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CORRECTION

Correction to: A Single-Molecule Magnet Tetranuclear $[Mn_3^{III}Mn^{IV}O_3Cl]$ Complex with Bis(diisopropylphosphinyl)imide Ligands

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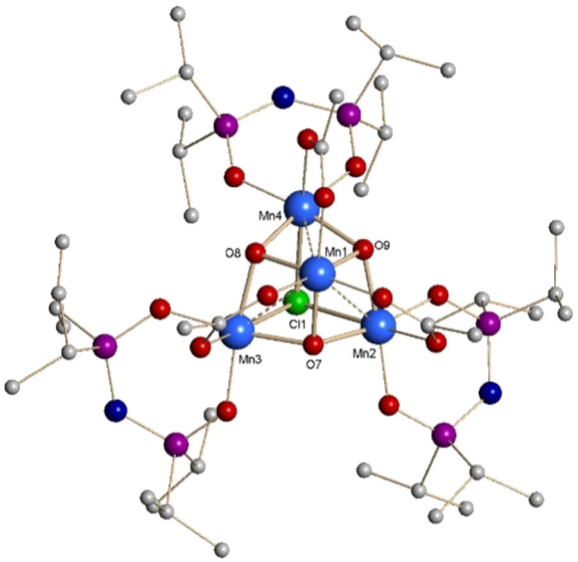
Abstract

Treatment of $Mn(OAc)_2{\cdot}4H_2O$ with 3 equiv. $HN(^iPr_2PO)_2$ in the presence of $^nBu_4NCl$ in acetonitrile resulted in isolation of a tetranuclear $\{Mn_3^{III}Mn^{IV}\}$ complex $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(\mu\text{-}OAc)_3\{N(^iPr_2PO)_2\}_3]{\cdot}0.5C_6H_{14}$ (**1**·0.5C_6H_{14}), and a dinuclear $\{Mn^{II}Mn^{III}\}$ complex $[Mn_2\{N(^iPr_2PO)_2\}\{\mu_2,\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2\{\eta^2\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2]{\cdot}0.25AcOH$ (**2**·0.25AcOH) with bis(diisopropylphosphinyl)imide ligands. The solid-state structures of two complexes have been established by single-crystal X-ray crystallography. The tetranuclear complex **1** belongs to a family of $\{Mn_3^{III}Mn^{IV}O_3X\}$ single-molecule magnets (SMMs) with effective energy barrier of 14.3 K by some further magnetic characterizations.

Graphic Abstract

A $[Mn_3^{III}Mn^{IV}O_3Cl]$-core-based $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(OAc)_3\{N(^iPr_2PO)_2\}_3]$ complex with single-molecule magnet (SMM) behaviour was prepared and structurally characterized, exhibiting slow magnetization relaxation under zero applied dc field with $\tau_0 = 9.0 \times 10^{-7}$ s and relaxation barrier $U_{eff} = 14.3$ K.



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The original version of this article unfortunately published with wrong author group. It also had several mistakes in text. The author group and errors are corrected in this erratum article.

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Keywords Manganese · Single molecule magnet (SMM) · Mixed-valence manganese complex · Bis(diisopropylphosphinyl)imide ligand · Crystal structure

Introduction

As the smallest known units with the ability of storing up to one bit of information per molecule at cryogenic temperatures, single-molecule magnets (SMMs) have attracted intense research activity in past decades [1–6]. Synthetic chemists have been able to synthesize various polynuclear manganese complexes, such as the dodecanuclear complexes $[Mn_{12}O_{12}(O_2CR)_{16}(H_2O)_x]$ (R = Me, Et, Ph; x = 3, 4) [7, 8], the polyoxometalate-based $\{Mn_2^{IV}Mn_6^{III}Mn_4^{II}\}$ complexes [9], a family of cationic oxime-based Mn_6^{III} complexes [10], and the high nuclearity mixed-valence $\{Mn_{32}\}$ complex [11], which show SMM features at the very lowest temperatures, but examples of homometallic Mn-based SMMS with blocking temperatures above 1 K are still relatively unusual. Among them is a family of tetranuclear mixed-valence manganese complexes with general formula $[Mn_4O_3X(OAc)_3(dbm)_3]$ and the core $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}X)]^{6+}$, where X = Br^-, Cl^-, OAc^- and F^-, OAc^- is the acetate ion and dbm^- is the anion of dibenzoylmethane [12]. These $\{Mn_3^{III}Mn^{IV}\}$ complexes possess a well-isolated ground-state spin of $S = {}^9/_2$, displaying hysteresis loops containing Quantum Tunnelling of the Magnetization (QTM) steps and exhibiting SMM behavior with an energy barrier of the order of 14.5 K [13]. On the other hand, stronger electron-donating phosphinate and phosphonate derivatives could stabilize the mixed-valence $\{Mn^{III}Mn^{II}\}$ complexes with large spin and coercivity [14–16]. With this in mind, we report synthesis, structure and magnetic properties of a tetranuclear $\{Mn_3^{III}Mn^{IV}\}$ complex with bis(diisopropylphosphinyl)imide ligands, $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(OAc)_3\{N(^iPr_2PO)_2\}_3]$ (**1**), which belongs to the above-mentioned $\{Mn_4O_3X\}$ family. Results of magnetic measurements show that the tetranuclear complex **1** exhibits SMM behavior.

