



# Thermal, spectroscopic, X-ray and theoretical studies of metal complexes (sodium, manganese, copper, nickel, cobalt and zinc) with pyrimidine-5-carboxylic and pyrimidine-2-carboxylic acids

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## Abstract

The new 3d metal complexes of pyrimidine-2-carboxylic (2PCA) and pyrimidine-5-carboxylic (5PCA) acids were synthesized and characterized using thermal analysis (TG–DSC, TG–FTIR), X-ray, spectroscopic (IR, Raman) methods and theoretical (DFT) studies. In the complexes of pyrimidine-2-carboxylic acid of the general formula  $M(2PCA)_2 \cdot xH_2O$  (where 2PCA—pyrimidine-2-carboxylate; M = Mn(II), Co(II), Ni(II), Cu(II) and Zn;  $x = 0$  for Mn and Cu;  $x = 2$  for Co, Ni and Zn) coordination of metal ions occurs through nitrogen atom from pyrimidine ring and carboxylate oxygen atom. The complexes of pyrimidine-5-carboxylic acid of the general formula  $M(5PCA)_2 \cdot xH_2O$  (where 5PCA—pyrimidine-5-carboxylate; M = Mn(II), Co(II), Ni(II), Cu(II) and Zn;  $x = 6$  for Cu and 4 for remaining complexes) were obtained as monomeric isostructural compounds. Coordination of metal centers occurs through two nitrogen atom from different pyrimidine-5-carboxylate ligand and four oxygen atoms from water molecules. The IR and Raman spectra of free acids as well as obtained metal(II) complexes were described in detail. Aromaticity (HOMA, EN, GEO and  $I_6$ ) of complexes was determined and discussed. The investigated compounds decompose in air in two main stages connected with dehydration and decomposition/burning of anhydrous compounds to the suitable metal oxides. Thermal decomposition in nitrogen leads to the evolution of water, carbon oxides, ammonia and pyrimidine molecules.

**Keywords** Pyrimidine-2-carboxylic acid · Pyrimidine-5-carboxylic acid · 3d metal complexes · Aromaticity indices · Thermal analysis

## Introduction

Complexes of transition metals (Co, Mn, Ni, Zn, Cu) play an important role in the life processes of living organisms. They are, among others, components of enzymes, proteins and vitamins; they participate in oxidation–reduction

processes or in binding, storage and transport of oxygen [1]. Many years ago, it was noticed that some complex compounds of transition metals, including copper and zinc, may react with the superoxide anion radical in a manner reminiscent of the activity of native superoxide dismutases. Hence, they were called superoxide dismutase mimetics SOD [2]. Ligands that are heterocyclic nitrogen bases have a significant effect on increasing the reactivity of copper(II) complexes with superoxide anion radicals. Although copper(II) complexes show a weaker antioxidant effect than the SOD enzyme, they are important as potential superoxide dismutase mimetics due to the low molecular weight [3, 4]. Complexes of transition metals have also been used as medicines. In many cases, the ligand responsible for the therapeutic effect of the drug may increase efficacy action after the complexation by metal. Metal ions in complexes with antitumor active ligands play the role as modulators of activity [5]. The metal affects the redox activity of the

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ligand, its mode of action, therapeutic efficacy, solubility, stability and a number of other properties [6]. Complexation by the metals effects on the change in the distribution of electronic charge in the ligand, which translates into a change in the physicochemical and biological properties of the chemical compound [7–10]. Studies have shown that transition metals influence the stability of the electron system of complexes the aromatic acids [11–14]. This impact depends not only on the properties of metal, but also on the way the ligand–metal coordination. In complexes of heteroaromatic acids, the metal may be attached via a carboxyl anion and heteroatoms of the aromatic ring. These types of binding are formed inter alia in the pyridine-2-carboxylic, pyrazine-2-carboxylic or pyrimidine-2-carboxylic acid complexes [15–18]. The pyrimidine-5-carboxylic acid may form complexes by attaching a metal via a ring nitrogen atom without the participation of a carboxylic group [19].

In the pyridinecarboxylic acid complexes, various types of metal–ligand coordination depending on the position of the nitrogen atom relative to the carboxylic group were observed [20–28]. In the case of 2-pyridinecarboxylic acid (picolinic acid) in which the nitrogen atom is located in close proximity to the carboxyl group, the metal is coordinated through the carboxyl group and the nitrogen atom to form a chelate ring [20–22]. In the case of nicotinic acid, a type of coordination through the carboxylic group was observed [23–25]. In the case of isonicotinic acid, the formation of complexes with coordination of metal–nitrogen atom and metal–carboxylic group was noticed [26–28]. It was found that the method of coordination affects the electronic structure of the aromatic system as well as the thermal properties of the complex [29, 30].

Zinc complexes with pyridinecarboxylic acids (nicotinic, picolinic, isonicotinic) show different thermal stability. Studies have shown that the complex with picolinic acid (coordination O, N) is less thermally stable than complexes with nicotinic acid and isonicotinic acid [29, 30].

In the present work, complexes of pyrimidinecarboxylic acids were investigated. Carboxylic acids derived from diazines (Fig. 1) due to the number of potential chelating sites can form various coordination structures in relationships with metal ions [18, 19, 31, 32]. The pyrimidine derivatives are found in living organisms where they play a fundamental role in many life processes. The most important naturally occurring pyrimidines are the pyrimidine uracil bases, thymine and cytosine, which are the basic building blocks of nucleosides in deoxyribonucleic acid (DNA) and ribonucleic acid (RNA).

Many of the pyrimidine derivatives show a wide range of biological activities and are known for their anticancer (fluorouracil) [33], antimicrobial, analgesic, antiviral (also

in the treatment of HIV, hepatitis B), antimalarial and anesthetic [34] properties. Some of the pyrimidine derivatives have been used as plant protection products [35]. Orotic acid is commonly known as a precursor in pyrimidine biosynthesis [36]. As a component of the daily diet is found in milk and other dairy products. In the body, it is transformed into uridine, which then participates in the pyrimidine recovery pathway in the kidneys, liver and erythrocytes [37].

In order to improve the bioavailability of orotic acid, its combinations with organic cations (e.g., choline, carnitine) or metal ions ( $Mg^{2+}$ ,  $Zn^{2+}$ ,  $Ca^{2+}$ ,  $K^+$ ) are used. Such combinations may be used in the therapy of metabolic syndromes and in bodybuilding [38].

The aim of the work was to obtain complexes of transition metals 3d with carboxylic acid pyrimidine derivatives. A series of complexes with ligands having different carboxylic acid positions relative to the nitrogen atoms in the aromatic ring (pyrimidine-2-carboxylic acid and pyrimidine-5-carboxylic acid) were synthesized. The influence of the metal on the electron system of the ligand and the thermal properties of the obtained complexes depending on the type of metal–ligand coordination were evaluated.

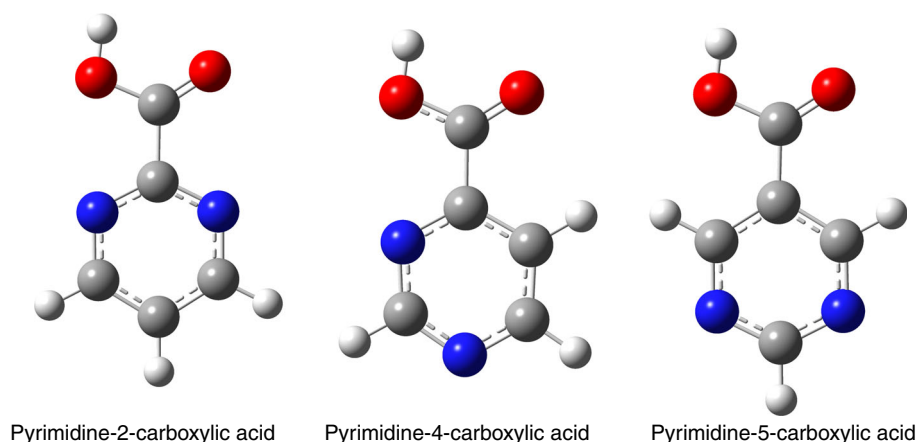
## Experimental

### Synthesis of complexes

The complexes of manganese(II), cobalt(II), nickel(II), zinc and copper(II) with pyrimidine-2-carboxylic (2PCA) acid complexes were prepared as follows: 49.6 mg of acid was added to the stoichiometric quantity of sodium hydroxide ( $4\text{ cm}^3$ ;  $0.1\text{ mol dm}^{-3}$ ). The suspension was placed in an ultrasonic bath at  $70\text{ }^\circ\text{C}$  until the acid dissolved. Then,  $2\text{ cm}^3$  of metal(II) chloride solution ( $0.1\text{ mol dm}^{-3}$ ) was added. The solution was shaken on a shaker for 1 h at room temperature, and the complexes were allowed to stand. The obtained precipitates/or crystals were washed with water to rinse the chlorides under the control of  $AgNO_3$  solution. Analogously, metal(II) pyrimidine-5-carboxylate complexes were obtained.

### Elemental analysis

The elementary analysis was carried out using the CHN 2400 PerkinElmer Analyzer. The water content was determined from the thermogravimetric curves. Table 1 presents the results of elemental analysis of pyrimidine-2-carboxylic acid complexes, and Table 2 presents the results of elemental analysis for pyrimidine-5-carboxylic acid complexes.

**Fig. 1** Structures of pyrimidinecarboxylic acids**Table 1** Elemental analysis for pyrimidine-2-carboxylates

Compound name	Formula	Content C/%		Content H/%		Content N/%		Yield/%	Color
		Calc.	Exp.	Calc.	Exp.	Calc.	Exp.		
Manganese pyrimidine-2-carboxylate	Mn(2PCA) <sub>2</sub>	39.89	39.32	2.01	1.99	18.61	18.14	80	Yellow
Cobalt pyrimidine-2-carboxylate	Co(2PCA) <sub>2</sub> ·H <sub>2</sub> O	35.21	35.42	2.95	2.97	16.42	16.14	75	Orange
Nickel pyrimidine-2-carboxylate	Ni(2PCA) <sub>2</sub> ·2H <sub>2</sub> O	35.23	35.11	2.96	2.83	16.43	16.18	70–80	Green
Zinc pyrimidine-2-carboxylate	Zn(2PCA) <sub>2</sub> ·2H <sub>2</sub> O	34.55	34.51	2.96	2.81	16.43	16.27	75	White
Copper pyrimidine-2-carboxylate	Cu(2PCA) <sub>2</sub>	38.78	38.63	1.95	1.91	18.09	18.16	70	Blue

2PCA = C<sub>5</sub>H<sub>3</sub>N<sub>2</sub>O<sub>2</sub>**Table 2** Elemental analysis for pyrimidine-5-carboxylates

Compound name	Formula	Content C/%		Content H/%		Content N/%		Yield/%	Color
		Calc.	Exp.	Calc.	Exp.	Calc.	Exp.		
Manganese pyrimidine-5-carboxylate	Mn(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	32.18	32.20	3.78	3.69	15.01	14.55	80	Yellow
Cobalt pyrimidine-5-carboxylate	Co(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	31.84	31.51	3.74	3.68	14.85	14.44	75–80	Orange
Nickel pyrimidine-5-carboxylate	Ni(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	31.86	30.84	3.74	3.77	14.86	14.02	80	Green
Zinc pyrimidine-5-carboxylate	Zn(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	31.31	31.42	3.68	3.55	14.60	14.34	85	White
Copper pyrimidine-5-carboxylate	Cu(5PCA) <sub>2</sub> ·6H <sub>2</sub> O	28.75	28.57	4.34	4.25	13.41	13.30	70–80	Blue

5PCA = C<sub>5</sub>H<sub>3</sub>N<sub>2</sub>O<sub>2</sub>

## Thermal analysis

Thermal analyses of the investigated complexes were carried out by the thermogravimetric (TG) and differential scanning calorimetry (DSC) methods using the Setsys 16/18 analyzer (Setaram). The samples (about 5–8 mg) were heated in alumina crucibles in the range of 30–1000 °C at a heating rate of 10 °C min<sup>-1</sup> in flowing air atmosphere ( $v = 0.75 \text{ dm}^3 \text{ h}^{-1}$ ). The TG-FTIR coupled measurements performed on Q5000 (TA) apparatus coupled with the Nicolet 6700 spectrophotometer. The samples of about 20–30 mg were heated in platinum crucibles from

room temperature up to 700 °C at a heating rate of 20 °C min<sup>-1</sup> in flowing nitrogen atmosphere (25 cm<sup>3</sup> min<sup>-1</sup>).

## FTIR and Raman study

The FTIR spectra were recorded with an Alfa (Bruker) spectrometer within the range of 400–4000 cm<sup>-1</sup>. Samples in the solid state were measured in KBr matrix pellets and ATR technique. FT-Raman spectra of solid samples were recorded in the range of 400–4000 cm<sup>-1</sup> with a MultiRam (Bruker) spectrometer.

## Theoretical calculations

To calculate optimized geometrical structures of pyrimidinecarboxylic acids and copper(II) complexes, the quantum-mechanical method was used: density functional (DFT) hybrid method B3LYP with non-local correlation provided by Lee-Young-Parr expression. All calculations for acids were carried out in functional base 6-311++G(d,p) and for acids and copper complexes in LANL2DZ base. Calculations were performed using the Gaussian09 package [39]. Experimental spectra were interpreted in terms of calculated at DFT method in B3LYP/6-311++G(d,p) level and literature data [40, 41]. Theoretical wavenumbers were scaled according to the formula:  $v_{\text{scaled}} = 0.98 \cdot v_{\text{calculated}}$  for B3LYP/6-311++G(d,p) level method. Spectral assignments used the normal oscillation of the pyrimidine ring measured using B3LYP/6-311++G(d,p).

