Erratum: A Microcalorimetric Study of the Association of Hexacyanoferrate(II) Ions with Calcium and Magnesium Ions in Aqueous Solution

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Tables III and IV were inadvertently omitted from this article which appeared in J. Solution Chem. 10, 411 (1981). These tables should have appeared following page 415 of that article.

TABLE III
Results for the association of Ca^{2+} with $\text{Fe}(\text{CN})_6^{4-}$

Total ^a $\text{CaCl}_2 \times 10^2$ (mol dm ⁻³)	Total ^a $\text{K}_2\text{Fe}(\text{CN})_6 \times 10^2$ (mol·dm ⁻³)	^a $\text{CaFe}(\text{CN})_6^{2-} \times 10^2$ (mol·dm ⁻³)	Initial ^b $\text{KFe}(\text{CN})_6^{3-} \times 10^2$ (mol dm ⁻³)	Final ^a $\text{KFe}(\text{CN})_6^{3-} \times 10^2$ (mol dm ⁻³)	Q (mW)	I ^c (mol dm ⁻³)	$\Delta H^\circ \text{CaFe}(\text{CN})_6^d$ (kJ mol ⁻¹)	
							A	B
0.5102	0.5062	0.1410	0.2443	0.1718	0.1030	0.0478	12.9 ₀	12.4 ₆
					0.1025		12.8 ₅	12.4 ₂
0.7625	0.5062	0.1796	0.2443	0.1494	0.1263	0.0531	12.8 ₉	12.3 ₉
					0.1260		12.8 ₇	12.3 ₇
0.5102	0.7706	0.1433	0.4054	0.3239	0.1130	0.0679	13.3 ₅	12.9 ₅
					0.1123		13.2 ₉	12.8 ₉
0.7625	0.7706	0.1917	0.4054	0.2941	0.1473	0.0728	13.4 ₅	12.9 ₉
					0.1470		13.4 ₃	12.9 ₇
0.7625	0.4261	0.1705	0.1973	0.1109	0.1168	0.0474	12.7 ₂	12.2 ₁
					0.1163		12.6 ₈	12.1 ₈
0.5102	0.4261	0.1369	0.1973	0.1298	0.0975	0.0418	12.7 ₂	12.2 ₇
0.7625	0.3257	0.1525	0.1404	0.0682	0.1005	0.0405	12.4 ₆	11.9 ₄
					0.0998		12.4 ₀	11.8 ₈
0.5102	0.3257	0.1270	0.1404	0.0819	0.0858	0.0344	12.2 ₇	11.8 ₀
					0.0850		12.1 ₉	11.7 ₂
0.2540	0.3257	0.080 ^a	0.1404	0.1046	0.0580	0.0295	12.3 ₂	11.9 ₄
0.6019	0.8570	0.1618	0.4594	0.3644	0.1283	0.0762	13.4 ₂	13.0 ₀
0.3005	0.8570	0.0912	0.4594	0.4072	0.0748	0.0711	13.1 ₇	12.8 ₉
0.2540	0.8570	0.0785	0.4594	0.4147	0.0655	0.0705	13.2 ₁	12.9 ₇

^a Concentrations are those for solution III

^b Concentration is for solution IV

^c Ionic Strength of solution III

^d See text

TABLE IV
Results for the association of Mg^{2+} with $\text{Fe}(\text{CN})_6^{4-}$

Total ^a $\text{MgCl}_2 \times 10^2$ (mol dm ⁻³)	Total ^a $\text{K}_4\text{Fe}(\text{CN})_6 \times 10^2$ (mol dm ⁻³)	^a $\text{CaFe}(\text{CN})_6^{2-} \times 10^2$ (mol dm ⁻³)	Initial ^b $\text{KFe}(\text{CN})_6^{3-} \times 10^2$ (mol dm ⁻³)	Final ^a $\text{KFe}(\text{CN})_6^{3-} \times 10^2$ (mol dm ⁻³)	^c μ (mM)	^c I (mol dm ⁻³)	$\Delta H^\circ \text{MgFe}(\text{CN})_6$ (kJ mol ⁻¹)	
							A	B
0.7840	0.4395	0.2142	0.2051	0.1007	0.1473	0.0463	13.3 ₀	12.7 ₅
					0.1493		13.4 ₂	12.8 ₇
0.5245	0.4395	0.1734	0.2051	0.1227	0.1225	0.0409	13.2 ₇	12.7 ₃
0.2611	0.4395	0.1053	0.2051	0.1568	0.0768	0.0371	13.1 ₆	12.6 ₈
					0.0755		13.0 ₀	12.5 ₁
0.5245	0.5258	0.1810	0.2559	0.1666	0.1298	0.0472	13.4 ₄	12.8 ₉
0.2611	0.5258	0.1064	0.2559	0.2051	0.0795	0.0437	13.4 ₂	12.9 ₄
					0.0780		13.2 ₃	12.7 ₅
0.7840	0.5258	0.2296	0.2559	0.1398	0.1625	0.0521	13.5 ₉	13.0 ₂
0.5245	0.5470	0.1824	0.2686	0.1778	0.1303	0.0487	13.4 ₁	12.8 ₆
0.7840	0.5470	0.2327	0.2686	0.1500	0.1663	0.0536	13.6 ₉	13.1 ₂
0.2611	0.5470	0.1065	0.2686	0.2173	0.0788	0.0453	13.3 ₂	12.8 ₄
					0.0790		13.3 ₄	12.8 ₆
2.088	0.5470	0.3288	0.2686	0.0891	0.2140	0.0875	13.8 ₆	13.4 ₇
0.5246	0.9326	0.1901	0.5071	0.4022	0.1478	0.0777	14.1 ₆	13.5 ₇
0.2611	0.9326	0.1039	0.5071	0.4510	0.0838	0.0747	14.0 ₅	13.5 ₆
0.7841	0.9326	0.2584	0.5071	0.3618	0.1998	0.0816	14.4 ₀	13.7 ₇
					0.1990		14.3 ₆	13.7 ₃
0.5246	0.2470	0.1359	0.0979	0.0419	0.0908	0.0279	12.5 ₆	12.0 ₄
0.2611	0.2470	0.0938	0.0979	0.0606	0.0660	0.0226	12.5 ₃	12.0 ₆
0.7841	0.2470	0.1544	0.0979	0.0328	0.1008	0.0346	12.7 ₅	12.2 ₄

^a Concentrations are those for solution III

^b Concentration is for solution IV

^c Ionic Strength for solution III