Fig. 2—Typical acid sawed surfaces with .cm scale. A) After acid sawing. B) After acid sawing and electropolishing.

the crystal at the crystal temperature for optimum cutting.

Under optimum conditions there is a uniform chemical reaction (very tiny brownish bubbles) along the cutting surface. Cuts are then smooth and appear polished directly after sawing. Fig. 2 shows an acid sawed surface just after acid cutting A, and after 1 h of electropolishing, B, in a solution of 60 pct sulphuric acid and 40 pct distilled water at 6 V and 1.2 amp.

1. F. W. Young and R. R. Wilson: *Rev. Sci. Instr.*, 1961, vol. 32, p. 559.
2. A. J. Berry: *J. Australian Inst. Metals*, 1961, vol. 6, p. 289.
3. W. J. McG. Tegart: *The Electrolytic and Chemical Polishing of Metals*, p. 102, Pergamon Press, 1959.
4. P. S. Ayres: M.S. Met. E. Thesis, Purdue University, June, 1965.

Corrections to *Met. Trans.*, 1972, vol. 3

Mechanisms and Kinetics of Absorption and Desorption Reactions in Systems of Refractory Metals with Nitrogen, Oxygen or Carbon, by Gerhard Hörz, pp. 3069-3076

Page 3070

Interchange the two footnotes in the right hand column.

Page 3071

In Table II the rate law for oxygen absorption in the systems V-O, Nb-O, Ta-O should read:

$$k_1 p_{O_2} / (1 + k_2^0 \exp(Q/RT))$$

Page 3072

Line 6 from bottom of right hand column should read:

... $\Delta\rho$ at room temperature ...

Pages 3073 and 3075

The footnote appearing at the top of page 3075 belongs at the bottom of the right hand column of page 3073.

Page 3075

Table IX should read:

Table IX. Some Results of the Kinetics of Decarburization of Niobium in Flowing Oxygen (Thin Samples, Solid Solutions of Niobium with Carbon)			
Reaction Range	Rate Law	Q^* kcal/mole	Rate-determining
I	$k k_1 p_{O_2} / (1 + k_2^{\circ} \exp (Q / R T))$	17.8	Transport and chemisorption of oxygen Diffusion of carbon
II	$k' \exp (-t / \tau)$ $1 / \tau = (1 / \tau_0) \exp (-Q / R T)$	38	
*For range I see footnote p. 3071.			