The aromaticity indices (HOMA, GEO, EN,  $I_6$ ) were calculated for geometric structures (theoretical and calculated) of pyrimidine-2-carboxylates and pyrimidine-5-carboxylates. The HOMA index (harmonic oscillator model of aromaticity) differs from all other geometry-based ones by assuming another reference bond length. In this model, instead of the mean bond length a concept of the optimal bond length is applied [42, 43]:

$$\begin{aligned} \text{HOMA} &= 1 - \left[ \alpha (R_{\text{opt}} - R_{\text{av}})^2 + \frac{\alpha}{n} \sum (R_{\text{av}} - R_i)^2 \right] \\ &= 1 - \text{EN} - \text{GEO} \end{aligned}$$

within the confines of the HOMA model, it is possible to obtain two components which describe different contribution to decrease in aromaticity, i.e., (a) due to bond elongation (the EN component) and (b) due to bond length alternation (the GEO component). The value of HOMA index is equal 1 for the entire aromatic system; HOMA = 0 when the structure is non-aromatic and HOMA < 0 for anti-aromatic ring.

The value of the Bird's aromaticity index ( $I_5$ ,  $I_6$ ) describes the equation [44]:

$$I = 100 \{ 1 - (V/V_k) \}$$

where  $V_k$  is for the five-membered rings 35 and the six-membered 33.3, and  $V$  is calculated from the equation:

$$V = (100/n_{\text{av}}) \left[ \sum_{r=1}^n (n_r - n_{\text{av}})^2 / n \right]^{\frac{1}{2}}$$

where  $n_{\text{av}}$ -average binding order,  $n$ -bond order based on bond length:  $n = (a/R) - b$ ,  $a$  and  $b$ -parameters depending on the type of atoms in the bond.

## SOD activity and HOMO, LUMO energies

SOD activity for copper(II) complexes was determined theoretically on the basis of the energy calculations for geometrically optimized structures. As a theoretical parameter defining the radical scavenging activity by the studied complexes, HOMO and LUMO energy values and EA affinity value were selected:

TEE-Cu(I)-TEE-Cu(II), where TE E-Cu(I) is the total energy of the complex E-Cu(I) and TEE-Cu(II) with the total energy of the E-Cu(II) complex.

## X-ray analysis

Suitable crystals were sequentially mounted on a fiber loop and used for X-ray measurements. X-ray data were collected on the Oxford Diffraction SuperNova DualSource diffractometer with use of the monochromated CuK $\alpha$  X-ray source ( $\lambda = 1.54184$ ). The crystals were kept at 100 K during data collection. Data reduction and analytical absorption correction were performed with CrysAlis PRO [45]. Using Olex2 [46], the structures were solved with the ShelXS [47] structure solution program using Direct Methods and refined with the ShelXL [47] refinement package using Least Squares minimization.

The non-hydrogen atoms were refined anisotropically. The positions of OH hydrogen atoms were found on a Fourier difference map and refined. Hydrogen atoms of aromatic ring were introduced in calculated positions with idealized geometry and constrained using a rigid body model with isotropic displacement parameters equal to 1.2 of the equivalent displacement parameters of the parent atoms. Summaries of relevant crystallographic data are given in Tables 3 and 6 for pyrimidine-2-carboxylate and pyrimidine-5-carboxylate complexes. The supplementary crystallographic data for this paper can be obtained free of charge from The Cambridge Crystallographic Data Centre via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif).

## Results

### Crystal structure description

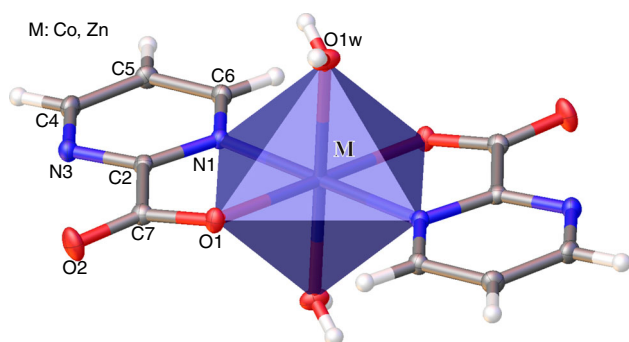
Isostructural copper(II) and zinc(II) pyrimidine-2-carboxylates M(2PCA) $_2 \cdot 2\text{H}_2\text{O}$  crystallize in the monoclinic  $P2_1/c$  space group (Table 3). The asymmetric unit is a half of the complex molecule, and the metal cations are located on a crystallographic inversion center. The cobalt and zinc metal cations are coordinated identically by two pyrimidine-2-carboxylate ligands in the equatorial plane and by two water molecules in the axial positions. The pyrimidine-

**Table 3** Crystal data and structure refinement of pyrimidine-2-carboxylates

Empirical formula	Co(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·2H <sub>2</sub> O	Zn(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·2H <sub>2</sub> O
Formula weight	341.15	347.59
Crystal system	Monoclinic	Monoclinic
Space group	<i>P</i> 2 <sub>1</sub> / <i>c</i>	<i>P</i> 2 <sub>1</sub> / <i>c</i>
<i>a</i> /Å	5.201(7)	5.234(1)
<i>b</i> /Å	11.713(2)	11.726(3)
<i>c</i> /Å	10.041(7)	9.793(8)
$\alpha$ /°	90	90
$\beta$ /°	96.448(4)	97.8387(12)
$\gamma$ /°	90	90
Volume/Å <sup>3</sup>	607.9(5)	595.4(8)
<i>Z</i>	2	2
$\rho_{\text{calc}}$ /g cm <sup>-3</sup>	1.864	1.939
Crystal size/mm <sup>3</sup>	0.312 × 0.091 × 0.091	0.317 × 0.153 × 0.109
Final <i>R</i> indexes [ <i>I</i> ≥ 2σ ( <i>I</i> )]	<i>R</i> <sub>1</sub> = 0.0308, <i>wR</i> <sub>2</sub> = 0.0794	<i>R</i> <sub>1</sub> = 0.0268, <i>wR</i> <sub>2</sub> = 0.0705
Final <i>R</i> indexes [all data]	<i>R</i> <sub>1</sub> = 0.0367, <i>wR</i> <sub>2</sub> = 0.0824	<i>R</i> <sub>1</sub> = 0.0272, <i>wR</i> <sub>2</sub> = 0.0708
Largest diff. peak/hole/e Å <sup>-3</sup>	0.25/− 0.56	0.33/− 0.70
CCDC no.	1,577,381	1,857,354

2-carboxylate ligands are bound to metal cations in a bidentate *N,O*-chelate mode, forming five-membered rings (Fig. 2). The bonds of the coordination sphere M–N1, M–O1 and M–O<sub>w</sub> are nearly equal in length (2.066(1)–2.123(2) Å), and the overall coordination sphere around cobalt and zinc cations can be described as an octahedral. However, the analysis of the valence angles O–M–O<sub>w</sub> and O–M–N reveal a small distortion of the coordination spheres, since the N1–Co–O1<sub>w</sub> and N1–Zn–O1 valence angles are not a right angles and equal 79.5(5) and 79.9(6) degrees, respectively (Table 4).

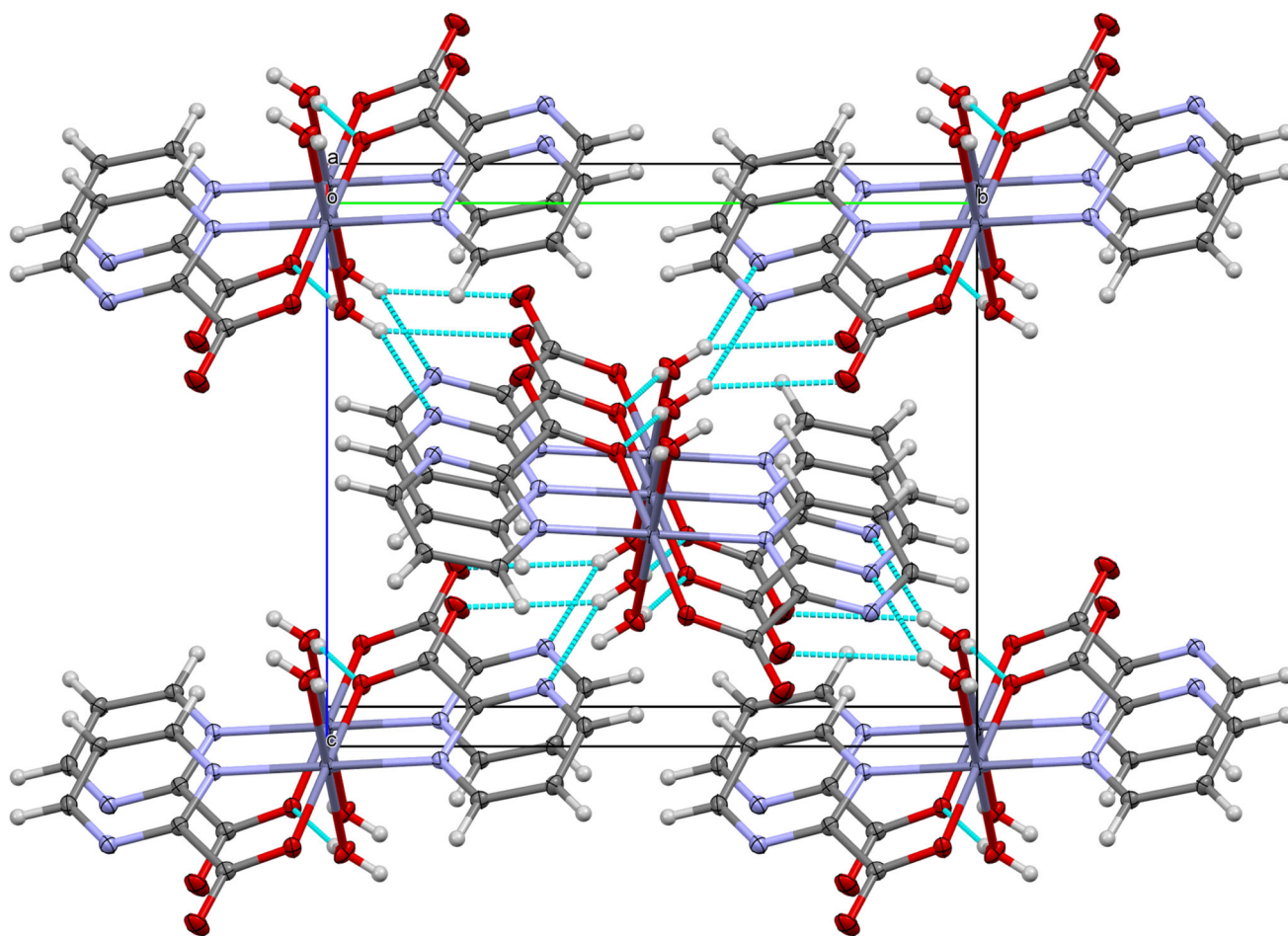
The hydrogen bonds patterns are marked in Fig. 3 as a dashed line. There are three intermolecular hydrogen bonds in the crystal structure of Co and Zn pyrimidine-2-carboxylate complexes. Two of them are O–H···O type and one is O–H···N type. Coordinated water molecules acts as a proton donor, while O1 (coordinated), O2 (uncoordinated) and N3 (uncoordinated) atoms play a role of the proton

**Fig. 2** The molecular structure of the Co(II) and Zn pyrimidine-2-carboxylates complexes**Table 4** Selected geometrical parameters of cobalt (II) and zinc pyrimidine-2-carboxylates complexes

Å, °	Co	Zn
N1–C2	1.337(2)	1.338(2)
C2–N3	1.324(2)	1.328(2)
N3–C4	1.337(3)	1.341(2)
C4–C5	1.381(3)	1.387(2)
C5–C6	1.380(3)	1.384(2)
C6–N1	1.335(3)	1.337(2)
C5–C7	1.528(3)	1.533(2)
C7–O1	1.275(2)	1.280(2)
C7–O2	1.222(3)	1.228(2)
M–O1	2.066(1)	2.082(2)
M–N1	2.114(1)	2.103(5)
M–O1 <sub>w</sub>	2.094(1)	2.123(2)
N1–C2–N3	125.8(2)	125.5(1)
C2–N3–C4	116.2(2)	116.4(1)
N3–C4–C5	122.4(2)	122.4(1)
C4–C5–C6	117.1(2)	116.8(1)
C5–C6–N1	121.1(2)	121.1(1)
C6–N1–C2	117.4(2)	117.7(1)
C2–C7–O1	115.4(2)	115.3(1)
C2–C7–O2	118.5(2)	118.0(1)
O1–M–O1 <sub>w</sub>	91.8(6)	87.3(5)
O1–M–N1	91.4(7)	79.9(6)
O1 <sub>w</sub> –M–N1	79.5(5)	87.1(4)

acceptors in hydrogen bonding. It is worth to emphasize that the hydrogen atom H1WB is a donor in a bifurcated (three-centered) hydrogen bond, where O2 and N3 atoms





**Fig. 3** Crystal structure of the Co(II) and Zn pyrimidine-2-carboxylates complexes. Hydrogen bonds are presented as a dashed lines

of the neighboring ligand molecule are acceptor centers. Selected geometrical parameters of the hydrogen bonds together with the symmetry codes (symm. op. 2) are presented in Table 5.

The complexes of manganese(II), cobalt(II) and zinc with pyrimidine-5-carboxylic acid of the formula  $M(5PCA)_2 \cdot 4H_2O$  crystallize in the monoclinic  $P2_1/c$  space group (Table 6). The crystal structure of copper complex  $M(5PCA)_2 \cdot 6H_2O$  contains an additional uncoordinated

water molecule, and the symmetry of the structure is significantly lower ( $P\bar{1}$ ).

