

Investigation of moisture stability and PL characteristics of terpineol-passivated organic–inorganic hybrid perovskite

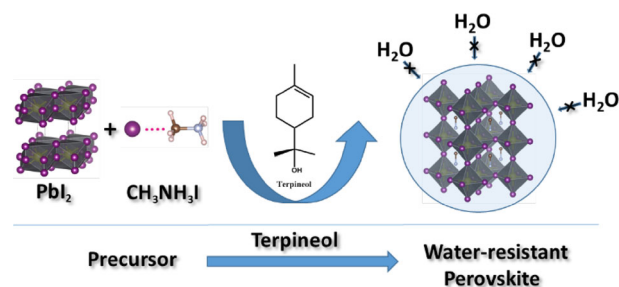
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Abstract This work presents a novel method for preparing perovskite films using a simple processing technique. Perovskite paste was prepared by dispersing an equimolar mix of PbI_2 and methyl ammonium iodide powders into terpineol with stirring. From these precursors, perovskite films were fabricated using doctor blading and drying for 24 h at room temperature. The prepared films were then placed into relative humidity (RH) levels of 30, 50, and 70 % to test the moisture stability. The crystal structure, phases, and morphology were investigated with XRD and SEM/EDX. These samples exhibited good stability against long time exposure to moisture for 70 days. The XRD results showed that samples stored at RH 70 % contained only a small amount of hydrate compound after 70 days storage, while in the sample stored at RH 50 %, the formation of PbI_2 was observed. The sample at RH 30 % manifested almost no change when stored for the same

storage period. We attribute the enhanced moisture stability, compared with the spin-coated samples, to a passivated surface of the perovskite film by terpineol which exhibits a hydrophobic moiety. Time-resolved photoluminescence measurements show that the passivation of surface defect states by the formation of either PbI_2 or hydrated compound leads to prolonged charge carrier recombination times.

Graphical Abstract



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Keywords Perovskite paste · Terpineol passivation · Moisture stability · Humidity degradation · PL decay lifetimes

Introduction

Over the last few years, solar cells based on organic–inorganic hybrid semiconductor materials with perovskite structure as light absorbers have been quickly progressing due to their excellent photovoltaic performance [1–3]. Perovskites were first utilized in dye sensitized solar cells with the perovskite being a replacement sensitizer for the standard molecular dyes. Kojima et al. reported the power

Table 1 PL decay lifetimes of fresh and aged perovskite films

Sample	Day	τ_1/ns	τ_2/ns
Fresh	0	–	16.9 ± 0.1
30 % RH	3	3.0 ± 0.1 (0.82)	13.0 ± 1.0 (0.18)
	10	4.2 ± 0.3 (0.78)	14.0 ± 1.0 (0.22)
	30	4.8 ± 0.3 (0.75)	15.0 ± 1.0 (0.25)
	30	4.2 ± 0.2 (0.82)	15.0 ± 1.0 (0.18)
50 % RH	10	3.7 ± 0.2 (0.87)	15.0 ± 1.0 (0.13)
	30	7.0 ± 0.2 (0.84)	36.0 ± 2.0 (0.16)
	30	5.1 ± 0.1 (0.81)	22.0 ± 1.0 (0.19)
70 % RH	3	3.8 ± 0.1 (0.89)	17.0 ± 2.0 (0.11)
	10	3.7 ± 0.1 (0.87)	18.0 ± 2.0 (0.13)
	30	5.1 ± 0.1 (0.81)	22.0 ± 1.0 (0.19)

conversion efficiency (PCE) of a $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite solar cell (PSC) of 3.81 %. However, incorporating perovskite into a solid-state architecture has greatly improved efficiencies, and to date, the highest certified PCE is 20.1 % [4]. This class of materials is advantageous for solar-cell applications, particularly with respect to their low processing cost, simple and various synthetic methods, and high power conversion efficiency [4–7]. These organic–inorganic hybrid semiconductors can be fabricated from relatively inexpensive materials by a simple solution process without annealing at high temperature, and their cost could be considerably reduced through mass production [8–11]. Moreover, solar cells based on metal halide perovskites have multiple advantages: excellent optical properties with tunable light absorption by managing the chemical composition, low exciton-binding energy with the excited state composed primarily of free carriers, fast photogeneration, ambipolar charge transport, excellent charge carrier mobility, and long electron–hole diffusion lengths [12–17].

