (Fig. 1A). The heatflow was dropped down ( $2.4 \mu\text{W}$ ) (Table 1) if  $12.8 \mu\text{M}$  GSN ( $1 \text{ mg mL}^{-1}$ ) was added to nanomolar  $\text{Ca}^{2+}$  containing buffer then increased slowly ( $1.7 \pm 0.3 \times 10^{-4} \mu\text{W s}^{-1}$ ) in the first 3 h (Fig. 1B). In case of  $46 \mu\text{M}$  F-actin ( $2 \text{ mg mL}^{-1}$ ) bound GSN in a ratio of 1: 3.6; under nanomolar  $\text{Ca}^{2+}$  concentration, there were two remarkable exothermic peaks. The first peak was increased quickly  $9.1 \pm 2.5 \text{ W}$  in the first 15 min by the rate of  $82 \pm 1.2 \times 10^{-4} \mu\text{W s}^{-1}$ , after a minor reduction it was followed by a second peak in 63 min. If GSN was formed complexes with  $1.6 \mu\text{M}$  PIP2 [49] in the presence of nanomolar  $\text{Ca}^{2+}$  and F-actin, the heatflow was increased quickly with multiple rates (totally  $4.6 \pm 0.7 \mu\text{W}$  with an average rate of  $17 \pm 2.5 \times 10^{-4} \mu\text{W s}^{-1}$ ) in 65 min. Just after the isothermal processes we performed denaturation with heating rate of  $0.3 \text{ K min}^{-1}$  (Fig. 1C). GSN in the presence of nanomolar  $\text{Ca}^{2+}$  shows a main transition at  $39.5 \text{ }^\circ\text{C}$  (Table 2). GSN and F-actin in the presence of nanomolar  $\text{Ca}^{2+}$  shows four transitions at 33, 40, 54 and  $60 \text{ }^\circ\text{C}$  and the GSN-PIP2 complex with F-actin show three transitions at 36, 56 and  $60.5 \text{ }^\circ\text{C}$ . PIP2 binding shifted the main peak from 33 to  $36 \text{ }^\circ\text{C}$ , however in case of free GSN the biggest

transition was observable at  $39.5 \text{ }^\circ\text{C}$ . Figure 1D shows that the isothermal heatflow of F-actin was dropped down (endothermic  $3.6 \mu\text{W}$ ) (Table 1) by mixing with millimolar  $\text{Ca}^{2+}$  containing buffer then it was increased by  $3.23 \mu\text{W}$  first in a quick step  $100 \pm 0.8 \times 10^{-4} \mu\text{W s}^{-1}$  then slowly  $6 \pm 1.5 \times 10^{-4} \mu\text{W s}^{-1}$  in the first hour. The previously described two exothermic peaks of GSN and F-actin under nanomolar  $\text{Ca}^{2+}$  were turned to be reduced in the presence of micromolar  $\text{Ca}^{2+}$ . Interestingly, we observed a minor drop (endothermic  $1.36 \mu\text{W}$ ) of the heatflow then it was followed by an exothermic change  $3.84 \pm 1.2 \mu\text{W}$  with the rate of  $37 \pm 0.8 \times 10^{-4} \mu\text{W s}^{-1}$ . If GSN was added to F-actin in the presence of millimolar  $\text{Ca}^{2+}$ , the heatflow showed a big drop  $9.7 \mu\text{W}$  at the beginning of the process then it was increased by  $7.7 \pm 2.8 \mu\text{W}$  with a relatively quick rate  $17 \pm 1.2 \times 10^{-4} \mu\text{W s}^{-1}$  in the first two hours. Figure 1E shows that the shape of thermal denaturation curves in each cases from panel C seems to be  $\text{Ca}^{2+}$  dependent. F-actin in the presence of millimolar  $\text{Ca}^{2+}$  shows a main transition at  $60 \text{ }^\circ\text{C}$  (Table 2). As we described above, the GSN and F-actin in the presence of nanomolar  $\text{Ca}^{2+}$  shows four transitions. However, in the presence of micromolar  $\text{Ca}^{2+}$ , there were multiple peaks at 45.5; 49; 53.5 and  $60 \text{ }^\circ\text{C}$ . Under millimolar  $\text{Ca}^{2+}$ , there was only one main transition at  $58 \text{ }^\circ\text{C}$  in case of GSN and F-actin.