However, in all cases the asymmetric unit is a half of the complex molecule and the metal cations are located on a crystallographic inversion center. The metal cations are coordinated identically by two pyrimidine-5-carboxylate ligands and two water molecules in the equatorial plane and by two water molecules in the axial positions. The pyrimidine-5-carboxylate ligands are bound to metal

**Table 5** Geometrical parameters of the hydrogen bonds of cobalt (II) and zinc pyrimidine-2-carboxylates complexes

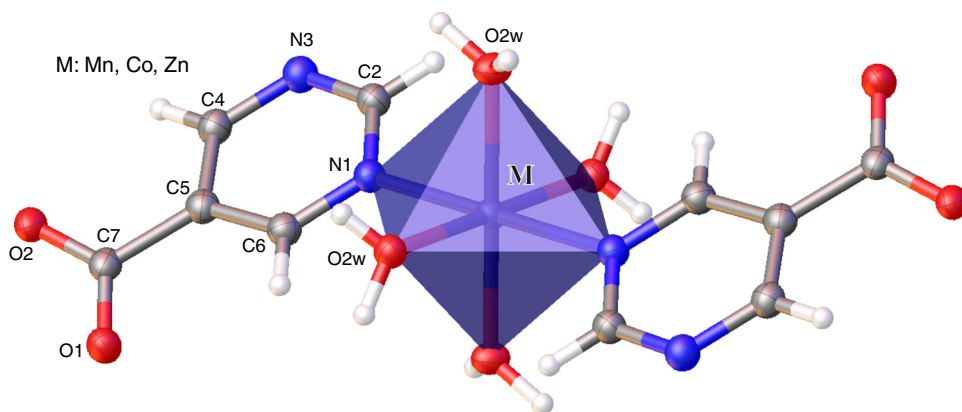
	D	H	A	d(D–H)/Å	d(H–A)/Å	d(D–A)/Å	D–H–A/°	Symm. op. 2
Co	O1 W	H1 WA	O1	0.79(4)	1.952	2.734(2)	169(3)	$1 - x, 1 - y, 1 - z$
	O1 W	H1 WB	O2	0.72(5)	2.601	3.229(3)	147(4)	$x, 1/2 - y, -1/2 + z$
	O1 W	H1 WB	N3	0.72(5)	2.272	2.883(2)	143(4)	$x, 1/2 - y, -1/2 + z$
Zn	O1 W	H1 WA	O1	0.78(3)	1.978	2.737(2)	164(3)	$-1 + x, y, z$
	O1 W	H1 WB	O2	0.74(3)	2.597	3.250(2)	147(3)	$1 - x, -1/2 + y, 1.5 - z$
	O1 W	H1 WB	N3	0.74(3)	2.236	2.858(2)	142(3)	$1 - x, -1/2 + y, 1.5 - z$

**Table 6** Crystal data and structure refinement of pyrimidine-5-carboxylates

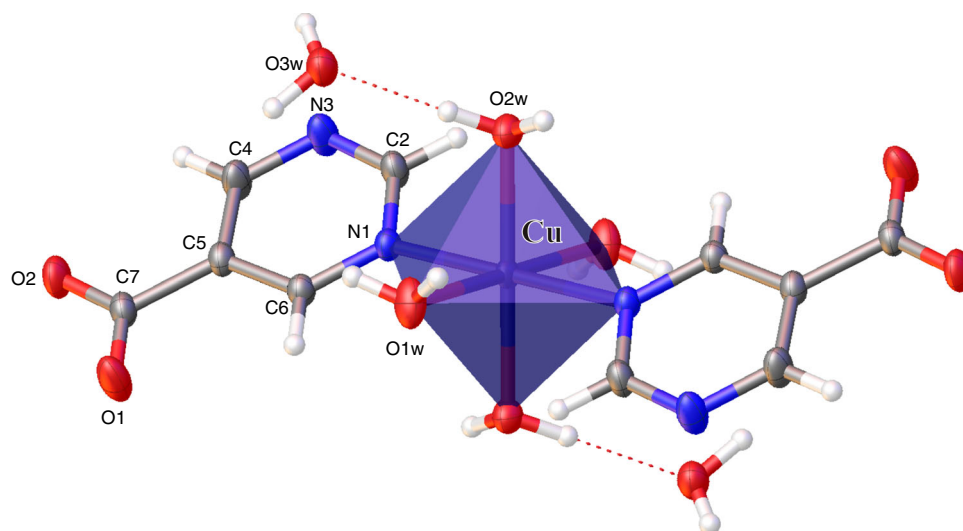
Empirical formula	Mn(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·4H <sub>2</sub> O	Co(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·4H <sub>2</sub> O	Zn(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·4H <sub>2</sub> O	Cu(C <sub>5</sub> H <sub>4</sub> N <sub>2</sub> O <sub>2</sub> ) <sub>2</sub> ·4H <sub>2</sub> O, 2H <sub>2</sub> O
Formula weight	373.19	377.18	383.62	417.82
Crystal system	Monoclinic	Monoclinic	Monoclinic	Triclinic
Space group	<i>P</i> 2 <sub>1</sub> / <i>c</i>	<i>P</i> 2 <sub>1</sub> / <i>c</i>	<i>P</i> 2 <sub>1</sub> / <i>c</i>	<i>P</i> – 1
<i>a</i> /Å	6.520(2)	6.514(6)	6.526(8)	6.189(1)
<i>b</i> /Å	12.455(3)	12.456(1)	12.470(5)	6.764(6)
<i>c</i> /Å	8.510(1)	8.508(3)	8.525(8)	9.521(2)
$\alpha$ /°	90	90	90	105.450(5)
$\beta$ /°	98.089(6)	98.12(2)	98.03(5)	92.93(7)
$\gamma$ /°	90	90	90	96.02(2)
Volume/Å <sup>3</sup>	684.24(2)	683.50(3)	687.12(4)	380.79(4)
<i>Z</i>	2	2	2	1
$\rho_{\text{calc}}$ /g cm <sup>-3</sup>	1.811	1.833	1.854	1.822
Crystal size/mm <sup>3</sup>	0.179 × 0.107 × 0.092	0.305 × 0.163 × 0.122	0.259 × 0.182 × 0.182	0.631 × 0.303 × 0.216
Final <i>R</i> indexes [ <i>I</i> ≥ 2σ( <i>I</i> )]	<i>R</i> <sub>1</sub> = 0.0364, <i>wR</i> <sub>2</sub> = 0.0966	<i>R</i> <sub>1</sub> = 0.0526, <i>wR</i> <sub>2</sub> = 0.1358	<i>R</i> <sub>1</sub> = 0.0282, <i>wR</i> <sub>2</sub> = 0.0718	<i>R</i> <sub>1</sub> = 0.0340, <i>wR</i> <sub>2</sub> = 0.0953
Final <i>R</i> indexes [all data]	<i>R</i> <sub>1</sub> = 0.0382, <i>wR</i> <sub>2</sub> = 0.0981	<i>R</i> <sub>1</sub> = 0.0583, <i>wR</i> <sub>2</sub> = 0.1451	<i>R</i> <sub>1</sub> = 0.0287, <i>wR</i> <sub>2</sub> = 0.0722	<i>R</i> <sub>1</sub> = 0.0387, <i>wR</i> <sub>2</sub> = 0.0970
Largest diff. peak/hole/e Å <sup>-3</sup>	0.39/– 0.57	0.81/– 1.07	0.32/– 0.78	0.40/– 0.69
CCDC no.	1,577,382	1,577,386	1,577,385	1,577,383

cations in a monodentate mode to the N1 nitrogen atom of the aromatic ring (Figs. 4 and 5). The analysis of the valence angles O1w–M–O2w, O1w–M–N1 and O1w–M–N1 does not indicate a distortion of the coordination sphere, since all three angles are close to the right angle. Additionally, the bond lengths: M–N1, M–O1w and M–O2w are nearly equal. Therefore, the coordination sphere can be described as an octahedral. However, in the case of the copper complex, the significant elongation of the Cu–O2w bonds is observed (Table 7). This is consistent with the Jahn–Teller effect (JTE) typical for d<sup>9</sup> electron configuration.

The hydrogen bond patterns for aforementioned pyrimidine-5-carboxylate complexes are marked in Fig. 6 for structures containing cobalt, zinc and manganese cations and in Fig. 7 for the structure of the copper cation separately. There are four intermolecular hydrogen bonds in the crystal packing of the Co, Zn and Mn pyrimidine-5-carboxylate complexes. Three of them are O–H···O type and one is O–H···N type. In case of the Mn, Co and Zn complexes, coordinated water molecules act as a proton donor, while O1 (uncoordinated), O2 (uncoordinated) and N3 (uncoordinated) atoms play a role of the proton acceptors in hydrogen bonding. However, due to the presence of an additional water molecule in the crystal structure of the Cu

**Fig. 4** The molecular structure of the Mn(II), Co(II) and Zn pyrimidine-5-carboxylates complexes

**Fig. 5** The molecular structure of the Cu(II) pyrimidine-5-carboxylates complex. The hydrogen bonds of (*x*, *y*, *z*) symmetry are presented as a dashed lines



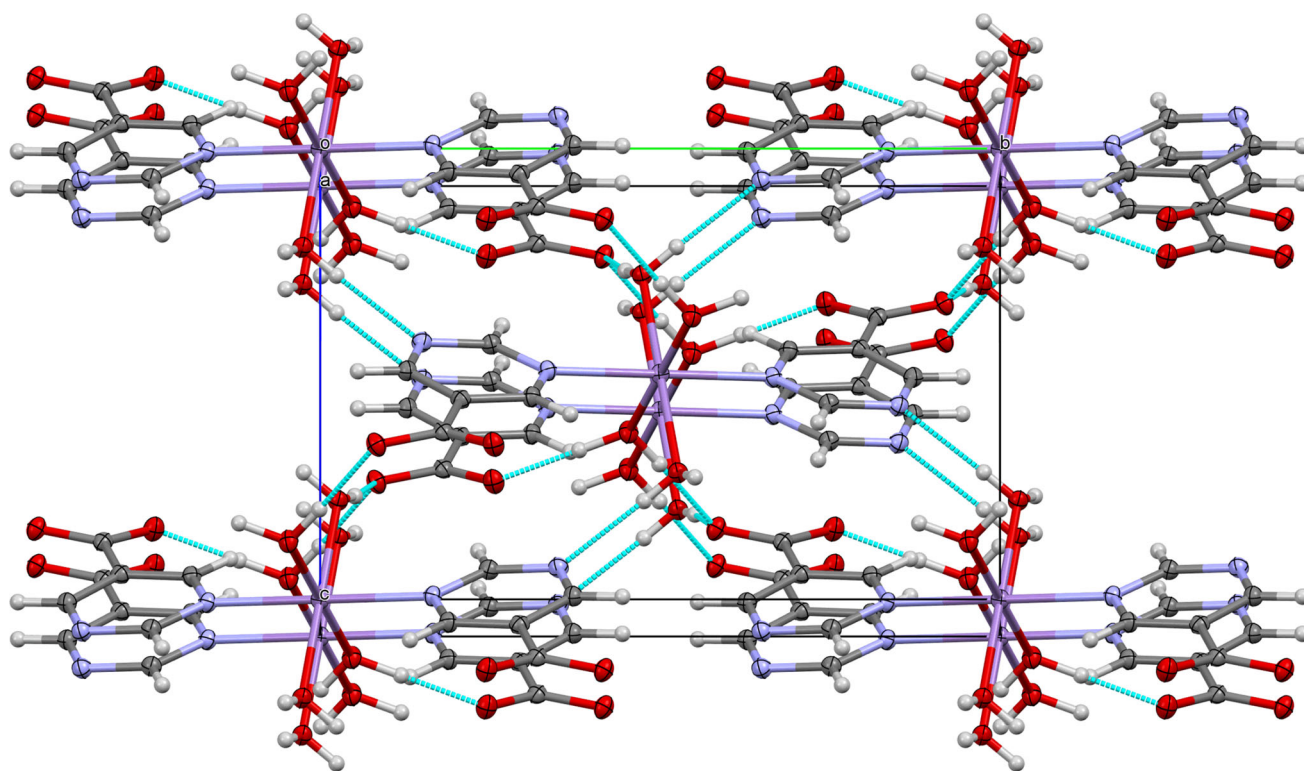
**Table 7** Selected geometrical parameters of manganese (II), cobalt (II), zinc and copper (II) pyrimidine-5-carboxylates complexes

Å, °	Co	Zn	Mn	Cu
N1–C2	1.344(3)	1.343(2)	1.348(3)	1.338(3)
C2–N3	1.331(3)	1.334(2)	1.333(3)	1.323(3)
N3–C4	1.346(4)	1.345(2)	1.342(3)	1.334(3)
C4–C5	1.388(4)	1.389(2)	1.386(4)	1.380(4)
C5–C6	1.389(3)	1.386(2)	1.387(3)	1.389(3)
C6–N1	1.337(3)	1.345(2)	1.342(3)	1.339(3)
C5–C7	1.511(3)	1.513(2)	1.512(3)	1.509(3)
C7–O1	1.247(3)	1.247(2)	1.246(3)	1.267(3)
C7–O2	1.259(3)	1.264(2)	1.262(3)	1.230(3)
M–N1	2.185(2)	2.194(1)	2.182(3)	2.036(5)
M–O1w	2.055(1)	2.079(5)	2.076(6)	2.000(8)
M–O2w	2.077(2)	2.070(8)	2.059(9)	2.345(5)
N1–C2–N3	125.7(2)	125.6(1)	125.6(2)	125.8(2)
C2–N3–C4	116.7(2)	116.6(1)	116.7(2)	116.0(2)
N3–C4–C5	122.3(2)	122.5(1)	122.6(2)	123.2(2)
C4–C5–C6	116.3(2)	116.2(1)	116.3(2)	116.6(2)
C5–C6–N1	122.4(2)	122.2(1)	122.4(2)	120.8(2)
C6–N1–C2	116.7(2)	116.8(1)	116.5(2)	117.6(2)
C5–C7–O1	118.7(2)	118.8(1)	118.9(2)	118.2(2)
C5–C7–O2	116.2(2)	115.9(1)	116.1(2)	116.3(2)
O1 W–M–O2 W	88.4(5)	88.8(4)	88.5(3)	88.1(9)
O1 W–M–N1	89.2(7)	91.3(2)	91.1(5)	91.8(3)
O2 W–M–N2	88.90(8)	90.9(1)	90.8(6)	87.0(6)