Although efficiency progress and photovoltaic performance of perovskite-based solar devices have been remarkable, the lead–halide perovskites have been suffering due to their instability when exposed to atmospheric moisture. The intercalated polar methyl ammonium ion, CH_3NH_3^+ , in the perovskite has a tendency to diffuse inside the perovskite framework with the aid of a polar solvent like water. The $\text{CH}_3\text{NH}_3\text{PbI}_3$ perovskite films will easily degrade into other chemical species when it is exposed to water (humidity), resulting in an intrinsic problem, regarding long-term stability [18–22]. Only limited research has been conducted on enhancing the stability and requires immediate attention if perovskite-based photovoltaic devices are to be used commercially.

It has been shown that a passivation layer is a good strategy to provide an enhanced stability for perovskites.

Recent published work [23] has also confirmed that PbI_2 with a bulk band gap of 2.3 eV can form a passivation layer in the unfilled grain boundaries of perovskite films. The slower injection rate and longer recombination times of photoinduced charge carriers demonstrate the effect of passivation. The post-modification of perovskite by aluminum oxide reported by Wang et al. [21] could both protect the perovskite sensitizer from degradation and effectively suppress electron recombination. Snaith et al. have also reported significantly reduced electron–hole recombination and longer carrier lifetime $\sim 2 \mu\text{s}$ within the $\text{CH}_3\text{NH}_3\text{PbI}_{3-x}\text{Cl}_x$ perovskite through the treatment of the crystal surfaces with the Lewis bases thiophene and pyridine [24]. It is found that there is an excess of charge building up by under-coordinated Pb atoms at crystal surfaces and grain boundaries. The electronic passivation by forming coordinate or dative bonds between unsaturated Pb ions and Lewis bases could effectively mitigate the charge build-up at defect sites and enhance the stability of the perovskite crystals along with power conversion efficiencies for the Lewis Base treated solar cells. A recent study by Karunadasa et al. demonstrated the enhanced moisture stability of a layered 2D hybrid perovskite [25], in which phenylethylamine (PEA) not only acts as a molecule for tuning the spacing between the perovskite layers toward higher bandgap, but also provides the perovskite absorber with a hydrophobic surface, which leads to increased durability. Yong et al. introduced a simple and effective passivation method to overcome water vulnerability for the perovskite solar cell [26]. By coating a hydrophobic polytetrafluoroethylene (PTFE) layer onto the top surface of the cell, water diffusion into the perovskite cell was effectively blocked and the solar cell exhibited a significantly enhanced stability under the ambient atmosphere conditions.

In this work, we introduce a new hydrophobic molecule in a new passivation process. By blading the perovskite paste, the modified perovskite film presented significantly enhanced moisture stability. Water molecule penetration into the perovskite crystal lattice was prevented chemically by the hydrophobic repulsion at the molecular level and physically by the layer encapsulation on a macroscopic scale. Through the hydrophobic passivation, a good carrier lifetime measured by time-resolved photoluminescence (TRPL) decay indicates less trap states formation and good electron transport with terpineol modification. We also present different degradation processes of samples at three different humidity levels, and traced it by the XRD analysis. SEM/EDX analysis was used to confirm morphology and compositional change during phase degradation.

Experimental methods

Materials

Methyl ammonium iodide (MAI) powders were synthesized by reacting a concentrated aqueous solution of hydroiodic acid (57 wt%, Aldrich) with methylamine (33 wt% in methanol) in an ice bath for 2 h with stirring. The mixture was dried at 65 °C for 30 min using a rotary evaporator. The precipitant obtained was washed three times with diethyl ether and dried in a vacuum oven at 60 °C for 24 h. Commercial powders of PbI_2 and α -Terpineol were used as received from Sigma-Aldrich.

