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ERRATUM

Erratum to: Dehydrogenation of 2-Phenyl-1-pyrroline with Palladium-Supported Catalysts: An Effective Route to the Synthesis of 2-Phenylpyrrole

Cláudia A. Figueira[1] · Pedro T. Gomes[1]

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The original version of this article unfortunately contained an error in Table 1. The unit mentioned in the header of last column in Table 1 (TOF) should be changed to min^{-1} instead of s^{-1}. The correct table is given below.

The online version of the original article can be found under doi:10.1007/s10562-014-1469-0.

✉ Pedro T. Gomes
pedro.t.gomes@tecnico.ulisboa.pt

[1] Centro de Química Estrutural, Departamento de Engenharia Química, Instituto Superior Técnico, Universidade de Lisboa, Av. Rovisco Pais, 1, 1049-001 Lisboa, Portugal

Table 1 Reaction data (conversions, yields and selectivities) obtained for the catalytic dehydrogenation of 2-phenyl-1-pyrroline to 2-phenylpyrrole with Pd-supported catalysts

Reaction no.	Catalytic Support[a]	Supplier	Pd[b] (mol%)	Time (h)	Conv.[c] (%)	Yield[d] (%)	Select.[e] (%)	TOF[f] (min^{-1})
C1[g]	C	Merck	10.00	115	42	42	100	0.001
C2	C	Merck	10.00	3.0	94	81	86	0.056
C3	C	Acrös	1.25	5.8	92	78	85	0.230
C4	C	Merck	1.00	4.5	94	81	86	0.370
C5	C	Acrös	0.40	28	86	77	90	0.149
C6	C	Acrös	0.10	43	52	49	93	0.388
Al1	Al_2O_3	Fluka	1.00	1.5	97	79	82	1.111
Al2	Al_2O_3	Fluka	0.40	1.5	98	90	93	2.778
Al3	Al_2O_3	Fluka	0.10	57	76	64	84	0.292

Reaction conditions: $n_{pyrroline}$ = initial number of moles of 2-phenyl-1-pyrroline (13.8 mmol, 2 g); V_{xylene} = 40 ml; T = 150 °C

[a] Supports: activated carbon (C) and alumina (Al_2O_3)

[b] mol % Pd = $[n_{Pd}/n_{pyrroline}] \times 100$; n_{Pd} = moles of Pd in the supported catalyst

[c] Conversion of 2-phenyl-1-pyrroline (determined by ^{1}H NMR spectroscopy)

[d] Yield in 2-phenylpyrrole (determined by ^{1}H NMR spectroscopy)

[e] Selectivity in 2-phenylpyrrole (determined by ^{1}H NMR spectroscopy)

[f] TOF = turnover frequency = moles of 2-phenyl-1-pyrroline converted/(moles Pd × time)

[g] Reaction conditions: $n_{pyrroline}$ = initial number of moles of 2-phenyl-1-pyrroline (6.89 mmol, 1 g); $V_{toluene}$ = 10 ml; T = 120 °C