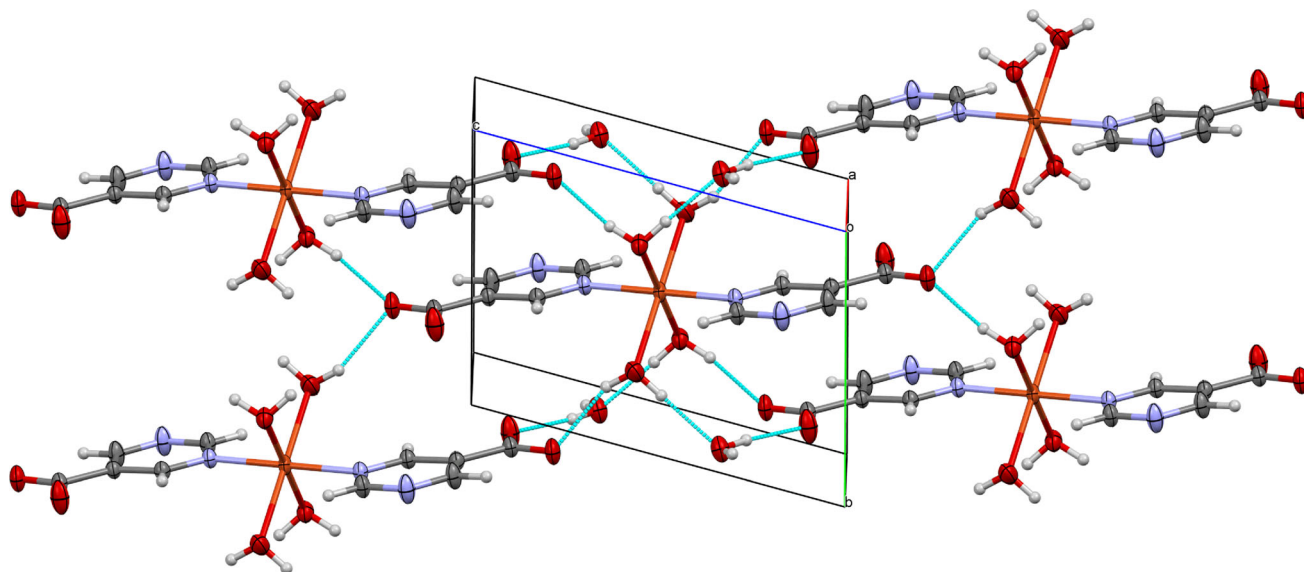
pyrimidine-5-carboxylate complex, the hydrogen bonds pattern differs. Six intermolecular hydrogen bonds of different symmetry codes can be observed here. Similarly to the structures described above, the water molecules play a role of the proton donors. However, all hydrogen bonds are O–H···O type only. The coordinated water molecules are non-covalently bonded to a two O1 atoms of the two different ligand molecules and to a two different O3W oxygen

atoms of the uncoordinated water molecules. The uncoordinated water molecule is a donor in a two additional hydrogen bonds O3W–H3WA···O2 and O3W–H3WB···O1, where O1 and O2 atoms are the oxygen atoms of different molecules. Selected geometrical parameters of the hydrogen bonds together with the symmetry codes (symm. op. 2) are presented in Table 8.





**Fig. 6** Crystal structure of the Mn(II), Co(II) and Zn pyrimidine-5-carboxylates complexes. Hydrogen bonds are presented as a dashed lines



**Fig. 7** Crystal structure of the Cu(II) pyrimidine-5-carboxylates complex. Hydrogen bonds are presented as a dashed lines

### Aromaticity

The aromaticity indexes (Homa and  $I_6$  (Bird index) based on the length of bonds in the aromatic ring allow to assess the effect of the metal on the electronic system of the molecule. The values of the indexes of aromaticity have been calculated based on the values of the length of bonds

determined experimentally for the 3d metal complexes with 5PCA acid. HOMA and  $I_6$  indexes for metal complexes have lower values than for ligand. Complexing by the metals (zinc, cobalt, copper, nickel and manganese) causes a decrease in the aromaticity of the pyrimidine ring (Table 9). Due to the lack of geometry data for 2PCA acid, the effect of 3d metals on the aromatic system of this

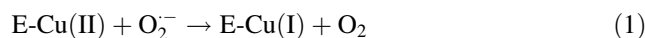
**Table 8** Geometrical parameters of the hydrogen bonds of manganese(II), cobalt(II), zinc and copper(II) pyrimidine-5-carboxylates complexes

	D	H	A	d(D–H)/Å	d(H–A)/Å	d(D–A)/Å	D–H–A/Å	Symm. op. 2
Mn	O1 W	H1 WA	N3	0.82(4)	2.046	2.835(3)	163(3)	$x, 1/2 - y, -1/2 + z$
	O1 W	H1WB	O2	0.92(4)	1.756	2.670(2)	172(4)	$-x, -1/2 + y, 1/2 - z$
	O2 W	H2 WA	O2	0.88(4)	1.849	2.720(2)	171(4)	$-1 - x, -1/2 + y, 1/2 - z$
	O2 W	H2WB	O1	0.96(4)	1.679	2.642(2)	177(4)	$-1 + x, y, z$
Co	O1 W	H1 WA	O1	0.76(4)	1.884	2.643(3)	174(4)	$1 - x, -y, -z$
	O1 W	H1WB	O2	0.96(5)	1.781	2.725(2)	169(6)	$x, 1/2 - y, -1/2 + z$
	O2 W	H2 WA	O2	0.76(4)	1.921	2.669(3)	167(4)	$-1 + x, 1/2 - y, 1/2 + z$
	O2 W	H2WB	N3	0.84(5)	2.028	2.832(3)	160(5)	$-x, -1/2 + y, 1/2 - z$
Zn	O1 W	H1 WA	N3	0.76(3)	2.104	2.8409(18)	164(3)	$x, 1.5 - y, -1/2 + z$
	O1 W	H1WB	O2	0.79(3)	1.89	2.6720(16)	172(3)	$-x, -1/2 + y, 1/2 - z$
	O2 W	H2 WA	O1	0.81(3)	1.837	2.6496(16)	175(3)	$-1 + x, y, z$
	O2 W	H2WB	O2	0.87(3)	1.863	2.7242(16)	172(3)	$1 - x, -1/2 + y, 1/2 - z$
Cu	O1 W	H1 WA	O1	0.87	2.114	2.799(2)	135.3	$1 - x, -y, -z$
	O1 W	H1WB	O3 W	0.87	1.978	2.837(3)	169.3	$1 - x, -y, 1 - z$
	O2 W	H2 WA	O1	0.87	1.921	2.713(2)	150.3	$x, y, -1 + z$
	O2 W	H2WB	O3 W	0.87	1.852	2.724(2)	178.2	$x, y, z$
	O3 W	H3 WA	O2	0.89(4)	1.815	2.699(3)	168(3)	$1 - x, -y, -z$
	O3 W	H3WB	O1	0.81(5)	2.02	2.835(3)	176(4)	$-x, -y, -z$

ligand can not be assessed using aromaticity indices. A comparison of the indexes for the complexes indicates that the 2PCA cobalt and copper complexes have lower aromaticity than the metal complexes of 5PCA, while the zinc complex shows higher aromaticity. The values of the aromatic indexes were also compared for the theoretically modeled structures in B3LYP/LANL2DZ for pyrimidine-2-carboxylic and pyrimidine-5-carboxylic acids and their complexes with copper (Fig. 8). The data indicate that copper may stabilize or destabilize the pyrimidine aromatic system depending on the position of the carboxyl group in the ring. In the case of 2PCA, index values indicated an increase in aromaticity after complexing with copper, while in the case of 5PCA, a decrease in the aromaticity system was observed (Table 10).

### SOD activity of copper complexes

SOD activity for the tested compounds was determined by theoretical methods. It is known that Cu, Zn-SOD dismutase and other compounds trap the superoxide anion in the following reactions:

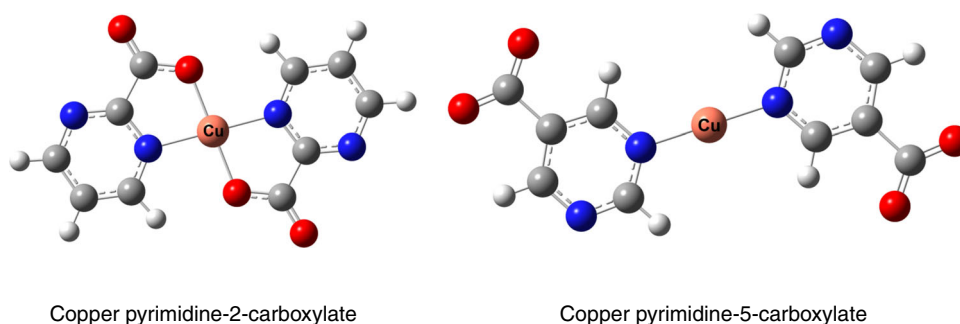


As a theoretical parameter defining the radical scavenging activity by the studied complexes, the HOMO and LUMO energy values and the EA electron affinity value defined as TEE-Cu(I)–TEE-Cu(II) were selected in which TE E-Cu(I) is the total energy of the complex E-Cu(I) and TEE-Cu(II) with the total energy of the E-Cu(II) complex.

**Table 9** Aromaticity (HOMA, EN, GEO and  $I_6$ ) of pyrimidine-2-carboxylates and pyrimidine-5-carboxylates (experimental structures)

Aromaticity index	Pyrimidine-5-carboxylic	Pyrimidine-5-carboxylate					Pyrimidine-2-carboxylate		
		Manganese complex	Cobalt complex	Copper complex	Zinc complex	Nickel complex [19]	Cobalt [18]	Copper [18]	Zinc complex
HOMA	0.998	0.993	0.995	0.996	0.993	0.994	0.994	0.978	0.997
GEO	0.002	0.004	0.003	0.004	0.003	0.005	0.005	0.019	0.002
EN	0.001	0.003	0.002	0.001	0.004	0.001	0.002	0.002	0.001
$I_6$	80.62	80.74	82.64	81.77	79.65	80.70	80.94	83.10	81.23

**Fig. 8** The coordination modes of copper complexes with pyrimidinecarboxylic acids calculated in B3LYP/LANL2DZ



**Table 10** Aromaticity (HOMA, EN, GEO and  $I_6$ ) of copper pyrimidine-2-carboxylates and copper pyrimidine-5-carboxylates (calculated in B3LYP/LANL2DZ)

Aromaticity index	Pyrimidine-2-carboxylic acid	Copper complex	Pyrimidine-5-carboxylic acid	Copper complex
HOMA	0.936	0.940	0.915	0.912
GEO	0.002	0.005	0.005	0.010
EN	0.062	0.055	0.079	0.077
$I_6$	80.32	82.15	83.68	77.29

The lower the EA value, the higher the electron transfer rate. Electronic affinity (EA) is a good descriptor for characterizing the superoxide radical scavenging activity [48, 49].

The compounds with the lowest EA value have the highest electron transfer capacity, thus giving the highest SOD activity [49]. Calculations of energy values of molecules and HOMO and LUMO molecular orbital have been made for optimized structures of copper complexes with diazines-derived acids. Geometric optimization was performed using the B3LYP density functional method using the LAN2LDZ calculation base.

Figure 9 presents HOMO and LUMO orbitals for copper complexes with pyrimidinecarboxylic acids calculated in B3LYP/LANL2DZ. In Table 11, the values of EA affinity and HOMO and LUMO energy and other parameters characterizing energy values are presented. For comparison, literature data on the energy of the active center of Cu, Zn-SOD enzyme were summarized [49]. EA calculations for copper complexes with diazines-derived acids have shown that the highest SOD activity is exhibited by the pyrimidine-5-carboxylic acid complex (it has the lowest EA energy value of  $-108,875 \text{ kcal mol}^{-1}$ ), while the pyrimidine-2-carboxylic acid complex has the lowest activity  $-100,021 \text{ kcal mol}^{-1}$ . On the basis of the calculated values of EA energy and comparison with the value of this energy for the active center of the enzyme dismutase, it can be concluded that copper complexes with pyrimidine-derived acids behave like superoxide dismutase mimetics. One of the stability criteria for aromatic systems

is the energy difference between HOMO and LUMO levels [50].

The higher the HOMO level energy of the  $\pi$ -reductant compounds, the more easily they can be oxidized, while the  $\pi$ -deficits compounds they are easier to reduce the easier the lower the energy has the LUMO orbital level.