Synthesis of MAPbI_3

MAPbI_3 perovskite film samples were prepared by doctor blading the perovskite paste. The mixture of PbI_2 and MAI with a molar ratio of 1:1 was first dispersed in α -Terpineol to form a 2.5:1 paste. The perovskite paste was finally produced by magnetic stirring for 3 days. The bladed MAPbI_3 perovskite films were dried at room temperature for 24 h.

Moisture stability of MAPbI_3 film

The stability of the bladed perovskite thin films on the glass substrates was tested by placing samples at three different RH levels: 30, 50, and 70 % for 70 days. The degraded conditions were tracked in 3, 10, 30, and 70 days using X-ray diffraction and scanning electron microscopy (SEM). The carrier lifetime of the 3, 10, and 30 days samples at each RH level was measured by time-resolved photoluminescence spectroscopy.

Materials characterization

X-Ray diffraction (XRD) patterns were obtained with a Rigaku MiniFlex instrument Cu $K\alpha$ beam ($\lambda = 1.54 \text{ \AA}$). SEM images and energy dispersive X-ray spectroscopy (EDX) spectra were acquired by Phenom ProX. Thermo Nicolet NEXUS 870 FT-IR spectrometer was used to collect the FT-IR spectral data for terpineol-modified perovskite films in the $4000\text{--}500 \text{ cm}^{-1}$ range.

TRPL measurements

Samples were excited using a Clark-MXR CPA 2001 laser that outputs 780 nm light with a pulse width of 150 fs at a rate of 1000 pulses/second. This light is passed through a Light Conversion TOPAS to generate the 480 nm excitation wavelength. TRPL kinetics was obtained using an Optronis streak camera.

Results and discussions

Moisture stability of terpineol-modified perovskite film

The moisture stability of the perovskite film at RH 70 % for 3, 10, 30, and 70 days was investigated by tracking the phase changes using the XRD analysis. The XRD patterns of the sample at five different times were labeled for a comparison with the calculated XRD pattern using the MDI Jade software, as shown in Fig. 1.

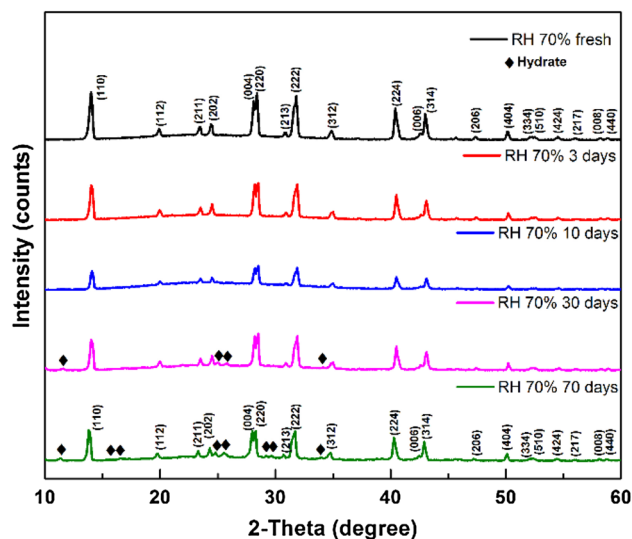


Fig. 1 XRD patterns of fresh perovskite film and stored at RH 70 % for 3, 10, 30, and 70 days. The hydrate phase is present for the film at 30 days and increases slowly for 70 days storage

