

Photocatalytic activity of biogenic silver nanoparticles synthesized using yeast (*Saccharomyces cerevisiae*) extract

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Received: 23 November 2014 / Accepted: 17 December 2014 / Published online: 31 December 2014
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Abstract Synthesis of metallic and semiconductor nanoparticles through physical and chemical route is quiet common but biological synthesis procedures are gaining momentum due to their simplicity, cost-effectivity and eco-friendliness. Here, we report green synthesis of silver nanoparticles from aqueous solution of silver salts using yeast (*Saccharomyces cerevisiae*) extract. The nanoparticles formation was gradually investigated by UV–Vis spectrometer. X-ray diffraction analysis was done to identify different phases of biosynthesized Ag nanoparticles. Transmission electron microscopy was performed to study the particle size and morphology of silver nanoparticles. Fourier transform infrared spectroscopy of the nanoparticles was performed to study the role of biomolecules capped on the surface of Ag nanoparticles during interaction. Photocatalytic activity of these biosynthesized nanoparticles was studied using an organic dye, methylene blue under solar irradiation and these nanoparticles showed efficacy in degrading the dye within a few hours of exposure.

Keywords Biogenic silver nanoparticles · Yeast (*Saccharomyces cerevisiae*) extract · UV–Vis spectroscopy · XRD · TEM · Photocatalytic activity

Introduction

Synthesis of metallic and semiconductor nanoparticles through various routes is a primary step for nano-research. The size, shape and morphology of the nanoparticles can be varied and modified to obtain enhanced property for suitable applications (Li et al. 2011a, b). The nano-sized particles possess large number of surface atoms that contribute to large surface area. Due to high surface area, these particles exhibit improved electronic (Löberg et al. 2013; Zaniewski et al. 2013), optical (Kelly et al. 2003) and catalytic (Bastús et al. 2014) properties than bulk materials. The noble metallic nanoparticles may be incorporated for numerous applications in different fields like electronics (Tran et al. 2013), microscopy (Lee and El-Sayed 2006), biomedicines (Huang et al. 2011; Satyavani et al. 2011) and textile (Zhang et al. 2009). In textile and paper industry, recently silver nanoparticles are used to degrade the organic dyes as they exhibit enhanced photocatalytic property for degrading organic dyes under solar radiation (Kumar et al. 2013). Recent reports suggest that the removal of organic dyes using Ag nanoparticles is a better choice than the common dye removal techniques like redox treatment (Kim and Hensley 1997), electro-coagulation (Szyrkowicz 2005), carbon sorption (Gomez et al. 2007) and UV photodegradation (Comparelli et al. 2005; Sohrabi and Ghavami 2008).

In this paper, we studied the photocatalytic degradation of methylene blue dye in presence of biogenic silver nanoparticles synthesized using yeast (*Saccharomyces cerevisiae*) extract. The biological route was chosen over common chemical and physical procedures because it is an easy, fast and eco-friendly option and does not involve any costly instrument or hazardous by-product as well (Awwad et al. 2013). In recent times, plant-mediated (Banerjee et al.

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2014; Roy et al. 2013; Phanjom et al. 2012) and microbe-mediated (Sudha et al. 2013; Malarkodi et al. 2013; Honary et al. 2013) synthesis procedures are frequently used for nanoparticle synthesis. But in particular, the microbe-mediated synthesis methods involve various isolation and culture techniques which are multi-step complicated procedures and time-consuming (Li et al. 2012; Bonde et al. 2012). For biosynthesis of Ag nanoparticles, we used the powder of yeast (*Saccharomyces cerevisiae*) extract which is readily available in market and contains the bioactive molecules that are able to reduce silver ions and stabilize the colloidal particles in the medium (Mosser et al. 2011; Li et al. 2011a, b). Here, we have successfully synthesized silver nanoparticles from AgNO₃ solution using yeast (*Saccharomyces cerevisiae*) extract and these nanoparticles have shown the efficiency in degrading methylene blue under solar irradiation.

Materials and methods

Materials

Extract powder of yeast (*Saccharomyces cerevisiae*) was purchased from Merck Specialties Pvt. Ltd. India. Pure silver nitrate (AgNO₃) and methylene blue (powder) used here were the products of Merck India Ltd.