From the HOMO and LUMO energy calculations for the copper complexes of pyrimidinecarboxylic acids, it appears that the 2PCA copper complex it is easier to reduce than the 5PCA complexes. The hardness of the molecule can be determined based on the HOMO and LUMO energy values. Absolute hardness is defined as follows: [51]:

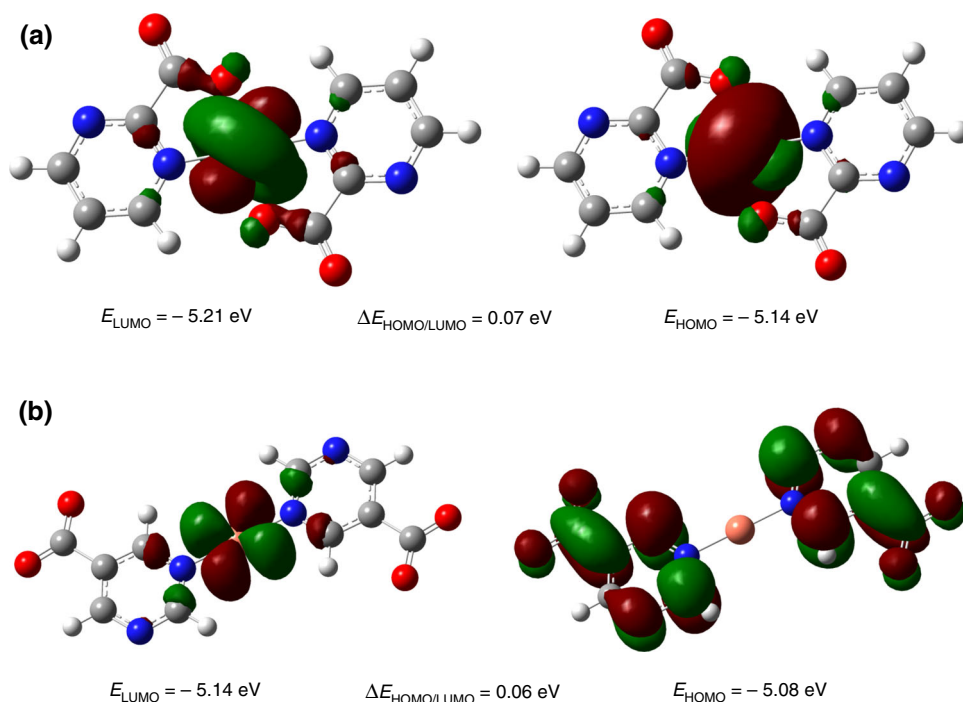
$$\eta = (e_{\text{LUMO}} - e_{\text{HOMO}})/2$$

The high value of absolute hardness  $\eta$  is a measure of the thermodynamic stability of the system. The value of chemical hardness for copper pyrimidine-5-carboxylate is slightly lower than of copper pyrimidine-2-carboxylate.

## IR and Raman spectra

The pyrimidine-2-carboxylic and pyrimidine-5-carboxylic acids spectra include characteristic bands associated with the oscillation of the carboxyl group and the vibration of the C–N bonds of the aromatic ring. The position of these bands in the spectrum of acid depends on the relative position of the carboxyl group and the nitrogen atoms in the aromatic ring. The bands derived from the vibrational stretching  $\nu_{\text{C=O}}$  in the infrared spectrum of pyrimidine-2-carboxylic acid are located at  $1743 \text{ cm}^{-1}$  ( $1738 \text{ cm}^{-1}$  in

**Fig. 9** HOMO and LUMO orbitals for copper complexes with pyrimidinecarboxylic acids calculated in B3LYP/LANL2DZ (**a**: copper pyrimidine-2-carboxylate, **b**: copper pyrimidine-5-carboxylate)



**Table 11** Geometrical parameters calculated in B3LYP/LANL2DZ

	Copper pyrimidine-2-carboxylate	Copper pyrimidine-5-carboxylate
Dipole moment (D)	0.00003	0.0013
TECu(II) (hartree)	- 1100.6524	- 1100.5447
TECu(I) (hartree)	- 1100.8118	- 1100.7182
EA (kcal mol <sup>-1</sup> )	- 100.021	- 108.875
HOMO (Hartree)	- 0.19142	- 0.18880
LUMO (Hartree)	- 0.18883	- 0.18662
HOMO (eV)	- 5.21	- 5.14
LUMO (eV)	- 5.14	- 5.08
Energy gap	0.07	0.06
Ionization potential. $I = - E_{\text{HOMO}}$	5.21	5.14
Electron Affinity. $A = - E_{\text{LUMO}}$	5.14	5.08
Electronegativity. $\chi = \frac{I+A}{2}$	5.18	5.11
Chemical potential. $\mu = -\frac{I+A}{2}$	- 5.18	- 5.11
Chemical hardness. $\eta = \frac{I-A}{2}$	0.035	0.030
Chemical softness. $S = \frac{1}{2\eta}$	14.29	16.67
Electrophilicity index. $\omega = \frac{\mu^2}{2\eta}$	383.32	435.20

the IR spectrum of ATR) and 1719 cm<sup>-1</sup> and at 1721 cm<sup>-1</sup> in the Raman spectrum (Table 12).

In the case of pyrimidine-5-carboxylic acid, the band is located at 1714 cm<sup>-1</sup> in the IR spectrum of KBr, 1704 cm<sup>-1</sup> IR ATR and at 1708 cm<sup>-1</sup> in the Raman spectrum. In the spectra of both acids, the bands originating from the vibrations of the carboxylic group have different

positions (Table 13). The close vicinity of the nitrogen atoms of the aromatic ring causes the vibrational numbers of the pyrimidine-2-carboxylic acid carboxyl group to have a higher value than in the pyrimidine-5-carboxylic acid spectrum. In the pyrimidine-5-carboxylic acid spectrum, the band is located at a lower wavenumber. Similar dependencies were noted by comparing the spectra of

**Table 12** Wavenumbers ( $\text{cm}^{-1}$ ), intensities and assignments of bands occurring in the IR (KBr, ATR and DFT) and Raman spectra of pyrimidine-2-carboxylic acid and sodium, manganese (II), nickel (II), cobalt (II), copper (II), zinc pyrimidine-2-carboxylates

Pyrimidine-2-carboxylic acid											Pyrimidine-2-carboxylate											Assignments						
Sodium						Manganese			Zinc			Cobalt			Nickel			Copper										
IR KBr	IR ATR	Raman	IR theor	Inten		IR KBr	IR ATR	Raman	IR KBr	IR ATR	IR KBr	IR ATR	Raman	IR KBr	IR ATR	IR KBr	IR ATR	IR KBr	IR ATR	IR KBr	IR ATR		IR KBr	IR ATR	IR KBr	IR ATR		
3426m	3460vw		3690	90.27	3434m	3463w										3451s	3468m	3440vs	3436w							vOH		
			3138	7.55			3174s	3154vw																			vCH	
3107s	3106vw	3108m	3096	14.63		3100m	3106s	3088s	3096m	3102w	3097s	3097s	3088s			3087m	3089w	3086s	3082vw								vCH	
3080s		3063m	3093	12.81	3048m	3084w	3068s																				vCH	
1743vs	1738m	1721s	1778	308.31																							vC = O	
1719vs																											vCH	
1646s		1638vw																									vCH	
1576vs	1573s	1561vs	1574	31.07	1639vs	1636s	1626vs	1621w	1664vs	1663s	1668vs	1663s	1621w	1664vs	1663s	1637vs	1637s	1653vs	1647vs								vasCOO <sup>-</sup>	
1572vs		1534vw	1569	79.59	1565s	1566m	1593m	1578m	1582s	1581m	1582vs	1581s	1578m	1582s	1581m	1590vs	1592vs	1588s	1589w								vCC, vCN, $\beta$ CH, $\gamma$ CH, $\gamma$ OH	
1459s					1569s	1564m	1561w	1559m											1534m	1528w							vCC, vCN, $\beta$ CH, $\gamma$ CH, $\gamma$ OH	
1442s	1449w	1459w	1435	5.33	1443m	1449w	1469s	1444s	1439m	1439w	1445s	1442w	1444s	1439m	1439w	1457w	1456w	1474m	1476vw								vCC, $\beta$ CH, $\gamma$ OH	
1408vs			1423	19.11		1434w	1420s				1422m							1416s	1416w								vCH, vCN, $\beta$ CH, $\gamma$ OH	
1381vs	1382m	1382vw	1354	99.24	1379vs	1378vs	1381w	1389vs	1358vs	1357vs	1357vs	1355vs	1389vs	1358vs	1357vs	1391s	1389s	1351vs	1351vs								$\beta$ OH	
1300vs		1301w																									vsymCOO <sup>-</sup>	
1279vs	1279s	1282m	1278	2.3	1279w	1281vw	1275s	1299vw	1275m	1275w	1275s	1275w	1299vw	1275m	1275w	1285w	1285w	1292s	1294w								$\beta$ OH	
			1223	32.45		1260vw	1232s																				$\beta$ CH	
1197s		1200w	1217	0.16	1219m	1225vw	1200s	1219vw	1207m	1209w	1209s	1208w	1209w	1207m	1209w	1209w	1211w	1183s									$\beta$ CH, vCC, vCN	
1174vs	1170vs		1128	257.13		1132s	1137vw	1200vw										1174s									$\beta$ CH, vCH	
1168vs																											$\beta$ CH, $\beta$ OH	
1102s	1105w	1106w	1092	44.10	1100w	1102w	1089s	1120vw	1098m	1096m	1096m	1096m	1120vw	1098m	1096m	1044w	1059s	1162s	1170m								$\beta$ CH	
1079s		1081s	1066	5.91	1079m	1079m	1072s	1072s	1072s	1072s	1072s	1072s	1072s	1072s	1072s	1072s	1072s	1059s	1056m								$\beta$ CH, $\beta_{\text{ring}}$	
		1033vw	992	2.81	1004w	1010w	1031s	1014w	1023m	1024w	1023s	1024w	1015vs	1023m	1024w	1044w	1059s	1056m									$\beta_{\text{ring}}$	
1004s	1003w	1005vs	989	0.07	1004w	1006w	1010w	1001vw										996m	996m								$\gamma$ CH	
940s	940m	977vw	979	0.01		997m																					$\gamma$ CH	
843s	845w	901vw	846	4.42	845m	843m	858w	889vs	853m	853m	855s	855m	852m	853m	853m	862m	861w	858s	860m								$\gamma$ CH	
807s			786	2.80																							$\beta_{\text{sym}}\text{COO}^-$	
814s	810m	811m	716	91.51		843w	744vs	718w	821m	819w	820s	817w	718w	821m	819w	731m	729m	796s	796w								$\beta$ OH	
		704vw	656	72.38	719s	716vw	668vs	651w	705s	702s	705vs	704s	718w	705s	702s	704s	704s	796s									$\gamma$ CH	
700vs	695s	637w	630	11.57	639m	637s	649m	638vw	656m	654s	656s	651m	651w	656m	654s	651m	630m	661m									$\beta_{\text{ring}}$	
637vs	637s		583	64.16																							$\beta_{\text{ring}}$	
528m																											$\gamma$ ring	
																												$\gamma$ OH
																												$\beta_{\text{as}}\text{COO}^-$



Table 12 (continued)

Pyrimidine-2-carboxylic acid			Pyrimidine-2-carboxylate												Assignments				
IR KBr	ATR	Raman	IR theor	Inten	Sodium			Manganese			Zinc		Cobalt		Nickel		Copper		
					IR KBr	ATR	Raman	IR KBr	ATR	IR	Raman	IR KBr	ATR	IR KBr	ATR	IR KBr	ATR	IR KBr	ATR
418w			489	1.15	422w	423w	463vs	436w	442m	443s	460m							$\gamma$ ring	
			411	2.08	403w	403w	377vs	422vw										$\gamma$ CH, $\gamma$ ring	24
			398	0.22	345vw	345vw												$\gamma$ CH, $\gamma$ ring	23

v: stretching;  $\beta$ : in plane;  $\gamma$ : out of plane deformations; s: strong; vs: very strong; m: medium; w: weak; vw: very weak; as—symmetric, sym—symmetric

pyridinecarboxylic acids (picolinic, nicotinic and isonicotinic acid) [52]. In the analyzed spectra, differences in the position of bands associated with the vibration of the aromatic system in both acids were observed. Spectral assignments used the normal oscillation of the pyrimidine ring (Fig. 10) measured using B3LYP/6-311++G(d,p).

For example, the band denoted no. 5 is located at  $1576\text{ cm}^{-1}$  in the pyrimidine-2-carboxylic acid IR spectrum and at  $1593\text{ cm}^{-1}$  in the pyrimidine-5-carboxylic acid spectrum.

Also bands 7 and 8 are located at higher values of wave numbers for 5PCA acid than in the 2PCA spectrum ( $1484\text{ cm}^{-1}$ ,  $1439\text{ cm}^{-1}$  and  $1442\text{ cm}^{-1}$ ,  $1408\text{ cm}^{-1}$ , respectively). In the case of other bands, e.g., 12, 13, 17, there are also significant differences in the position of the bands for each acid. Some of the bands are observed in the pyrimidine-2-carboxylic acid spectrum (e.g., band 14), and they are absent in the pyrimidine-5-carboxylic acid spectrum. Some of the bands are observed in the pyrimidine-5-carboxylic acid spectrum (e.g., 10, 15 bands), but they are absent in the pyrimidine-2-carboxylic acid spectrum.

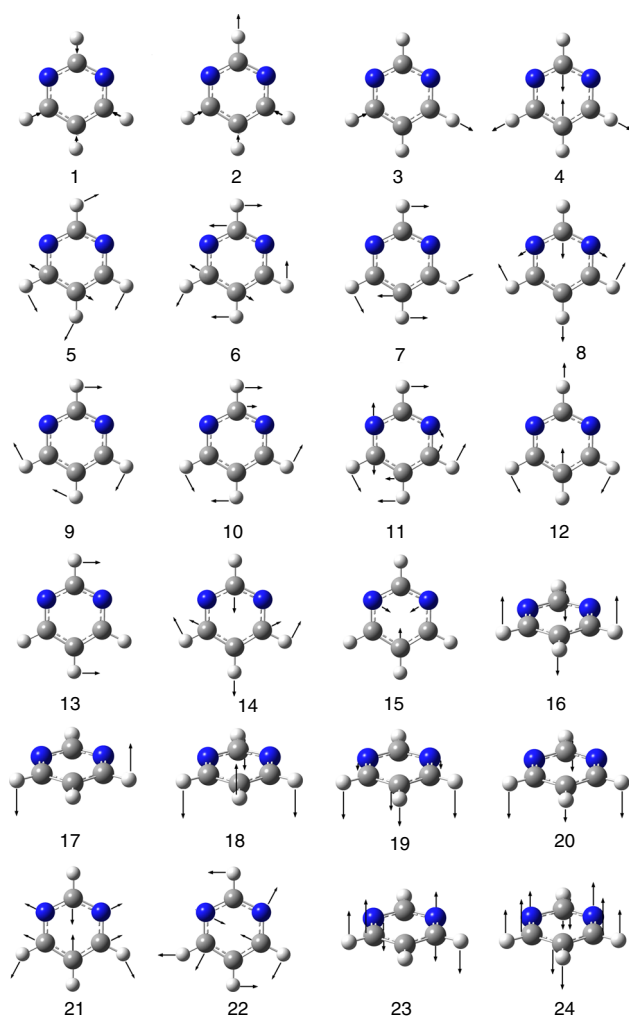
The vibrations of the carboxylate anion are present in the spectra of the pyrimidinecarboxylic acids metal complexes (Figs. 11 and 12). The analysis of position of the asymmetric and symmetrical stretching bands of the carboxylate anion in comparison with position of the bands of carboxylate anion of the sodium salt allows the determination of the metal–ligand coordination mode. Table 14 contains the values of wave numbers of the carboxylate anion vibrations for the 2PCA acid complexes as well as differences in the values of these vibrations. According to Nakamoto, this allows the assessment of the metal–ligand coordination mode [53]. In the case of complexes (Mn, Co, Cu, Zn) of pyrimidine-2-carboxylic acid, the difference in wavenumbers of tensile vibrations of asymmetric and asymmetric carboxylate anions ( $\Delta\nu = \nu_{\text{as}}\text{ COO}^- - \nu_{\text{sym}}\text{ COO}^-$ ) for 3d metal complexes of 2PCA acid is higher than for sodium salts indicates a monodentate mode of metal–ligand coordination by a carboxyl group, which agrees with the X-ray diffraction results for single crystals. Additional metal–ligand binding takes place via the nitrogen atom of the pyrimidine ring, which is observed in changes in the infrared spectra (the disappearance of the tensile bands  $\nu_{\text{CN}}$  in the spectra of the complexes).