Showing that typical peaks at 14.10° , 23.47° , 28.42° , and 30.89° corresponding to miller planes (110), (211), (220), and (213) of the tetragonal perovskite phase are present in all the cases. The peak at 28.15° also becomes visible which corresponds to (004). Before 10 days, there are almost no impurity phases detected, indicating the strong moisture resistance of the terpineol-modified perovskite film. There are also unidentified peaks at 11.34° , 16.70° , 17.32° , 24.74° , 25.56° , etc. that showed up in 30 and 70 days. The unidentified peaks did not match either the pure PbI_2 or MAI tetragonal phase, which is stable at room temperature. This indicated a possible formation of the hydrate intermediate compound and by comparing with the calculated XRD data of $(\text{CH}_3\text{NH}_3)_4\text{PbI}_6 \cdot 2\text{H}_2\text{O}$, those peaks could be identified [33]. This suggests that the water molecules could penetrate into the crystal lattice under high vapor pressure and produce the hydrate compound at 70 % humidity level. A slow growth of the hydrate phase inside the film reflects not only the moisture stability of the terpineol-modified perovskite but also the hydrate phase at the sub-surface can reach equilibrium with both the water vapor at the top surface and the inner perovskite phase resulting in difficulty of the further growth of hydrate compound. For comparison to perovskite films prepared using the more common spin-coating technique, samples were prepared by spin-coating perovskite on to glass and were stored under RH 70 %. In stark contrast, the spin-coated sample placed under RH 70 % was already degraded into the PbI_2 phase and hydrate phase in only 1 day and almost completely degraded in 3 days, as shown in Figure S2.

The second set of experiments was the moisture stability of the perovskite film under RH 50 % for the same time scale as above. The XRD patterns of the sample in five different times are shown in Fig. 2 and were labeled for a comparison with the calculated XRD patterns of perovskite phases.

It can be found that almost no impurity phases existed before 10 days, indicating the strong moisture resistance of the terpineol-modified perovskite film. After that, the lead iodide phase appears only after 30 and 70 days storage at 50 % humidity level. The formation of the PbI_2 phase inside the film originates from the diffusion of intercalated polar CH_3NH_3^+ away from the inorganic PbI_2 cage with the aid of water molecules. The fact that the PbI_2 phase grows slowly inside the film which also provides evidence for the moisture stability of the terpineol-modified perovskite.

The perovskite film exhibited excellent stability when stored at 30 % RH in the same time scale. The XRD patterns of the sample stored for five different times are shown in Fig. 3, displaying that no detectable impurity phases existed inside the perovskite film after storage for 70 days.

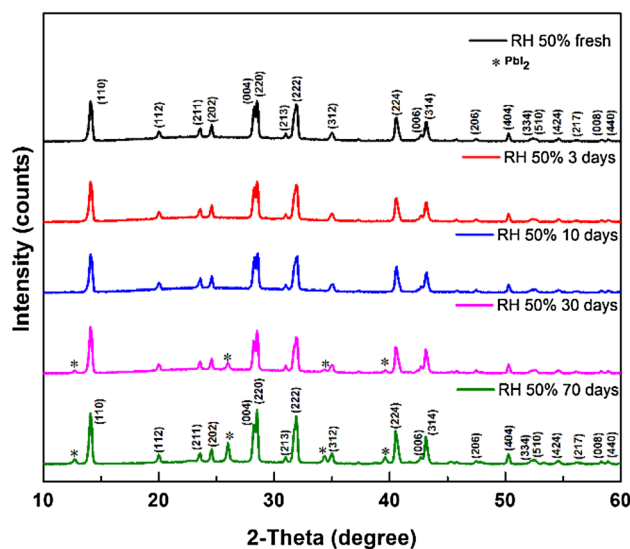


Fig. 2 XRD patterns of fresh perovskite film and stored at RH 50 % for 3, 10, 30, and 70 days. The PbI_2 phase is present in the film stored for 30 days and increases slowly after 70 days storage