Methods

Synthesis of silver nanoparticles

To prepare yeast (*Saccharomyces cerevisiae*) extract solution of 1 g/l concentration, 0.1 g dry powder of the extract was dissolved in 100 ml de-ionized (DI) water. 0.34 g silver nitrate was dissolved in 100 ml DI water to prepare 20 mM stock solution of AgNO₃ for further use. To prepare silver nanoparticles, the silver ions present in the AgNO₃ solution was reduced using solution of yeast (*Saccharomyces cerevisiae*) extract. The 100 ml solution of yeast extract was added dropwise to the 100 ml aqueous solution of AgNO₃ resulting into a half diluted mixture (10 mM) of silver nitrate with exposure to the fungal extract. Formation of silver nanoparticles at room temperature was visually detected by gradual change in the color of the mixture (shown in Fig. 1) and later investigated by UV–Vis spectroscopy. When reaction completed (after 24 h of observation), the mixture was centrifuged at 10,000 rpm for 30 min to separate colloidal particles from the mixture. After centrifugation, the supernatant was decanted and the obtained precipitation was dissolved in 20 ml DI water for further centrifugation at 5,000 rpm for 10 min. In this way, maximum biomass residue was

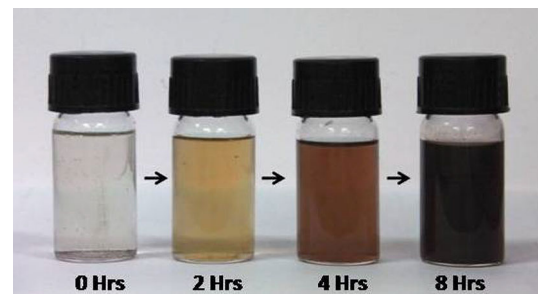


Fig. 1 Color change with reaction time

removed from the surface of nanoparticles and the level of purity of these silver nanoparticles was enhanced. At last, the pellet of nanoparticles formed at the bottom of centrifuge tube was collected and dried inside a vacuum dryer to obtain dry powder of biosynthesized Ag nanoparticles.

Characterization of silver nanoparticles

The formation of silver nanoparticles in the mixture was studied by measuring the absorption spectra of the mixture at regular interval using Perkin Elmer UV–Vis spectrometer (USA). The XRD analysis of dry silver nanoparticles was performed by Rigaku Ultima-III X-ray diffractometer ($\lambda = 0.154$; operating voltage 40 kV). FTIR study of these biosynthesized nanoparticles was carried out using IR-Prestige FTIR spectroscope (Shimadzu). To prepare sample grid for TEM analysis, the dry powder of silver nanoparticles was suspended in small amount of DI water keeping a concentration of 50 $\mu\text{g/ml}$. A drop of this suspension was placed on carbon-coated copper grid and the grid was dried before scanning under high-resolution JEOL-2010 TEM (operating voltage 200 kV).

Study of photocatalytic activity of Ag nanoparticles

The photocatalytic activity of these biosynthesized Ag nanoparticles was studied by degradation of methylene blue under sunlight irradiation. Initially, the dye solution was prepared by dissolving 1 mg powder of methylene blue in 100 ml DI water (keeping 10 mg/l concentration). Around 10 mg of silver nanoparticles was added to 50 ml dye solution of methylene blue and the mixture was stirred magnetically for 45 min in darkness before exposing to sunlight. A control was prepared and kept under the similar condition for comparing any change in color of the dye solution. The colloidal suspension was then put under sunlight irradiation with constant stirring. The average temperature of the ambience during the experiment was found to be around 27 °C with 6 h mean shine duration (Kolkata: 22°32'N, 88°19'E; Date: 12th June, 2014). At frequent intervals (every 2 h), 2 ml suspension was taken

from the colloidal mixture and centrifuged at 5,000 rpm for 15 min to obtain clean supernatant soup of the tested dye. The soup was then scanned at different wavelength from 350 to 850 nm using the Perkin Elmer UV–Vis spectroscope (USA) to study the dye degradation in presence of Ag nanoparticles.