In the case of the nickel complex, the difference  $\Delta\nu$  in the spectrum is smaller than in the sodium salt, which suggests a bidentate chelation coordination mode. The mode of metal–ligand coordination by the nitrogen atom without the participation of a carboxylic group is more likely due to the characteristic changes in the bands originating from vibrations of the aromatic system (disappearance of bands in the nickel complex) and the appearance of an unbound carboxylate anion with aligned





wave numbers of many bands originating from the vibrations of the aromatic system was observed (5, 7, 8, 11, 13, 20, 21). There are also vibration-derived bands that were absent in the spectrum of the ligand (e.g., band No. 15). For some of the bands, decreasing of the wavenumber value in the complex spectrum with respect to the ligand or disappearance of the bands (6, 9, 12, 16) was observed. This indicates an increase in the aromaticity of the 2PCA acid system under the influence of complexation. In the spectra of pyrimidine-5-carboxylic acid metal complexes, greater changes were observed with respect to the ligand than in the spectra of the pyrimidine-5-carboxylic acid complexes. Decreased values of wave numbers or disappearance of bands derived from vibrations of the pyrimidine carboxylic acid ring were observed. These are bands marked with 5, 6, 7, 8, 9, 12, 13, 14. An increase in wavenumbers was observed only for some bands (10, 11, 15). A comparative analysis of the spectra of pyrimidine carboxylate complexes and pyrimidinecarboxylic acids has shown that



**Fig. 10** Normal modes of pyrimidine ring calculated in DFT/B3LYP/6-311++(d,p)

transition 3d metals have a greater impact on the 2PCA acid aromatics system than 5PCA acid. In the case of 2PCA, the transition metals attached via the carboxyl group and the nitrogen atom of the aromatic ring stabilize the electron system, whereas in the case of pyrimidine-5-carboxylic acid complexes, a disorder of the aromatic ligand system is observed. Direct attachment of the metal to the aromatic pyrimidine ring causes an increase in the disturbance of the aromatic system.

## Thermal analysis

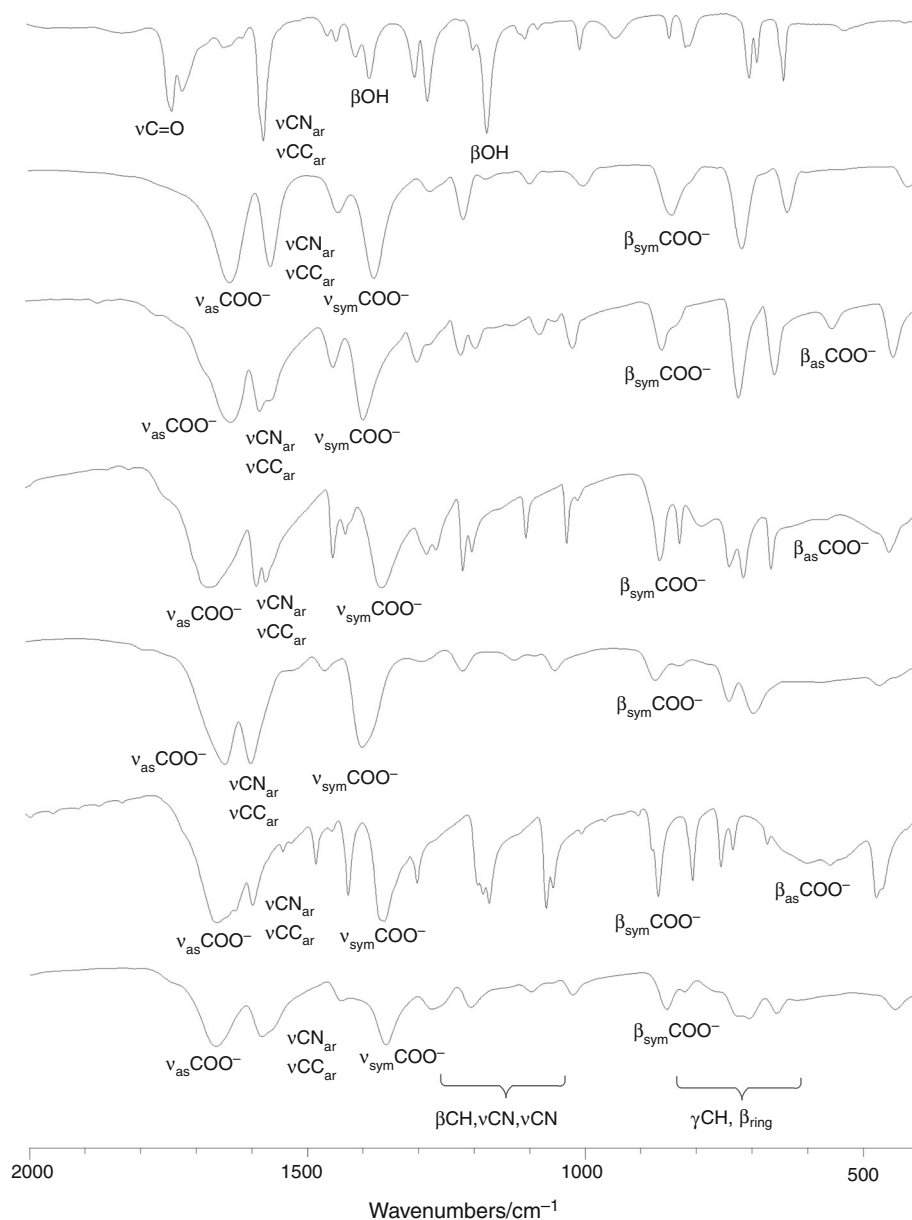
Different methods of thermal analysis (TG–DSC, TG–FTIR) were used to determine thermal behavior of two series of metal(II) complexes during heating in air and nitrogen atmospheres. Thermal analysis methods allow to demonstrate that the same metal ions with the isomers of pyrimidinecarboxylic acid give complexes with a different number of solvent molecules. The pyrimidine-2-carboxylate complexes contain in the structure a smaller number of water molecules compared to the corresponding complexes with pyrimidine-5-carboxylic acid.

Thermal decomposition of the investigated complexes in air atmosphere proceeds in two main stages connected with dehydration and decomposition/burning processes (Tables 15, 16, Figs. 13–17).

Manganese(II) pyrimidine-2-carboxylate ( $MnL_1$ ) as an anhydrous complex exhibits high thermal stability up to about 250 °C, while manganese(II) pyrimidine-5-carboxylate tetrahydrate ( $MnL_2 \cdot 4H_2O$ ) is stable to 110 °C (Fig. 13). Further heating leads to the one-stage removal of all four water molecules associated with 18.85% mass loss in the temperature range 110–140 °C accompanied by a strong endothermic effect at 122 °C observed on the DSC curve. The corresponding molar enthalpy of dehydration process is equal to 153.5 kJ mol<sup>-1</sup>. An anhydrous form of manganese(II) pyrimidine-5-carboxylate is thermally stable up to about 350 °C. Further heating of both manganese(II) complexes results in distinct mass losses observed on the TG curves connected with one-stage decomposition and combustion of organic ligands. Very strong exothermic effects appear on the DSC curves. Manganese(III) oxide is the final decomposition product formed at about 470–480 °C. During heating up to about 940 °C,  $Mn_2O_3$  transforms into  $Mn_3O_4$  [56, 57].

The first mass losses on the TG curves of cobalt complexes  $CoL_1 \cdot 2H_2O$  and  $CoL_2 \cdot 4H_2O$  are a result of the dehydration process (Fig. 14). The release of water molecules in the  $CoL_1 \cdot 2H_2O$  complex takes place in the temperature range 150–220 °C. The other complex is less stable, and loss of solvent molecules is observed at lower temperature 123–205 °C. The shapes of the DSC curves of both complexes  $CoL_1 \cdot 2H_2O$  and  $CoL_2 \cdot 4H_2O$  indicate that

**Fig. 11** IR spectra of pyrimidine-2-carboxylates in the range from 500 to 2000  $\text{cm}^{-1}$  (from above: pyrimidine-2-carboxylic acid, sodium complex, manganese(II) complex, cobalt(II) complex, nickel(II) complex, copper(II) complex and zinc complex)



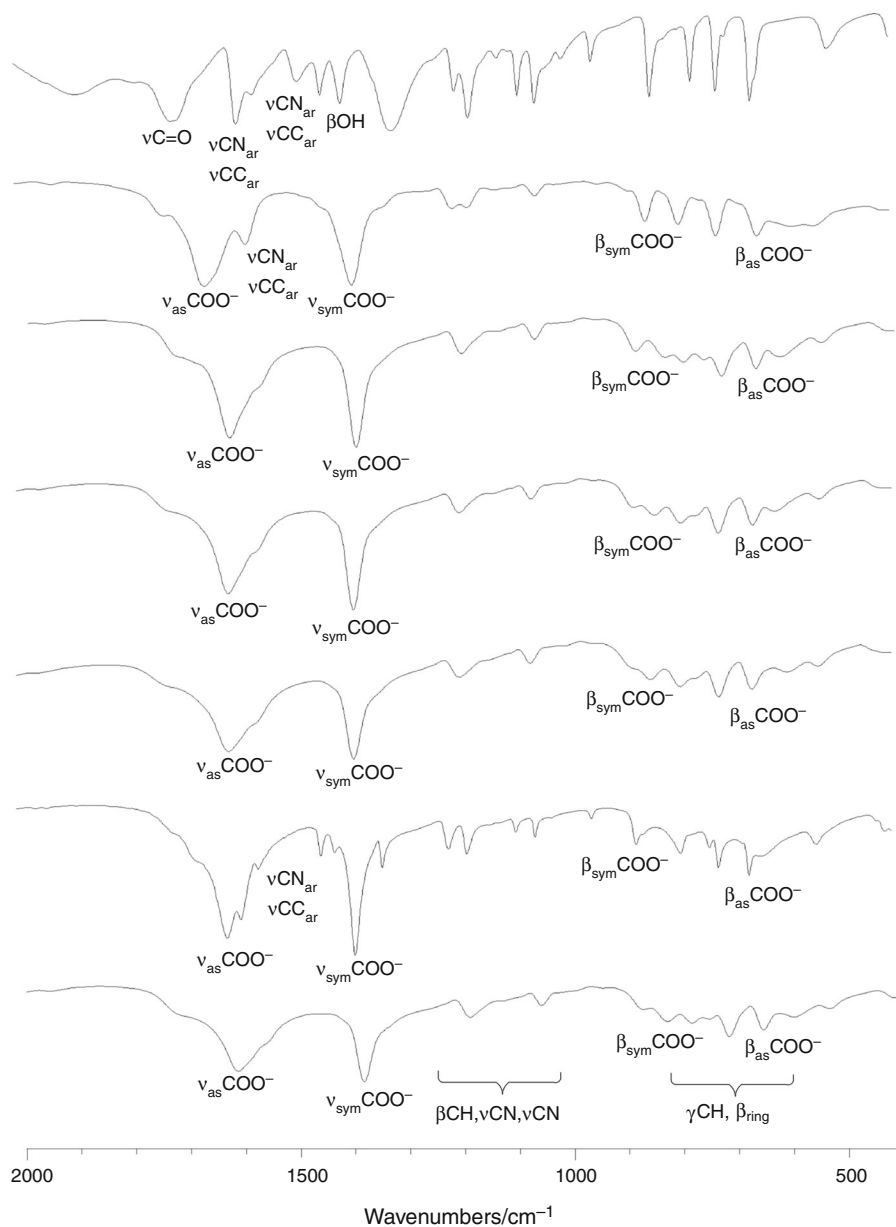
removal of water molecules occurs in overlapping stages with the overall molar dehydration enthalpy values of 78.0  $\text{kJ mol}^{-1}$  and 177.9  $\text{kJ mol}^{-1}$ , respectively. The anhydrous complexes  $\text{CoL}_1$  and  $\text{CoL}_2$  are stable in the narrow temperature ranges 220–307 and 205–260  $^{\circ}\text{C}$ , respectively. At a higher temperature, two mass losses appear on the TG curves as a result of the decomposition and burning of the anhydrous form of complexes. These stages are accompanied by very strong exothermic effects. As the final solid product of thermal decomposition,  $\text{Co}_3\text{O}_4$  is formed. Further heating of such compound leads to the transformation into the  $\text{CoO}$  at about 940  $^{\circ}\text{C}$  [58, 59].