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Experimental

Materials and Measurements

All the reagents were commercially purchased and used without further purification. The $HN(^iPr_2PO)_2$ ligand was synthesized according to the literature [17]. Elemental analyses were carried out using a Perkin-Elmer 2400 CHN analyzer.

Synthesis of $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(\mu\text{-}OAc)_3\{N(^iPr_2PO)_2\}_3]\cdot 0.5C_6H_{14}$ (1·0.5C_6H_{14}) and $[Mn_2\{N(^iPr_2PO)_2\}\{\mu_2,\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2\{\eta^2\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2]\cdot\cdot 0.25AcOH$ (2·0.25AcOH)

To a suspension of $Mn(OAc)_2{\cdot}4H_2O$ (95 mg, 1.0 mmol) in acetonitrile (10 mL) was added $HN(^iPr_2PO)_2$ (84 mg, 3.0 mmol) and $^nBu_4NCl$ (70 mg, 0.25 mmol) in the air. The reaction mixture was stirred overnight at room temperature. The volatile was removed by rotation evaporation. The brown residue was extracted with hexane and then filtered. Evaporation of the hexane solution afforded complex **1**·0.5C_6H_{14} as red plate-shaped crystals and complex **2**·0.25AcOH as pink prism-shaped crystals, both suitable for crystallography. Note that yields have not been optimized. Two complexes may be separated by a physical method in light of their different colors and shapes of crystalline products. Yield of **1**·0.5C_6H_{14}: 20 mg, 6% (based on Mn). Anal. Calcd. for $C_{45}H_{100}O_{15}N_3ClP_6Mn_4$: C, 39.62; H, 7.39; N, 3.08%. Found: C, 39.66; H, 7.42; N, 3.12%. Yield of **2**·0.25AcOH: 28 mg, 4% (based on Mn). Anal. Calcd. for $C_{48.50}H_{117}O_{14.50}N_5P_{10}Mn_2$: C, 40.96; H, 8.29; N, 4.92%. Found: C, 40.93; H, 8.26; N, 4.90%.

Crystal Structure Determinations

A summary of crystallographic data and experimental details for complexes **1**·0.5C_6H_{14} and **2**·0.25AcOH are listed in Table 1. Intensity data were collected on a Bruker

Table 1 Crystallographic data and experimental details for $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)\text{-}(\mu\text{-}OAc)_3\{N(^iPr_2PO)_2\}_3]\cdot 0.5C_6H_{14}$ (**1**·0.5C_6H_{14}), and $[Mn_2\{\text{-}N(^iPr_2PO)_2\}\{\mu_2,\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2\{\eta^2\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2]\cdot 0.25AcOH$ (**2**·0.25AcOH)

Complex	**1**·0.5C_6H_{14}	**2**·0.25CH_3CO_2H
Empirical formula	$C_{45}H_{100}ClMn_4N_3O_{15}P_6$	$C_{48.50}H_{117}Mn_2N_5O_{14.50}P_{10}$
Formula weight	1364.31	1422.04
Crystal system	triclinic	triclinic
Space group	$P\bar{1}$	$P\bar{1}$
a (Å)	12.1543 (15)	13.4952 (13)
b (Å)	20.683 (2)	14.5487 (14)
c (Å)	27.320 (3)	20.661 (2)
α (°)	107.965 (2)	110.344 (2)
β (°)	92.601 (2)	98.465 (2)
γ (°)	95.963 (2)	102.756 (2)
V (Å^3)	6476.0 (13)	3596.6 (6)
Z	4	2
D_{calcd} (g cm^{-3})	1.399	1.313
Temperature (K)	100 (2)	100 (2)
F(000) (*e*)	2872	1518
μ(Mo-Kα) (mm^{-1})	1.010	0.631
Refl. total	34,280	19,299
Refl. unique/R_{int}	22,366/0.0707	12,399/0.0594
Ref. parameters	1333	724
$R1$[a], $wR2$[b] [$I > 2\sigma(I)$]	0.0592, 0.0962	0.0600, 0.1174
$R1$[a], $wR2$[b] (all data)	0.1449, 0.1121	0.1215, 0.1345
GoF[c]	1.012	0.880
$\Delta\rho_{fin}$ (max/min) (*e* Å^{-3})	+ 0.703/− 0.548	+ 0.966/− 0.459