## Discussion and conclusions

The aim of our present work is to show thermodynamical changes due to the activation and inactivation of gelsolin. Heat flow follows the structural entropy change as increases with structural stabilization.

Under nanomolar  $\text{Ca}^{2+}$  concentration, GSN responded with a minor isothermal change to the mixing of sample. GSN addition to F-actin resulted a strong exothermic change linked to the reduction of their structural entropy. Possibly, free actin monomers were depleted by GSN binding [62, 63] which mechanism can be described as the Type I.  $\text{Ca}^{2+}$  binding sites [67] are partially saturated, and thus initialized a structural change of GSN for a cooperative binding to actin monomer. Here we found that the actin binding leads to a lower thermodynamical stability of GSN. The PIP2 binding was stabilized the structure of GSN similar to the free form with slow and weak response to the actin addition (Fig. 2). In good agreement with our previous results, since the intrinsic tryptophan of GSN responds with increased emission to PIP2 addition, the response was a less flexible structural dynamics of GSN [49]. At the end of the isotherm assays, the majority of actin was denatured around  $60 \text{ }^\circ\text{C}$ , it can be interpreted by the hydrolyzation of ATP due to the long term processes [68].



**Fig. 1** Isothermal severing and capping of F-actin by GSN. **A** Schematic curve of isothermal heat flow change by the time with the mixing effect and the investigated parameters. **B** Time dependent change of isothermal heat flow in case of GSN with (light brown line /GF) or without (blue line /G) F-actin in the presence of nanomolar  $\text{Ca}^{2+}$ . The response of GSN-PIP2 complex in the presence of F-actin (green line /GFP). **C** All cases from panel B were terminated with

thermal denaturation DSC scans. **D** Isothermal heat flow kinetics of F-actin in the absence (cyan line /F) or in the presence of GSN under different  $\text{Ca}^{2+}$  concentrations, nanomolar (light brown line), micromolar (magenta line) or millimolar (dark brown line). **E** Heat denaturation curves of processes from panel D. The scans are the average of three measurements, and endotherm effect is deflected downwards which is valid with all panels

**Table 1** Isothermal kinetic data achieved by exponential fit to the heat flow change between the minimum and maximum value of the curves. The values are the average of three measurements

Isothermal kinetic parameters			
	$k (\times 10^{-4} \mu\text{W s}^{-1})$	$\Delta Q_{\text{initial}} / \mu\text{W}$	$\Delta Q_{\text{top}} / \mu\text{W}$
nM			
GSN	$1.7 \pm 0.3$	- 2.4	$2.2 \pm 0.5$
GSN + F-actin	$82 \pm 1.2$	- 2.2	$9.1 \pm 2.5$
GSN + F-actin + PIP2	$17 \pm 2.5$	- 1.2	$4.6 \pm 0.7$
$\mu\text{M}$			
GSN + F-actin	$37 \pm 0.8$	- 1.36	$3.84 \pm 1.2$
mM			
GSN + F-actin	$17 \pm 1.2$	- 9.7	$7.7 \pm 2.8$
F-actin	$100 \pm 0.8; 6 \pm 1.5$	- 3.6	$3.23 \pm 0.9$