Result and discussions

UV–Vis spectroscopy

As the yeast extract powder is a rich source of functional organic molecules, it has the ability to reduce silver cations in the reacting solution. After a few hours of addition of yeast extract to the AgNO_3 solution, the color of the solution began to change from colorless to yellow indicating the formation of Ag nanoparticles in the mixture. The color of the solution gradually intensified with incubation time and turned into dark brown after 4 h as shown in Fig. 1. This change of solution color may be attributed to the surface plasmon resonance of Ag nanoparticles (Prakasha et al. 2013). The production of nanoparticles was monitored by scanning the mixture using a UV–Vis spectrometer at regular interval. The maximum absorbance noticed at 450 nm further confirmed the formation of nanoparticles in the mixture. The UV–Vis spectra obtained at different time intervals (2, 4, 8 and 16 h) are shown in Fig. 2. The maximum absorbance was found to be increasing with incubation time and its variation with time is shown in Fig. 2 (inset). It is clear from the inset figure that the absorbance continued to increase almost linearly up to 8 h of incubation perhaps due to formation of more number of nanoparticles in the reacting mixture. After 16 h

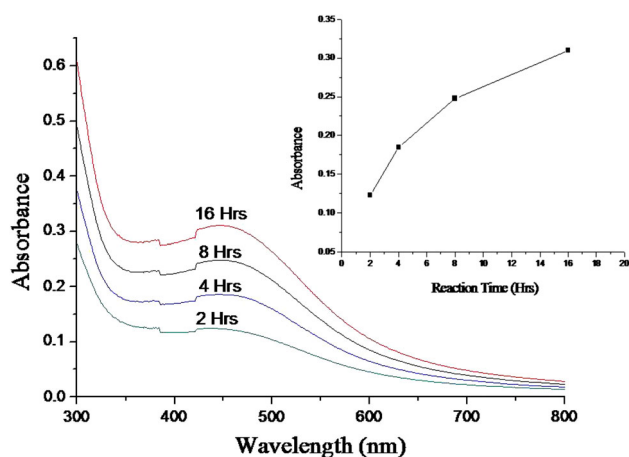


Fig. 2 UV spectra of the reacting mixture at different time intervals. Inset curve indicates the variance of maximum absorbance with incubation time

of incubation, the formation rate saturated indicating the completion of reaction.

XRD analysis

Crystallinity and different phases of biogenic nanoparticles were identified by analyzing X-ray diffraction pattern. The XRD pattern (shown in Fig. 3) consists of six noticeable peaks at $2\theta = 27.62^\circ$, 32.04° , 46.1° , 54.74° , 57.38° and 76.82° which, respectively, correspond to (220), (122), (231), (331), (241) and (311) planes of Ag (as correlated to JCPDS: File No. 4-783). This result shows that these nanoparticles are crystalline and face-centered cubic in nature.

TEM analysis

The size, shape and morphology of the biogenic silver nanoparticles were studied by transmission electron microscopy (TEM). The high-resolution TEM images (Fig. 4) show that the particles are almost spherical in shape with an average diameter of 10 nm. The lattice fringes denote high crystallinity with an interplanar spacing of 0.28 nm that can be attributed to the (122) planes of FCC phase of biosynthesized silver nanoparticles.

FTIR spectroscopy

Figure 5 shows the FTIR spectrum of the silver nanoparticles synthesized using yeast (*Saccharomyces cerevisiae*) extract. This characterization was done to detect the functional organic molecules which cause the production and capping of biogenic Ag nanoparticles in the mixture. The spectrum (in absorbance mode) consists of six