[a] $R1 = \Sigma||F_o|-|F_c||/\Sigma|F_o|$

[b] $wR2 = [\Sigma w(F_o^2-F_c^2)^2/\Sigma w(F_o^2)^2]^{1/2}$, $w = [\sigma^2(F_o^2) + (AP)^2 + BP]^{-1}$, where $P = (\mathrm{Max}(F_o^2, 0) + 2F_c^2)/3$

[c] $\mathrm{GoF} = S = [\Sigma w(F_o^2-F_c^2)^2/(n_{obs}-n_{param})]^{1/2}$

SMART APEX 2000 CCD diffractometer using graphite-monochromated Mo-Kα radiation ($\lambda = 0.71073$ Å) at 100(2) K. The data were corrected for absorption using the program SADABS [18]. Structures were solved by the direct methods and refined by full-matrix least-squares on F^2 using the SHELXTL software package [19, 20]. All non–hydrogen atoms were refined anisotropically except for the solvent molecules due to disorder. Hydrogen atoms in the phenyl and other organic moieties were treated as idealized contributions (C_{sp3}–H = 0.96, C_{sp2}–H = 0.93 Å, and N–H 0.86 Å). CCDC reference numbers 1844319 and 1844320 contain the supplementary crystallographic data for this paper. Copies of the data can be obtained free of charge on application to CCDC, 12 Union Road, Cambridge CB2 1EZ, UK [fax: (+44)1233-336-033; e-mail: deposit@ccdc.cam.ac.uk].

Magnetic Measurements

Magnetic data (ac and dc) were recorded on a Quantum Design MPMS-XL SQUID magnetometer equipped with a 7 T magnet and capable of operating in the 1.8–300 K temperature range. Ac magnetic susceptibility measurements were performed in an oscillating ac field of 3.5 G and a zero dc field. The oscillation frequencies were in the 10–1000 Hz range. Diamagnetic corrections to the observed susceptibilities were applied using Pascal's constant.

Magnetic studies below 1.8 K were carried out on the single-crystals using a micro-SQUID apparatus operating down to 0.04 K [21]. The samples were cooled at 10 K/min under 1 atm of helium gas. Temperature control was achieved within the variable-flow cryostat belonging to a Quantum Design PPMS system. A well-formed plate-like crystal was indexed, and mounted with the applied magnetic field normal to the dominant (010) face. The two independent cluster molecules have their respective

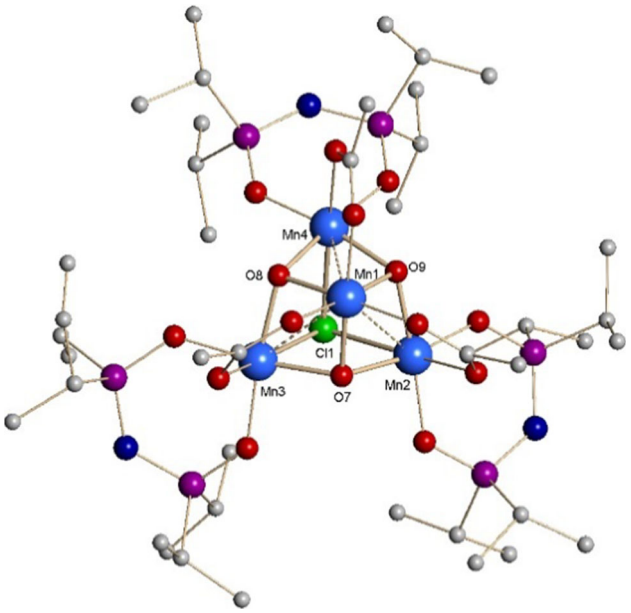

Fig. 1 Structure of one of the two Mn_4 molecules in **1**·0.5C_6H_{14}. Selected bond lengths (Å) and angles (deg): Mn1···Mn2, 2.8053(12); Mn1···Mn3, 2.8003(13); Mn1···Mn4, 2.8023(13); Mn2···Mn3, 3.2160(12); Mn3···Mn4, 3.2734(12); Mn2···Mn4, 3.2344(12); Mn2–Cl1, 2.6256(17); Mn3–Cl1, 2.6130(17); Mn4–Cl1, 2.6434(17); O7–Mn(1,2,3), 1.846(4), 1.920(4), 1.931(4), respectively; O8–Mn(1,3,4), 1.847(4), 1.936(4), 1.959(4), respectively; O9–Mn(1,2,4), 1.853(4), 1.941(4), 1.916(4), respectively; Mn2–Cl1–Mn3, 75.75(5); Mn2–Cl1–Mn4, 75.74(5); Mn3–Cl1–Mn4, 77.03(5); Mn1–O7–Mn2, 96.29(16); Mn1–O7–Mn3, 95.68(15); Mn2–O7–Mn3, 113.24(19); Mn1–O8–Mn3, 95.48(15); Mn1–O8–Mn4, 94.78(15); Mn3–O8–Mn4, 114.37(18); Mn1–O9–Mn2, 95.31(16); Mn1–O9–Mn4, 96.04(15); Mn2–O9–Mn4, 113.97(19)

threefold axes close to co-parallel, with both inclined by ca. 45° to the normal to the (010) face. Thus all the molecular threefold axes subtended an angle of ca. 45° to the applied magnetic field.