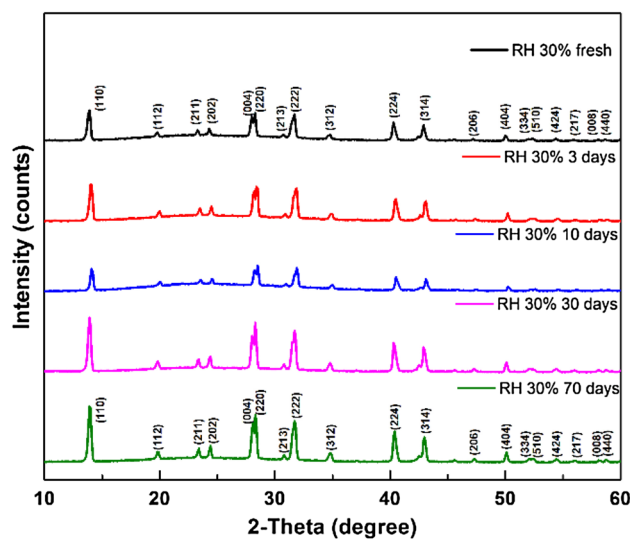


Fig. 3 XRD patterns of fresh perovskite film and stored at RH 30 % for 3, 10, 30, and 70 days. Almost no impurity phases are present in the film in all cases

The XRD patterns of samples under RH30, 50, 70 % for 10, 30, and 70 days, respectively, were shown in Figure S3–S5.

The SEM and EDX can also confirm the XRD analysis. The SEM images and EDX spectrum of the fresh sample are shown in Fig. 4. The perovskite grains were observed to be polygons with sharp outlines and facets in different sizes. The EDX spectrum of the overview region shows the composition of perovskite. There are slightly larger deviation errors with the concentration of C and N probably due to the light element and small electron scattering cross section.

The SEM and EDX elemental mapping of perovskite film after 70 days storage at RH70 % are shown in Fig. 5. The morphology of the fresh perovskite grains changed from sharp outlines to a more rounded shape like melting after long storage time under 70 % moisture atmosphere. The grain size on average also becomes smaller. Through a square scan of the EDX mapping on the bulk segregation, it can be found that the oxygen exists in the film which is further evidence for the presence of the hydrate compound detected in the XRD measurement.

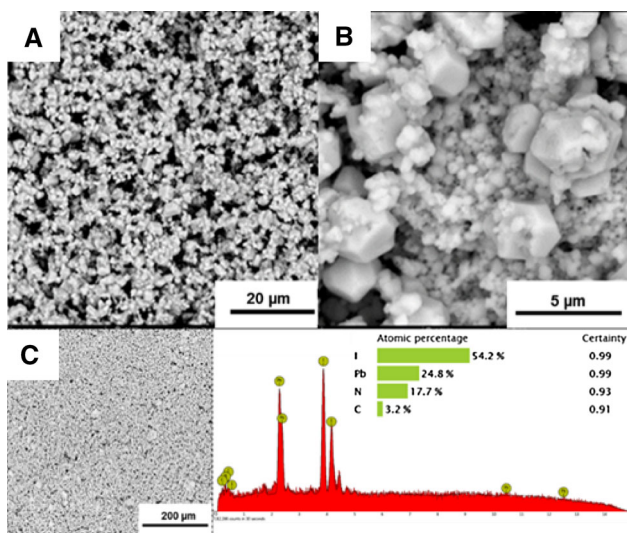


Fig. 4 SEM images and EDX spectrum of the fresh perovskite film, a at 20 μm scale, and b at 5 μm scale, and c EDX spectrum acquired by overview scanning of image c, scale bar is 200 μm