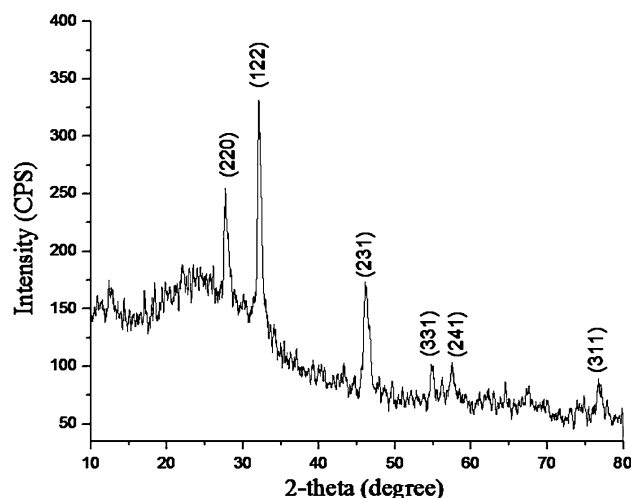


Fig. 3 X-ray diffraction curve of biosynthesized Ag nanoparticles

Fig. 4 TEM images of biogenic silver nanoparticles

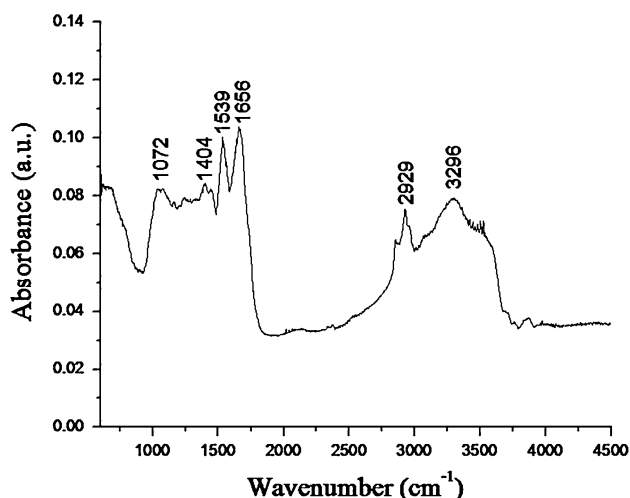
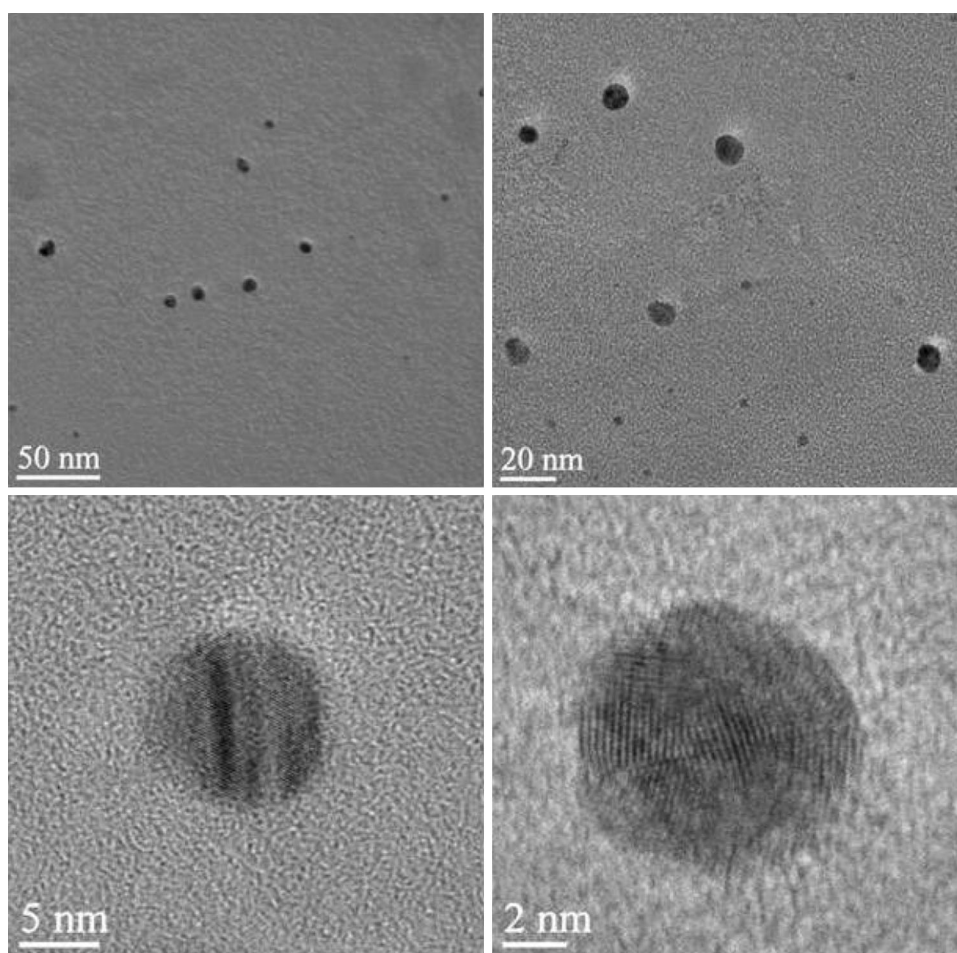