Results and Discussion

Syntheses and Crystal Structures

Treatment of $Mn(OAc)_2·4H_2O$ in acetonitrile with ca. 3 equiv. $HN(^iPr_2PO)_2$ in the presence of $^nBu_4NCl$ resulted in isolation of the tetranuclear $\{Mn_3^{III}Mn^{IV}\}$ complex $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(\mu\text{-}OAc)_3\{N(^iPr_2PO)_2\}_3]$·0.5$C_6H_{14}$ (**1**·0.5C_6H_{14}) as red crystals and the dinuclear $\{Mn^{III}Mn^{II}\}$ complex $[Mn_2\{N(^iPr_2PO)_2\}\{\mu_2,\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2\{\eta^2\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2]$·0.25AcOH (**2**·0.25AcOH) as pink crystals in 5–10% non-optimized yields. Complexes **1**·0.5C_6H_{14} and **2**·0.25AcOH may be physically separated according to their obviously different colors and shapes. The previously reported and structurally-similar complex $[Mn_4O_3Cl(OAc)_3(dbm)_3]$ was synthesized by treatment of $[Mn_4O_2(OAc)_6(py)_2(dbm)_2]$ with $^nBu_4NCl$ or Me_3SiCl in dichloromethane (py = pyridine) [22]. Single-crystal X-ray diffraction analysis reveals that structure of the tetranuclear **1**·0.5C_6H_{14} (red crystals) consists of a distorted cubane core $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)]$ with Mn1 as the Mn^{IV}. Bond valence sum analysis [25, 26] resulted in calculated valences of 4.05 and 4.11 for Mn1A and Mn1B, with the remaining Mn sites all having valences in the range 3.11–3.22 (Fig. 1). The asymmetric unit contains two independent cluster molecules, but these are closely isostructural and that containing Mn1 will be described here. Three AcO^- groups bridge each $\{Mn^{III}Mn^{IV}\}$ pair, and three chelating $[N(^iPr_2PO)_2]^-$ ions on each Mn^{III} complete the peripheral ligation. The $\{Mn_4O_3Cl\}$ core of complex **1**·0.5C_6H_{14} is essentially superimposable on those of other $\{Mn_4O_3Cl\}$ complexes [13, 22]. For example, the $Mn^{III}\cdots Mn^{III}$ separations [3.2288(12)–3.2891(12) Å] and $Mn^{III}\cdots Mn^{IV}$ distances [2.8002(13)–2.8048(12) Å] in complex **1**·0.5C_6H_{14} are similar to those in $[Mn_4O_3Cl(OAc)_3(dbm)_3]$ [$Mn^{III}\cdots Mn^{III}$ = 3.237(5)–3.264(4) Å, $Mn^{III}\cdots Mn^{IV}$ = 2.792(5)–2.797(4) Å] [22], showing that the complete exchange of oxygen-ligating ligands ($[N(^iPr_2PO)_2]^-$) for the previously present peripheral ligands (py, imidazole, Cl^-, dbm^-) has little structural effect. The Mn^{IV}–O bond lengths of 1.846(4)–1.853(3) Å are a little shorter than those of Mn^{III}–O bond lengths [1.916(4)–1.959(3) Å] in complex **1**·0.5C_6H_{14}. The P–O lengths [1.531(4)–1.541(4) Å] and P–N lengths [1.582(5)–1.599(5) Å] of the three $[N(^iPr_2PO)_2]^-$ groups in complex **1**·0.5C_6H_{14} indicate a good π-electron delocalized ligand system [23, 24].

The structure of $[Mn_2\{N(^iPr_2PO)_2\}\{\mu_2,\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)\}_2\{(^iPrPO_2)NH(^iPr_2PO)\}_2]$·-0.25AcOH (**2**·0.25AcOH) is a dinuclear $\{Mn^{II}Mn^{III}\}$ complex bridged by two $[\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)]^-$ ligands which were probably generated by partial hydrolysis of the

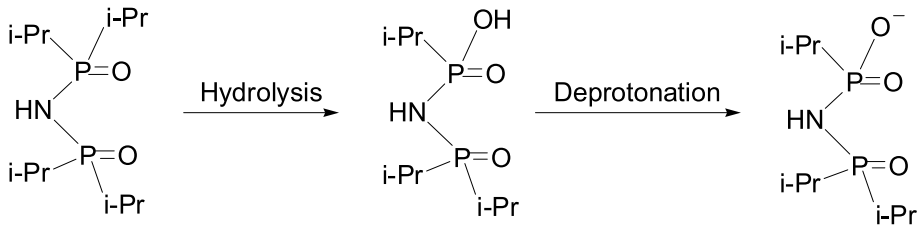

Fig. 2 Partial hydrolysis of the ligand $HN(^iPr_2PO)_2$ and its possible deprotonated form