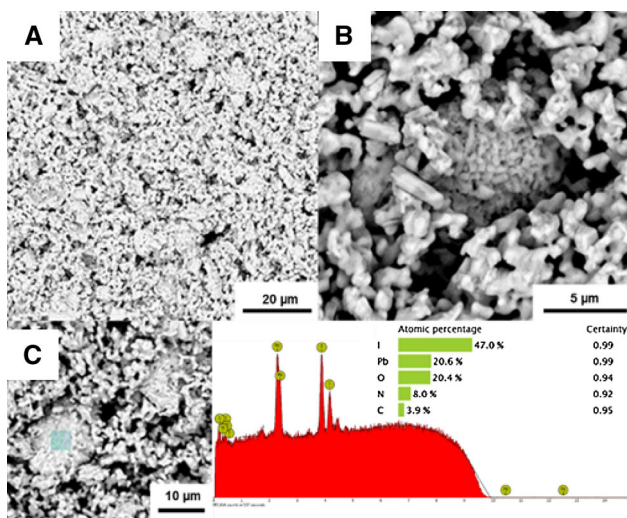


Fig. 5 SEM images and EDX spectrum of the degraded regions in the perovskite film stored at RH 70 % for 70 days, a at 20 μm scale, and b at 5 μm scale, and c EDX spectrum acquired in a square area of image c, scale bar is 10 μm

Similar observation of the sample stored at RH 50 % for 70 days is shown in Fig. 6. The morphology of the perovskite grains changed from sharp outlines for the fresh state to a less defined shape after long storage time under 50 % moisture atmosphere. Through a rectangle scan of the EDX mapping on the honeycomb bulk body, only elements Pb and I can be found that well confirm the presence of surface rich PbI_2 phase detected in the XRD measurement.

The SEM images and EDX spectrum of the sample stored at RH 30 % for 70 days are shown in Fig. 7. There

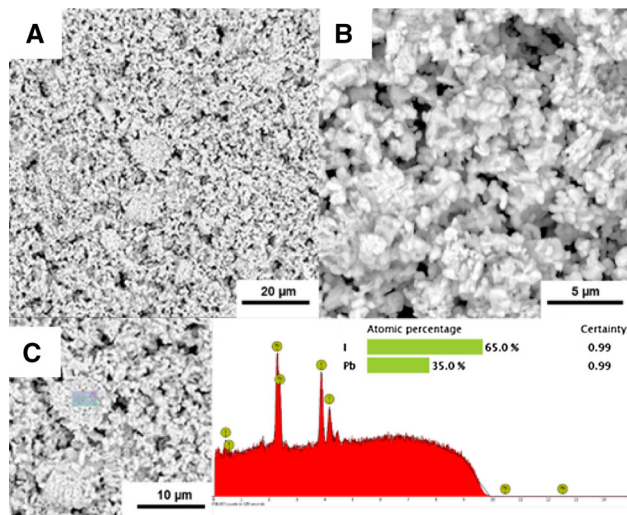


Fig. 6 SEM images and EDX spectrum of the degraded regions in the perovskite film stored at RH 50 % for 70 days, a at 20 μm scale, and b at 5 μm scale, and c EDX spectrum acquired in a square area of image c, scale bar is 10 μm

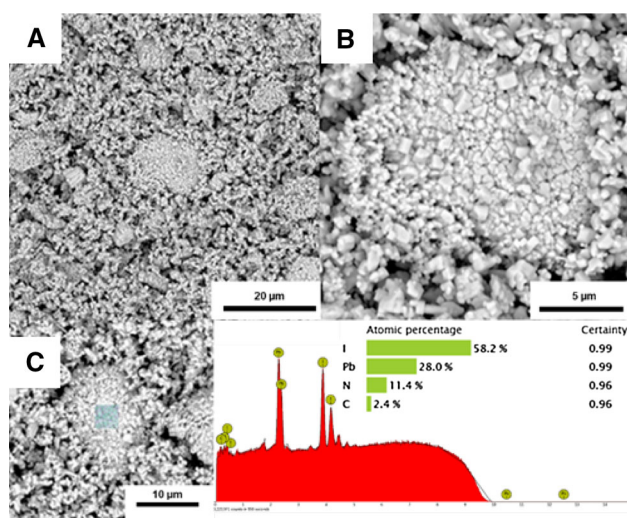


Fig. 7 SEM images and EDX spectrum of the perovskite film stored at RH 30 % for 70 days, a at 20 μm scale, and b at 5 μm scale, and c EDX spectrum acquired in a square area of image c, scale bar is 10 μm

are no major differences in the grain morphology observed between the sample stored at RH 30 % and the fresh one.