Fig. 5 FTIR spectrum of the biosynthesized silver nanoparticles

noticeable peaks between 500 and 4,500 cm^{-1} . Bands noticed at 1,656 and 1,404 cm^{-1} may be assigned to the C=O stretching vibration of amides and C–H bending of alkanes, respectively (Smith 1999). A band at 1,072 cm^{-1} may indicate stretching of C–N bonds that exist in amines

whereas the band at 3,296 cm^{-1} denotes the O–H stretching of aromatic compounds like phenol. (Isaac et al. 2013). The band observed at 2,929 cm^{-1} may correspond to the C–H stretching of aldehydes and a band at 1,539 cm^{-1} indicates the bending of C–H bonds present in hydrocarbons as well (Sathyavathi et al. 2010). From this study, it is clear that the functional molecules like aldehydes, amides and phenols present in yeast extract could possibly reduce the Ag^+ ions and stabilize the biosynthesized nanoparticles in the mixture.

Analysis of photocatalytic activity

Photocatalytic activity of the biosynthesized silver nanoparticles was evaluated by degradation of methylene blue under solar irradiation. Dye degradation was visually detected by gradual change in the color of the dye solution from deep blue to colorless (shown in Fig. 6). The characteristic absorption peak for methylene blue was noticed at 660 nm. The degradation of the dye in presence of biogenic Ag nanoparticles was verified by the decrease of the peak intensity (at 660 nm) during 6 h of exposure in sunlight (shown in Fig. 7).

Fig. 6 Gradual change in dye color from *blue* to *colorless* denotes dye degradation

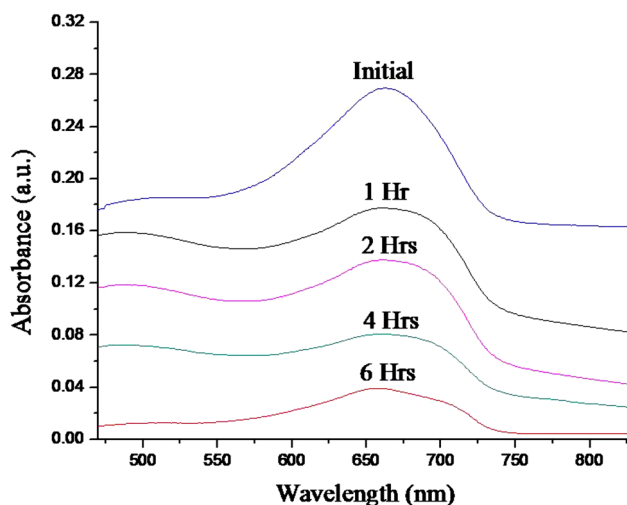
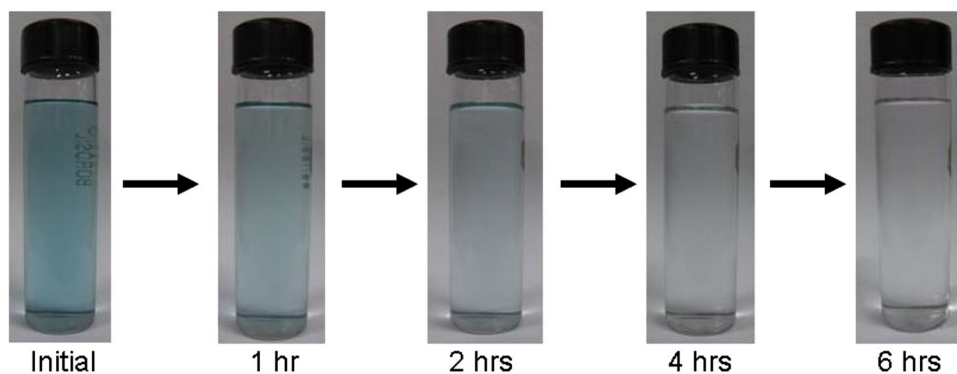