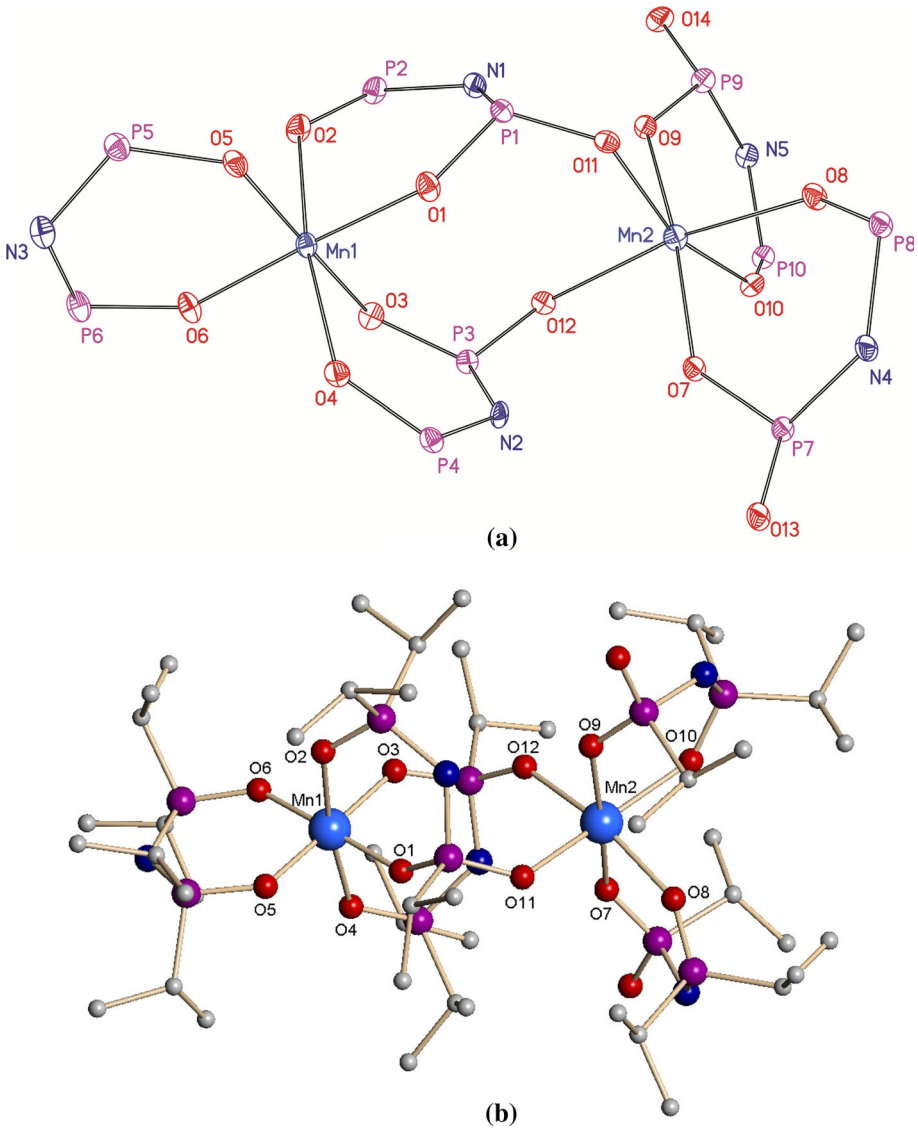

Fig. 3 **a** Framework of complex **2**·0.25CH_3CO_2H. Carbons, hydrogens an—d the acetic acid molecule were omitted. Selected bond lengths (Å) and angles (deg): Mn1–O(1,2,3,4,5,6), 1.898(3), 2.210(3), 1.901(3), 2.198(3), 1.899(3), 1.908(3), respectively; Mn2–O(7,8,9,10,11,12), 2.177(3), 2.181(3), 2.187(3), 2.160(3), 2.175(3), 2.202(3), respectively; P1–O(1,11), 1.511(3), 1.489(3), respectively; P2–O2, 1.496(3); P3–O(3,12) 1.527(3), 1.475(3), respectively; P4–O4, 1.490(3); P5–O5, 1.535(3); P6–O6, 1.525(3); P7–O(7,13), 1.518(3), 1.485(3), respectively; P8–O8, 1.490(3); P9–O(9,14), 1.507(3), 1.497(3), respectively; P10–O10, 1.495(3); O1–Mn1–O2, 86.57(12); O3–Mn1–O4, 87.11(12); O5–Mn1–O6, 92.33(13); O7–Mn2–O8, 82.62(11); O9–Mn2–O10, 82.51(11). **b** Structure of complex **2**·0.25CH_3CO_2H. Hydrogens and the acetic acid molecule were omitted