The constituents of the paste evolved during the stirring process before the final formation of perovskite phase. From the XRD patterns in Figure S6, the black perovskite paste is not pure and contains the intermediate phase and PbI_2 in the first few hours. The intermediate phase might be the coordination compound of terpineol and MAI. With the longer stirring time, the tetragonal perovskite phase co-existed with only PbI_2 phase in 1 day and became very pure with almost no PbI_2 observed after 3 days of stirring.

The presence of terpineol in the perovskite film and the interaction between them are further confirmed from the Fourier transform-infrared (FT-IR) spectrum. Figure 8 shows the IR spectra for pure terpineol solvent and the freshly prepared perovskite film made from the perovskite paste. It can be clearly observed that terpineol molecules were contained within the perovskite film, while the typical peaks of N–H and C–H vibrations for perovskite can also be identified. It is worth noting that the stretching vibration of O–H in terpineol appeared at 3374 cm^{-1} is shifted to lower wavenumber of 3350 cm^{-1} . The shift of the O–H peak of terpineol is due to an interaction between the –OH group and Lewis acids of MA^+ ions in the perovskite film [27]. In the harmonic approximation using a diatomic model, the frequency of vibration is proportional to the square root of the force constant [28]. Thus, the red shift of

O–H stretching frequency indicates a lower force constant, which is due to a decreased bond strength between oxygen and hydrogen as a consequence of the adduct formation and interactions with MA^+ or I^- ions. This observation further suggests the effective passivation of terpineol in this system.

Transient PL measurement for carrier lifetime

The carrier lifetime is an important parameter for developing perovskite-based optoelectronic devices. It has a direct relation to the defect formation and carrier transport properties, and thus reflects the materials quality. To prove that the terpineol-passivated perovskite film has a good stability without a compromise of good carrier lifetimes, time-resolved photoluminescence measurements were employed. Figure 9 shows the kinetic decay curves monitored at λ_{PL} of 780 nm. The PL decay of the fresh perovskite film could be fitted with a single exponential decay equation and a PL lifetime of 16.9 ns was attained. All samples that were exposed to humidity were fit with a multi-exponential decay equation. The fact that two time constants were required to fit the decay curve suggests that there is some decay in these films after being exposed to moisture. However, based on the structural analysis above, this degradation should be minimal. All samples stored under a humid atmosphere, except for two, which show relatively the same time constants. The fast component ranged from 3 to 5 ns and the slower component ranged from 13 to 18 ns. The longer time constants increased with increased percentages of humidity. The two exceptions to the lifetime trend above are the films stored under 50 and

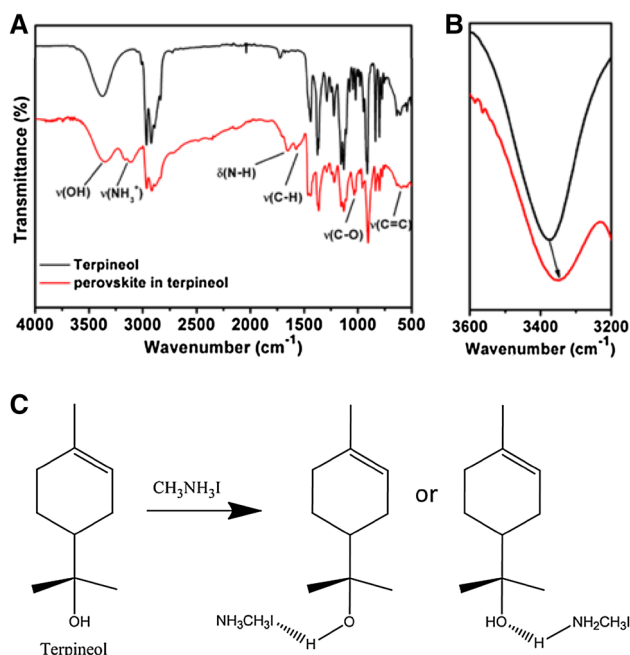


Fig. 8 Fourier transform-infrared spectra (FT-IR) of **a** pure terpineol and perovskite film made from the perovskite paste, **b** expanded fingerprint region for the O–H vibrations, and **c** schematic of possible interactions between MAI and terpineol along with their chemical structures