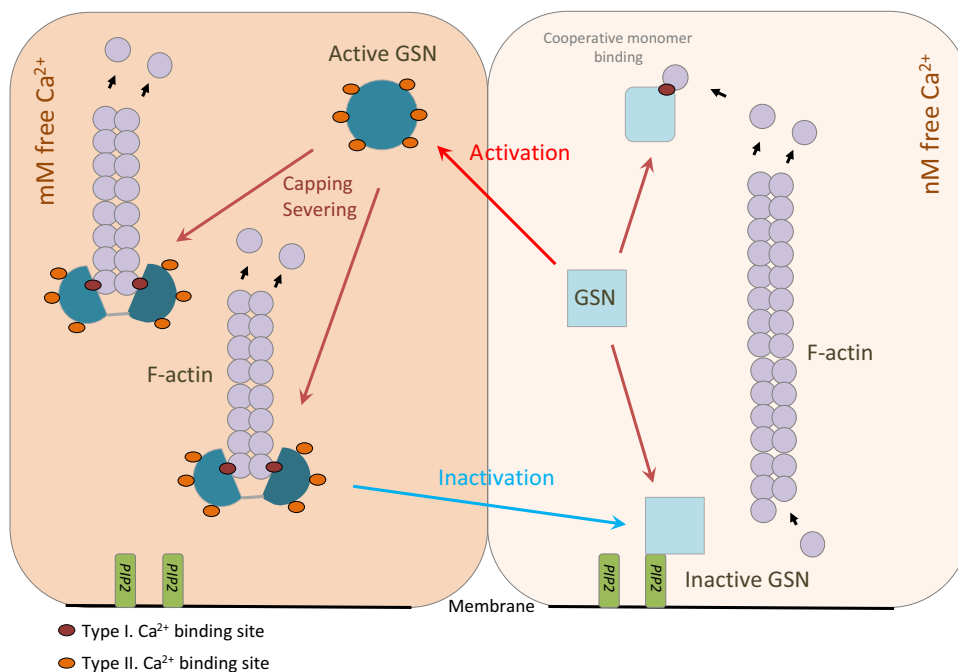
**Table 2** Heat denaturation parameters were collected after the end of the isothermal processes

Denaturation parameters		
	$T_m / ^\circ\text{C}$	$\Delta H_{\text{cal}} / \text{J g}^{-1}$
nM		
GSN	39.5	$0.058 \pm 0.0006$
GSN + F-actin	33; 40; 54; 60	$0.045 \pm 0.003$
GSN + F-actin + PIP2	36; 56; 60.5	$0.04 \pm 0.0032$
$\mu\text{M}$		
GSN + F-actin	45.5; 49; 53.5; 60	$0.054 \pm 0.0065$
mM		
GSN + F-actin	58	$0.04 \pm 0.0036$
F-actin	60	$0.038 \pm 0.0019$

Under micromolar  $\text{Ca}^{2+}$  concentration, the  $\text{Ca}^{2+}$  binding to Type II. binding sites [67] were initialized structural change of GSN and the mixing of sample can make fragments of long actin filaments and all newly exposed barbed end can be quickly capped by the partially active GSN. However, a slow structural transition in all capped actin filaments was observable.

Under millimolar  $\text{Ca}^{2+}$  concentration, the fully active GSN can do very fast severing within the mixing of the sample and induced a relatively big endotherm drop of the heat flow at the beginning of the process. Then subsequently, there was a relatively slow exotherm change which can be explained by the capping and within caused slow structural rearrangement of short filaments. However, if we mixed F-actin with buffer only it shows a very fast then a very slow isothermal process, possibly it was a recovery after certain fragmentations. As Nag et al.

**Fig. 2** Schematic model of calcium induced GSN activation in the near of the plasma membrane



published in 2009 [69] the transition temperature of GSN domains was increased from 35 to 60 °C with the change of nanomolar to millimolar  $\text{Ca}^{2+}$  concentration. Here we found that the heat denaturation of actin-GSN shows two main components nearby 30 °C and 60 °C in the presence of nanomolar, multiple peaks between 30 and 60 °C in the presence of micromolar and one main component nearby 60 °C in the presence of millimolar  $\text{Ca}^{2+}$ . What we can interpret as the GSN domains were turned to be more compact by the calcium binding which were connected together by flexible linkers seems like beads on string.

Here we have demonstrated that PIP2 binding stabilized the whole structure of gelsolin and possibly reduced its actin monomer binding activity under nanomolar calcium as the typical cytoplasmic calcium concentration of resting cells. If gelsolin-capped filaments point at the plasma membrane, the local concentration of gelsolin increases by forming PIP2 clusters [70]. All the conditions which help the binding between gelsolin and PIP2, and hence, filament uncapping in case of resting cells. The total inactivation of gelsolin is required PIP2.

We can presume that the low free calcium concentration keeps on the structure of gelsolin which is able to bind actin within the cooperativity of actin bound calcium (Fig. 2). Gelsolin can help to manage monomer pool far from the membrane and it can be linked to a basic sensory mechanism which influences the direction of filament growth in the near of the membrane.