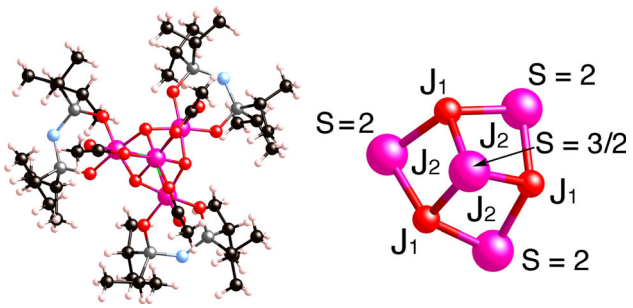

Fig. 4 **a** The structure of **1**·0.5C_6H_{14} viewed along the pseudo-C_3 axis. The Mn ions are shown in pink with the Mn^{IV} center. **b** Schematic representation of the exchange coupling in **1**·0.5C_6H_{14} with the principal exchange pathways J_1 and J_2

$HN(^iPr_2PO)_2$ ligand in wet solution (Fig. 2). The assignment of the oxidation states for the manganese ions were once again confirmed by bond valence sum calculations. The valence sums [25, 26] for Mn1 (3.09) and Mn2 (1.95) are in good agreement with the expected valences of +3 and +2. The Mn1 is six-coordinated by two oxygen atoms from the $[N(^iPr_2PO)_2]^-$ ligand and four oxygen atoms from two bridging $[\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)]^-$ ligands, while the Mn_2 is six-coordinated to two oxygen atoms from two bridging $[\eta^3\text{-}(^iPrPO_2)NH(^iPr_2PO)]^-$ ligands and four oxygen atoms from two chelating $[\eta^2\text{-}(^iPrPO_2)NH(^iPr_2PO)]^-$ ligands (Fig. 3). The average Mn_1–O and Mn_2–O lengths in **2**·0.25AcOH are 2.002(3) and 2.180(3) Å, respectively, which are consistent with the typical bond lengths of Mn^{III}–O and Mn^{II}–O, respectively [9]. The N–P bond lengths ranging from 1.649(4) to 1.684(4) Å in the $[(^iPrPO_2)NH(^iPr_2PO)]^-$ ligands are longer than those in the $[N(^iPr_2PO)_2]^-$ ligand (1.579(4) and 1.596(4) Å) in complex **2**·0.25AcOH, suggesting characteristic of single bond as

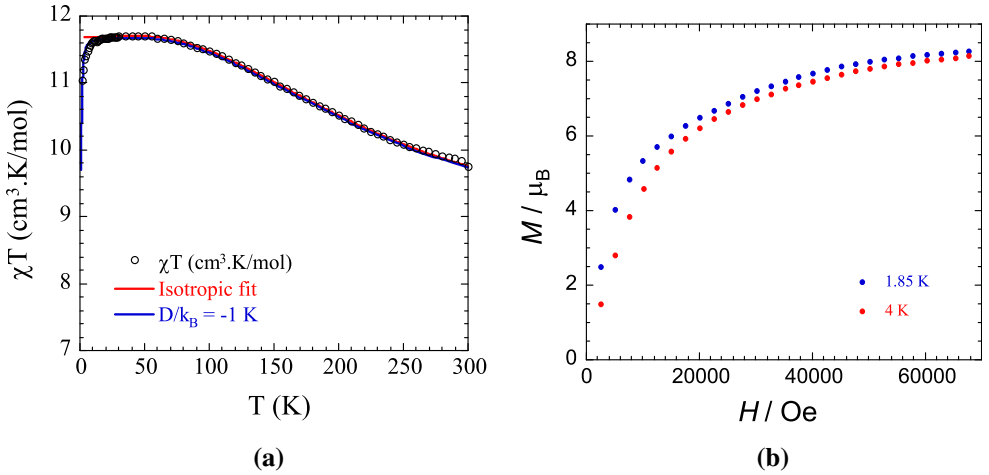

Fig. 5 **a** Plots of χT versus T for complex **1**·0.5C_6H_{14} in an applied field of 1000 Oe in the 2.0–300 K range. **b** Field dependence of the magnetization at 1.85 K and 4 K for complex **1**·0.5C_6H_{14}