The nickel complexes  $\text{NiL}_1 \cdot 2\text{H}_2\text{O}$  and  $\text{NiL}_2 \cdot 4\text{H}_2\text{O}$  exhibit thermal stability up to 164 and 132  $^{\circ}\text{C}$ , respectively (Fig. 15). During heating of nickel(II) pyrimidine-2-

carboxylate dihydrate, the removal of both water molecules occurs in the temperature range 164–258  $^{\circ}\text{C}$ . The first mass loss connected with the dehydration process is 11.58% (calc. 10.57%). As can be deduced from the shape of the DSC curve, this process probably proceeds in two stages (temperatures of the endothermic peaks are 201  $^{\circ}\text{C}$  and 229  $^{\circ}\text{C}$ ) with the overall enthalpy of 83.6  $\text{kJ mol}^{-1}$ . Nickel(II) pyrimidine-5-carboxylate tetrahydrate loses all water molecules in one stage in the temperature range 132–208  $^{\circ}\text{C}$  with the corresponding molar dehydration enthalpy value of 192.3  $\text{kJ mol}^{-1}$  (the maximum of the endothermic peak at 200  $^{\circ}\text{C}$ ). Anhydrous compounds are stable in narrow temperature ranges and heating leads to the two-stage decomposition connected with burning of the organic ligand. Very strong exothermic effects can be



**Fig. 12** IR spectra of pyrimidine-5-carboxylates in the range from 500 to 2000  $\text{cm}^{-1}$  (from above: pyrimidine-5-carboxylic acid, sodium complex, manganese(II) complex, cobalt(II) complex, nickel(II) complex, copper(II) complex and zinc complex)



**Table 14** Spectral parameters (vibrations stretching the carboxylate anion) from the IR spectra of pyrimidine-2-carboxylates and proposed type of coordination the metal to carboxylic group of the ligand

Compound	$\nu_{\text{asym}}(\text{COO}^-)$ Wavenumber/ $\text{cm}^{-1}$	$\nu_{\text{sym}}(\text{COO}^-)$ Wavenumber/ $\text{cm}^{-1}$	$\Delta\nu^a$ Wavenumber/ $\text{cm}^{-1}$	Proposed coordination
Sodium salt	1639	1379	260	
Manganese complex	1626	1311	315	Monodentate
Cobalt complex	1668	1357	311	Monodentate
Nickel complex	1637	1391	246	Bidentate chelation
Copper complex	1653	1351	302	Monodentate
Zinc complex	1664	1358	306	Monodentate

<sup>a</sup> $\Delta\nu = \nu_{\text{asym}}(\text{COO}^-) - \nu_{\text{sym}}(\text{COO}^-)$

observed on the DSC curves. Nickel(II) oxide is formed at about 480 °C as the final solid product of thermal decomposition nickel complexes [60].

The copper(II) compound with pyrimidine-2-carboxylic acid as an anhydrous complex displays great thermal stability. Two-stage decomposition occurs in the temperature range of 260–390 °C. As the final solid product of decomposition, copper(II) oxide is formed. Copper(II) pyrimidine-5-carboxylate hexahydrate is stable to 80 °C. During further heating, the dehydration process occurs in two hardly separated stages. Mass losses of 24.30% and 2% observed on the TG curve are associated with the removal of 5.5 and 0.5 water molecules. The first stage of dehydration is accompanied by a very strong endothermic effect with the molar enthalpy of dehydration equal to 200.7 kJ mol<sup>-1</sup> (temperature of the peak maximum is 95.6 °C). The other endothermic effect was recorded at 150 °C. The dehydrated form of complex is stable to 245 °C. Next, the two-step decomposition takes place in the range 286–384 °C (Fig. 16). Also, CuO is formed as stable solid residue [58, 61].

The zinc complexes ZnL<sub>1</sub>·2H<sub>2</sub>O and ZnL<sub>2</sub>·4H<sub>2</sub>O are thermally stable up to 110 and 98 °C, respectively (Fig. 17). At higher temperature, mass losses on the TG curves are observed due to the dehydration process. As can be concluded from the TG and DSC curves of the ZnL<sub>1</sub>·2H<sub>2</sub>O complex, release of water molecules takes place in two hardly distinguishable stages. The mass loss of 10.97% is observed in the temperature range 110–190 °C. The dehydration process proceeds with the two overlapping endothermic effects with the maxima at 144 °C and 156 °C and total molar dehydration enthalpy of 83.6 kJ mol<sup>-1</sup>. In the other zinc complex, four water molecules are released in one stage in the narrow temperature range 110–140 °C. This process is accompanied by a very strong endothermic effect with the molar dehydration enthalpy 174.8 kJ mol<sup>-1</sup>. The dehydration process resulted in the formation of anhydrous compounds which are stable in the

relatively wide temperature ranges: 190–370 °C and 152–316 °C. Similarly to the above-described compounds, further heating leads to the two-stage decomposition resulting from degradation and burning of anhydrous form of complex. Zinc oxide is formed as a final solid product of complexes decomposition [58, 60].

### TG–FTIR analysis

Volatile products of thermal decomposition of the CoL<sub>1</sub>·2H<sub>2</sub>O and CoL<sub>2</sub>·4H<sub>2</sub>O complexes were investigated by the coupled TG–FTIR method as representatives of two series of investigated complexes. The infrared spectra of gaseous products released during complexes heating in nitrogen atmosphere are given in Fig. 18. The CoL<sub>1</sub>·2H<sub>2</sub>O complex is stable in the nitrogen temperature to 178 °C. Next, in the temperature range 180–240 °C, water molecules are evolved. The FTIR spectra exhibit characteristic bands in the wavenumber ranges 4000–3400 cm<sup>-1</sup> and 1800–1200 cm<sup>-1</sup> corresponding to the stretching and deformation vibrations of water molecules [57]. As shown in Fig. 18, the dehydrated form of the complex is stable up to 330 °C. At higher temperature, carbon dioxide molecules are released. The FTIR spectra show very diagnostic characteristic bands in the region 2359–2310 cm<sup>-1</sup> and that at 669 cm<sup>-1</sup> derived from the stretching and deformation vibrations of carbon dioxide (Fig. 18). Above 450 °C besides the bands from the carbon dioxide molecules, also a double band with the maxima at 2182 and 2094 cm<sup>-1</sup> is observed due to evolved carbon monoxide [62]. Moreover, along with carbon oxides weak bands derived from the ammonia and pyrimidine molecules are observed. These gases give the most intense bands at 440–500 °C. Identification of pyrimidine molecules can be made based on the bands in the region 3100–3000 cm<sup>-1</sup> which can be assigned to the stretching vibrations of CH groups. The bands recorded at 1569, 1540 cm<sup>-1</sup> and those at 1436, 1339 cm<sup>-1</sup> can be ascribed to the stretching

**Table 15** Results of thermal decomposition of Ni(II), Co(II), Cu(II), Zn(II) and Mn(II) complexes with pyrimidine-2-carboxylic acid in air

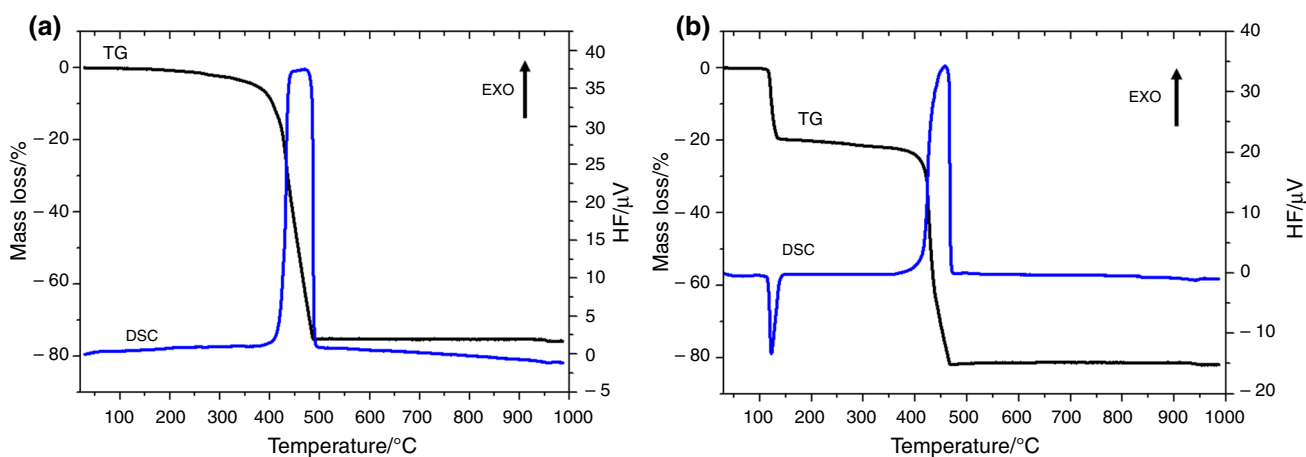
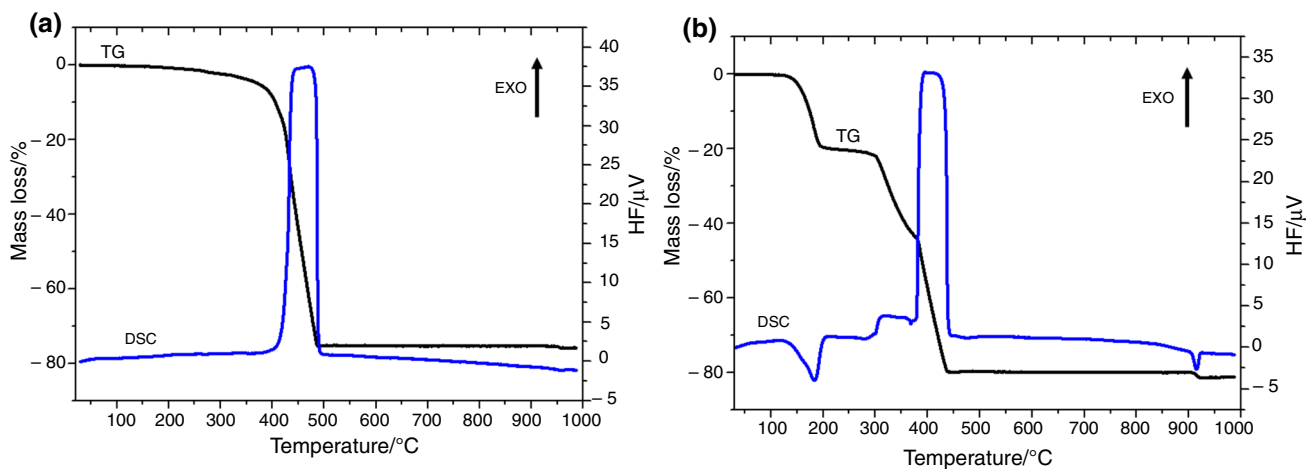
Complex	$\Delta T_1/^\circ\text{C}$	$\Delta H/\text{kJ mol}^{-1}$ ( $T_{\text{peak}}/^\circ\text{C}$ )	Mass loss/%		$\Delta T_2/^\circ\text{C}$	Mass loss/%		Residue
			Found	Calc.		Found	Calc.	
Co(2PCA) <sub>2</sub> ·2H <sub>2</sub> O	150–220	78.0 (183, 198)	12.39	10.55	307–480*	78.60	77.07	Co <sub>3</sub> O <sub>4</sub>
Ni(2PCA) <sub>2</sub> ·2H <sub>2</sub> O	164–258	83.6 (201, 229) (201, 229)	11.58	10.57	320–480	78.72	78.08	NiO
Zn(2PCA) <sub>2</sub> ·2H <sub>2</sub> O	110–190	83.6 (144, 156)	10.97	10.36	370–530	77.67	76.56	ZnO
Mn(2PCA) <sub>2</sub>	–	–	–	–	250–489**	74.15	73.78	Mn <sub>2</sub> O <sub>3</sub>
Cu(2PCA) <sub>2</sub>	–	–	–	–	260–390	74.57	74.30	CuO

2PCA—C<sub>5</sub>H<sub>3</sub>N<sub>2</sub>O<sub>2</sub> (pyrimidine-2-carboxylate);  $\Delta T_1$ —temperature range of dehydration;  $\Delta T_2$ —temperature range of degradation of anhydrous complex to suitable oxide; \*—further transformation of Co<sub>3</sub>O<sub>4</sub> oxide into CoO oxide is observed; \*\*—further transformation of Mn<sub>2</sub>O<sub>3</sub> oxide into Mn<sub>3</sub>O<sub>4</sub> oxide is observed

**Table 16** Results of thermal decomposition of Ni(II), Co(II), Cu(II), Zn(II) and Mn(II) complexes with pyrimidine-5-carboxylic acid in air

Complex	$\Delta T_1/^\circ\text{C}$	$(\Delta H/\text{kJ mol}^{-1} (T_{\text{peak}}/^\circ\text{C}))$	Mass loss/%		$\Delta T_2/^\circ\text{C}$	Mass loss/%		Residue
			Found	Calc.		Found	Calc.	
Co(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	123–205	177.9 (183)	19.95	19.10	260–449*	79.73	78.71	Co <sub>3</sub> O <sub>4</sub>
Ni(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	132–208	192.3 (200)	20.30	19.11	300–480	79.96	80.17	NiO
Zn(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	98–152	174.8 (133)	19.23	18.78	316–572	82.09	78.76	ZnO
Mn(5PCA) <sub>2</sub> ·4H <sub>2</sub> O	110–140	153.5 (122)	19.85	19.29	350–470**	80.75	78.84	Mn <sub>2</sub> O <sub>3</sub>
Cu(5PCA) <sub>2</sub> ·6H <sub>2</sub> O	80–167	200.7 (95.6)	26.13	25.86	245–384	82.95	80.94	CuO

