

Addendum to “A simple kinetic method for the determination of the reaction model from non-isothermal experiments”

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To the Editor,

In 2010, J. Farjas, N. Butchosa, and P. Poura at the University of Girona published an article in the *Journal of Thermal Analysis and Calorimetry* titled “A simple kinetic method for the determination of the reaction model from non-isothermal experiments” [1]. The article describes an approach to determining reaction order from the shape of the reaction exothermic peak using the time at full-width-at-half-maximum (Δt_{FWHM}) as the metric. Equation 24 of that publication relates the value of the parameter $\Delta t'_{\text{FWHM}}$ to reaction order n for n th order reactions, through third-order polynomials in n . Because the purpose of the article is to obtain n from $\Delta t'_{\text{FWHM}}$ a relationship where $\Delta t'_{\text{FWHM}}$ is the independent variable would be more useful to users. In correspondence, Dr. Farjas has re-examined the data and offers the following relationship for this purpose.

$$n = -0.39374 + 0.36063\Delta t'_{\text{FWHM}} + 0.08934\Delta t'^2_{\text{FWHM}}$$

The difference between the exact and approximate values of n in this equation is below 0.04 for $0.11 < n < 3.96$ covering the complete practical range for n .

$\Delta t'_{\text{FWHM}}$ can be determined from the value of time's half width at half maximum at infinite heating rate, $\Delta t'_{\text{FWHM}}/A$, where A is the pre-exponential term of the rate constant. Both A and $\Delta t'_{\text{FWHM}}/A$ can be obtained from the y-intercept of the Kissinger and $\ln(\Delta t_{\text{FWHM}})$ versus the reciprocal of peak temperature plots, respectively.

Readers of the original article may well wish to add this relationship to the original article as a useful expansion.

Reference

1. Farjas J, Butchosa N, Roura P. A simple kinetic method for the determination of reaction model from non-isothermal experiments. *J Therm Anal Calorim.* 2010;102:615–25.

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