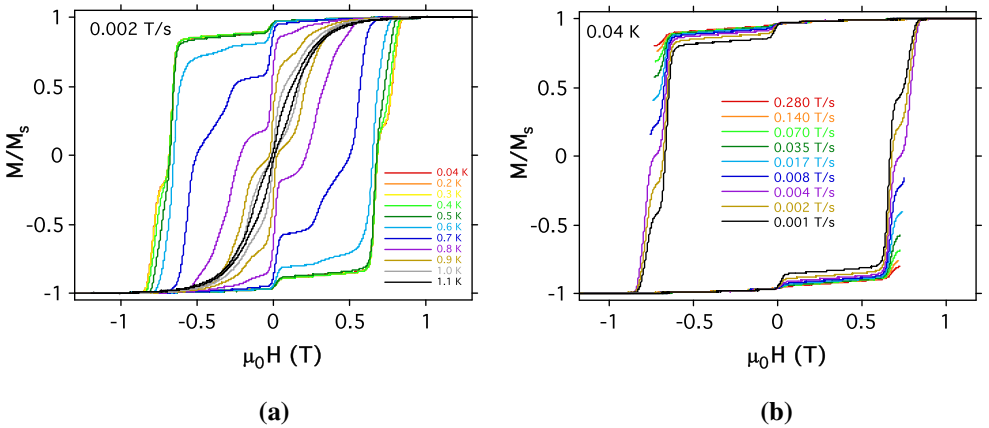

Fig. 6 Magnetization (M) versus dc field hysteresis loops for **1**·0.5C_6H_{14} at the indicated temperatures and a fixed field sweep rate of 0.002 T/s (**a**), and at the indicated field sweep rates and a fixed temperature 0.04 K (**b**) by the micro-SQUID technique with the magnetic field applied normal to the (010) face and at ca. 45° to the molecular threefold axes. The magnetization is normalized to its saturation value, M_S

shown in Chart 1. The Mn1···Mn_2 separation is ca. 5.327 Å in the non-planar eight-membered ring $Mn_2O_4P_2$.

Magnetic Properties

Based on the structure of tetranuclear complex **1**, the molecular point symmetry is approximately $C3v$ with the $C3$ axis passing through the Mn^{IV} and Cl^- ions (Fig. 4). The average angle (α) between the three JT elongation axes and the molecular anisotropy axis (i.e., the virtual C_3 axis) is ∼ 45.4° like in $[Mn_4O_3Cl(OAc)_3(dbm)_3]$ (45.1°) [22]. To model the magnetic properties using an isotropic Heisenberg model in the weak field approximation, we simulated the χT product taking into account only two isotropic intra-cluster magnetic interactions between Mn^{III} ($S = 2$) and Mn^{IV} ($S = {}^3/_2$) sites and neglecting the low temperature data (below 20 K) to avoid the influence of the anisotropy and weak inter-cluster interactions. The following spin-Hamiltonian is considered [22]:

$$H = -2J_1(S_1 \cdot S_2 + S_1 \cdot S_3 + S_2 \cdot S_3) - 2J_2(S_4 \cdot (S_1 + S_2 + S_3))$$

with $S_1 = S_2 = S_3 = 2$ for the Mn^{III} ions and $S_4 = 3/2$ for the Mn^{IV} ion, respectively. J_1 and J_2 are the Mn(III)-Mn(III) and Mn(III)-Mn(IV) exchange coupling constants, respectively. The fitting procedure using the deduced susceptibility gave J_1/k_B = 13.5(2) K, J_2/k_B = − 53.1(2) K and g = 1.94, the magnitude of the magnetic interactions are significantly higher than that in the related complex $[Mn_4O_3Cl(OAc)_3(dbm)_3]$ (11.9 K (8.3 cm^{-1}), − 40.8 K (− 28.4 cm^{-1}), 1.98) [22]. The value of χT for complex **1**·0.5C_6H_{14} increases from 9.86 cm^3 mol^{-1} K at 300 K to a maximum of 11.72 cm^3 mol^{-1} K at 25 K and then decreases rapidly to 11.00 cm^3 mol^{-1} K at 2.0 K (Fig. 5a).

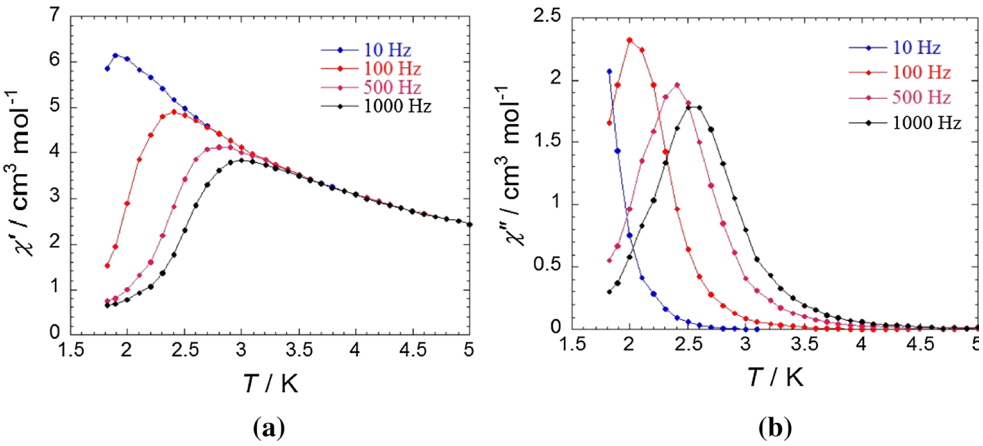

Fig. 7 Ac susceptibility of complex **1**·0.5C_6H_{14} in a 3.5 G field oscillating at 10, 100, 500, and 1000 Hz: in-phase signal (χ') and out-of-phase signal (χ'') versus T

