Additions and Corrections

Sir,

The account of some of our work reported⁽¹⁾ in *TMC Literature Highlights-13* conflates a number of papers and significant errors have crept in. This unfortunately detracts from an interesting review.

GeH_4 reacts⁽²⁾ with $[\text{Co}_2(\text{CO})_8]$ to form $[\text{GeCo}_4(\text{CO})_{14}]$ A, not $[\text{Ge}_2\text{Co}_6(\text{CO})_{20}]$ B. A consists of two GeCo_2 triangles sharing the *spiro*-Ge. A reacts⁽³⁾ with $\text{Co}(\text{CO})_4^-$ to give $[\text{GeCo}_5(\text{CO})_{16}]^-$ anion C, which does contain pentaligated Ge shared between one GeCo_2 triangle and one GeCo_3 pyramid.

B was formed⁽⁴⁾ from Ge_2H_6 or Ge_3H_8 and probably has an extended structure of GeCo_2 triangles linked alternatively by Ge apices and Co-Co edges. B loses CO to yield $[\text{Ge}_2\text{Co}_6(\text{CO})_{19}]$ D which contains a second form of five-coordinate Ge with two $\text{GeCo}(\text{CO})_4$ groups capping both faces of a $\text{Co}_4(\text{CO})_{11}$ quadrilateral.

We know of no germanium compound containing 'one $\text{Co}(\text{CO})_4$ group and two $\text{Co}_2(\text{CO})_6$ groups'. C and D represent the only known forms of Ge pentaligated to cobalt.

References

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- ⁽²⁾ R. F. Gerlach, K. M. Mackay, B. K. Nicholson and W. T. Robinson, *J. Chem. Soc., Dalton Trans.*, 80 (1981).
- ⁽³⁾ R. A. Croft, D. N. Duffy and B. K. Nicholson, *J. Chem. Soc., Dalton Trans.*, 1023 (1982).
- ⁽⁴⁾ S. P. Foster, K. M. Mackay and B. K. Nicholson, *Inorg. Chem.*, **24**, 909 (1985); cited in: *Transition Met. Chem.*, **11**, 426 (1986).

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