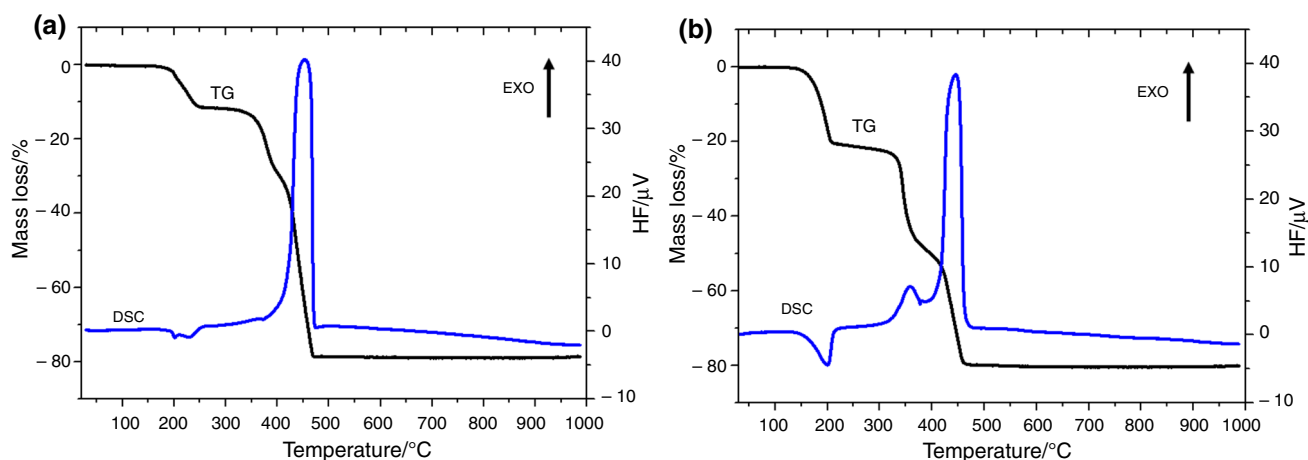
5PCA—C<sub>5</sub>H<sub>3</sub>N<sub>2</sub>O<sub>2</sub> (pyrimidine-5-carboxylate);  $\Delta T_1$ —temperature range of dehydration;  $\Delta T_2$ —temperature range of degradation of anhydrous complex to suitable oxide; \*—further transformation of Co<sub>3</sub>O<sub>4</sub> oxide into CoO oxide is observed; \*\*—further transformation of Mn<sub>2</sub>O<sub>3</sub> oxide into Mn<sub>3</sub>O<sub>4</sub> oxide is observed

**Fig. 13** TG and DSC curves of manganese(II) complexes with: **a** pyrimidine-2-carboxylic acid; **b** pyrimidine-5-carboxylic acid**Fig. 14** TG and DSC curves of cobalt(II) complexes with: **a** pyrimidine-2-carboxylic acid; **b** pyrimidine-5-carboxylic acid

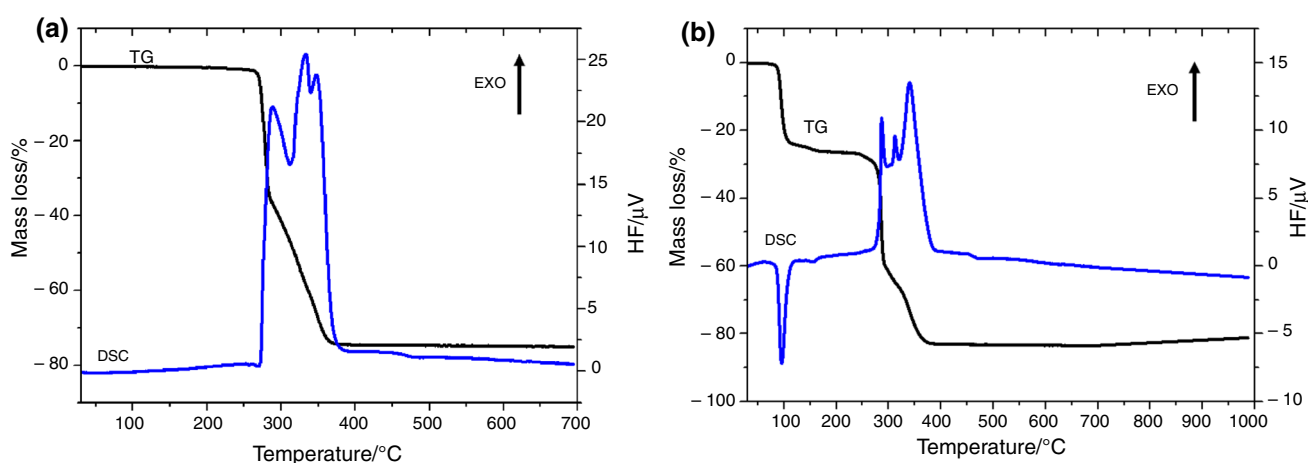
vibrations of aromatic ring  $\nu(\text{CC})$  and  $\nu(\text{NC})$  and deformation vibrations of CH moieties [63]. Ammonia molecules were detected among the gaseous products of complex thermal decomposition due to the presence of very diagnostic double peak bands at 967 and 933  $\text{cm}^{-1}$ .

At higher temperature, only carbon oxides and traces of water are evolved.

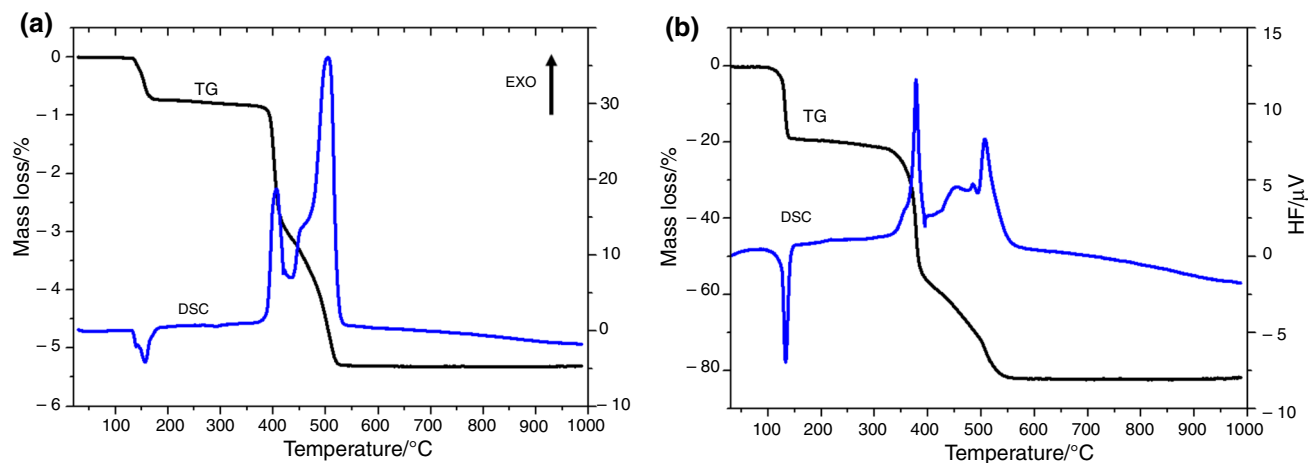
Cobalt(II) pyrimidine-5-carboxylate tetrahydrate is stable in the inert atmosphere to 127 °C. Next, the dehydration process occurs with water molecules evolution in



**Fig. 15** TG and DSC curves of nickel(II) complexes with: **a** pyrimidine-2-carboxylic acid; **b** pyrimidine-5-carboxylic acid



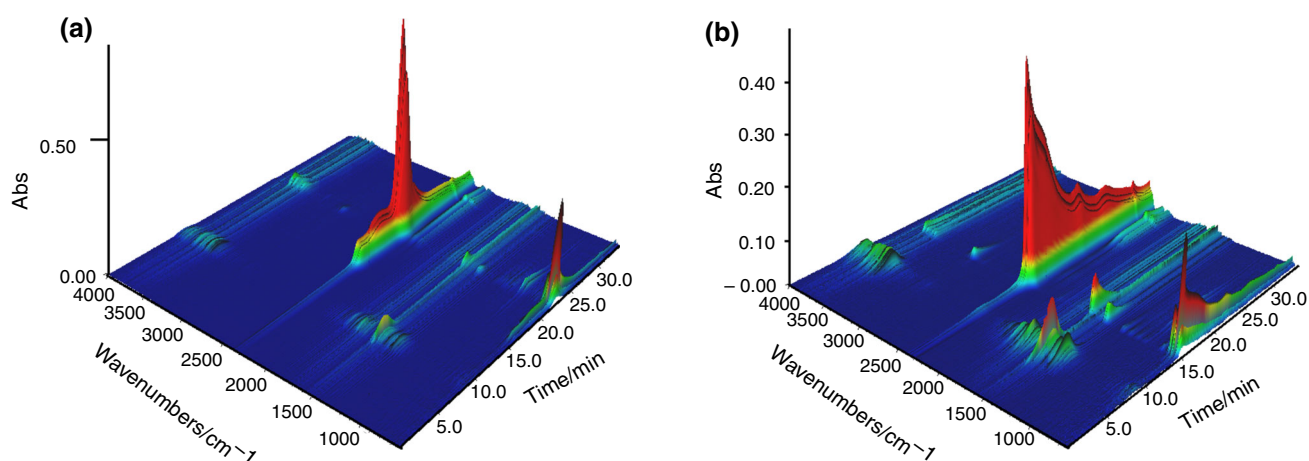
**Fig. 16** TG and DSC curves of copper(II) complexes with: **a** pyrimidine-2-carboxylic acid; **b** pyrimidine-5-carboxylic acid



**Fig. 17** TG and DSC curves of zinc(II) complexes with: **a** pyrimidine-2-carboxylic acid; **b** pyrimidine-5-carboxylic acid

the range 130–210 °C. The FTIR spectra exhibit diagnostic bands derived from the stretching and deformation vibrations of water (Fig. 18). The anhydrous compound is

stable to 310 °C. Further heating leads to the degradation of the pyrimidine-5-carboxylate ligand with the evolution of carbon dioxide, ammonia and pyrimidine molecules at



**Fig. 18** FTIR spectra of gaseous products evolved during the decomposition of: **a**  $\text{Co}(\text{2PCA})_2 \cdot 2.5\text{H}_2\text{O}$ ; **b**  $\text{Co}(\text{5PCA})_2 \cdot 4\text{H}_2\text{O}$

the same time. Compared with the previously described thermal decomposition of cobalt(II) pyrimidine-2-carboxylate, the bands derived from the pyrimidine molecules are more intense. In the wavenumber region  $3100\text{--}3000\text{ cm}^{-1}$ , there can be distinguished three peaks at  $3086$ ,  $3055$  and  $3040\text{ cm}^{-1}$  from the stretching vibrations of CH. In the lower wavenumber region, there are observed several double bands with the maxima at:  $1574$ ,  $1507\text{ cm}^{-1}$ ;  $1436$ ,  $1387\text{ cm}^{-1}$ ;  $1230$ ,  $1209\text{ cm}^{-1}$  and  $1151$ ,  $1140\text{ cm}^{-1}$  assigned to the stretching and different modes of deformation vibrations of CC, CN and CH moieties [63]. Carbon monoxide is evolved above  $600\text{ }^\circ\text{C}$ .

## Conclusions

Two series of novel metal(II) complexes with pyrimidine-2-carboxylic and pyrimidine-5-carboxylic acid were obtained and characterized. The work allowed to assess the influence of selected metals on the electronic structure and antioxidant properties of the pyrimidinecarboxylic acids. The influence of metals on the above properties was evaluated depending on the location of the carboxyl group relative to the heteroatoms in the aromatic ring. As a result of synthesis, some complexes were obtained in a crystalline form, which allowed for precise determination of the structure using X-ray diffraction. In addition, theoretical calculations of some structures were made, which allowed for a broader interpretation of the obtained results. Two ligands (pyrimidine-2-carboxylic acid and pyrimidine-5-carboxylic acid) were selected for the experimental tests.

Taking into account obtained results several conclusions can be drawn:

1. The metal–ligand coordination mode in complexes depends on the position of the carboxylic group in

relation to positions of the heteroatoms in the aromatic rings of acids. In the case of metal complexes with pyrimidine-2-carboxylic acid, the metal ions are linked to the ligand via the monodentate carboxylate group and the nitrogen atom of the aromatic ring. In the case of metal complexes with pyrimidine-5-carboxylic acid, the metal ions are only attached to the ligand via the nitrogen atom. These conclusions were confirmed by single-crystal X-ray diffraction method as well as spectroscopic (IR, Raman) investigations.

2. Water molecules in hydrated form of complexes play role of monodentate ligands. Only in copper(II) pyrimidine-5-carboxylate, two additional water molecules are in inner coordination sphere.
3. The analysis of spectroscopic data (IR, Raman) and theoretical calculations (aromaticity indexes) showed that transition 3d metals cause changes in the electronic charge distribution of the ligands. This effect is the greatest for the pyrimidine-5-carboxylates in where the metal ions are coordinated directly to the aromatic ring (by nitrogen atom). The metals can stabilize the electron system of pyrimidine-2-carboxylate ligand, while in the case of pyrimidine-5-carboxylate ligand decrease in aromaticity is observed.
4. Theoretical calculations (the SOD activity) have shown that the copper complex of pyrimidine-2-carboxylic acid has lower antioxidant activity than the copper complexes of pyrimidine-5-carboxylic acids.
5. The metal coordination mode also affects the thermal stability of the studied complexes. Thermal analysis has shown that the metal complexes with pyrimidine-2-carboxylic acid (with the exception of the manganese complex) exhibit a higher thermal stability compared to the metal pyrimidine-5-carboxylates. The final products of thermal degradation in the air atmosphere of studied complexes are suitable metal oxides. The



research shows that complexes of transition metals with pyrimidine carboxylic acids are more thermally stable when the metal is coordinated through the carboxyl group and the nitrogen atom of the pyrimidine ring than complexes in which the metal is coordinated only through the nitrogen atom of the pyrimidine ring.

- Investigated metal complexes decompose in nitrogen atmosphere with releasing of water, carbon oxides, ammonia and pyrimidine molecules.

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