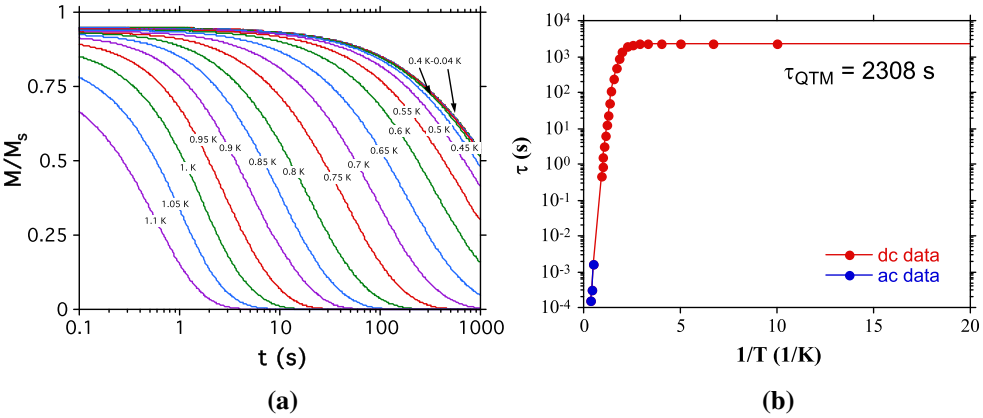

Fig. 8 **a** Relaxation of the magnetization below 1.1 K normalized to the initial magnetization at 0.04 K by the micro-SQUID technique with the magnetic field applied normal to the (010) face and at ca. 45° to the molecular threefold axes. **b** Relaxation time τ versus T^{-1} plot in zero ac and dc field

The maximum χT for **1**·0.5C_6H_{14} is slightly below the spin-only (g = 2) value expected for a complex with an $S = {}^9/_2$ ground state (12.38 cm^3 mol^{-1} K), and this is consistent with $g < 2.0$, as expected for Mn [27]. The fit of the temperature dependence of the χT product allows the molecular zero-field splitting parameter D to be determined for **1**·0.5C_6H_{14}. The accuracy of this parameter is rather low; we estimated $D = -1.0(4)$ K. This is comparable to values from other Mn4O3X SMMs (typically D/kB = − 0.7 K) [2, 27]. The simulation with the fitting with D/kB = −1 K is shown in Figure 5a. It is known that a compound with a large negative D value will show a much lower magnetization at a given field than for Brillouin behaviour. As shown in Fig. 5b, the magnetization value at 7 T of 8.30 μ_B for complex **1**·0.5C_6H_{14} at 1.85 K is less than the expected saturation magnetization value of 9 μ_B.

Hysteresis loops are observed below 1.2 K, confirming the negative value for D (easy-axis behavior) (Fig. 6). These data are strongly temperature and sweep-field rate dependent as expected for the SMM. Nevertheless below 0.5 K, the hysteresis loops at a given scan rate become temperature independent as expected for a Quantum Tunnelling regime. Taking the data at the lowest temperature available: 0.04 K, steps are clearly observed indicative a Quantum Tunnelling of the Magnetization (QTM, in zero-field). Slow relaxation of the magnetization is also observed by ac susceptibility measurements. Traditional feature of the SMM is observed with strongly frequency dependence in-phase and out-of-phase ac susceptibilities (Fig. 7). Using these data in addition to the direct relaxation of the magnetization, the relaxation time can be deduced. The same shape of M/M_s versus t curves (where M/M_s is the magnetization normalized to its saturation value) was obtained for the whole temperature range (0.04–1.1 K) (Fig. 8a) and therefore the data were scaled into a single master curve [8]. The relaxation time was extracted simply taking the time when M/M_s reaches the value $1/e$. As indicated by the magnetization measurements, a pure QTM regime is observed below 0.4 K, similar to that of 0.3 K for $[Mn_4]_2$ dimer [2] and 0.6 K for $[Mn_4O_3Cl(OAc)_3(dbm)_3]$ [22], with a characteristic time τ_{QTM} = 2308 s (Fig. 8b). The effective energy gap of the thermally activated regime (that can also be influenced by QTM) is about 14.3 K and the τ_0 is about 9×10^{-7} s.

In summary, two mixed-valence Mn_2 and Mn_4 complexes with bis(diisopropylphosphinyl)-imide ligands were

synthesized and isolated with the crystalline state. The Mn_4 complex $[Mn_4(\mu_3\text{-}O)_3(\mu_3\text{-}Cl)(OAc)_3\{N(^iPr_2PO)_2\}_3]$ presents the typical feature of SMM and its D/k_B value can be estimated at $-1.0(4)$K. Isolation of the tetranuclear mixed-valence Mn_4 complex may provide a new route for assembly of novel mixed-valence manganese complexes based on Mn_4 cores and new type of phosphinyl-surrounded Mn_4 SMM materials.

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