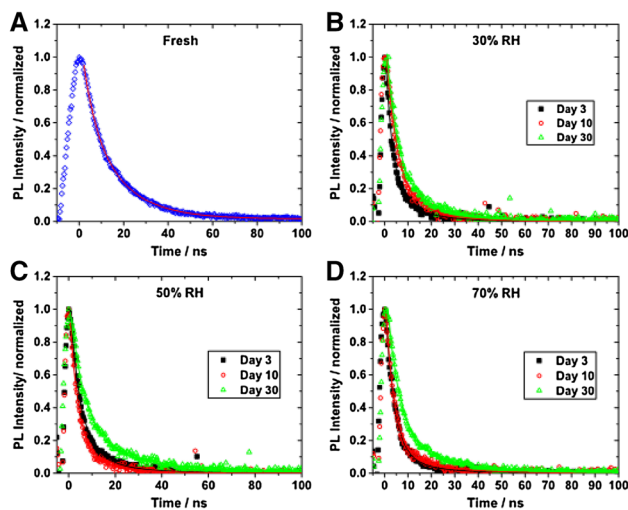


Fig. 9 **a** TRPL decay dynamics of the freshly prepared perovskite film. **b–d** TRPL decay dynamics of perovskite films prepared under 30, 50, and 70 % relative humidity, respectively. Films were stored for 3 (black squares), 10 (red circles), and 30 (green triangles) days. PL decay lifetimes are reported in Table 1

70 % humidity for 30 days. The sample stored at 50 % humidity for 30 days shows the evidence of degradation in the XRD and this leads to the formation of PbI_2 on the surface of the perovskite grains. PbI_2 has been shown to act as a passivating layer that can prolong charge carrier recombination [23]. Here, we find that the lifetime is extended from 15 to 36 ns when PbI_2 is present in the film. The sample stored at 70 % RH for 30 days shows longer charge carrier recombination times, as well. This result in conjunction with the XRD analysis indicates that the hydrate phase also acts as passivating layer which leads to an increase of the PL decay time from 17 ns to 22 ns when the hydrate is present. This effect may also explain why an increase of the PL lifetime is observed when the perovskite films are stored under more humid atmospheres. Typically, the surface modifications of perovskite are performed during the solution preparation or annealing step [29, 30]. However, the findings of this work show that when minimal surface modifications occur due to degradation, surface states can become passivated leading to prolonged charge carrier recombination lifetimes. The corresponding time-resolved photoluminescence contours are shown in Figure S7 and S8.

Conclusions

We demonstrated a novel method to produce perovskite paste by dispersing equimolar mixed PbI_2 and MAI powders into the terpeneol by way of stirring. The perovskite film prepared by blading the paste showed very good moisture stability. By placing the samples at three different moisture levels, it can be found that all three samples exhibit good stability at long time exposure to moisture for 70 days. The XRD and SEM/EDX results showed that there is a small amount of hydrate compound in the sample stored at RH 70 % for 70 days, while the PbI_2 as the degraded phase was observed in the sample stored at RH 50 %. However, the sample stored at RH 30 % exhibited almost no change in the meantime. The modified perovskite film by hydrophobic terpeneol molecules presented significantly enhanced moisture stability. Water diffusion into perovskite crystal lattice was effectively blocked chemically by the hydrophobic repulsion at the molecular level and physically by the layer encapsulation on a macroscopic scale. Time-resolved photoluminescence spectroscopy revealed that there was some degradation of perovskite films stored under the humid conditions. However, XRD and SEM reveal that this degradation is quite minimal. In the most humid atmospheres, prolonged charge carrier recombination lifetimes due to the passivation of surface states by the formation of either a PbI_2 or a hydrated perovskite rich surface are observed.

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