Fig. 7 UV spectra indicate photocatalytic degradation of methylene blue with reaction time

The control exhibited no change in coloration during exposure in sunlight. The dye degradation (%) was calculated using the following rule and its variance with the time of solar exposure is shown in Fig. 8.

$$\text{Dye degradation (\%)} = [(C_0 - C_t) / C_0] \times 100$$

Where C_0 is the initial concentration of the methylene blue solution and C_t is the concentration of the dye solution after t hours of exposure in solar irradiation. All dye concentrations were measured by taking the value of absorbance at 660 nm in the UV–Vis spectra as the concentration is directly proportional to the absorbance value.

The solar light was found to be more effective than other irradiation techniques for degrading dyes as reported by the previous studies (Kumar et al. 2013; Kansal et al. 2008). During exposure in sunlight, when the photons hit the nanoparticles present in the colloidal mixture, the electrons at the particle surface are excited (Yu et al. 2012). The dissolved oxygen molecules in the reacting medium accept the excited electrons from particle surface and are converted into oxygen anion radicals. These radicals break the

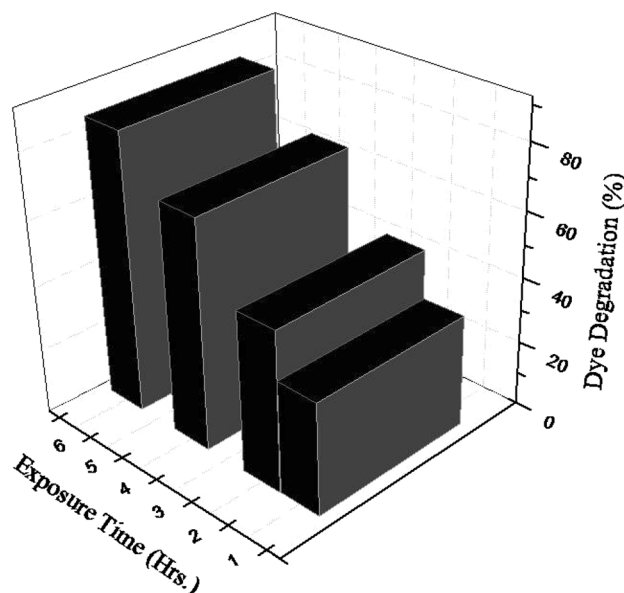


Fig. 8 3D bar graph shows the variation of dye degradation at different exposure time

organic dye into simpler organic molecules leading to the rapid degradation of the dye (Houas et al. 2001; Ameta et al. 2013). Therefore, the biosynthesized Ag nanoparticles may act as a stable and efficient photocatalyst for degradation of methylene blue under visible light irradiation (Vanaja et al. 2014).

Conclusion

Biological synthesis of metallic nanoparticles has emerged as a simpler and better option than physical and chemical procedures as it is fast, clean and eco-friendly alternative that does not involve any costly instruments as well. Here, we synthesized silver nanoparticles from silver nitrate solution using yeast (*Saccharomyces cerevisiae*) extract and the biosynthesized Ag nanoparticles were found to be almost spherical in shape with an average diameter of 10 nm. Crystalline nature of the particles was verified by

XRD, whereas the FTIR study detected the functional biomolecules that contributed to the reduction and capping of the Ag nanoparticles during interaction. The result of the photocatalytic study concludes that these biogenic silver nanoparticles have efficiency to degrade methylene blue under solar irradiation. Therefore, they can find applications in textile industry and water treatment plants.

Acknowledgments One of the authors (Kaushik Roy) thanks and acknowledges CSIR for financial assistance (Senior Research Fellowship).

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