## Materials and methods

### Proteins

His-tagged human wild-type gelsolin was expressed in *E. coli* strain Rosetta2 (DE3) pLysS cells from a pSY5 plasmid [71]. The protein was subjected to Ni-NTA affinity chromatography, HRV 3C protease cleavage, followed by gel filtration (Superdex 200, GE Healthcare) in 10 mM MOPS, 150 mM NaCl, pH 8. Traces of calcium were removed by dialysis (4 mM MOPS, 1 mM EGTA, pH 7.4, overnight). Rabbit skeletal muscle actin was prepared from acetone powder by a modified protocol from Spudich and Watt [72]. Actin was stored in buffer A (4 mM MOPS, 0.2 mM ATP (ATP disodium trihydrate, Sigma-Aldrich), 0.1 mM  $\text{CaCl}_2$ , pH 7.4). We applied 2 mM EGTA then 2 mM  $\text{MgCl}_2$ . Actin polymerization process was initialized by addition of 100 mM KCl following the same protocol as in our previous study [64]. 12.8  $\mu\text{M}$  gelsolin was incubated under physiological salt conditions (100  $\mu\text{M}$   $\text{CaCl}_2$ , 100 mM KCl, 1 mM  $\text{MgCl}_2$ , 0.2 mM ATP, 4 mM MOPS, pH 7.4) supplemented with EGTA or  $\text{CaCl}_2$  to vary the free calcium

levels (calculated with MaxChelator Stanford); for nM: 2 mM EGTA; for  $\mu\text{M}$ : 100  $\mu\text{M}$  EGTA; for mM: 1 mM  $\text{CaCl}_2$ . We used 1.6  $\mu\text{M}$  PIP2 (PtdIns-(3,5)-P2(1,2-Dihexanoyl), Cayman Chemicals) which is below the CMC of PIP2 [73–75], incubated with F-actin in the presence of 2 mM EGTA.

### DSC measurements

The actin samples were freshly prepared before all measurements. The isotherm kinetics analysis was made by a SETARAM Micro-DSCII calorimeter on room temperature, 22 °C for 16 h. After the third hour of measurements, we observed the continuous dropping of heat flow which can be described by the decay of ATP and actin filaments. We used Hastelloy designed pair of “mixing batch” vessels for isotherm measurements ( $V_{\text{lower}} = 500 \mu\text{L}$ ,  $V_{\text{upper}} = 200 \mu\text{L}$ ) in the reference vessel the GSN was mixed with buffer only and in the sample one with F-actin in the presence of nanomolar, micromolar or millimolar free  $\text{Ca}^{2+}$  (calculated by MaxChelator, <https://somapp.ucdmc.ucdavis.edu>) and after it subsequently undergone for denaturation measurements. Kinetics measurements were carried out with an isotherm program then finally denatured by heating up them with 0.3 K/min up to 80 °C. MOPS buffer was used as a reference. The reference and sample vessels were equilibrated with a precision of  $\pm 0.1$  mg; this way we did not need to do any correction between vessels' heat capacity. The first point of our analysis was the minimum of the heat flow change ( $\Delta Q_{\text{initial}}$ ) after a single sharp peak which shows the mixing effect, and thus the curve was increased slowly to the maximum point ( $\Delta Q_{\text{top}}$ ) (Fig. 1A). The exponential fitting resulted the time rate ( $k$ ) of the peaks, the measurements were started in 5 min after reaching the thermal equilibrium between the vessels, but we used for the analysis the part of the curve after the mixing effect. Actin and gelsolin based molecular events were happened only in the first three hours then all curves decreased slowly possibly by the decay of ATP thus the most reasonable time duration of kinetics was analyzed in the first 10,800 s. All kinetic processes were terminated with heat denaturation assays. With the help of a two-point SETARAM peak integration setting, the denaturation calorimetric enthalpy was calculated from the area under the heat absorption curve, and then, the results [denaturation or melting temperature ( $T_m$ ) and calorimetric enthalpy ( $\Delta H_{\text{cal}}$ ) data of samples] were compared. This method is identical with the protocol as we applied in our previous study [64].

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**Author's contribution** DS: rising the problem, sample preparation and handling, data analysis, manuscript writing. DL: corresponding author, principal investigator, DSC experiments, data analysis, manuscript writing.

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**Data availability** There are no additional available data to upload.

## Declarations

**Conflict of interest** The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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